We thank the referees for their relevant and constructive comments and we appreciated the quality of the observations, which allowed us to improve the first version of our paper.

The referee comments are reported below with the answers we provided and the changes we did in the revised manuscript, attached at the end of this document.

Anonymous referee#1

This study investigates the hydrogen dynamics in soil organic matter to quantify processes such as the preservation of organic matter and microbial biosynthesis. Seems to be that this research is potentially useful to understand the fate of tritium (3H) in ecosystems. The approach described in this paper for determining the fate of hydrogen in soil systems using three types of labelling experiments (substance, substance/water, only water) is an original approach. However, my major comment is related to the assumptions used for hydrogen exchangeability, which were poorly explained. I believe this manuscript needs significant explanation about the hydrogen isotope analyses and modelling. I therefore recommend publication only after major revisions.

1. *H exchangeability – Soil organic matter could be a heterogeneous material in terms of hydrogen exchangeability. Uncontrolled isotopic exchange between sample and laboratory ambient vapour can introduce bias in δ2H measurements. The authors did not explicitly account for H exchangeability in their analysis by using the Comparative Equilibration method or the aid of devices that allows vapour equilibration before analysis.*

   Answer: We agree with the referee that some details of the methodology are missing in this work, especially to deal with the exchange of hydrogen. In this study, we did not use devices for equilibration but we compared the isotopic composition of unlabeled and highly labeled samples of soils after equilibration of the dried samples with laboratory atmosphere. No standards were chosen in the comparison but we assumed that unlabeled and labeled samples exposed to the same atmosphere before δ^2H measurement reach the same concentration in deuterium in the labile hydrogen pool. Then calculating the difference in ^2H abundance between unlabeled and labeled samples allows eliminating the contribution of labile hydrogen (short-time exchange) in the final isotopic calculation. This difference represents the isotopic composition of hydrogen that did not exchange during the length of equilibration. It is called non-exchangeable hydrogen in this paper. See also comments 11.

2. *Moreover, bulk soil samples without lipid extraction was conducted. As the authors pointed out, lipids do not usually exchange with atmospheric vapour because of the C-H bonds in their main structure. However, differential lipid content in bulk soil might bias the δ2H measurements as well.*

   Answer: In our study, lipids are extracted to provide, through their excess ^{13}C/excess ^2H, a proxy of the ratio of organic HNE biosynthesis/ C biosynthesis. In these soils, extractable lipids (by conventional CHCl3-MeOH method) accounts for ca 1.5 % of soil C, whereas total C-H (sum of aliphatic,

sugar C-H and aromatic C account for more than 50% of soil organic C (as determined from CP-MAS 13C NMR of one of the soils). We therefore consider that differences in lipids content between soils would not affect the results more than differences in bulk organic carbon content. Moreover, the isotopic results of the labeled lipids are always corrected from the extracted unlabeled lipids to interpret only the excess ^{13}C/ excess^{2}H.

3. *In this study, non-exchangeable standards of non-similar matrix to the samples were run for calibration and hydrogen exchangeability seem to be corrected by measuring labeled and unlabeled samples at the same time. In theory, this could be a reasonable way to deal with this issue, but the authors should provide more details.*

   Answer: We agree and improved it: See comments 11

Specific comments

4. *p4, line 106: Residual soil moisture is of great relevance when determining H isotope measurements because it would be a reservoir of H in the sample to be analyzed. Was it estimated once at the beginning of the experiment?*

   Answer: The residual soil moisture was estimated once before adding water and substrate.

   Action: The residual soil moisture is added to the manuscript section 2.2.1. p.4

5. *Was performed after collection from incubators and freeze-drying? This step is crucial to eliminate any 'contamination' of residual moisture from the experiments.*

   Answer: We determined the end of freeze-drying when the weight of the sample reached its initial dry weight. See also comments 8.

   Action: Dry residual soil moisture were added in the section 2.2.1, p. 4

6. *p4, line 109: Please confirm amounts of water added.*

   Action: We completed in line 112/113 p. 4 the amount of added water to reach the humidity required for the incubation.

7. *p4, line 108: Provide uncertainty associated with this value.*

   Action: The uncertainty of the isotopic composition is added line 111 p. 4 ($\delta^{2}H = - 63.8 \pm 0.5$ ‰).

8. *p4, lines 134-140: One striking thing is the incubation experiment protocols. The authors opened the incubation systems every two days during the first three weeks and then every week. I understand this is important to keep*

aerobic conditions along the experiment. Would this compromise the ^{2}H abundance of the water? Further explanation is required here.

Answer: By taking the value of the saturation vapor pressure at 28°C (28 g/m^3), the amount of water contained in the headspace jar (0.17 dm^3) was 4.8 mg. The proportion of labeled water vapor lost by the renewal of incubation jar headspace was estimated at 0.7 % the first months and 2 % at one year. The impact of the atmosphere renewal on the isotopic composition was therefore neglected.

Actions: We provided information in section 2.2.3, p. 4/5

9. *p5, line 147: For how long were samples freeze-dried? Again, this step is further relevant to eliminate possible contamination of 'deuterated' moisture in the sample to be analyzed. Previous investigations with organic materials have found that long periods of drying are needed.*

Answer: We agree this is an important question. The efficiency of freeze drying is highly dependent on sample volume and geometry, gas flows, gas pressures, and sample temperature. We optimized these four factors. Samples were freeze-dried for 28 hours. We conducted tests that revealed a constant weight after 16 hours in our conditions. Final sample temperature was 24°C and final vapor pressure < 0.1 mbar.

Action: We completed the text line 156 p.5

10. *p5, lines 153-162: Needs a more detailed description of the analyses. For example, a merit of precision using this method based on the standards measured is needed. How the ^{2}H abundance of water was measured?*

Answer: The ^{2}H abundance of labeled water was calculated not measured. The isotopic composition of the deuterated solution added in the soil was calculated and adjusted during the dilution step.

