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Model reactions and natural occurrence of furans from hypersaline environments

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Abstract

5 Volatile organic compounds like furan and its derivatives are important for atmospheric properties and reactions. In this paper the known abiotic formation of furan from catechol under Fenton-like conditions with Fe^{3+} sulphate was revised by the use of a bispidine Fe^{2+} complex as a model compound for iron with well-known characteristics. While total yields were comparable to those with the Fe^{3+} salt, the turnover numbers of the active iron species increased. Additionally, the role of iron and pH will be discussed during furan formation from model compounds and in natural sediment and water samples collected from the Dead Sea and several salt lakes in Western Australia. Various alkylated furans and even traces of halogenated furans (3-chlorofuran and 3-bromofuran) were found in these samples. Furthermore, the emission of furans is compared to the abundance of several possible precursors such as isoprene and aromatic hydrocarbons as well as to the related thiophenes.

10 It is assumed that the emissions of volatile organic compounds such as furans contribute to the formation of ultra fine particles in the vicinity of salt lakes and are therefore important for the local climate.

1 Introduction

20 Furans are heterocyclic aromatic compounds with a five-membered ring consisting of four carbons and one oxygen atom and are closely related to the sulphur containing thiophenes. With manifold natural and anthropogenic sources furans are an important class of volatile organic compounds (VOC).

25 Furan and several mono- and dialkylated furans were found in forest soils applying pyrolysis GC/MS (Hempfling and Schulten, 1990) and furfural as well as 2-pentylfuran in soil samples containing plant litter (Leff and Fierer, 2008; Jordan et al., 1993). Plant material was confirmed as a source of furans by several studies including living

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plants (Isidorov et al., 1985), leaf litter (Isidorov and Jdanova, 2002) and solid woods (Risholm-Sundman et al., 1998).

Huber et al. (2010) showed that soils are also an abiotic source for furans and presented a formation pathway producing furans from catechol and similar precursors using Fenton-like reaction conditions.

In this study the biomimetic nonheme bispidine $[\text{Fe}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$ complex (Fig. 1) was used as model compound for natural iron complexes to investigate abiotic furan formation under Fenton conditions. Bispidine complexes have a defined coordination sphere due to the rigid adamantane backbone (Comba et al., 2010) and are well studied in their coordination with catechol (Comba et al., 2011) as well as their redox reaction with hydrogen peroxide (Bukowski et al., 2006). Aside from that, such chelating agent can be used to enhance the reaction and to impede side reactions.

Further sources of furans are biomass burning (Christian et al., 2003) and food processing (Rogge et al., 1991). Atmospheric oxidation reactions of 1,3-butadiene and 1,3-pentadiene yield furan and 2-methylfuran (Ohta, 1984), respectively, while isoprene, the most abundant volatile biogenic compound, is oxidised to 3-methylfuran (Atkinson et al., 1989). Moreover, anthropogenic benzene (Berndt and Boge, 2001), toluene (Dumdei et al., 1988) and xylenes (Shepson et al., 1984) are degraded to furans by OH-radical reactions in the atmosphere.

According to the Maximum Incremental Reactivity (MIR) scale furans play an important role in the formation of ground-level ozone (Carter, 2009). In addition, furans have an impact on the oxidative budget of the lower troposphere reacting with NO_3 -radicals at nighttime (Kind et al., 1996; Berndt et al., 1997) and with OH-radicals (Bierbach et al., 1995) or ozone (Atkinson et al., 1983) at daytime. The atmospheric lifetime of furan is about 4 h in the presence of OH-radicals and decreases with higher alkylation (Bierbach et al., 1995; Gu et al., 1985). The reaction rate of 3-methylfuran with OH-radicals is similar if not even higher than of its precursor isoprene (Atkinson et al., 1989; Gu et al., 1985). Furthermore, photooxidation of furans with nitrous acid (HONO) as a precursor for OH-radicals confirmed the formation of secondary organic aerosol (SOA)

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(Gómez Alvarez et al., 2009). Such newly formed particles at nanometre scale are of significant importance in the troposphere altering cloud microphysics. They act as additional cloud condensation nuclei (Pierce and Adams, 2007), increasing the amount of new small droplets in clouds while reducing the size of old droplets leading to a change in the radiative properties and to rain suppression (Khain et al., 2005).

Another atmospheric sink for furans, mainly attributed to the marine boundary layer, is the reaction with Cl-radicals being even faster compared to OH-radicals (Villanueva et al., 2009; Cabañas et al., 2005) and possibly a source for halogen-induced organic aerosol (XOA) (Ofner et al., 2013). Reactive chlorine species are also abundant in the vicinity of various salt lakes in arid and semi-arid environments (Stutz et al., 2002). These terrestrial sources for atmospheric chlorine and other halogens become even more important with the expansion of salt lakes and salt influenced soils around the globe due to anthropogenic salinisation (Williams, 1999). One of the most prominent salt lakes is the Dead Sea being part of this study on the natural occurrence of furans. The Dead Sea is of particular interest because of the high atmospheric abundance of reactive bromine (Matveev et al., 2001) and iodine species (Zingler and Platt, 2005). Additional sample locations were various salt lakes in the Western Australian wheat belt. These salt lakes are of peculiar interest because of their geochemical properties and their atmospheric impact. On the one hand, chemical parameters such as iron content and pH vary significantly between adjacent lakes (Bowen and Benison, 2009) and on the other hand, the formation of ultra fine particles was measured in the vicinity of the salt lakes which are assumed to be responsible for decreasing rainfall in the area over the last decades (Junkermann et al., 2009).

Supplementary, the investigated salt lakes are situated in regions favouring high seasonal temperatures and elevated solar radiation. The latter is important for a constant supply of H₂O₂ in surface waters and top soils (Cooper et al., 1988) which is essential in Fenton reactions including abiotic furan formation. H₂O₂ is also formed in the atmosphere and is transported to the surface by rain events, especially during thunderstorms (Deng and Zuo, 1999; Gunz and Hoffmann, 1990).

The natural occurrence of furans in salt lake sediments including the abiotic pathway provides the basis to understand the atmospheric impact of saline environments with regard to particle formation and local climate changes.

The first part of the paper concerns the abiotic furan formation using a bispidine Fe^{2+} complex in a Fenton reaction with catechol and the second part deals with the abundance and the underlying conditions of furan emissions from natural salt lake sediments and waters.

