Jean-Pierre Gattuso,

We are pleased to resubmit our manuscript with minor revisions made. In this document we detail how we responded to each reviewer suggestion/comment. After, we provide a complete copy of our manuscript with tracked changes. In our responses to reviewers, our responses are in plain text, and are written among italicized text from reviewers. We indicate text from our revised manuscript in bold.

In addition, we now thank Dr. Judith Hauck by name in the Acknowledgements section. We requested and received her permission to be named there through E-mail.

Sincere thanks,

-Brendan and coauthors

Referee #1

This paper is much improved in the re-submitted version. Shorter, easier to read, and a good use of the heavy armament in the Princeton-GFDL complex. The history of the field is a bit abbreviated, but that would only lengthen the paper. I only have one comment for curiosity. The authors note the strong signals in the Red Sea. I suspect that the Persian Gulf has perhaps stronger riverine signals overlaid due to the significant river input at the northern end, and older but good quality alkalinity data are available (e.g. Brewer, P. G. and D. Dyrssen (1985). Chemical Oceanography of the Persian Gulf. Prog. Oceanog.,14,41-55.) The Persian Gulf influence is small on the oceanic scale, but it may offer an interesting end member for the discussion.

The data presented in that paper do indeed suggest the Persian Gulf has a similar *Alk** distribution to the Red Sea. We now note as much:

Brewer and Dyrssen (1985) provide seawater chemistry measurements from the neighboring Persian Gulf that suggest strong calcium carbonate formation results in low *Alk there as well (< –240 µmol kg^{-1} along the Trucial Coast).**

Referee #3

The authors have responded to all of the comments I made. They have not been ignored.

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Although I do not completely agree with all of their responses to my comments (for instance, a lack of certainty about errors in the Alk values does not in my opinion justify omitting statistical hypothesis testing), I think the most serious and obvious mistakes have been corrected. However, perhaps not surprisingly given the scale of the changes, other errors have now crept in as detailed below. There are also grammatical errors, references to the wrong figures or figures that no longer exist, etc. I recommend that the manuscript needs to be checked through rigorously one more time by the authors before it can be accepted.*

Line 22: The numbers are wrong. 0.5 – 1.6 Gt C y-1 (in Table 3 of Berelson 2007) converts (multiply by 1000, divide by 12 then multiply by 2) to 80-270 TMol alkalinity y-1. The "x 10" should also be removed.

We removed the "x 10" as rightly suggested. However, we didn't change our numbers because we are using a different estimate from that paper than the one Reviewer 3 supposed. The range pointed out by the reviewer is for overall calcium carbonate production. We are referring to "net calcium carbonate production" or export… which is listed as 0.4 to 1.8 GtC yr^{-1}… in the abstract and in Table 3.

Line 50: "interested calcium" ?

We added "**in**" between these two words.

Line 105 should be changed to read "...plus the small (in most places) residual variation..." It is negligible across most of the ocean but in fact rather important in several regions.

Added.

Line 203: there is no figure 6d anymore.

We removed this reference.

Line 211: Only Fig 7b is relevant and it only shows the Bay of Bengal.

We changed this to 6b after swapping the order of figures 6 and 7.

Line 223: Fig 7 no longer shows this.

We swapped the order of figures 6 and 7.

Line 231: Fig 8 no longer shows data from the Red Sea.

This now reads "**Fig. 6.**"

Line 238: better to say "abiotic carbonate formation" (I presume that is the intent) because inorganic carbon(ate) is PIC.

We changed "inorganic" to "abiotic" in several places in that section.

Additional Referee…

Review of Carter et al: Processes determining the marine alkalinity and calcium carbonate saturation distributions

Text to editor and authors:

The paper improved a lot, it is much more concise, well readable and the figures are of good quality. I can now follow why this tracer might be a valuable concept. I can recommend this manuscript for publication in BG after a relatively small number of edits, mostly in the abstract and introduction, that the authors can probably clarify quickly.

Judith Hauck

Normal font is text from the manuscript, suggested edits are in bold, and my comments are in italics.

I italicized all the review text and removed bold formatting.

Comments:

Line 19 give reference after carbon uptake (e.g. Orr et al, 2005, Nature)

Done.

L. 20: give definition of OmegaC. This is used later, but never defined. As this is the main topic of the paper, a short introduction to Omega should be given.

Done:

With marine calcite saturation states decreasing due to anthropogenic carbon uptake (Orr et al., 2005), it is important to understand the degree to which carbonate cycling impacts the calcium carbonate saturation state. Carbonate saturation state is a measure of how supersaturated seawater is with respect to a given mineral form of calcium carbonate. It is expressed as the ratio Ω_C between the product of Ca^{2+} and CO_3^{2-} ion concentrations and the calcite thermodynamic equilibrium solubility product. Values of Ω_C greater than one

indicate calcite precipitation is favored thermodynamically over calcite dissolution, and the reverse is true for values less than one.

L. 36: some calculations → specify which ones use gridded data set or which ones bottle data whatever is easier.

We removed all calculations that used bottle data in previous revisions, so this now reads simply:

We use our gridded dataset in our calculations…

L. 80: give reference for empirical ratio (again)

L. 151-53: „In Fig. 5 we provide ... discrete surface Alk ... . Figure 5 also provides ... gridded Alk* ... . „: I can only see one measure of Alk* on the y-axis – are these discrete samples or gridded data?*

We provide the gridded mean value in writing on the plots themselves. We are now more clear about this:

Figure 5 also indicates a single volume-weighted mean gridded *Alk for each basin (in writing).**

L. 188: Alk = -16.5 µmol/kg : give number for other regions, too.*

Done:

Mean *Alk is higher in the Pacific than the Atlantic and Indian, even when neglecting the region north of 40°N as we do for the Atlantic (*Alk** = −16.5 µmol kg^{-1} when omitted vs. −22.9 µmol kg^{-1} for the Atlantic and −22.2 µmol kg^{-1} for the Indian).**

L. 200 from here on Figure numbering is erroneous: Figure 4d → Figure 3d

Fixed.

L. 205: Alk → AT. I assume that Gascard et al refer to AT, not Alk* as you just introduce this concept here.*

We opted to speak more generally in this revision:

Gascard et al. (2004a, b) suggest that waters along the coast of Norway are part of the Norwegian Coastal Current, and originate in the Baltic and North Seas where there are also strong riverine inputs (Thomas et al., 2005).

Switch Figures 6 and 7 as you describe Figure 7 first and then Figure 6

Switched.

L. 242ff: high-magnesium calcite: your hypothesis only works if you can give a reference that high-Mg calcite is less soluble than biogenic aragonite. According to Morse et al 2006 (doi:10.1016/j.gca.2006.08.017) aragonite is more soluble than synthetic high-Mg-calcite, but less soluble than biogenic high-Mg calcite. So, your hypothesis is possible according to Morse et

al if we assume that inorganically precipitated high-Mg calcite in the open ocean behaves rather like a synthetic high-Mg calcite.

We added text to address this line of thought:

Morse et al. (2006) find that synthetic high magnesium calcite—unlike biogenic high magnesium calcite—is less soluble than aragonite, so this substitution is favored thermodynamically if the abiotic mineral forms similarly to the synthetic mineral.

L. 361-363: Don't give results here of another study (which was not published) which are not part of this manuscript:

AT varies substantially in response to freshwater cycling, so Alk trends may be able to be detected sooner and more confidently attributed to changes in calcium carbonate cycling than trends in AT. Preliminary explorations of Earth System Model output suggest time of trend emergence for the alkalinity trends discussed by Ilyina et al. (2009) could be reduced by as much as a factor of 5. → AT varies substantially in response to freshwater cycling, so Alk* trends may be able to be detected sooner and more confidently attributed to changes in calcium carbonate cycling than trends in AT (Ilyina et al. , 2009).*

We followed this suggestion verbatim.

References: add dois where available.

Added dois to 25 references.

Caption Figure 2: „gridded CARINA, PACIFICA and GLODAP bottle data product" → gridded or bottle data? Or did you make your own gridded product from all the bottle data?

We did, so we try to be more clear about this in the caption:

Global (a) total alkalinity A_T, (b) salinity, (c) *Alk, and (d) phosphate distributions at the surface (10 m depth surface) from our gridded CARINA, PACIFICA, and GLODAP bottle data product detailed in Supplementary Materials document SA. Areas with exceptionally poor coverage in the data used to produce the gridded product are blacked out.**

Technical corrections:

Title and multiple times throughout the manuscript: calcium carbonate saturation → calcium carbonate saturation state

We added "state" in many places.

*Line 2: composite tracer for the marine system, Alk**

Added **for the marine system**

L. 4: We estimate that the...

Added **that**

L. 5: 8 → eight

Changed

L.9: We use the Alk distribution to estimate the variability of the calcite saturation state*

Changed as suggested

L. 10: We show that regional

Added **that**

L. 11: calcite saturation state (throughout the manuscript, will not point to any more occurrences, or use Omega after introducing it)

Changed to Omega in several places and added "state" in many others.

L. 12: The variations in net calcium carbonate cycling revealed by Alk play a comparatively minor role in determining the calcium carbonate saturation state.*

Changed as suggested

L. 20: impacts the calcium carbonate saturation state

Changed as suggested

L. 22: cycling → sources and sinks

Changed as suggested

L. 28: estimate that hydrothermal

Changed as suggested

L. 50: interested in calcium

Changed as suggested

L. 191: smaller Pacific mean net calcium carbonate formation (or larger CaCO3 dissolution).

We added "formation" but omitted the parenthetical statement because it is implied by "net."

L. 201: Figure 3 → Figure 2

We changed figure numbers here and elsewhere

L. 203: Fig. 6d → Fig. 5d

We changed figure numbers here and elsewhere

L. 211: Figure 7 provides Alk depth sections for both areas. → Figure 7b provides an Alk* depth section for this area. You only introduced one area so far.*

Fixed

L. 212: Bay of Bengal (Fig. 7b)

We fixed the figure number after rearranging figures

L. 214: Ganges (Fig. 7a)

We did not separate these two rivers as suggested since they both feed into the Bay of Bengal (6b)

L. 219: largest AT discharge → largest AT concentration

Discharge is correct. The largest A_T concentration I'm aware of is the Daugava, though I'd guess there are larger.

L. 223: Figure 7 → is actually Figure 6, might be 7 again after switching.

Addressed

L. 231: Fig 8 → Fig. 7

Addressed

L. 252: sea ice that Dieckmann

Added **that**

L. 261 and 262: calcium carbonate cycling (and calcite saturation state)

*Alk** can be affected by magnesium carbonate cycling as well, so we didn't think adding "calcium" would be making the statement more correct.

L. 315: cycling as compared to section 4.1 is

Added a variant on this:

The increased importance of freshwater cycling compared to section 4.1 is because freshwater dilutes C_T by more than the equilibrium C_T decreases from A_T dilution, so carbon uptake tends to follow freshwater precipitation and carbon outgassing follows evaporation.

L. 340/1: and to change

Added "**to**"

L. 568: then and cancel → then cancel

Deleted "**and**"

In Figure 8: surface calcite saturation → surface calcite saturation state

Changed to "surface Ω_C "

Table A1 and A3: Line 5 in this table is misleading, Add another line after this line: „Process i bold numbers." In the upper line write „dOmegaC/dXj" and then give bold numbers, in the new/lower line write „Process" and „i" as headers for the columns below

The suggested changed left some blank space, so we added a little bit more (e.g.):

Table A1. $\frac{\partial \Omega_C}{\partial X_j}$ (bold text) and $\frac{\partial X_{j,i}}{\partial R_i}$ (italic text) terms used in Eq. (A5) for atmospherically isolated mean seawater from all ocean depths. These terms are specific to the j = 7 (columns) properties we use to calculate Ω_C and i = 6 (rows) processes we consider. Units for $\frac{\partial \Omega_C}{\partial X_j}$ are the inverse of the listed X_j units. Units for $\frac{\partial X_{j,i}}{\partial R_i}$ are the X_j units divided by the R_i units given in Table 1.

Properties		Pressure	Temp	Salinity	Phos.	Silicate	A_T	C_T
X_j *units*		db	°C		µmol/kg	µmol/kg	µmol/kg	µmol/kg
j		1	2	3	4	5	6	7
Mean seawater values		2235	3.7	34.71	2.15	49.0	2362	2254
$\frac{\partial \Omega_C}{\partial X_j}$		**–0.00028**	**0.014**	**–0.011**	**–0.0085**	**–0.00012**	**0.0082**	**–0.0079**
Process	*i*	$\frac{\partial X_{1,i}}{\partial R_i}$	$\frac{\partial X_{2,i}}{\partial R_i}$	$\frac{\partial X_{3,i}}{\partial R_i}$	$\frac{\partial X_{4,i}}{\partial R_i}$	$\frac{\partial X_{5,i}}{\partial R_i}$	$\frac{\partial X_{6,i}}{\partial R_i}$	$\frac{\partial X_{7,i}}{\partial R_i}$
Carbonate cycling	1	-	-	-	-	-	*1*	*0.5*
Org. matter cycling	2	-	-	-	*1*	-	*–20.16*	*117*
Freshwater cycling	3	-	-	*1*	*0.062*	*1.4*	*68*	*65*
Sinking / shoaling	4	*1*	*0.00010*	-	-	-	-	-
Warming / cooling	5	-	*1*	-	-	-	-	-
Denit ./ nit. fix.	6	-	-	-	-	-	*–1.26*	-

Processes determining the marine alkalinity and calcium carbonate saturation state distributions

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Abstract

We introduce a composite tracer for the marine system, *Alk**, that has a global distribution primarily determined by $CaCO_3$ precipitation and dissolution. *Alk** is also affected by riverine alkalinity from dissolved terrestrial carbonate minerals. We estimate that the Arctic receives approximately twice the riverine alkalinity per unit area as the Atlantic, and ~~8~~eight times that of the other oceans. Riverine inputs broadly elevate *Alk** in the Arctic surface and particularly near river mouths. Strong net carbonate precipitation results in low *Alk** in subtropical gyres, especially in the Indian and Atlantic Oceans. Upwelling of dissolved $CaCO_3$ rich deep water elevates Northern Pacific and Southern Ocean *Alk**. We use the *Alk** distribution to estimate the variability of the calcite saturation state ~~variability~~ resulting from $CaCO_3$ cycling and other processes. We show that regional differences in surface calcite saturation state are due primarily to the effect of temperature differences on CO_2 solubility and, to a lesser extent, differences in freshwater content and air-sea disequilibria. The variations in net calcium carbonate cycling revealed by *Alk** play a comparatively minor role in determining the calcium carbonate saturation state.

1. Introduction

Our goal is to use high-quality total alkalinity (A_T) observations to examine the effects of calcium carbonate cycling on marine A_T and calcium carbonate saturation states. This study is motivated in part by ocean acidification. With marine calcite saturation states decreasing due to anthropogenic carbon uptake (Orr et al., 2005), it is important to understand the degree to which carbonate cycling impacts the calcium carbonate saturation state ~~calcite saturations~~.

Carbonate saturation state is a measure of how supersaturated seawater is with respect to

a given mineral form of calcium carbonate. It is expressed for calcite as the ratio Ω_C between the product of Ca^{2+} and CO_3^{2-} ion concentrations and the calcite thermodynamic equilibrium solubility product. Values of Ω_C greater than one indicate calcite precipitation is favored thermodynamically over calcite dissolution, and the reverse is true for values less than one.

Field Code Changed

Field Code Changed

Field Code Changed

Marine calcium carbonate cycling includes both internal and external calcium carbonate ~~cycling~~sources and sinks. Internal cycling refers to net formation of 67-300~~×10~~ Tmoles A_T yr^{-1} worth of calcium carbonate ~~(Berelson et al., 2007)~~ in the surface ocean (Berelson et al., 2007) and net dissolution of most of this calcium carbonate at depth. External marine carbonate cycling refers to inputs of carbonate minerals dissolved in rivers, sediment pore waters, hydrothermal vent fluids, and submarine groundwater discharge, and to loss due to biogenic carbonate mineral burial and authigenic mineralization in sediments. Rivers add 33 Tmoles A_T yr^{-1} worth of dissolved bicarbonate to the ocean (Cai et al., 2008). Wolery and Sleep (1988) estimate that hydrothermal vents add an additional 6.6 Tmoles A_T yr^{-1}, though deVilliers (1998) argues the hydrothermal contribution may be as high as 30 Tmoles A_T yr^{-1}. Submarine groundwater discharge is poorly constrained, but is thought to exceed riverine inputs in some areas (Moore, 2010).

We investigate calcium carbonate cycling using the global A_T distribution in a dataset we created by merging the PACIFICA (Suzuki et al., 2013), GLODAP, and CARINA discrete data products (Key et al. 2004; 2010; Velo et al., 2009). We have combined and gridded these data products using methods detailed in Supplementary Materials document SA. We use our gridded dataset in ~~some~~ our calculations to limit sampling biases and to enable us to make volume-weighted mean property estimates.

Dickson (1981) defines total alkalinity as the concentration excess "of proton acceptors

formed from weak acids ($pK \leq 4.5$) relative to proton donors (weak bases with $pK > 4.5$)" at a reference temperature, pressure, and ionic strength. A_T can be thought of as a measure of how well buffered seawater is against changes in pH. This operational definition gives A_T (expressed in mol kg^{-1}) several properties that make it an especially useful carbonate system parameter for examining carbonate cycling:

1. It mixes conservatively,
2. … and is therefore diluted and concentrated linearly by evaporation and precipitation.
3. It responds in predictable ways to calcium carbonate cycling.
4. … as well as organic matter formation and remineralization.
5. It is not changed by air-sea exchange of heat or carbon dioxide.
6. It is however affected by anaerobic redox reactions (Chen, 2002).

We are primarily interested in calcium carbonate cycling, item 3 in our list. In section 2 of this paper we therefore define a tracer we call *Alk** that removes the majority of the influences of organic matter cycling (item 4), freshwater cycling (item 2), and non-sedimentary anaerobic redox reactions (item 6) while still mixing conservatively, remaining insensitive to gas exchange, and responding to calcium carbonate cycling. In section 3 we discuss processes that govern the *Alk** distribution globally, by ocean basin, and regionally. In section 4 we define a metric to quantify the influence of various processes on the marine calcite saturation state. We use this metric with our gridded dataset and *Alk** to determine the relative importance of the various controls on calcite saturation state in the ocean and at the ocean surface. We summarize our findings in section 5.

2. The *Alk** tracer

In defining *Alk**, we take advantage of the potential alkalinity (Brewer et al., 1975) concept to remove the majority of the influence of organic matter cycling and denitrification, and use a specific salinity normalization scheme (Robbins, 2001) to remove the influence of freshwater cycling. We detail the *Alk** definition and the reasoning behind it in this section.

The influence of organic matter cycling on A_T is due primarily to the biologically-driven marine nitrogen cycle. Nitrate uptake for anaerobic denitrification and the production of amino acids occurs in a~~n~~ ~1:1 mole ratio with the release of molecules that increase A_T (Chen, 2002). Similarly, nitrate from fixation of nitrogen gas and remineralization of amino nitrogen is released in a 1:1 mole ratio with acids that titrate away A_T (Wolf-Gladrow et al., 2007). This observation led Brewer et al. (1975) to propose the idea of "potential alkalinity" as the sum of A_T and nitrate with the aim of creating a tracer that responds to the cycling of calcium carbonates without changing in response to organic matter cycling. Feely et al. (2002) since used a variant that relies on the empirical relationship between dissolved calcium concentrations, A_T, and nitrate determined by Kanamori and Ikegami (1982). This variant has the advantage of implicitly accounting for the A_T changes created by the exchange of numerous other components of marine organic matter besides nitrate (e.g. sulfate and phosphate). We thus use the ratio found by Kanamori and Ikegami (1982) to define potential alkalinity (A_P).

$$A_P = A_T + 1.26 * [NO_3^-] \quad (1)$$

While the empirical Kanamori and Ikegami (1982) ~~empirical~~ ratio of 1.26 may be specific to the elemental ratios of the North Pacific, Wolf-Gladrow et al. (2007) provide a theoretical derivation from Redfield ratios and obtain a similar value of 1.36.

The sensitivity of the A_T distribution to freshwater cycling is due primarily to the dilution or concentration of the large background A_T fraction that does not participate in carbonate

cycling on timescales of ocean mixing. This background fraction behaves conservatively, so we call it conservative potential alkalinity (A_P^C) and estimate it directly from salinity as:

$$A_P^C \equiv S \frac{\overline{A_P}}{\overline{S}} \quad (2)$$

Here, terms with a bar are reference values chosen as the mean value for those properties in the top 20 meters of the ocean. We obtain a volume-weighted surface $\overline{A_P}$ (2305 µmol kg^{-1}) to $\overline{S}$ (34.71) ratio of 66.40 µmol kg^{-1} from our gridded dataset. The mean surface values are chosen in an effort to best capture the impact of freshwater cycling where precipitation and evaporation occur.

Robbins (2001) showed that subtracting an estimate of the conservative portion of a tracer, such as A_P^C, produces a salinity-normalized composite tracer that mixes conservatively. This scheme also retains the 2:1 change of A_T to dissolved inorganic carbon (C_T) with carbonate cycling. We follow this approach in our definition of *Alk**. In Supplementary Materials document SB we estimate this approach removes 97.5% of the influence of freshwater cycling on potential alkalinity and reduces the influence of freshwater cycling on *Alk** to less than 1% of the *Alk** variability. In Supplementary Materials document SC we demonstrate that *Alk** mixes conservatively, and briefly contrast *Alk** to traditionally normalized potential alkalinity which does not mix conservatively (Jiang et al., 2014).

In total, we define *Alk** as the deviation of potential alkalinity from A_P^C,

$$Alk^* \equiv A_P - A_P^C \quad (3)$$

$$\equiv A_P - \frac{\overline{A_P}}{\overline{S}} S \quad (4)$$

$$\equiv A_P - 66.4 \times S \quad (5)$$

where *Alk** has the same units as A_T (µmol kg^{-1}). The *Alk** distribution is attributable primarily

to carbonate cycling plus the small (in most places) residual variation due to freshwater cycling that is not removed by subtracting A_P^C. However, hydrothermal vent fluid and non-denitrification anaerobic redox chemistry may substantively affect alkalinity distributions in certain marine environments, and *Alk** distributions could not be attributed purely to internal and external calcium carbonate cycling in these locations.

Mean global surface *Alk** is zero by definition, and thus *Alk** can have negative as well as positive values. For reference, more than 95% of our gridded *Alk** dataset falls between –35 and 220 µmol kg^{-1}. Comparing gridded *Alk** to *Alk** from measurements suggests a standard disagreement of order 10 µmol kg^{-1}. We adopt this number as an estimate of standard gridded *Alk** error despite noting there are reasons to suspect that this value could be either an underestimate (correlated errors) or an overestimate (we are directly comparing instantaneous point measurements to estimates for annual averages for a grid cell).

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3. *Alk** distributions

We consider *Alk** distributions globally, by ocean basin, and regionally in the context of sources and sinks of the tracer both globally and regionally. We pay special attention to riverine *Alk** because it is easily identified where it accumulates near river mouths.

*3.1 Global distribution of Alk**

Figure 1 maps surface *Alk** (top 50 m) at the measurement stations. We provide this figure to show where we have viable *Alk** estimates and to demonstrate that our gridded data product adequately captures the measured *Alk** distribution. Figure 2 maps gridded global surface A_T, salinity, *Alk**, and phosphate distributions and masks the regions that are lacking data

in Fig. 1.

The similarity of the A_T (Fig. 2a) and salinity (Fig. 2b) distributions demonstrates the strong influence of freshwater cycling on the surface marine A_T distribution (see also: Millero et al. 1998, Jiang et al., 2014). The dissimilarity between *Alk** (Fig. 2c) and salinity (Fig. 2b) suggests *Alk** removes the majority of this influence. The phosphate (Fig. 2d) and *Alk** (Fig. 2c) distributions are similar at the surface. They are also similar at depth: Figures 3 and 4 show zonally-averaged gridded depth sections of *Alk** and phosphate. *Alk** and phosphate concentrations are low in the deep Arctic Ocean (Figs. 3d, and 4d), intermediate in the deep Atlantic Ocean (Figs. 3a and 4a), and high in the deep North Pacific (Figs. 3b and 4b) and deep North Indian (Figs. 3c and 4c) Oceans. *Alk** and phosphate distributions are similar because similar processes shape them: the hard and soft tissue pumps transport A_T and phosphate, respectively, from the surface to depth. The "oldest" water therefore has the highest net phosphate and *Alk** accumulation. High surface phosphate and *Alk** in the Southern Ocean and North Pacific in Figs. 2, 3, and 4 are due to upwelled old deep waters.