2 Experimental section

2.1 Chemicals

The following chemicals were used: catechol (99 %; Sigma-Aldrich), 4-ethylcatechol (95 %; Aldrich), hydrogen peroxide (30 %; Merck), $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$ (Börzel et al., 2002), Fe^{3+} -sulphate (99 %; Fe 21–23 %; Riedel-de Haën), furan (99 %; Fluka), 2-methylfuran (99 %; Sigma-Aldrich), 2-ethylfuran (97 %; Aldrich), 2-propylfuran (97 %; abcr), 2-butylfuran (98 %; abcr), 2-pentylfuran (97 %; Aldrich), 2,3-dimethylfuran (99 %; Aldrich), 2,5-dimethylfuran (99 %; Aldrich), 3-chlorofuran (Zhang et al., 2005), 3-bromofuran (97 %, Aldrich), thiophene (99 %; Aldrich), 2-methylthiophene (98 %; Aldrich), 3-methylthiophene (99 %; Aldrich), Isoprene (99 %; Aldrich), EPA 624 calibration mix B (analytical standard; Supelco), sulphur mix (analytical standard; Spectrum Quality Standards), potassium chloride (99.5 %; Merck), sodium hydroxide (99 %; Aldrich), sulphuric acid (96 %; Riedel-de Haën), nitric acid (65 %; Bernd Kraft), calcium carbonate (99 %; LECO). Type I ultrapure water ($> = 18 \text{ M}\Omega \text{ cm}^{-1}$) from a Purelab UHQ System by ELGA LabWater was used in all experiments.

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2.2 Sediment samples

49 samples from Australia in 2011, 30 samples from Australia in 2012 and 13 samples from the Dead Sea, Israel, in 2012 were analysed for various chemical parameters and the natural abundance of furanoic compounds. The samples were taken from salt lakes or salt influenced sediments from various depths and were stored cooled for transportation. The soil samples were freeze-dried for long time storage and in order to remove residues of volatile compounds in a Lyovac GT2 by Steris. Afterwards, the samples were ground using a Pulverisette 5 planetary mill by Fritsch to a grain size below 315 μm for homogenisation.

2.3 Water samples

The water samples comprised eight samples from Australia in 2011, eight samples from Australia in 2012 and seven samples from the Dead Sea in 2012. For analysis, triplicates of 10 mL lake water were transferred into 20 mL headspace vials, closed with air-tight caps on-site and stored at 4 °C until measurement. Additionally, water samples for iron analysis were sampled and stabilised by acidification.

2.4 Instrumentation

For the analysis of VOC (b.p. -24 °C to 200 °C) a Varian gas chromatograph GC 3400 linked to a Varian Saturn 4D with electron impact ionisation and ion trap mass spectrometer was used. The GC was equipped with a J&W Scientific DB-5 (60 m; 0.32 mm i.d.; 1 μm film thickness) capillary column by Agilent Technologies. Dynamic injection and preconcentration was performed by a customised purge and trap system (Mulder et al., 2013).

The total iron content of the sediment samples was measured by X-ray fluorescence using an energy-dispersive miniprobe multielement analyser (EMMA) (Cheburkin et al., 1997) while dissolved iron of the water samples was quantified with an ICP-OES Vista

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MPX system by Varian. The C_{org} content of the sediments was calculated by subtraction of inorganic carbon content, analysed with a carbonate bomb (Müller and Gastner, 1971), from total carbon content analysed with a SC-144DR carbon/sulphur analyser by Leco. Dissolved organic carbon of the water samples from Australia was determined using a TOC-5000 by Shimadzu. The pH measurements were conducted with a Sentix 41 electrode from WTW calibrated on the free hydrogen scale with buffers by Merck at pH 4.01 and 7.00.

2.5 Experimental procedure

Model reactions, emissions from natural samples and calibration measurements were conducted in 20 mL air-tight headspace glass vials with 10 mL aqueous phase and 10 mL gas phase connected via two stainless steel needles to the GC/MS system. Before analysis the vials were shaken on a MTS 2 rotary board by IKA at 500 rpm in the dark at a controlled temperature in a B10 incubator by Memmert for a defined time. Preceding GC, the analytes from the headspace vials were focussed and pre-concentrated on a glass lined tubing at -190°C . The purge flow of the helium carrier gas was 7 mL min^{-1} during the optimisation steps of the model reactions and 15 mL min^{-1} during the analysis of natural samples. After 7.5 min the analytes were transferred from the purge and trap system by helium (15 psi precolumn pressure) onto the GC column. The oven program during the measurements of furan formation in model reactions was: 35°C held 8.30 min, 35°C to 150°C at $5.5^{\circ}\text{C min}^{-1}$, 150°C held 5 min, 150°C to 210°C at $30^{\circ}\text{C min}^{-1}$, 210°C held 15 min. For the natural samples and the Fenton-like reaction with 4-ethylcatechol the GC program was modified: 30°C held 15.5 min, 30°C to 114°C at $5^{\circ}\text{C min}^{-1}$, 114°C to 210°C at $30^{\circ}\text{C min}^{-1}$, 210°C held 8 min. All GC/MS measurements were done at least in triplicates.

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2.5.1 Model reactions

As in previous studies, catechol, being a model compound for natural organic matter, was used in the following oxidation reactions with iron and hydrogen peroxide to comprehend abiotic furan formation processes. In all experiments the catechol concentration amounted to 5 μmol in a 10 mL (0.5 mM) aqueous phase while the concentration of iron, H_2O_2 , KCl as well as reaction time, pH value and temperature were changed to yield an optimal furan formation.

2.5.2 Optimisation of $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$

To determine the optimal iron concentration for furan formation, 10 mL aqueous solutions containing 0.5 mM catechol, 1 mM H_2O_2 and 5 mM KCl were assayed with 0 to 0.1 mM bispidine Fe^{2+} complex. The solution had a pH of around 4.6 and was incubated for 30 min at 40 $^\circ\text{C}$.

2.5.3 Optimisation of hydrogen peroxide

Optimal H_2O_2 concentration for the reaction was evaluated between 0 and 8 mM while the other parameters were fixed at 12.5 μM $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$, 5 mM KCl, pH of 4.6 and 30 min incubation time at 40 $^\circ\text{C}$.

2.5.4 Time dependence

With the optimal concentrations for Fe^{2+} at 12.5 μM and H_2O_2 at 2 mM and the initial settings of the solution pH of 4.6 as well as the incubation temperature of 40 $^\circ\text{C}$ the reaction time was tried in turn between 0 and 24 h with 5 mM KCl.

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2.5.5 Optimisation of pH

The pH was adjusted with sulphuric acid or sodium hydroxide. The optimal reaction conditions were evaluated between a pH of 2.1 and 8.7 with concentration of 12.5 μM $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$, 2 mM H_2O_2 and 5 mM KCl. Incubation was accomplished at 40 °C for 30 min.

2.5.6 Temperature dependence

The temperature was changed between 20 and 70 °C while the other parameters were held at 12.5 μM $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$, 2 mM H_2O_2 and 5 mM KCl, a solution pH of 4.6 and an incubation time of 30 min. The furan calibration was adjusted to the varying temperatures to account for the change in furan vapour pressures.

2.5.7 Optimisation of chloride

With optimal concentrations of 12.5 μM $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$, 2 mM H_2O_2 , a solution pH of 4.6 and an incubation time of 30 min at 40 °C the KCl concentration was varied between 0 and 50 mM.

2.5.8 Reaction with 4-ethylcatechol

Furan formation with longer alkyl chains was tested under Fenton-like conditions using 50 μM Fe^{3+} sulphate with 0.35 mM H_2O_2 and 0.5 mM 4-ethylcatechol as the precursor.

2.5.9 Sediment samples

1 g milled sediment was suspended with 10 mL of water in a headspace vial and incubated at 40 °C for 24 h in the dark on the rotary board prior to measurements with GC/MS.

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2.5.10 Water samples

The on-site sampled and at 4 °C stored 10 mL lake waters were incubated in their respective headspace vials for 1 h at 40 °C in the dark on the rotary board.

2.6 Identification and quantification

5 The analytes were identified by comparison of the mass spectra with the National Institute of Technology MS library and with those obtained from commercial standards including their retention times. External multipoint calibrations were performed for all commercially available analytes and the synthesised 3-chlorofuran. In case of 3-methylfuran and 2,4-dimethylfuran recovery rates analogous to their isomers were assumed. Additionally, the known ionisation energies of 3-methylfuran and 2-methylfuran are comparable with 8.64 eV and 8.56 eV, respectively (Spilker and Grutzmacher, 10 1986).

3 Results and discussion

15 In the following the results of the model reactions and from the natural samples are presented and discussed. Geographical locations, geochemical properties and emission data from the natural samples listed in tabular form can be found in the Supplement.

3.1 Model reactions

20 In the course of the optimisation steps on furan formation from catechol an optimal $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$ concentration of 12.5 μM was determined (Fig. 2a). In the absence of iron no detectable amount of furan was formed expressing the significance of iron in this reaction. Higher concentration led to a moderate depletion of furan concentrations levelling off in a steady plateau.

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The optimal $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$ concentration accounts to a molar ratio of 1 : 40 compared to 0.5 mM catechol deployed in the experiments. In previous studies with Fe^{3+} sulphate by Huber et al. (2010) the optimal molar ratio of iron to catechol was 1 : 5. Hence, the turnover numbers are eight times higher than for the applied bispidine Fe^{2+} complex. This demonstrates not only the feasibility of the Fe^{2+} complex but also its superiority over Fe^{3+} sulphate. While the bispidine Fe^{2+} complex is stabilised in solution by its ligands, the Fe^{3+} salt is susceptible to precipitation and is removed from the reaction cycle. This cycle includes the oxidation of Fe^{2+} by H_2O_2 in the Fenton reaction to Fe^{3+} (Remucal and Sedlak, 2011) and forming a highly reactive OH-radical and a hydroxyl anion (Haber and Weiss, 1932). Fe^{3+} on the other hand can be either reduced by catechol, which is a one to two electron donor, or by the Fenton-like reaction with H_2O_2 yielding in both cases Fe^{2+} . In the former redox reaction catechol is either oxidised to 1,2-semiquinone in an one electron transfer or to 1,2-benzoquinone in a two electron transfer. Additionally, the Fenton-like reaction yields a superoxide radical and a proton.

Following the iron optimisation step, H_2O_2 was the next important factor in this reaction. Again, as in the case of iron, the abundance of H_2O_2 is essential for the furan formation in this reaction. At high concentrations the furan yield declined, probably due to the promotion of side- and follow-up reactions leading to higher oxidised products (Studenroth et al., 2013). The optimal concentration was determined at 2 mM (Fig. 2b) compared to 0.5 mM catechol giving a molar ratio of 4 : 1 while in the previous study with Fe^{3+} sulphate the molar ratio was at 1 : 1.4 (Huber et al., 2010).

Reaction time under these conditions with a prominent optimal furan yield was 30 min (Fig. 2c). Afterwards a decline can be accounted for by an equilibrium of furan between the gas and aqueous phase. In the latter, furan can further react with hydroxyl radicals comparable to reactions in the gas phase (Gómez Alvarez et al., 2009; Atkinson et al., 1983).

The reaction was susceptible to pH variations of the solution with an optimum at 4.6 (Fig. 2d). At lower pH the reaction of $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$ and H_2O_2 is inhibited due

to slower reaction rates between iron and hydrogen peroxide while at higher pH iron precipitation takes place and OH-radical formation is inhibited (Remucal and Sedlak, 2011).

The initial temperature of 40 °C for the experiments was also the optimal temperature (Fig. 2e). Lower temperatures are accompanied by lower energy input to the reaction while higher temperatures again could promote side and follow-up reactions.

While chloride had seemingly no influence on the reaction of Fe³⁺ sulphate with H₂O₂ and catechol (Huber et al., 2010), the reaction with the bispidine Fe²⁺ complex is inhibited (Fig. 2f) due to the radical scavenger traits of chloride (Grebel et al., 2010). Furan yields decreased linearly with increasing amounts of chloride. Highest furan yields were attained in the absence of chloride.

Overall, the bispidine Fe²⁺ complex is a feasible model compound with furan yields of 3.17 nmol (2.15 µg) from 5 µmol catechol under optimal conditions. This accounts for a turnover rate of 0.6 % from catechol and is comparable with the overall yield of the previous study using Fe³⁺ sulphate (Huber et al., 2010).

In the former studies only methylated catechols or phenols were used in Fenton-like reactions leading to mono- or dimethylated furans depending on the educts (Huber et al., 2010). Due to the fact that furans with extended alkyl chains were found in natural environments the Fenton-like reaction was extended to 4-ethylcatechol as a commercially available and soluble catechol derivative. In the Fenton-like reaction of Fe³⁺ sulphate with 4-ethylcatechol the presumed 3-ethylfuran was obtained.

3.2 Sediment samples

The model reactions were accompanied by a survey spanning over two years in which 92 sediment samples from different salt lakes in Western Australia and of the Dead Sea in Israel were collected and analysed for the emission of furanoic compounds. Additionally, the emission data were correlated to iron concentration and organic carbon content, as well as pH to compare the results from these natural samples with the model reactions. Fe²⁺ as used in these investigations was abundant in the Australian

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5 samples identified by the chelating agent 2,2'-bipyridine. Besides black sediment layers containing pyrite the wet sediments, ground and surface waters contained soluble Fe^{2+} (Shand et al., 2008; Degens and Shand, 2010). On the one hand, total iron content ranged from 0.46 to 3.6 w%, organic carbon from below 0.1 w% up to 5.9 w% and pH from 4.2 to 8.4 for the Australian samples displaying a wide divergence, while for the Dead Sea samples on the other hand, total iron content was between 1.2 to 1.8 w%, organic carbon concentration was below 0.1 w% up to 0.56 w% and pH was between 7.7 and 8.1 (Fig. 3a).

10 As expected from a previous work by Huber et al. (2011), furan as well as 2-methylfuran and 3-methylfuran were emitted from most sediment samples. Additionally, 2,3- 2,4- and 2,5-dimethylfuran, 2-ethylfuran, 2-propylfuran, 2-butylfuran and 2-pentylfuran were identified from various sediments. Most of these compounds are known to be products from food processing, plants and woods but are rarely mentioned in connection with soils and sediments. Furan and the two methylfuran isomers are the most prominent furan species in the samples followed by further monoalkylated furans, while the concentrations of dimethylated furans were significantly lower (Fig. 3b).

15 From all above mentioned furan derivatives, furan shows the most prominent dependency to iron and organic carbon content as well as pH (Fig. 4a–c). Hence, emissions tend to be higher with elevated iron and carbon concentrations at lower pH which is in good agreement with the model reactions where iron is needed in the redox reactions taking place as well as the low pH between 4 and 5.