Several qualitative differences between *Alk** and phosphate distributions are visible in Figs. 2c, 2d, 3, and 4. Surface phosphate is low in the Bay of Bengal and high in the Arabian Sea (Fig. 2d), while the opposite is true for *Alk** (Fig. 2c). Also, *Alk** reaches its highest surface concentration in the Arctic (Figs. 2c and 3d) where phosphate is not greatly elevated (Figs. 2d and 4d). These surface differences are due to regional riverine *Alk** inputs (Section 3.3). Another difference is that *Alk** reaches a maximum below 2000 m in all ocean basins except the Arctic, while phosphate maxima are above 2000 m. We attribute the deeper *Alk** maxima to deeper dissolution of calcium carbonates than organic matter remineralization. Finally, *Alk** values are higher in the deep Indian Ocean than in the deep Pacific. This is likely due to elevated

biogenic carbonate export along the coast of Africa and in the Arabian Sea (Sarmiento et al., 2002; Honjo et al., 2008).

3.2 Alk by ocean basin*

In Fig. 5 we provide 2-D color histograms of discrete surface *Alk** and salinity measurements for the five major ocean basins. Figure 5 also ~~provides~~ indicates a single volume-weighted mean gridded *Alk** for each basin (in writing). We attribute the decrease in *Alk** as salinity increases—especially visible in the low-salinity bins in the Arctic Ocean (Fig. 5d)—to mixing between high-*Alk** low-salinity river water and low-*Alk** high-salinity open ocean water. Net precipitation in the tropics and net evaporation in the subtropics widens the histograms across a range of salinities and alkalinities without affecting *Alk** in Figs. 5a, 5b, and 5c. The *Alk** elevation associated with upwelled water is most visible in Fig. 5e where Upper Circumpolar Deep Water upwelling near the Polar Front results in high-frequency (i.e. warm colored) histogram bins at high-*Alk**. Similarly, the high-frequency *Alk** bins in Fig. 5b with salinity between 32.5 and 33.5 are from the North Pacific Subpolar Gyre, and are due to upwelled old high-*Alk** water (cf. the Si* tracer in Sarmiento et al. (2004)). River water contributions can be most easily seen in a scattering of low-frequency (cool colored) high-*Alk** and low-salinity bins in the Arctic Ocean.

The surface Southern Ocean has the highest *Alk** followed by the Arctic and the ~~,~~ Pacific~~, and t~~. The Indian and Atlantic have similar and low mean *Alk**. The high mean Southern Ocean *Alk** is due to upwelling. The high mean Arctic surface *Alk** is due to riverine input. The Atlantic and the Arctic together receive ~65% of all river water (Dai and Trenberth, 2002). We construct a budget for terrestrial A_T sources to the various surface ocean basins using the

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following assumptions:

1. the A_T of 25 large rivers are as given by Cai et al. (2008),
2. the volume discharge rates of 200 large rivers are as given by Dai and Trenberth (2002),
3. groundwater and runoff enter each ocean in the same proportion as river water from these 200 rivers,
4. the A_T of all water types that we do not know from assumption 1. is the 1100 µmol kg^{-1} global mean value estimated by Cai et al. (2008), and
5. 40°N is the boundary between the Atlantic and the Arctic and 40°S is the boundary between the Southern and the Atlantic Oceans (based upon the region of elevated surface phosphate in Fig. 2d),

Our detailed budget is provided as Supplementary Materials file SD. We estimate 40% of continentally derived A_T enters the Atlantic, 20% enters the Arctic, and 40% enters all remaining ocean basins. These ocean areas represent 17%, 5%, and 78% of the total surface ocean area in our gridded dataset respectively, so the Artic receives approximately twice as much riverine A_T per unit area as the Atlantic, and 8 times the rest of the world ocean. The Atlantic has the lowest open-ocean surface *Alk** value and ~~the~~ low~~est~~ basin mean surface *Alk** despite the large riverine sources. The large riverine A_T input must therefore be more than balanced by strong net calcium carbonate formation. The Indian Ocean has comparably low mean surface *Alk** to the Atlantic, but a smaller riverine source. Mean *Alk** is higher in the Pacific than the Atlantic and Indian, even when neglecting the ~~upwelling~~ region north of 40°N as we do for the Atlantic (*Alk** = −16.5 µmol kg^{-1} when omitted vs. −22.9 µmol kg^{-1} for the Atlantic and −22.2 µmol kg^{-1} for the Indian). The difference between the Pacific and the other basins is significant when considering

the large number of grid cell *Alk** values averaged (> 6000 in the Atlantic), and the small estimated uncertainty for each value (~10 µmol kg^{-1}). Considering the weak Pacific riverine input, this suggests that, relative to other ocean basins, there are either larger *Alk** inputs from exchange with other basins and deeper waters or smaller Pacific basin mean net calcium carbonate formation.

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3.3 Riverine Alk regionally*

For river water with negligible salinity, *Alk** equals the potential alkalinity. This averages around 1100 µmol kg^{-1} globally (Cai et al., 2008), but is greater than 3000 µmol kg^{-1} for some rivers (Beldowski et al., 2010). Evidence suggests that riverine A_T is increasing due to human activities (Kaushal et al., 2013).

The most visible riverine *Alk** signals are in the Arctic due to the large riverine runoff into this comparatively small basin and the confinement of this low-density riverine water to the surface (Jones et al., 2008; Yamamoto-Kawai et al., 2009; Azetsu-Scott et al., 2010). Figure ~~4~~3d shows the high Arctic *Alk** plume is confined to the top ~200 m. Figure ~~3~~ 1 shows that these high *Alk** values extend along the coast of Greenland and through the Labrador Sea. *Alk** decreases with increasing salinity in this region (Fig. 5~~6~~d) due to mixing between the fresh high *Alk** surface Arctic waters and the salty lower *Alk** waters of the surface Atlantic. Gascard et al. (2004a, b) suggest that ~~high *Alk** ~~ waters along the coast of Norway are part of the Norwegian Coastal Current, and originate in the Baltic and North Seas where there are also strong riverine inputs (Thomas et al., 2005).

Elevated *Alk** can also be seen in the Bay of Bengal with surface values ~100 µmol kg^{-1} higher than those in the central Indian Ocean. This bay has two high A_T rivers that join and flow

into it, the Brahmaputra (A_T = 1114 µmol kg^{-1}) and the Ganges (A_T = 1966 µmol kg^{-1}) (Cai et al., 2008). Figure ~~7~~ 6b provides an *Alk** depth section for this region~~s for both areas~~. The riverine *Alk** plume can be clearly seen in the top 50 m ~~of the Bay of Bengal~~. No similar increase is seen in the Arabian Sea (Fig. 6a) where the Indus River (1681 µmol kg^{-1}) discharges only ~1/10th of the combined volume of the Brahmaputra and the Ganges.

The Amazon River is the largest single riverine marine A_T source. This river has low A_T (369 µmol kg^{-1} (Cai et al., 2008)), but has the largest water discharge volume of any river, exceeding the second largest—the Congo—by a factor of ~5 (Dai and Trenberth, 2002). Consequently, the Amazon discharges approximately 50% more A_T per year than the river with the second largest A_T discharge, the Changjiang (Cai et al., 2008). The Amazon's influence can be seen as a region of abnormally low salinity and A_T in Fig. 2a and b. Despite the high discharge volume, the influence is only barely visible as a region of elevated *Alk** in Fig. 2c due to the comparatively low Amazon *Alk**. However, the influence of the Amazon on *Alk** can be seen in the seasonal *Alk** cycle in the Amazon plume. Figure 7 provides a map of *Alk** for this region scaled to show the influence of this low *Alk** river in the Northern Hemisphere (a) winter and (b) summer months. The higher *Alk** found for summer months is consistent with Amazon discharge and A_T seasonality (Cooley et al., 2007) and Moore et al.'s (1986) radium isotope based finding that Amazon River outflow comprises 20-34% of surface water in this region in July compared to only 5-9% in December.

3.4 Regional ~~inorganic~~ abiotic carbonate cycling

The Red Sea portion of Fig. ~~8~~ 6a is strongly depleted in *Alk**, and contains the lowest single *Alk** measurement in our dataset, –247 µmol kg^{-1}. The GEOSECS expedition Red Sea

alkalinity measurements (Craig and Turekian, 1980) predate alkalinity reference materials (Dickson et al., 2007), but are supported by more recent measurements (Silverman et al., 2007). Like Jiang et al. (2014), we attribute low Red Sea *Alk** to exceptionally active calcium carbonate formation. Brewer and Dyrssen (1985) provide seawater chemistry measurements from the neighboring Persian Gulf that suggest strong calcium carbonate formation results in low *Alk** there as well (< –240 µmol kg^{-1} along the Trucial Coast).

The Red Sea is one of the only regions where Ω_C ~~calcium carbonate saturation~~ is sufficiently high for ~~inorganic~~ abiotic carbonate precipitation to significantly contribute to overall carbonate precipitation (Milliman et al. 1969; Silverman et al., 2007). Notably, saturation state remains high at depth in the Red Sea (see Section 4.2). In this region, biogenic aragonitic corals and pteropod shells are progressively removed with depth in sediments, and pores left behind are filled in with high-magnesium calcite cement (Gevirtz and Friedman, 1966; ~~Despite this, carbonate sediments in the modern Red Sea are mostly biogenic aragonitic corals and pteropod shells (Gevirtz and Friedman, 1966). However, in this region, pores in sediments are filled in with high-magnesium calcite cement (~~Almogi-Labin et al., 1986). We hypothesize biogenic carbonates are dissolved by CO_2 from sedimentary organic matter remineralization, as occurs elsewhere (e.g. Hales and Emerson, 1997; Hales, 2003; Boudreau, 2013), ~~but~~ and that high deep Red Sea Ω_C ~~calcium carbonate saturation~~ leads to ~~inorganic~~ abiotic re-calcification in sediment pores. Morse et al. (2006) find that synthetic high magnesium calcite—unlike biogenic high magnesium calcite—is less soluble than aragonite, so this substitution is favored thermodynamically if the abiotic mineral forms similarly to the synthetic mineral.

~~Inorganic c~~Calcium carbonate has recently been found as metastable ikaite (a hydrated ~~calcium carbonate~~ mineral with the formula $CaCO_3 \cdot 6H_2O$) in natural sea ice (Dieckmann et al.,

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2008). Ikaite cycling provides a competing explanation for the high Arctic surface *Alk** values if high A_T low-salinity ikaite-rich ice melt becomes separated from low A_T high-salinity rejected brines. However, riverine A_T inputs better explain the magnitude of the feature: The ~5 mg ikaite L^{-1} sea ice that Dieckmann et al. (2008) found in the Antarctic could only enrich A_T of the surface 100 m by ~1 µmol kg^{-1} for each meter of ice melted, and Arctic surface 100 m *Alk** is elevated by 59 µmol kg^{-1} relative to the deeper Arctic in our gridded dataset. By contrast, Jones et al. (2008) estimate a ~5% average riverine end-member contribution to the shallowest 100 m of this region, which accounts for ~55 µmol kg^{-1} *Alk** enrichment. Also, surface *Alk** in the Southern Ocean—which has sea ice but lacks major rivers—is not similarly elevated relative to surface phosphate (Fig. 2) or deep *Alk** (Fig. 3).