The concentrations of alkylated furans tend to increase with the organic carbon content but seem to be rather independent of pH or iron concentration.

25 For one sample a time dependency for the emission of VOCs was established between 1 and 96 h incubation time (Fig. 5a). It shows an exponential increase of the furan concentration over time revealing that the formation of compounds with low boiling points takes place during incubation rather than being remnants in the samples after freeze-drying.

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Similar to catechol further aromatic compounds like phenols and other aromatic hydrocarbons react with hydroxyl radicals forming furans. From atmospheric research it is known that toluene is a precursor for 2-methylfuran (Shepson et al., 1984) and benzene for furan (Berndt and Boge, 2001), respectively. It can be assumed that ethyl benzene is a precursor for 2-ethylfuran. From data plots a correlation between benzene concentrations and furan emissions can be deduced (Fig. 5c). A coherence between toluene abundance and 2-methylfuran emission is ambiguous (Fig. 5d) and there is no correlation between ethylbenzene and 2-ethylfuran (Fig. 5e).

The oxidation of isoprene with hydroxyl radicals in the gas phase yields small amounts of 3-methylfuran (Atkinson et al., 1989). While isoprene was detected in the headspace with concentrations up to 111 ng g^{-1} the data plot shows no obvious relation between isoprene and 3-methylfuran (Fig. 5f). Additionally, the concentrations of 3-methylfuran and 2-methylfuran after incubation display a linear connection demonstrating a possible mutual origin (Fig. 5b).

Furthermore, the emissions of other heterocyclic aromatics such as thiophene, 2- and 3-methylthiophene plotted against the respective furans show a linear dependency between these compounds (Fig. 6a–c). Hence, this suggests a mutual precursor and a competition between thiol and hydroxyl groups during the reaction.

Additionally, several furans not mentioned before in the study of Huber et al. (2010) were found in the Australian and Dead Sea samples, including a series of alkylated furans from ethylfuran to pentylfuran. While the 2-alkylfuran isomers were abundant, the 3-alkylfuran isomers were found in none of the samples.

Special attention deserves the finding of 3-chlorofuran (up to 0.37 ng g^{-1}) in three of the Australian samples. 3-chlorofuran was a compound found by Huber et al. (2010) during the reaction of Fe^{3+} with hydrogen peroxide and chlorinated catechols and phenols. Hypothetical atmospheric chlorination by chlorine atoms and sorption on the sediment matrix prior sampling are unlikely due to the unfavourable energy states of the formed intermediate (Zhang and Du, 2011). Hence, the source should be terrestrial in nature.

The Dead Sea samples show an overall lower emission compared to the Australian samples which is in good compliance with lower organic carbon and iron content as well as higher pH.

Even though formation processes are more complex in sediments due to the nature of the matrix and to competing redox reactions than in model reactions, the furan formation shows a clear dependency on iron content and pH which were also important parameters in the model reactions with catechol.

3.3 Water samples

During the three campaigns 23 water samples were collected in air tight vials on-site. Dissolved iron content for the Australian samples ranged from below the detection limit (0.2 mg L^{-1}) to 173 mg L^{-1} correlating with a pH between 8.7 and 2.4. In all of the Dead Sea samples the iron content was below the detection limit (0.2 mg L^{-1}) while the pH was between 5.1 and 6.5 (Fig. 7a).

Water analysis showed the abundance of the different furan species that were also found in the sediment samples (Fig. 7b) including traces of 3-chlorofuran with concentrations up to 0.35 ng mL^{-1} in four samples. Additionally, the homologous 3-bromofuran was found in three samples with concentrations up to 0.03 ng mL^{-1} . Until now this compound was only known to be a product of onion bulbs incubated with the proteobacterium *Erwinia carotovora* now called *Pectobacterium carotovorum* (Prithiviraj et al., 2004).

Furan emission was less pronounced in relation to the methylated furans while higher alkylated furan concentrations were on par with the dimethylated furans. One reason for the discrimination of higher alkylated furans could be the short incubation time of one hour compared to 24 h for the sediments. The water samples confirm the natural abundance of furan and its derivatives in the studied salt lakes. Likewise to the emissions from sediments furan and 3-methylfuran concentrations are higher with elevated dissolved iron content and lower pH values (Fig. 8a–d). However, 2-methylfuran nei-

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ther shows a correlation with iron content and pH value nor with the concentration of 3-methylfuran as was demonstrated for the sediments.

4 Conclusions

The previously established model reaction by Huber et al. (2010) for an abiotic furan formation from catechol was verified using a bispidine Fe^{2+} complex. Furthermore, the formation of ethyl furan from ethyl catechol was assayed and the emission of several furan derivatives was observed from sediment and water samples of hypersaline environments.

Total yields of 0.6 % of furan from catechol using the bispidine Fe^{2+} complex are comparable to previous studies using Fe^{3+} sulphate, but the turnover numbers of the Fe^{2+} complex are eight times higher than those of the Fe^{3+} salt. The main advantage of the bispidine Fe^{2+} complex for this kind of model reaction is the defined and well known coordination sphere which can be used in theoretical studies to calculate possible intermediates and reaction pathways.

Important parameters of the model reactions like iron concentration and pH can be projected on natural sediment and water samples as higher iron concentration and lower pH favoured furan emissions. Additionally, many geochemical parameters of Australian salt lakes and the Dead Sea were acquired and the emission of furan derivatives from these environments was discussed for the first time. These include various methylated and higher alkylated furans as well as 3-chlorofuran and 3-bromofuran. To our knowledge it is the first finding of 3-chlorofuran in natural samples while bromofuran was only reported so far from onion bulbs incubated with *Erwinia carotovora* (Prithiviraj et al., 2004).

Furthermore, our results show that the wet sediments of salt lakes are a natural source for furan derivatives which can contribute to the atmospheric formation of ultra fine particles observed around the Australian salt lakes and hence have an impact on the local climate (Junkermann et al., 2009).

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References

- Atkinson, R., Aschmann, S. M., and Carter, W. P. L.: Kinetics of the reactions of O₃ and OH radicals with furan and thiophene at 298 ± 2 K, *Int. J. Chem. Kinet.*, 15, 51–61, doi:10.1002/kin.550150106, 1983.
- Atkinson, R., Aschmann, S. M., Tuazon, E. C., Arey, J., and Zielinska, B.: Formation of 3-methylfuran from the gas-phase reaction of OH radicals with isoprene and the rate constant for its reaction with the OH radical, *Int. J. Chem. Kinet.*, 21, 593–604, doi:10.1002/kin.550210709, 1989.
- Berndt, T. and Böge, O.: Gas-phase reaction of OH radicals with benzene: products and mechanism, *Phys. Chem. Chem. Phys.*, 3, 4946–4956, doi:10.1039/B106667F, 2001.
- Berndt, T., Böge, O., and Rolle, W.: Products of the gas-phase reactions of NO₃ radicals with furan and tetramethylfuran, *Environ. Sci. Technol.*, 31, 1157–1162, doi:10.1021/es960669z, 1997.
- Bierbach, A., Barnes, I., and Becker, K. H.: Product and kinetic study of the OH-initiated gas-phase oxidation of furan, 2-methylfuran and furanaldehydes at ≈ 300 K, *Atmos. Environ.*, 29, 2651–2660, doi:10.1016/1352-2310(95)00096-H, 1995.
- Börzel, H., Comba, P., Hagen, K. S., Lampeka, Y. D., Lienke, A., Linti, G., Merz, M., Pritzkow, H., and Tsymbal, L. V.: Iron coordination chemistry with tetra-, penta- and hexadentate bispidine-type ligands, *Inorg. Chim. Acta*, 337, 407–419, doi:10.1016/S0020-1693(02)01100-3, 2002.