4. Controls on the calcite saturation state

The *Alk** tracer provides an opportunity to estimate the impact of carbonate cycling on Ω_C ~~the calcite saturation~~. In addition to (1) carbonate cycling, Ω_C ~~calcite saturation~~ is affected by (2) organic matter cycling, (3) freshwater cycling, (4) pressure changes on seawater, (5) heating and cooling, and (6) A_T changes from nitrogen fixation and denitrification. For each of these six processes, we estimate the standard deviation of the net influence of the process globally by considering the standard deviation of a "reference" tracer R_i for the process, "σ_{R_i}", where R_i is *Alk** for $CaCO_3$ cycling, phosphate for organic matter cycling, salinity for freshwater cycling, pressure for pressure changes, temperature for heating and cooling, and *N** (Gruber and Sarmiento, 1997) for nitrogen fixation and denitrification. We use the standard deviation of the reference tracer as a measure of the oceanic range of the net influence of the corresponding process. We measure the impact of this range on Ω_C ~~calcite saturation~~ using a

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metric *M*, which we define as:

$$M_i = \sigma_{R_i} \left| S_{R_i} \right| \quad (6)$$

where S_{R_i} is the Ω_C ~~calcite saturation~~ sensitivity to a unit process change in R_i, which we estimate in Appendix A. We are interested in the relative importance *I* of our 6 processes, so we also calculate the percentage that each metric value estimate contributes to the sum of all 6 metric value estimates:

$$I_i = 100\% \times \frac{M_i}{\sum_{i=1}^{6} M_i} \quad (7)$$

We derive and estimate our metric and its uncertainty in Appendix A. We carry out our analysis for the full water column assuming it to be isolated from the atmosphere (section 4.1), and also for just the top 50 m of the water column assuming it to be well-equilibrated with the atmosphere (section 4.2). Finally, we consider how equilibration with an atmosphere with a changing pCO_2 alters surface Ω_C ~~calcite saturation~~.

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4.1 Process importance in atmospherically-isolated mean seawater from all ocean depths

Our metric M_i is an estimate of the standard deviation of the global distribution of Ω_C resulting from the *i*th process. Our relative process importance metric I_i is an estimate of the percentage of overall variability of the Ω_C distribution that can be attributed to that process. We provide *M* and *I* values for mean seawater from the full water column alongside the R_i, S_{R_i}, and σ_{R_i} values used to estimate them in Table 1. These calculations assume that the seawater is isolated from the atmosphere.

Relative process importance estimates *I* indicate organic matter cycling (48%) is the

dominant process controlling Ω_C ~~calcite saturation~~ for mean seawater. Changing pressure (28%) is the second most important process, followed by calcium carbonate cycling (17%), temperature changes (4%), nitrogen fixation and denitrification (1.21%), and freshwater cycling (0.78%).

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4.2 Process importance in well-equilibrated surface seawater

In Table 2 we provide M_i values for well-equilibrated seawater in the top 50 m of the ocean alongside the R_i, σ_{R_i}, S_{R_i} used to estimate them. These surface seawater M_i values are calculated assuming the water remains equilibrated with an atmosphere with 400 µatm pCO_2. We test the validity of this assumption by also estimating M for the observed global pCO_2 variability in the Takahashi et al. (2009) global data product. This test reveals transient air-sea disequilibria are indeed important for surface ocean Ω_C ~~calcite saturation~~, but only as a secondary factor when considered globally. Despite this, it is important to recognize that air-sea equilibration following a process is not instantaneous, and that the S_{R_i} value estimates in section 4.1 ~~will~~ may be better for estimating short term changes following fast acting processes such as spring blooms (e.g. Tynan et al., 2014) or upwelling events (e.g. Feely et al., 1988). We omit the disequilibrium M value estimate from the denominator of Eq. (7) to allow I values for surface seawater to be compared to I values from mean seawater globally.

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Warming and cooling are the dominant processes controlling Ω_C for well-equilibrated surface seawater (76%). The large increase in M for warming and cooling relative to the value calculated for mean seawater is due to lower equilibrium C_T at higher temperatures. Freshwater cycling is the second most important process (13%), followed by carbonate cycling (8%), organic matter cycling (2%), pressure changes (1%), and denitrification and nitrogen fixation

(0.4%). The increased importance of freshwater cycling compared to section 4.1 is because freshwater dilutes C_T by more than the equilibrium C_T decreases from A_T dilution, so carbon uptake tends to follow freshwater precipitation and carbon outgassing follows evaporation. Carbonate cycling is less important because A_T decreases with carbonate precipitation lead to lower C_T at equilibrium. Organic matter cycling is much less important because atmospheric re-equilibration mostly negates the large changes in C_T. Pressure changes are negligible because we only consider water in the surface 50 m. Our air-sea disequilibrium *M* estimate suggests surface disequilibria are comparably important to freshwater cycling for surface ~~-~~ Ω_C ~~calcite saturation,~~ but substantially less important than temperature changes (this would correspond to an *I* value of ~14%).

The dominance of warming and cooling and freshwater cycling over carbonate cycling is most evident in the Red Sea where high temperatures (>25 °C) and high salinities (>40) lead to surface Ω_C ~~calcite saturations~~ exceeding 6 despite extremely low *Alk** (<–200 µmol kg^{-1}). The deep Red Sea is also unusual for having deep water that was warm when it last left contact with the atmosphere (the Red Sea is >20 °C at >1000 m depth). This provides high initial deep Ω_C ~~calcite saturation~~ that—combined with decreased influence of pressure changes at higher temperatures—keeps deep Red Sea $\Omega_C > 3$. Similarly, the lowest surface Ω_C values ~~saturation states~~ are in the Arctic where there are low temperatures, low salinity, and high *Alk** from riverine inputs. The importance of warming and cooling is also suggested by the correlation between global surface Ω_C ~~calcite saturation~~ and the surface temperature ($R^2 = 0.96$). These properties are plotted ~~shown for our gridded dataset~~ in Fig. 8.

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5. Conclusions

*Alk** isolates the portion of the A_T signal that varies in response to calcium carbonate cycling and exchanges with terrestrial and sedimentary environments from the portion that varies in response to freshwater and organic matter cycling. The salinity normalization we use has the advantage over previous salinity normalizations that it allows our tracer to mix linearly and to change in a 2:1 ratio with C_T in response to carbonate cycling. We highlight the following insights from *Alk**:

(1) Alk distribution*: The *Alk** distribution clearly shows the influence of biological cycling including such features as the very low *Alk** in the Red Sea due to the high calcium carbonate precipitation there. We also find evidence of strong riverine A_T sources in the Bay of Bengal and in the Arctic. We show river inputs likely dominate over the small influences of ikaite cycling on the Arctic alkalinity distribution.

(2) Influence of calcium carbonate cycling on marine calcite saturation state: *Alk** allows us to quantify the net influence of calcium carbonate cycling on marine Ω_C ~~calcite saturation~~. For well-equilibrated surface waters, carbonate cycling is less influential for Ω_C ~~calcite saturation~~ than gas exchange driven by warming and cooling and freshwater cycling. At depth, the carbonate cycling signal is smaller than the signal from organic matter cycling and from pressure changes. Temperature is the dominant control on Ω_C ~~calcite saturation~~ of surface waters in equilibrium with the atmosphere. This accounts for the low calcite saturation states in the cold surface of the Arctic and Southern Oceans despite high regional *Alk**, and high Ω_C ~~calcite saturations~~ in the warm subtropics despite low regional *Alk**.

We intend to use *Alk** for two future projects. First, *Alk** is superior to A_T for monitoring

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and modeling changes in marine chemistry resulting from changes in carbonate cycling with ocean acidification. A_T varies substantially in response to freshwater cycling, so *Alk** trends may be able to be detected sooner and more confidently attributed to changes in calcium carbonate cycling than trends in A_T. ~~Preliminary explorations of Earth System Model output suggest time of trend emergence for the alkalinity trends discussed by (~~Ilyina et al.~~(~~, 2009~~) could be reduced by as much as a factor of 5~~. Secondly, we will estimate global steady state *Alk** distributions using *Alk** sources and sinks from varied biogeochemical ocean circulation models alongside independent water mixing and transport estimates (e.g. Khatiwala et al., 2005; Khatiwala, 2007). We will interpret findings in the context of two hypotheses proposed to explain evidence for calcium carbonate dissolution above the aragonite saturation horizon: (1) that organic matter remineralization creates undersaturated microenvironments that promote carbonate dissolution in portions of the water column which are chemically supersaturated in bulk, and (2) that high-magnesium calcite and other impure minerals allow chemical dissolution above the saturation horizon.

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References

407 Almogi-Labin, A., B. Luz, and J. Duplessy (1986), Quaternary paleo-oceanography, pteropod
408 preservation and stable-isotope record of the Red Sea, *Palaeogeogr., Palaeoclimateol.,*
409 *Palaeoecol.*, *57*, 195-211, ~~DOI~~doi: 10.1016/0031-0182(86)90013-1.
410 Anderson, L. A. and J. L. Sarmiento (1994), Redfield ratios of remineralization determined by
411 nutrient data analysis, *Global Biogeochem. Cycles*, *8*, 65-80, doi: 10.1029/93GB03318.
412 Azetsu-Scott, K., A. Clarke, K. Falkner, J. Hamilton, E. P. Jones, C. Lee, B. Petrie, S.
413 Prinsenberg, M. Starr, and P. Yeats (2010), Calcium carbonate saturation states in the waters
414 of the Canadian Arctic Archipelago and the Labrador Sea, *J. Geophys. Res. Oceans*, *115*,
415 C11. doi: 10.1029/2009JC005917.
416 Beldowski, J., A. Löffler, B. Schneider, and L. Joensuu (2010), Distribution and biogeochemical
417 control of total CO_2 and total alkalinity in the Baltic Sea, *J. Mar. Sys.*, *81*, 252-259.
418 doi:10.1016/j.jmarsys.2009.12.020.
419 Berelson, W. M., W. M. Balch, R. Najjar, R. A. Feely, C. Sabine, and K. Lee (2007), Relating
420 estimates of $CaCO_3$ production, export, and dissolution in the water column to measurements
421 of $CaCO_3$ rain into sediment traps and dissolution on the sea floor: A revised global
422 carbonate budget, *Global Biogeochem. Cycles*, *21*, GB1024, doi: 10.1029/2006GB002803.
423 Boudreau, B. P. (2013), Carbonate dissolution rates at the deep ocean floor, *Geophys. Res. Lett.*,
424 *40*, 1-5, doi: 10.1029/2012GL054231.
425 Brewer, P. G., and D. Dyrssen (1985), Chemical oceanography of the Persian Gulf. *Prog.*
426 *Oceanogr.*, *14*, 41-55, doi:10.1016/0079-6611(85)90004-7.
427 Brewer, P. G., G. T. F. Wong, M. P. Bacon, D. W. Spencer (1975), An oceanic calcium
428 problem? *Earth and Planet. Sci. Lett.*, *26* (1), 81-87, doi: 10.1016/0012-821X(75)90179-X.
429 Cai, W.-J. X. Guo, C. A. Chen, M. Dai, L. Zhang, W. Zhai, S. E. Lohrenz, K. Yin, P. J. Harrison,

Y. Wang (2008), A comparative overview of weathering intensity and HCO_3^- flux in the world's major rivers with emphasis on the Changjiang, Huanghe, Zhujiang (Pearl) and Mississippi Rivers, *Contental Shelf Res.*, *28*, 1538-1549, doi:10.1016/j.csr.2007.10.014.

Chen, C.-T. A. (2002), Shelf-vs. dissolution-generated alkalinity above the chemical lysocline, *Deep Sea Res. II*, *49* (24–25), 5365-5375, doi: 10.1016/S0967-0645(02)00196-0.

Cooley, S. R., V. J. Coles, A. Subramaniam, and P. P. Yager (2007), Seasonal variations in the Amazon plume-related atmospheric carbon sink, *Global Biogeo. Chem. Cycles.*, *21* (3), GB3014, doi: 10.1029/2006GB002831.

Craig, H., and K.K. Turekian (1980), The GEOSECS program 1976-1979, *Earth Planet. Sci. Lett.*, *49*, 263-265, doi: 10.1016/j.bbr.2011.03.031.

Dai, A. and K. E. Trenberth (2002), Estimates of freshwater discharge from continents: latitudinal and seasonal variations, *J. Hydrometeorology*, *3*, 660-687, doi: http://dx.doi.org/10.1175/1525-7541(2002)003<0660:EOFDFC>2.0.CO;2.

Dickson, A.G. (1981), An exact definition of total alkalinity and a procedure for the estimation of alkalinity and total inorganic carbon from titation data, *Deep Sea Res. A, 28 (6)*, 609-623, doi: 10.1016/0198-0149(81)90121-7.

Dickson, A. G. and F. J. Millero (1987), A comparison of the equilibrium constants for the dissociation of carbonic acid in seawater media, *Deep-Sea Res. A*, *34*, 1733-1743, doi:10.1016/0198-0149(87)90021-5.

Dieckmann, G.S., G. Nehrke, S. Papadimitriou, J. Göttlicher, R. Steininger, H. Kennedy, D. Wolf-Gladrow, and D. N. Thomas (2008), Calcium carbonate as ikaite crystals in Antarctic sea ice, *Geophys. Res. Lett.* , *35*, LO8051, doi:10.1029/2008GL033540.

Feely, R. A., C. L. Sabine, K. Lee, F. J. Millero, M. F. Lamb, D. Greeley, J. L. Bullister, R. M.

Key, T. H. Peng, and A. Kozyr (2002), In situ calcium carbonate dissolution in the Pacific Ocean~~.~~. *Global Biogeochem. Cycles*, *16*, 1144, doi: 10.1029/2002GB001866.

Feely, R. A., R.H. Byrne, J. G. Acker, P. R. Betzer, C. A. Chen, J. F. Gendron, and M. F. Lamb (1988), Winter-summer variations of calcite and aragonite saturation in the northeast Pacific. *Mar. Chem.* 25, *3*, 227-241, doi:10.1016/0304-4203(88)90052-7.~~.~~

Fofonof, N. P., and R. C. Millard (1983), Algorithms for computations of fundamental properties of seawater. UNESCO Technical Papers in Marine Science No. 44, 53 pp.

Gascard, J. C., G. Raisbeck, S. Sequeira, F. Yiou, and K. Mork (2004), Correction to 'The Norwegian Atlantic Current in the Lofoten basin inferred from hydrological and tracer data(I-129) and its interaction with the Norwegian Coastal Current'. *Geophys. Res. Lett.*, *31*(8), doi: 10.1029/2003GL018303.~~.~~

Gascard, J. C., G. Raisbeck, S. Sequeira, F. Yiou, and K. Mork (2004), Correction to 2003GL01803, *Geophys. Res. Lett.*, *31*, L08302, doi:10.1029/2004GL020006, 2004.

Gevirtz, J. L., and G. M. Friedman (1966), Deep-Sea carbonate sediments of the Red Sea and their implications on marine lithification, *J. Sed. Petrol.*, *36*, 143-151.

Gruber, N., and J. L. Sarmiento (1997), Global patterns of marine nitrogen fixation and denitrification, *Global Biogeochem.~~ical~~ Cycles*, *11*(2), 235-266, doi: 10.1029/97GB00077.~~.~~

Hales, B. (2003), Respiration, dissolution, and the lysocline. *Paleoceanogr.*, *18*(4), 1099, doi: 10.1029/2003PA000915.~~.~~

Hales, B., and S. Emerson (1997), Calcite dissolution in sediments of the Ceara Rise: In situ measurements of porewater O_2, pH, and CO_2 (aq). *Geochim. Cosmochim. Acta*, *61*(3), 501-514, doi:10.1016/S0016-7037(96)00366-3.~~.~~

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475 Honjo, S., S. J. Manganini, R. A. Krishfield, and R. Francois (2008), Particulate organic carbon
476 fluxes to the ocean interior and factors controlling the biological pump: A synthesis of global
477 sediment trap programs since 1983, *Prog. Oceanogr.*, *76*(3), 217-285,
478 doi:10.1016/j.pocean.2007.11.003.
479 Ilyina, T. R. E. Zeebe, E. Maier-Reimer, and C. Heinze (2009), Early detection of ocean
480 acidification effects on marine calcification, *Global Biogeochem. Cycles*, *23*, GB1008, doi:
481 10.1029/2008GB003278.
482 Jiang, Z. P., T. Tyrrell, D.J. Hydes, M. Dai, and S.E. Hartman (2014), Variability of alkalinity
483 and the alkalinity-salinity relationship in the tropical and subtropical surface ocean. *Global*
484 *Biogeochem. Cycles*, *28*(7), 729-742, doi: 10.1002/2013GB004678.
485 Jones, E. P., L. G. Anderson, S. Jutterström, L. Mintrop, and J. H. Swift (2008), Pacific
486 freshwater, river water and sea ice meltwater across Arctic Ocean basins: Results from the
487 2005 Beringia Expedition, *J. Geophys. Res. Oceans*, *113*(C8), doi: 10.1029/2007JC004124.
488 Kanamori, S. and H. Ikegami (1982), Calcium-alkalinity relationship in the North Pacific, *J.*
489 *Oceanogr.*, *38*, 57-62, doi: 10.1007/BF02110291.
490 Kaushal, S. S., G. E. Likens, R. M. Utz, M. L. Pace, M. Grese, and M. Yepsen (2013), Increased
491 river alkalinization in the Eastern U.S., *Envi. Sci. Tech.*, *47*, 10302-10311, doi:
492 10.1021/es401046s.
493 Key, R. M., A. Kozyr, C. L. Sabine, K. Lee, R. Wanninkhof, J. L. Bullister, R. A. Feely, F. J.
494 Millero, C. Mordy, and T. H. Peng (2004), A global ocean carbon climatology: Results from
495 Global Data Analysis Project (GLODAP), *Global Biogeochem. Cycles*, *18*, GB4031, doi:
496 10.1029/2004GB002247.
497 Key, R. M., T. Tanhua, A. Olsen, M. Hoppema, S. Jutterström, C. Schirnick, S. van Heuven, X.

Lin, D. Wallace and L. Mintrop (2009), The CARINA data synthesis project: Introduction and overview, *Earth Sys. Sci. Data*, *2(1)*, 579-624, doi:10.5194/essdd-2-579-2009.

Khatiwala, S., M. Visbeck, and M. A. Cane, (2005), Accelerated simulation of passive tracers in ocean circulation models, *Ocean Modelling*, *9*(1), 51-69, doi:10.1016/j.ocemod.2004.04.002.

Khatiwala, S. (2007), A computational framework for simulation of biogeochemical tracers in the ocean, *Global Biogeochemical Cycles*, *21*(3), GB3001, doi: 10.1029/2007GB002923.

Mehrbach, C., C. H. Culberson, J. E. Hawley, and R. M. Pytkowicz (1973), Measurement of the apparent dissociation constants of carbonic acid in seawater at atmospheric pressure, *Limnol. Oceanogr.*, *18*, 897-907.

Millero, F. J., K. Lee, and M. Roche, (1998), Distribution of alkalinity in the surface waters of the major oceans, *Marine Chemistry*, *60*, 111-130, doi:10.1016/S0304-4203(97)00084-4.

Milliman, J. D., D. A. Ross, and T. L. Ku (1969), Precipitation and lithification of deep-sea carbonates in the Red Sea, *J. Sed Res.*, *39*(2), 724-736, doi: 10.1306/74D71CFD-2B21-11D7-8648000102C1865D.

Morse, J. W., A. J. Andersson, F. T. Mackenzie (2006), Initial responses of carbonate-rich shelf sediments to rising atmospheric pCO_2 and "ocean acidification": Role of high Mg-calcites. *Geochim. Cosmochm. Acta.*, *70* (23), 5814-5830, doi:10.1016/j.gca.2006.08.017.

Moore, W. S. (2010), The effect of submarine groundwater discharge on the ocean. *Marine Sci.*, *2*, 59-88, doi: 10.1146/annurev-marine-120308-081019

Moore, W. S., J. L. Sarmiento, and R. M. Key (1986), Tracing the Amazon component of surface Atlantic water using ^{228}Ra, salinity, and silica, *J. Geophys. Res.*, *91* (*C2*), 2574-2580, doi: 10.1029/JC091iC02p02574.

Orr, J. C., V. J. Fabry, O. Aumont, L. Bopp, S. C. Doney, R. A. Feely, A. Gnanadesikan, N.

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Gruber, A. Ishida, F. Joos, R. M. Key, K. Lindsay, E. Maier-Reimer, R. Matear, P. Monfray, A. Mouchet, R. G. Najjar, G. Plattner, K. B. Rodgers, C. L. Sabine, J. L Sarmiento, R. Schlitzer, R. D. Slater, I. J. Totterdell, M. Weirig, Y. Yamanaka, and A. Yool, (2005), Anthropogenic ocean acidification over the twenty-first century and its impact on calcifying organisms. *Nature*, *437*, 681-686, doi:10.1038/nature04095.

Robbins, P. E. (2001), Oceanic carbon transport carried by freshwater divergence: Are salinity normalizations useful?. *J. Geophys. Res*, *106*(C12), 30939-30,~~.~~ doi: 10.1029/2000JC000451.

Sarmiento, J. L., J. Dunne, A. Gnanadesikan, R.M. Key, K. Matsumoto, R. Slater (2002), A new estimate of the $CaCO_3$ to organic carbon export ratio. *Global Biogeochem. Cy.*, *16*(4), 1107, doi: 10.1029/2002GB001919.~~.~~

Sarmiento, J. L., N. Gruber, M. A. Brzezinski, and J. P. Dunne (2004), High-latitude controls of thermocline nutrients and low latitude biological productivity. *Nature*, *427*(6969), 56-60, doi:10.1038/nature02127.~~.~~

Silverman, J. B. Lazar, and J. Erez (2007), Effect of aragonite saturation, temperature, and nutrients on the community calcification rate of a coral reef, *J. Geophys. Res.*, *112*, CO05004, ~~DOI~~doi: 10.1029/2006JC003770.

Suzuki, T., M. Ishii, M. Aoyama, J. R. Christian, K. Enyo, T. Kawano, R. M. Key, N. Kosugi, A. Kozyr, L. A. Miller, A. Murata, T. Nakano, T. Ono, T. Saino, K. Sasaki, D. Sasano, Y. Takatani, M. Wakita and C. Sabine (2013), PACIFICA Data Synthesis Project. ORNL/CDIAC-159, NDP-092. Carbon Dioxide Information Analysis Center, Oak Ridge National Laboratory, U.S. Department of Energy, Oak Ridge, Tennessee. doi:10.3334/CDIAC/OTG.PACIFICA_NDP092.

Takahashi, T., S. C. Sutherland, R. Wanninkhof, C. Sweeney, R. A. Feely, D. W. Chipman, B.

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544 Hales, G. Friederich, F. Chavez, C. Sabine, A. Watson, D. C. E. Bakker, U. Schuster, N.
545 Metzl, H. Yoshikawa-Inoue, M. Ishii, T. Midorikawa, Y. Nojiri, A. Körtzinger, T. Steinhoff,
546 M. Hoppema, J. Olafsson, T. S. Arnarson, B. Tilbrook, T. Johannessen, A. Olsen, R.
547 Bellerby, C. S. Wong, B. Delille, N. R. Bates, and J. W. deBarr, (2009), Climatological mean
548 and decadal change in surface ocean pCO_2, and net sea–air CO_2 flux over the global
549 oceans, *Deep Sea Res. II*, *56*(8), 554-577, doi:10.1016/j.dsr2.2008.12.009.

550 Thomas, H., Y. Bozec, H. J. De Baar, K. Elkalay, M. Frankignoulle, L. S. Schiettecatte, G.
551 Kattner, and A. V. Borges (2005), The carbon budget of the North Sea. *Biogeosci.*, *2*(1), 87-
552 96, doi: 10.5194/bg-2-87-2005.

553 Tynan, E., T. Tyrrell, and E. P. Achterberg (2014), Controls on the seasonal variability of
554 calcium carbonate saturation states in the Atlantic gateway to the Arctic Ocean. *Mar. Chem.*
555 158 (2014) 1-9., doi: 10.1016/j.marchem.2013.10.010.

556 van Heuven, S., D. Pierrot, E. Lewis, and D. Wallace (2009), MATLAB Program developed for
557 CO_2 system calculations, *ORNL/CDIAC-105b, Carbon Dioxide Information Analysis Center,*
558 *Oak Ridge National Laboratory, US Department of Energy, Oak Ridge, Tennessee*.

559 Velo, A., F. F. Perez, P. Brown, T. Tanhua, U. Schuster, and R. M. Key (2009), CARINA
560 alkalinity data in the Atlantic Ocean, *Earth Syst. Sci. Data*, *1*, 45-61, doi:10.5194/essd-1-45-
561 2009.

562 de Villiers, S. (1998), Excess dissolved calcium in the ocean: a hydrothermal hypothesis, *Earth*
563 *and Plan. Sci. Lett.*, *164(3-4)*, 624-641, doi:10.1016/S0012-821X(98)00232-5.

564 Wolery, T. J., and N. H. Sleep (1988), Interactions of geochemical cycles with the mantle. In:
565 Gregor, C. B., R. M. Garrels, F. T. Mackenzie, and J. B. Maynard (eds) Chemical cycles in
566 the evolution of the earth. Wiley, New York, 77-103.