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Bowen, B. B. and Benison, K. C.: Geochemical characteristics of naturally acid and alkaline saline lakes in southern Western Australia, *Appl. Geochem.*, 24, 268–284, doi:10.1016/j.apgeochem.2008.11.013, 2009.

Bukowski, M. R., Comba, P., Lienke, A., Limberg, C., de Laorden, C. L., Mas-Balleste, R., Merz, M., and Que, L.: Catalytic epoxidation and 1,2-dihydroxylation of olefins with bispidine-iron(II)/H₂O₂ systems, *Angew. Chem. Int. Edit.*, 45, 3446–3449, doi:10.1002/anie.200504357, 2006.

Cabañas, B., Villanueva, F., Martín, P., Baeza, M. T., Salgado, S., and Jiménez, E.: Study of reaction processes of furan and some furan derivatives initiated by Cl atoms, *Atmos. Environ.*, 39, 1935–1944, doi:10.1016/j.atmosenv.2004.12.013, 2005.

Carter, W. P. L.: Updated Maximum Incremental Reactivity Scale and Hydrocarbon Bin Reactivities for Regulatory Applications, California Air Resources Board Contract 07–339, 2009.

Cheburkin, A. K., Frei, R., and Shoty, W.: An energy-dispersive miniprobe multielement analyzer (EMMA) for direct analysis of trace elements and chemical age dating of single mineral grains, *Chem. Geol.*, 135, 75–87, doi:10.1016/S0009-2541(96)00105-2, 1997.

Christian, T. J., Kleiss, B., Yokelson, R. J., Holzinger, R., Crutzen, P. J., Hao, W. M., Saharjo, B. H., and Ward, D. E.: Comprehensive laboratory measurements of biomass-burning emissions: 1. Emissions from Indonesian, African, and other fuels, *J. Geophys. Res.-Atmos.*, 108, 4719, doi:10.1029/2003JD003704, 2003.

Comba, P., Fukuzumi, S., Kotani, H., and Wunderlich, S.: Electron-transfer properties of an efficient nonheme iron oxidation catalyst with a tetradentate bispidine ligand, *Angew. Chem. Int. Edit.*, 49, 2622–2625, doi:10.1002/anie.200904427, 2010.

Comba, P., Wadepohl, H., and Wunderlich, S.: Oxidation versus dioxygenation of catechol: the iron-bispidine system, *Eur. J. Inorg. Chem.*, 2011, 5242–5249, doi:10.1002/ejic.201100802, 2011.

Cooper, W. J., Zika, R. G., Petasne, R. G., and Plane, J. M. C.: Photochemical formation of H₂O₂ in natural-waters exposed to sunlight, *Environ. Sci. Technol.*, 22, 1156–1160, doi:10.1021/es00175a004, 1988.

Degens, B. and Shand, P.: Assessment of Acidic Saline Groundwater Hazard in the Western Australian Wheatbelt: Yarra Yarra, Blackwood and South Coast, CSIRO, 2010.

Deng, Y. W. and Zuo, Y. G.: Factors affecting the levels of hydrogen peroxide in rainwater, *Atmos. Environ.*, 33, 1469–1478, doi:10.1016/S1352-2310(98)00239-8, 1999.

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- Dumdei, B. E., Kenny, D. V., Shepson, P. B., Kleindienst, T. E., Nero, C. M., Cupitt, L. T., and Claxton, L. D.: MS/MS analysis of the products of toluene photooxidation and measurement of their mutagenic activity, *Environ. Sci. Technol.*, 22, 1493–1498, doi:10.1021/es00177a017, 1988.
- 5 Gómez Alvarez, E., Borrás, E., Viidanoja, J., and Hjorth, J.: Unsaturated dicarbonyl products from the OH-initiated photo-oxidation of furan, 2-methylfuran and 3-methylfuran, *Atmos. Environ.*, 43, 1603–1612, doi:10.1016/j.atmosenv.2008.12.019, 2009.
- Grebel, J. E., Pignatello, J. J., and Mitch, W. A.: Effect of halide ions and carbonates on organic contaminant degradation by hydroxyl radical-based advanced oxidation processes in saline waters, *Environ. Sci. Technol.*, 44, 6822–6828, doi:10.1021/es1010225, 2010.
- 10 Gu, C. L., Rynard, C. M., Hendry, D. G., and Mill, T.: Hydroxyl radical oxidation of isoprene, *Environ. Sci. Technol.*, 19, 151–155, doi:10.1021/es00132a007, 1985.
- Gunz, D. W. and Hoffmann, M. R.: Atmospheric chemistry of peroxides – a review, *Atmos. Environ. A-Gen.*, 24, 1601–1633, 1990.
- 15 Haber, F. and Weiss, J.: Über die Katalyse des Hydroperoxydes, *Naturwissenschaften*, 20, 948–950, doi:10.1007/BF01504715, 1932.
- Hempfling, R. and Schulten, H. R.: Chemical characterization of the organic matter in forest soils by Curie point pyrolysis-GC/MS and pyrolysis-field ionization mass spectrometry, *Org. Geochem.*, 15, 131–145, doi:10.1016/0146-6380(90)90078-E, 1990.
- 20 Huber, S. G., Wunderlich, S., Schöler, H. F., and Williams, J.: Natural abiotic formation of furans in soil, *Environ. Sci. Technol.*, 44, 5799–5804, doi:10.1021/es100704g, 2010.
- Isidorov, V. and Jdanova, M.: Volatile organic compounds from leaves litter, *Chemosphere*, 48, 975–979, doi:10.1016/S0045-6535(02)00074-7, 2002.
- Isidorov, V. A., Zenkevich, I. G., and Ioffe, B. V.: Volatile organic compounds in the atmosphere of forests, *Atmos. Environ.*, 19, 1–8, doi:10.1016/0004-6981(85)90131-3, 1985.
- 25 Jordan, E. D., Hsieh, T. C. Y., and Fischer, N. H.: Volatiles from litter and soil associated with *Ceratiola ericoides*, *Phytochemistry*, 33, 299–302, doi:10.1016/0031-9422(93)85507-N, 1993.
- Junkermann, W., Hacker, J., Lyons, T., and Nair, U.: Land use change suppresses precipitation, *Atmos. Chem. Phys.*, 9, 6531–6539, doi:10.5194/acp-9-6531-2009, 2009.
- 30 Khain, A., Rosenfeld, D., and Pokrovsky, A.: Aerosol impact on the dynamics and microphysics of deep convective clouds, *Q. J. Roy. Meteor. Soc.*, 131, 2639–2663, doi:10.1256/qj.04.62, 2005.