567 Wolf-Gladrow, D. A., R. E. Zeebe, C. Klaas, A. Körtzinger, and A. G. Dickson (2007), Total
568 alkalinity: The explicit conservative expression and its application to biogeochemical
569 processes. *Marine chemistry*, *106*(1), 287-300, doi:10.1016/j.marchem.2007.01.006.
570 Yamamoto-Kawai, M., F. A. McLaughlin, E. C. Carmack, S. Nishino, and K. Shimada (2009),
571 Aragonite undersaturation in the Arctic Ocean; effects of ocean acidification and sea ice
572 melt, *Science*, *326*, 1098, doi:10.1126/science.1174190.

573 **Appendix A: Definition of the process importance metric M**

574 In simplest terms, our metric is the product of the Ω_C ~~calcite saturation~~ sensitivity to a

575 process and the variability of the net influence of the process globally. The difficulty in this

576 calculation lies in quantifying the "net influence of a process." We first show how we change

577 coordinates so we can use reference tracers as a proxy measurement for these net influences.

578 Our metric for Ω_C variability resulting from the ith process is expressed as M_i:

579 $$M_i = \sigma_{P_i} \left| \frac{\partial \Omega_C}{\partial P_i} \right| \quad \text{(A1)}$$

580 where P_i is an abstract variable representing the net process influence (that we will later factor

581 out), and $\frac{\partial \Omega_C}{\partial P_i}$ is the Ω_C ~~calcite saturation~~ sensitivity to the process. We expand $\frac{\partial \Omega_C}{\partial P_i}$ using the

582 chain rule to include a term for Ω_C sensitivity to changes in the reference tracer R_i (see section 4)

583 and a term $\frac{\partial R_i}{\partial P_i}$ representing changes in R_i resulting from the ith process:

584 $$\frac{\partial \Omega_C}{\partial P_i} = \frac{\partial \Omega_C}{\partial R_i} \frac{\partial R_i}{\partial P_i} \quad \text{(A2)}$$

585 In practice, we calculate Ω_C as a function of $j = 7$ properties: (1) pressure, (2)

586 temperature, (3) salinity, (4) phosphate, (5) silicate, (6) A_T, and (7) C_T for mean seawater and

587 pCO_2 for surface seawater, so we use the chain rule again to expand the $\frac{\partial \Omega_C}{\partial R_i}$ terms as follows:

588 $$\frac{\partial \Omega_C}{\partial R_i} = \sum_{j=1}^{7} \frac{\partial \Omega_C}{\partial X_j} \frac{\partial X_{j,i}}{\partial R_i} \quad \text{(A3)}$$

589 Here, the $\frac{\partial X_{j,i}}{\partial R_i}$ are assumed terms (assumptions detailed shortly) that relate the effect of the ith

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process on the *j*th property to the effect of the process on R_i, and the $\frac{\partial \Omega}{\partial X_j}$ terms reflect Ω_C ~~calcite saturation~~ sensitivity to changes in the *j* properties used to calculate it.

We make assumptions regarding the $\frac{\partial X_{j,i}}{\partial X_R}$ terms: we relate changes in temperature from sinking or shoaling to changes in pressure using the potential temperature (θ) routines of Fofonoff and Millard (1983); we assume freshwater cycling linearly concentrates A_T, C_T, phosphate, and silicate by the same ratio that it changes salinity; we relate C_T, phosphate, and A_T changes from organic matter formation to changes in phosphate using the remineralization ratios found by Anderson and Sarmiento (1994) and the empirical relationship of Kanamori and Ikegami (1982); we also use Kanamori and Ikegami (1982)'s constant to relate changes in A_T from nitrogen fixation and denitrification to changes in N* from these processes; and we assume that an increase in A_T from calcium carbonate dissolution equals the *Alk** increase, and that the corresponding increase in C_T equals half of this *Alk** increase. We neglect any changes in C_T from denitrification and nitrogen fixation because these changes are better thought of as organic matter cycling occurring alongside nitrogen cycling.

We estimate $\frac{\partial \Omega}{\partial X_j}$ property sensitivity terms as the differences between Ω_C calculated before and after augmenting *j*th property by 1 unit. Ω_C is calculated with the MATLAB CO2SYS routines written by van Heuven et al. (2009) using the carbonate system equilibrium constants of Mehrbach et al. (1973), as refit by Dickson and Millero (1987). Seawater pCO_2 is used in place of C_T for the surface seawater calculations (when $j = 7$) to calculate the change in Ω_C that remains after the surface seawater is allowed to equilibrate with the atmosphere.

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610 We assume that the distributions of our R_i reference properties are linearly related to the
611 P_i net activities of their associated processes. This assumption implies:

612 $$\sigma_P = \sigma_{R_i} \left| \frac{\partial P_i}{\partial R_i} \right| \quad \text{(A4)}$$

613 We can then substitute Eq. (A3) into Eq. (A2), and substitute this combined equation for $\frac{\partial \Omega_C}{\partial P_i}$
614 and (A4) into Eq. (A1). We then ~~and~~ cancel the $\frac{\partial P_i}{\partial R_i}$ and $\frac{\partial R_i}{\partial P_i}$ terms to obtain:

615 $$M_i = \sigma_{R_i} \left| \sum_{j=1}^{7} \frac{\partial \Omega_C}{\partial X_j} \frac{\partial X_{j,i}}{\partial R_i} \right| \quad \text{(A5)}$$

616 We then define ~~our~~ Ω_C ~~saturation~~ sensitivity S_{R_i} as:

617 $$S_{R_i} = \left| \sum_{j=1}^{7} \frac{\partial \Omega_C}{\partial X_j} \frac{\partial X_{j,i}}{\partial R_i} \right| \quad \text{(A6)}$$

618 where S_{R_i} is the Ω_C ~~saturation~~ sensitivity to a change in the *i*th process scaled to a unit change in
619 the reference variable for that process. We can then substitute Eq. (A6) into Eq. (A5) to obtain
620 Eq. 6. We use Eqn. (A6) to define S_{R_i} and Eqn. 6 to calculate *M*. We provide the $\frac{\partial \Omega_C}{\partial X_j}$ and
621 $\frac{\partial X_{j,i}}{\partial R_i}$ values we use to estimate S_{R_i} for atmospherically isolated seawater from all depths in
622 Table A1 and for well-equilibrated surface seawater in Table A3. We perform a sample *I* and *M*
623 calculation in Supplementary Materials document SE.

624 We use a Monte Carlo analysis to estimate variability and uncertainty in our metric *M*
625 and our percent relative process importance *I* calculations. We calculate the standard deviations,
626 σ_M and σ_I, of pools of 1000 *M* and *I* estimates calculated after adjusting the seawater properties

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627 X_i with a normally-distributed perturbation with a standard deviation equal to the property

628 standard deviation from the gridded dataset. We find $\frac{\sigma_I}{I}$ is typically much smaller than $\frac{\sigma_M}{M}$.

629 This is because Ω_C ~~calcite saturation~~ sensitivity is typically proportional to the Ω_C ~~calcite~~

630 ~~saturation~~ itself, so individual Monte Carlo *M* estimates vary with the initial Ω_C ~~initial calcite~~

631 ~~saturation~~ and one another. Our σ_M estimates are therefore better thought of as measures of the

632 ranges of sensitivities found in the modern ocean, while σ_I represent variability in the relative

633 importance of processes. We provide σ_M and σ_I for atmospherically isolated seawater globally

634 in Table A2, and for well-equilibrated surface seawater in Table A4.

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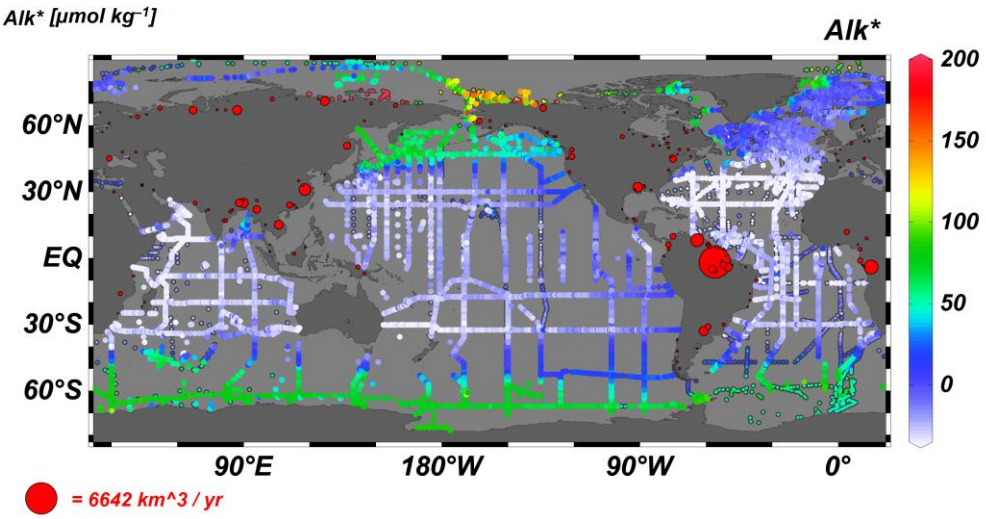

Figure 1. A map of station locations at which we use measurements to estimate *Alk** (in µmol kg^{-1}). Dot color indicates surface *Alk**. Points with black borders indicate that either A_T was measured prior to 1992 (i.e. before reference materials were commonly used) or that no nitrate value was reported (in which case a nitrate concentration of 5 µmol kg^{-1} is assumed). Red dots on land indicate the mouth locations and mean annual discharge volumes (indicated by dot size) of 200 large rivers, as given by Dai and Trenberth (2002).

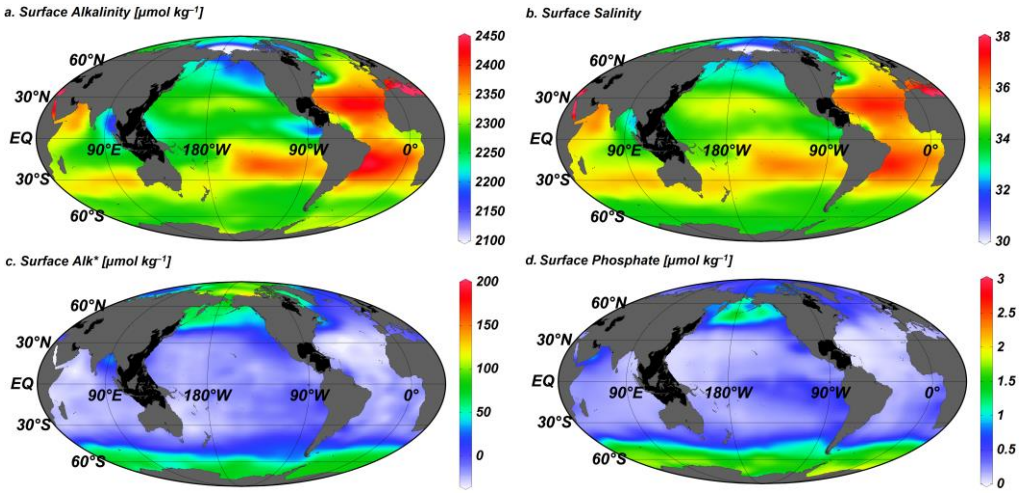

Figure 2. Global (a) total alkalinity A_T, (b) salinity, (c) *Alk**, and (d) phosphate distributions at the surface (10 m depth surface) from our gridded CARINA, PACIFICA, and GLODAP bottle data product detailed in Supplementary Materials document SA. Areas with exceptionally poor coverage in the data used to produce the gridded product are blacked out.