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Kind, I., Berndt, T., Böge, O., and Rolle, W.: Gas-phase rate constants for the reaction of NO₃ radicals with furan and methyl-substituted furans, Chem. Phys. Lett., 256, 679–683, doi:10.1016/0009-2614(96)00513-1, 1996.

Leff, J. W. and Fierer, N.: Volatile organic compound (VOC) emissions from soil and litter samples, Soil Biol. Biochem., 40, 1629–1636, doi:10.1016/j.soilbio.2008.01.018, 2008.

Matveev, V., Peleg, M., Rosen, D., Tov-Alper, D. S., Hebestreit, K., Stutz, J., Platt, U., Blake, D., and Luria, M.: Bromine oxide-ozone interaction over the Dead Sea, J. Geophys. Res.-Atmos., 106, 10375–10387, doi:10.1029/2000JD900611, 2001.

Mulder, I., Huber, S. G., Krause, T., Zetzsch, C., Kotte, K., Dultz, S., and Schöler, H. F.: A new purge and trap headspace technique to analyse low volatile compounds from fluid inclusions of rocks and minerals, Chem. Geol., 358, 148–155, doi:10.1016/j.chemgeo.2013.09.003, 2013.

Müller, G. and Gastner, M.: The “Karbonat-Bombe”, a simple device for the determination of carbonate content in sediment, soils, and other materials, Neues Jb. Miner. Monat., 10, 466–469, doi:10013/epic.27884, 1971.

Ofner, J., Kamilli, K. A., Held, A., Lendl, B., and Zetzsch, C.: Halogen-induced organic aerosol (XOA): a study on ultra-fine particle formation and time-resolved chemical characterization, Faraday Discuss., doi:10.1039/C3FD00093A, 2013.

Ohta, T.: Furan ring formation in OH-initiated photooxidation of 1,3-butadiene and cis-1,3-pentadiene, B. Chem. Soc. Jpn., 57, 960–966, 1984.

Pierce, J. R. and Adams, P. J.: Efficiency of cloud condensation nuclei formation from ultrafine particles, Atmos. Chem. Phys., 7, 1367–1379, doi:10.5194/acp-7-1367-2007, 2007.

Prithviraj, B., Vikram, A., Kushalappa, A. C., and Yaylayan, V.: Volatile metabolite profiling for the discrimination of onion bulbs infected by *Erwinia carotovora* ssp. *carotovora*, *Fusarium oxysporum* and *Botrytis allii*, Eur. J. Plant Pathol., 110, 371–377, doi:10.1023/B:EJPP.0000021058.81491.f8, 2004.

Remucal, C. K. and Sedlak, D. L.: The Role of iron coordination in the production of reactive oxidants from ferrous iron oxidation by oxygen and hydrogen peroxide, in: Aquatic Redox Chemistry, ACS Symposium Series, American Chemical Society, 177–197, 2011.

Risholm-Sundman, M., Lundgren, M., Vestin, E., and Herder, P.: Emissions of acetic acid and other volatile organic compounds from different species of solid wood, Holz Roh. Werkst., 56, 125–129, doi:10.1007/s001070050282, 1998.

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- Rogge, W. F., Hildemann, L. M., Mazurek, M. A., Cass, G. R., and Simoneit, B. R. T.: Sources of fine organic aerosol. 1. Charbroilers and meat cooking operations, *Environ. Sci. Technol.*, 25, 1112–1125, doi:10.1021/es00018a015, 1991.
- Shand, P., Degens, B., Western Australia Dept. of W., Western Australia Dept. of Agriculture and F., National Action Plan for Salinity and Water Q., and Cooperative Research Centre for Landscape Environments and Mineral E., Avon Catchment Acidic Groundwater: Geochemical Risk Assessment, Crc LEME, 2008.
- Shepson, P. B., Edney, E. O., and Corse, E. W.: Ring fragmentation reactions in the photooxidations of toluene and o-xylene, *J. Phys. Chem.*, 88, 4122–4126, doi:10.1021/j150662a053, 1984.
- Spilker, R. and Grutzmacher, H.-F.: Isomerization and fragmentation of methylfuran ions and pyran ions in the gas phase, *Org. Mass Spectrom.*, 21, 459–466, doi:10.1002/oms.1210210803, 1986.
- Studenroth, S., Huber, S. G., Kotte, K., and Schöler, H. F.: Natural abiotic formation of oxalic acid in soils: results from aromatic model compounds and soil samples, *Environ. Sci. Technol.*, 47, 1323–1329, doi:10.1021/es304208a, 2013.
- Stutz, J., Ackermann, R., Fast, J. D., and Barrie, L.: Atmospheric reactive chlorine and bromine at the Great Salt Lake, Utah, *Geophys. Res. Lett.*, 29, 18-11–18-14, doi:10.1029/2002GL014812, 2002.
- Villanueva, F., Cabañas, B., Monedero, E., Salgado, S., Bejan, I., and Martín, P.: Atmospheric degradation of alkylfurans with chlorine atoms: product and mechanistic study, *Atmos. Environ.*, 43, 2804–2813, doi:10.1016/j.atmosenv.2009.02.030, 2009.
- Williams, W. D.: Salinisation: a major threat to water resources in the arid and semi-arid regions of the world, *Lakes & Reservoirs: Research & Management*, 4, 85–91, doi:10.1046/j.1440-1770.1999.00089.x, 1999.
- Zhang, H.-Z., Kasibhatla, S., Kuemmerle, J., Kemnitzer, W., Ollis-Mason, K., Qiu, L., Crogan-Grundy, C., Tseng, B., Drewe, J., and Cai, S. X.: Discovery and structure–activity relationship of 3-aryl-5-aryl-1,2,4-oxadiazoles as a new series of apoptosis inducers and potential anti-cancer agents, *J. Med. Chem.*, 48, 5215–5223, doi:10.1021/jm050292k, 2005.
- Zhang, W. and Du, B.: Products and mechanism of the Cl-initiated atmospheric oxidation of furan: a theoretical study, *Comput. Theor. Chem.*, 963, 348–356, doi:10.1029/2004JD004993, 2011.

Zingler, J. and Platt, U.: Iodine oxide in the Dead Sea Valley: evidence for inorganic sources of boundary layer IO, J. Geophys. Res.-Atmos., 110, D07307, doi:10.1029/2004JD004993, 2005.

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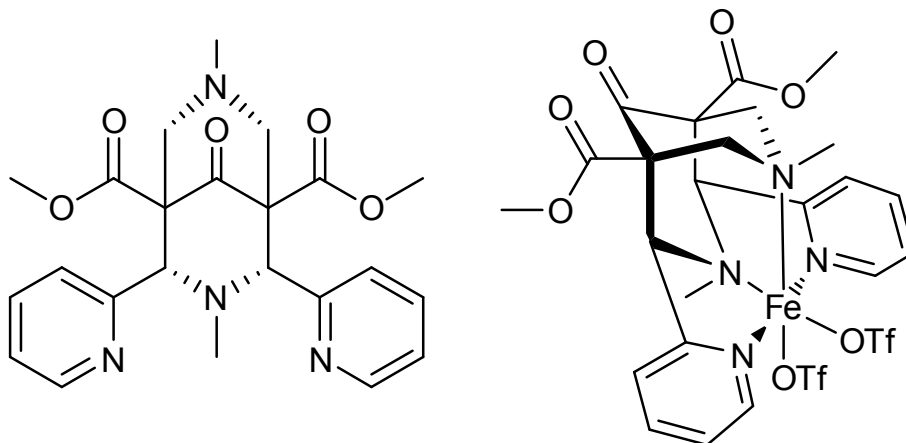


Fig. 1. Structure of the bispidine N_2Py_2 ligand (left) and geometry of the $[Fe^{II}(N_2Py_2)(OTf)_2]$ complex (right).

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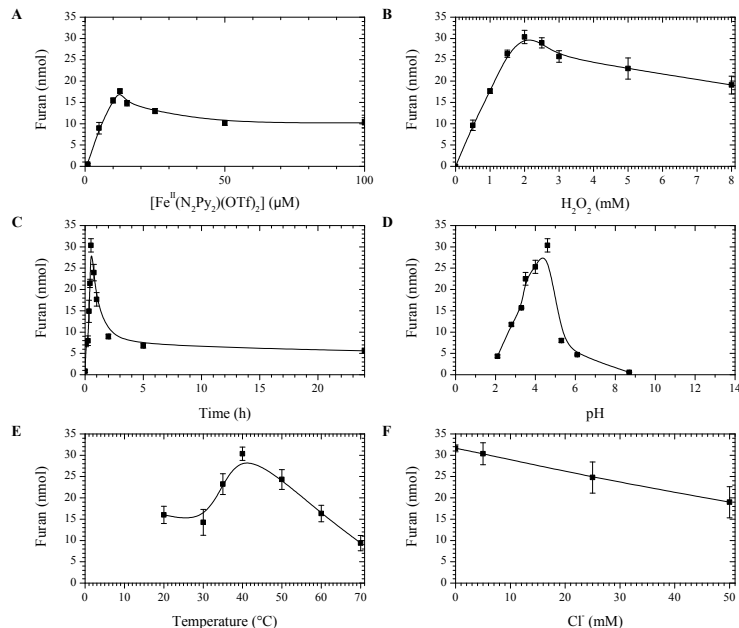


Fig. 2. Effects of different parameters on the furan formation from a 10 mL solution. **(A)** Variation of $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$: reaction of 0.5 mM catechol with 0–100 μM $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$, 1 mM H_2O_2 , 5 mM KCl at a pH of 4.6 and 40 $^\circ\text{C}$ after 30 min. **(B)** Variation of H_2O_2 : reaction of 0.5 mM catechol with 12.5 μM $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$, 0–8 mM H_2O_2 , 5 mM KCl at a pH of 4.6 and 40 $^\circ\text{C}$ after 30 min. **(C)** Time dependency: reaction of 0.5 mM catechol with 12.5 μM $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$, 2 mM H_2O_2 , 5 mM KCl at a pH of 4.6 and 40 $^\circ\text{C}$ after 0–24 h. **(D)** Variation of pH: reaction of 0.5 mM catechol with 12.5 μM $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$, 2 mM H_2O_2 , 5 mM KCl at a pH of 2.1–8.7 and 40 $^\circ\text{C}$ after 30 min. **(E)** Temperature dependency: reaction of 0.5 mM catechol with 12.5 μM $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$, 2 mM H_2O_2 , 5 mM KCl at a pH of 4.6 and 20–70 $^\circ\text{C}$ after 30 min. **(F)** Variation of chloride: reaction of 0.5 mM catechol with 12.5 μM $[\text{Fe}^{\text{II}}(\text{N}_2\text{Py}_2)(\text{OTf})_2]$, 2 mM H_2O_2 , 0–50 mM KCl at a pH of 4.6 and 40 $^\circ\text{C}$ after 30 min.

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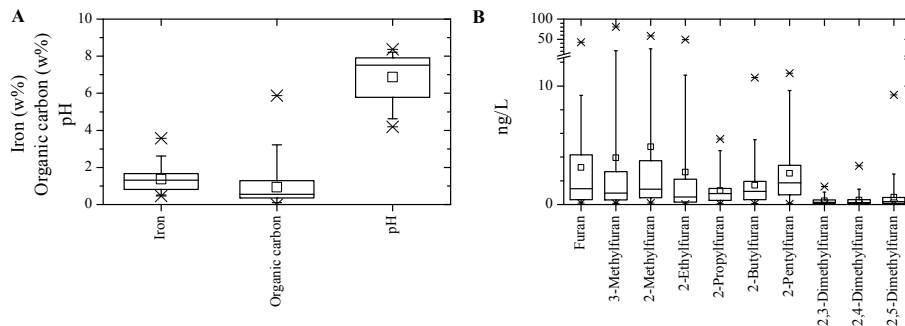


Fig. 3. Distribution of **(A)** iron and organic carbon concentration as well as pH and **(B)** of the furan emissions from the various Australian and Dead Sea sediment samples.

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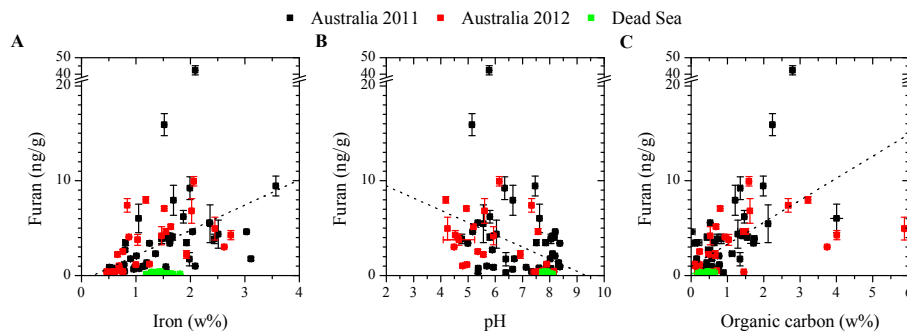


Fig. 4. Furan emissions dependent on **(A)** iron concentration, **(B)** pH and **(C)** organic carbon content of the sediments (Australia 2011: black; Australia 2012: red; Dead Sea 2012: blue).