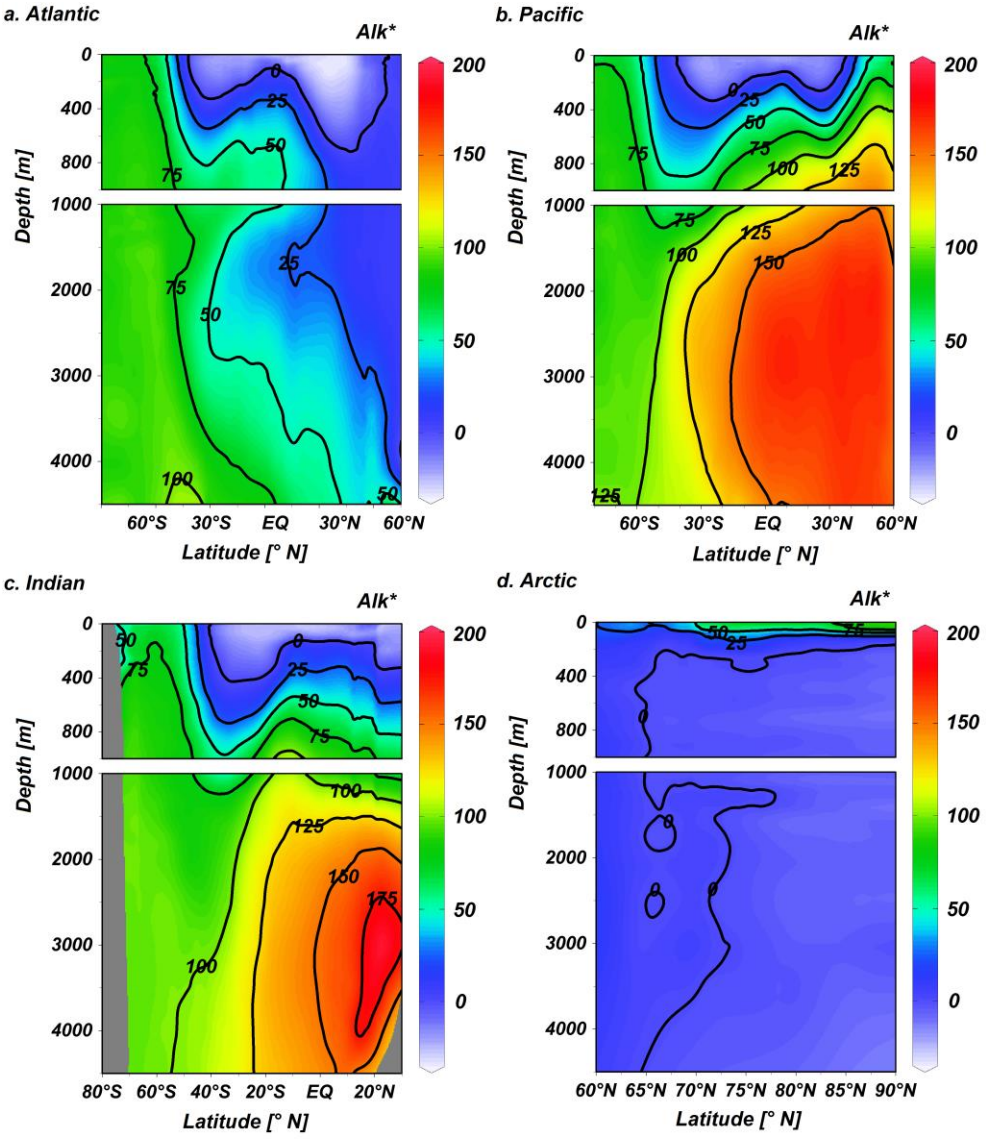

Figure 3. Zonal mean gridded *Alk** (in µmol kg^{-1}) in the (a) Atlantic, (b) Pacific, (c) Indian, and (d) the Arctic oceans plotted against latitude and depth.

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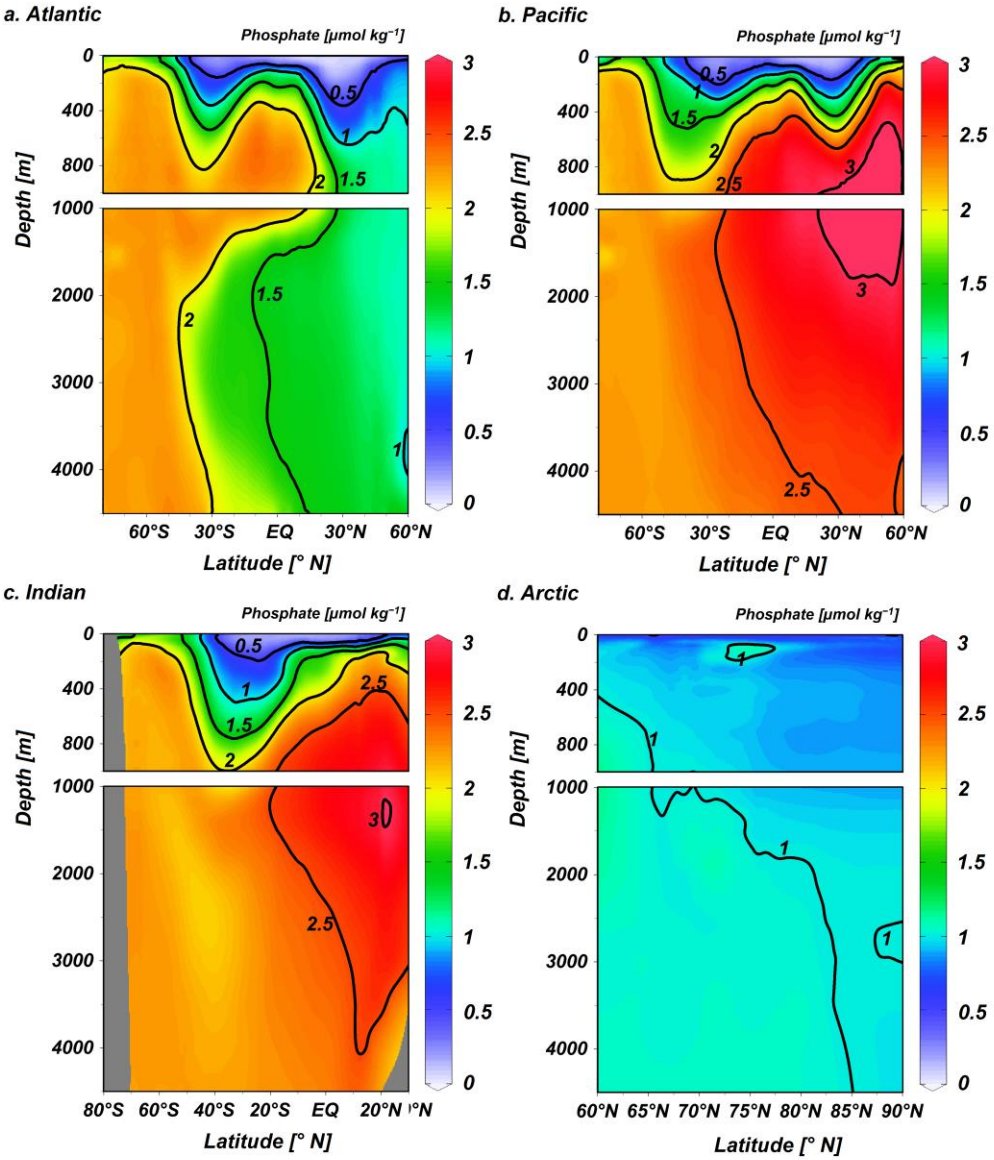

Figure 4. Zonal mean gridded phosphate (in µmol kg^{-1}) in the (a) Atlantic, (b) Pacific, (c) Indian, and (d) the Arctic oceans plotted against latitude and depth.

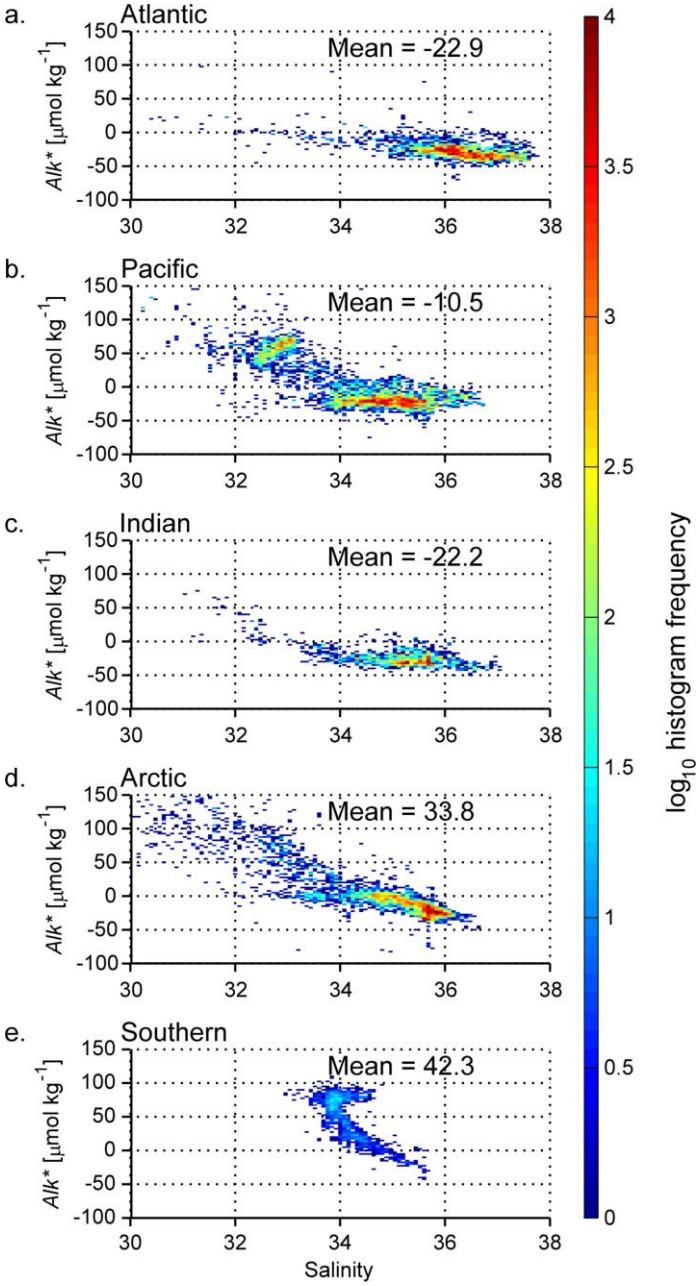

Figure 5. 2-D histograms indicating the log (base 10) of the number of measurements that fall within bins of *Alk** vs. salinity with color. Data are limited to the top 50 m of the (a) Atlantic, (b) Pacific, (c) Indian, (d) Arctic, and (e) Southern Oceans. Where basins connect, the boundary between the Atlantic and the Arctic oceans is 40°N, between the Atlantic and the Indian is 20° E,

between the Indian and the Pacific is 131° E, between the Pacific and the Atlantic is 70° W, and between the Southern Ocean and the other oceans is 40°S.
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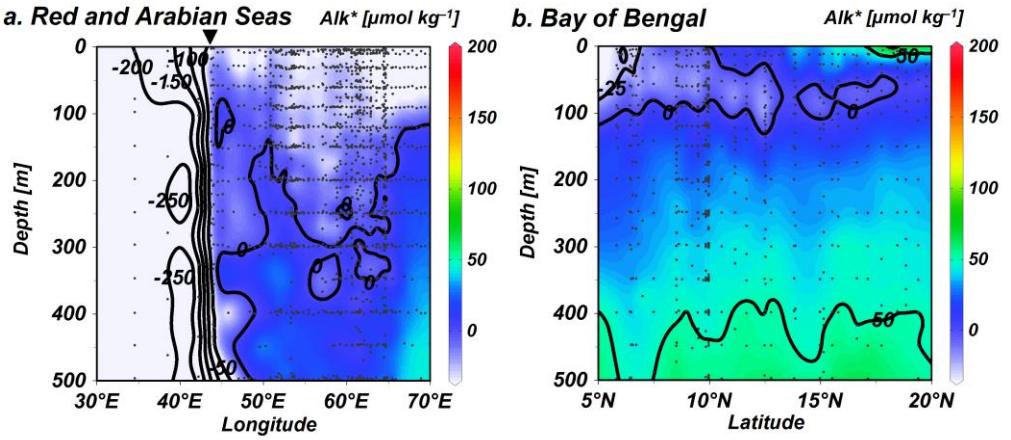

Figure ~~7~~6. *Alk** distributions (in µmol kg^{-1}) (a) between 5° and 30°N in the Red and Arabian Seas shown against longitude, and (b) between 75° and 100° E in the Bay of Bengal plotted against latitude. Small black dots indicate where data is present. The inverted triangle above (a) indicates the longitude of the mouth of the Red Sea.

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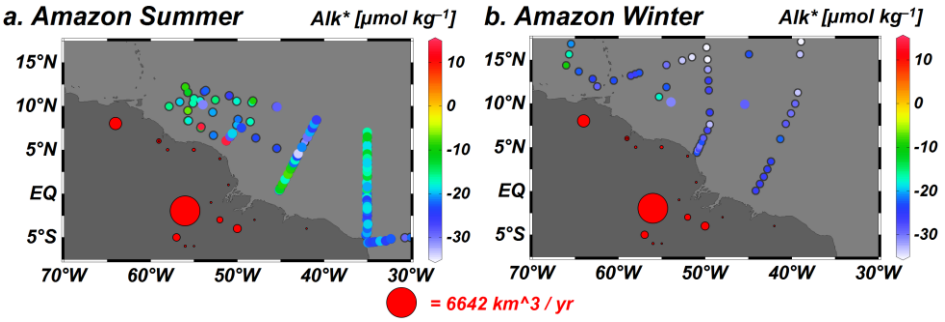

Figure ~~6~~7. *Alk** (in µmol kg^{-1}) in top 50 m of the ocean near the Amazon River outflow plotted in color, though with a narrower color scale than is used for all other plots. Panel (a) is limited to data collected in November through January, and in panel (b) is limited to measurements from May through July. Points with black borders indicate that either the A_T was measured prior to 1992 (before reference materials were commonly used) or that no nitrate value was reported (in which case a nitrate concentration of 5 µmol kg^{-1} is assumed). Red dots on land indicate the mouth locations and mean annual discharge volumes (indicated by dot size) of large rivers, as given by Dai and Trenberth (2002).

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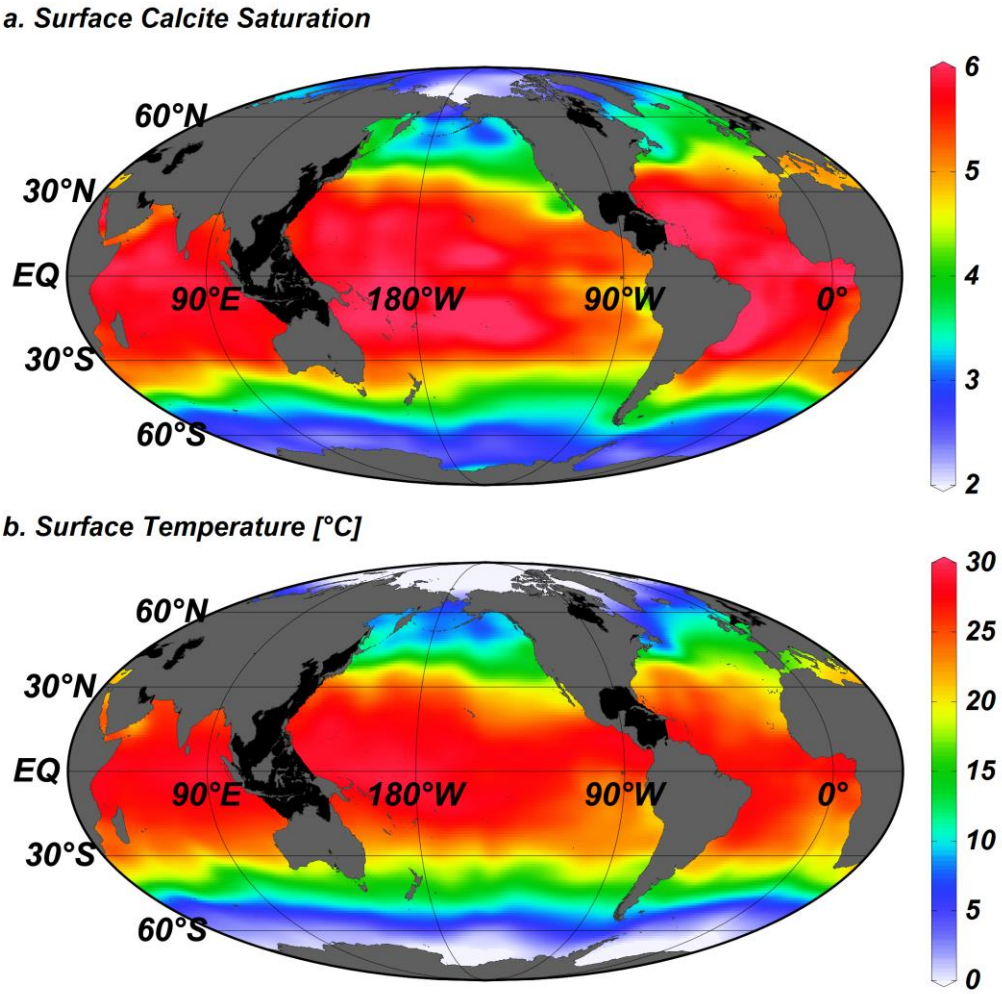

Figure 8. Gridded global (a) calcite saturation state Ω_C, and (b) temperature at the surface (10 m depth surface) of our gridded CARINA, PACIFICA, and GLODAP bottle data product. Areas with exceptionally poor coverage in the data used to produce the gridded product are blacked out.

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Table 1. Metric estimates M_i, relative process importance percentages I_i, Ω_C ~~calcite saturation~~ sensitivities S_{R_i} to unit changes in the R_i reference properties, and reference property standard deviations σ_{R_i} for the i = 6 processes in atmospherically isolated mean seawater from all ocean depths. See Appendix A for details on how these terms are estimated and explanation of how M_i and I_i uncertainties are obtained.

Process	i	R_i	S_{R_i}	σ_{R_i}	M_i	I_i
Carbonate cycling	1	*Alk**	0.0043	53.5 µmol/kg	0.23	17%
Org. matter cycling	2	Phosphate	−0.0069	0.60 µmol/kg	0.66	48%
Freshwater cycling	3	Salinity	0.032	0.27	0.011	0.78%
Sinking / shoaling	4	Pressure	−0.00028	1411 db	0.4	28%
Warming / cooling	5	Temp.	0.014	4.20 °C	0.06	4%
Denit./nit. fix.	6	N*	−0.010	1.6 µmol/kg	0.017	1.2%

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Table 2. Metric estimates M_i, relative process importance percentages I_i, Ω_C ~~calcite saturation~~ sensitivities S_{R_i} to unit changes in the R_i reference properties, and reference property standard deviations σ_{I_i} for the i = 6 processes in well-equilibrated surface seawater. See Appendix A for details on how these terms are estimated and explanation of how M_i and I_i uncertainties are obtained.

Process	i	R_i	S_{R_i}	σ_{R_i}	M_i	I_i
Carbonate cycling	1	*Alk**	0.0034	36.9 µmol/kg	0.13	7.8%
Org. matter cycling	2	Phosphate	−0.0045	0.51 µmol/kg	0.037	2.3%
Freshwater cycling	3	Salinity	0.20	0.86	0.22	13.2%
Sinking / shoaling	4	Pressure	−0.00083	15 db	0.011	0.70%
Warming / cooling	5	Temp.	0.14	8.8 °C	1.2	76%
Denit ./ nit. fix.	6	N*	−0.0043	1.5 µmol/kg	0.006	0.40%
pCO$_2$ disequilibria	†	pCO$_2$	−0.0086	27 µatm*	0.23	†

* standard deviation of the Takahashi et al. (2009) revised global monthly pCO$_2$ climatology

† the M value for disequilibria is only calculated to test our assumption of surface seawater air-sea equilibration, and is omitted from calculations of I_i for comparison with Table 1.

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Table A1. $\frac{\partial \Omega_C}{\partial X_j}$ (bold text) and $\frac{\partial X_{j,i}}{\partial R_i}$ (italic text) terms used in Eq. (A5) for atmospherically isolated mean seawater from all ocean depths. These terms are specific to the j = 7 (columns) properties we use to calculate Ω_C and i = 6 (rows) processes we consider. Units for $\frac{\partial \Omega_C}{\partial X_j}$ are the inverse of the listed X_j units. Units for $\frac{\partial X_{j,i}}{\partial R_i}$ are the X_j units divided by the R_i units given in Table 1.

Properties		Pressure	Temp	Salinity	Phos.	Silicate	A_T	C_T
X_j *units*		db	°C		µmol/kg	µmol/kg	µmol/kg	µmol/kg
j		1	2	3	4	5	6	7
Mean seawater values		2235	3.7	34.71	2.15	49.0	2362	2254
$\frac{\partial \Omega_C}{\partial X_j}$		**–0.00028**	**0.014**	**–0.011**	**–0.0085**	**–0.00012**	**0.0082**	**–0.0079**
Process	*i*	$\frac{\partial X_{1,i}}{\partial R_i}$	$\frac{\partial X_{2,i}}{\partial R_i}$	$\frac{\partial X_{3,i}}{\partial R_i}$	$\frac{\partial X_{4,i}}{\partial R_i}$	$\frac{\partial X_{5,i}}{\partial R_i}$	$\frac{\partial X_{6,i}}{\partial R_i}$	$\frac{\partial X_{7,i}}{\partial R_i}$
Carbonate cycling	1	-	-	-	-	-	*1*	*0.5*
Org. matter cycling	2	-	-	-	*1*	-	*–20.16*	*117*
Freshwater cycling	3	-	-	*1*	*0.062*	*1.4*	*68*	*65*
Sinking / shoaling	4	*1*	*0.00010*	-	-	-	-	-
Warming / cooling	5	-	*1*	-	-	-	-	-
Denit ./ nit. fix.	6	-	-	-	-	-	*–1.26*	-

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Table A2. Monte Carlo derived estimates for M_i variability (σ_{M_i}) and I_i variability (σ_{I_i}) for atmospherically-isolated mean seawater from all ocean depths.

Process	i	σ_{M_i}	σ_{I_i}
Carbonate cycling	1	0.09	1%
Org. matter cycling	2	0.2	3%
Freshwater cycling	3	0.006	0.08%
Sinking / shoaling	4	0.2	5%
Warming / cooling	5	0.02	2%
Denit. / nit. fix.	6	0.006	0.1%

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Table A3. $\frac{\partial \Omega_C}{\partial X_j}$ (bold text) and $\frac{\partial X_{j,i}}{\partial R_i}$ (italic text) terms used in Eq. (A5) for well-equilibrated surface seawater. These terms are specific to the $j = 7$ (columns) properties we use to calculate Ω_C and $i = 6$ (rows) processes we consider. Units for $\frac{\partial \Omega_C}{\partial X_j}$ are the inverse of the listed X_j units. Units for $\frac{\partial X_{j,i}}{\partial R_i}$ are the X_j units divided by the R_i units given in Table 2.

Properties		Pressure	Temp	Salinity	Phos.	Silicate	A_T	pCO_2
units		db	°C		µmol/kg	µmol/kg	µmol/kg	µatm
j		1	2	3	4	5	6	7
Mean seawater values		25	18.3	34.82	0.51	2.5	2305	350
$\frac{\partial \Omega_C}{\partial X_j}$		**–0.00084**	**0.14**	**–0.022**	**–0.0038**	**–0.00013**	**0.0034**	**–0.0086**
Process	*i*	$\frac{\partial X_{1,i}}{\partial R_i}$	$\frac{\partial X_{2,i}}{\partial R_i}$	$\frac{\partial X_{3,i}}{\partial R_i}$	$\frac{\partial X_{4,i}}{\partial R_i}$	$\frac{\partial X_{5,i}}{\partial R_i}$	$\frac{\partial X_{6,i}}{\partial R_i}$	$\frac{\partial X_{7,i}}{\partial R_i}$
Carbonate cycling	1	-	-	-	-	-	*1*	-
Org. matter cycling	2	-	-	-	*1*	-	*–20.16*	-
Freshwater cycling	3	-	-	*1*	*0.015*	*0.072*	*65.9*	-
Sinking / shoaling	4	*1*	*0.00010*	-	-	-	-	-
Warming / cooling	5	-	*1*	-	-	-	-	-
Denit./nit. fix.	6	-	-	-	-	-	*–1.26*	-

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Table A4. Monte Carlo derived estimates for M_i variability (σ_{M_i}) and I_i variability (σ_{I_i}) for well-equilibrated surface seawater.

Process	i	σ_{M_i}	σ_{I_i}
Carbonate cycling	1	0.03	0.8%
Org. matter cycling	2	0.01	0.2%
Freshwater cycling	3	0.04	0.5%
Sinking / shoaling	4	0.001	0.03%
Warming / cooling	5	0.2	1%
Denit. / nit. fix	6	0.002	0.04%
pCO$_2$ disequilibria	†	0.05	†

† disequilibria are included only as a test of our assumption of surface seawater air-sea equilibration, so these M_i values are omitted from calculations of I

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