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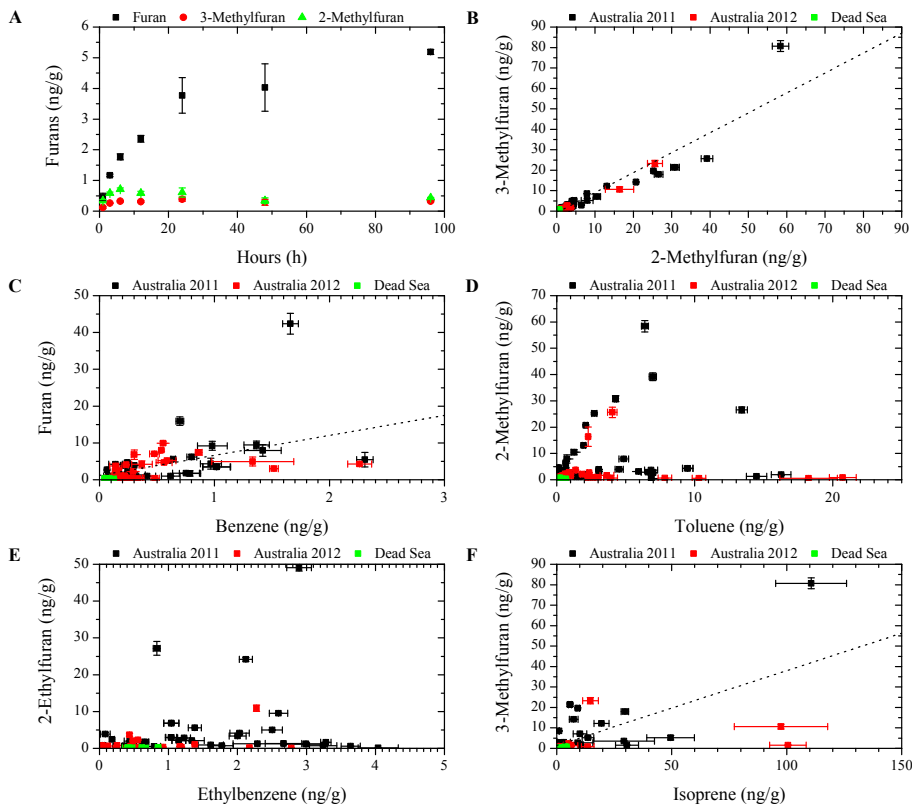


Fig. 5. (A) Evolution of the furan (black), 2-methylfuran (green) and 3-methylfuran (red) concentrations in the headspace vials from the Lake Boats 0–2 cm sample over time. Correlation between (B) 2-methylfuran and 3-methylfuran emissions, (C) furan and benzene, (D) 2-methylfuran and toluene, (E) 2-ethylfuran and ethylbenzene and (F) 3-methylfuran and isoprene from sediment samples (Australia 2011: black; Australia 2012: red; Dead Sea 2012: blue).

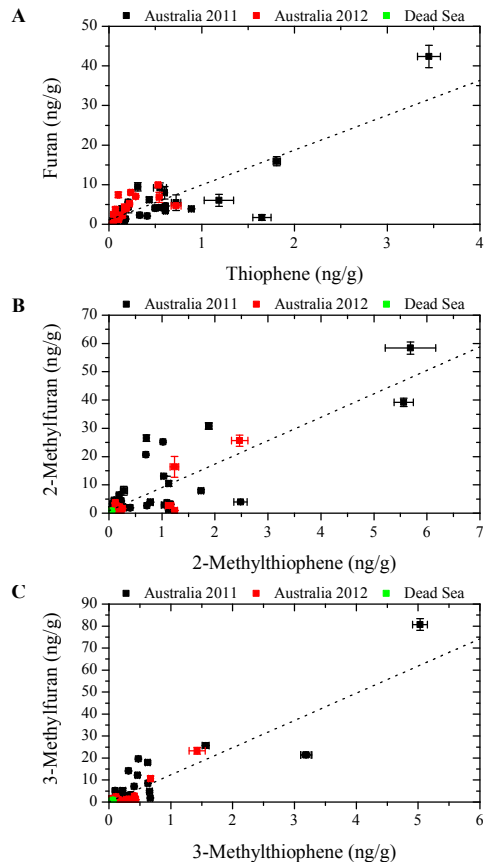


Fig. 6. Correlation between the emission of related heterocyclic aromatic compounds such as **(A)** furan and thiophene, **(B)** 2-methylfuran and 2-methylthiophene and **(C)** 3-methylfuran and 3-methylthiophene from sediment samples (Australia 2011: black; Australia 2012: red; Dead Sea 2012: blue).

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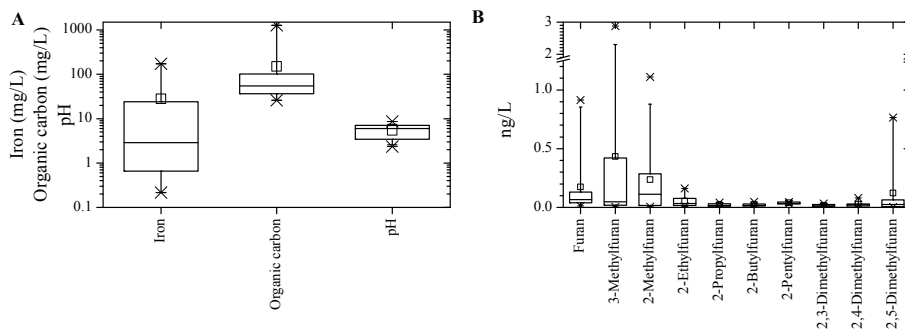


Fig. 7. Distribution of **(A)** iron and organic carbon concentration as well as pH and **(B)** of the furan emissions from the various Australian and Dead Sea water samples. Only the Australian water samples were analysed for organic carbon.

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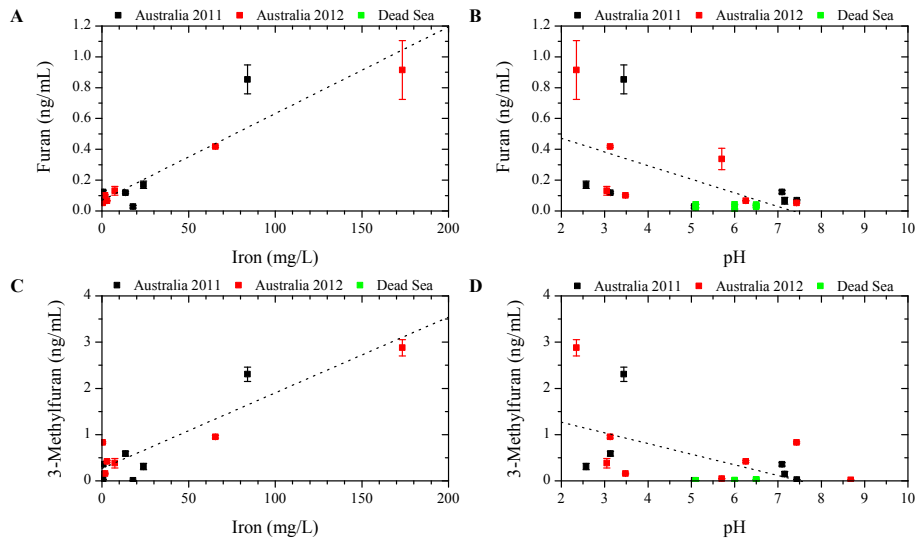


Fig. 8. Furan emissions dependent on **(A)** iron concentration, **(B)** pH and **(C)** organic carbon content in water samples (Australia 2011: black; Australia 2012: red; Dead Sea 2012: blue).

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