11. *More importantly, how the authors deal with the hydrogen exchangeability is quite reduced in the manuscript and relatively obscured to the reader. In the section 2.4, the authors only stated the following sentence: "Labeled and unlabeled samples were kept under the same atmosphere until the final $\delta^{2}H$ measurement." Would that mean that they conducted a comparative equilibration method? This method is extensively used in the literature, but mostly for natural abundance samples. Any modifications for labeled samples are required? How long were the samples left under the same atmosphere? Which atmosphere? Laboratory atmosphere? Or inside a desiccator and then opened to the laboratory ambient? In short, the authors need to provide more details in their methodological section.*

Answer: We totally agree with this comment. We were unclear and the equilibration is in itself the definition of NHE. Together with hydrogen bound

to carbon, non-exchangeable hydrogen may include other species with exchange rate depending on organo-mineral and mineral dynamics.
The definition of NEH depends on the method used for equilibration, from simple atmospheric equilibration (Wassenaar and Hobson, 2003) to high pressure and high temperature equilibration with water vapor (Schimmelmann, 1991). The definition of NHE in this study corresponds therefore to hydrogen in dry soil that didn't exchange with atmosphere during the equilibration phase.
We equilibrated unlabeled and labeled samples with the lab atmosphere for 2 hours after soil grinding (exchanges also occur during the grinding ≈ 20 min and during the evaporation by nitrogen flushing in the CM-CRDS introduction line).
The differences of δ^2H between unlabeled and labeled samples are a mean to eliminate the contribution of labile hydrogen (short-time exchange) in the final isotopic calculation. Unlabeled and labeled sample received exactly the same treatment. When compared to Comparative Equilibration method , the absolute δ^2H of NHE is not quantified, but it is calculated and is equal to the relative enrichment of labeled vs. unlabeled sample is similar (see mass balance calculation section 2.5.), because both are equilibrated with the same atmosphere. One advantage of our method is that no standard material is needed for NHE quantification.

Considering your question of adaptation to labeled samples, there is no theoretical consideration that would differentiate highly labeled samples. Considering the sensitivity of the method, in our case, the difference in δ^2H between the equilibration source (natural) and sample (labeled) is very high, so that the sensitivity would be better than using natural samples.

Action: We added a specific section to the attention of the reader in section 2.5.1 p. 6 to explain further the equilibration method we used and the definition of non-exchangeable hydrogen we consider.

12. *Another question I have is whether the use of two reference materials for calibration that cover a very small range of delta values (~2 per mil for $\delta_{13}C$ and ~20 for δ_2H) can adversely affect the accuracy of their measurements of labeled samples among runs. Do the authors consider the use of a labelled standard?*

    Answer: We did not use labeled standard. However, to validate the measurement of highly enriched samples, and the linearity, we calculated the theoretical abundance of the labeled samples at the initial condition, before degradation. Measured values are compared to the theoretical values using the ^{13}C- ^{2}H- labeled organic substrates (slope of 1.02 and $r^2 = 0.99$).

    Action: We added information in section 2.4. p.5

13. *Were other analytical issues such as memory effects considered? Previous published studies (i.e. Koehler and Wassenaar 2012 Anal Chem) that describe this type of technique for natural abundance samples (combustion*

+ laser spectroscopy) have shown some measurable memory issues for hydrogen, at least. I suspect that labeled samples should be further affected.

Answer: To deal with the ^{2}H memory effect often recorded with CM-CRDS, 5 repetitions were done for each sample. The last three were used for interpretation when standard deviation was less than 3 ‰ for natural samples and less than 10 ‰ for enriched samples. Moreover, we analyzed samples from the more depleted to the more enriched and ashes samples were removed from the combustion tube each 45 samples (maximum) to limit contamination.

Action: We added explanations line 175-179 p. 5/6.

14. *p8, lines 271-273: How was this amount calculated?*

Action: We added the calculation in supplementary material. Calculation is shown page vii of this document. By recalculating the recycling for the three soils, this amount reaches 7 % instead of 5 %. See line 323 p.9.

15. *p8, lines 282-283: "In the present experiment, we show that more than 70 % of the H-C bonds are broken". Is this correct in view of the H exchangeability concern during analysis?*

Answer: The conservation of carbon from the molecule is higher than the conservation of non-exchangeable hydrogen from the molecule during the length of incubation suggesting that the initial C-H bonds of the molecule are broken (fig. 1). The exchange of hydrogen is then possible.

Action: We added details lines 336-337 p.9

16. *Figures. During the whole manuscript I missed the results of $_{13}C$ and $_{2}H$ abundances of the bulk soil and lipids during the length of the labeled and unlabeled experiment. Specially, when the correction of hydrogen exchangeability seem to be performed by measuring the labeled and unlabeled samples.*

Answer: We agree that such data are missing. Because these would require 80 kinetic curves, that bring no more information, we put these curves as supplementary material. The curves are also shown p. vii to ix of this document.

Action: In the text, we indicated the magnitude of $\delta^{13}C$ and $\delta^{2}H$ signatures, for the reader to catch the high difference between labeled and natural samples at the beginning of the Results section: Lines 226-240 p. 7.

17. *Figure 2. In the hydrogen labeling experiments performed, there are two sources of hydrogen: substrate and water. In relation to the mineralization of labeled substrates is clear to me since a starting amount of molecule (day 0) became consumed along the experiment and the labeling signature is decreased. However, for the water, it is a different story. I believe the*

incubators used were filled with plenty of labeled water, which means the labeled signature never is consumed or decreased. I wonder if the trend of estimated H derived from water in this figure is based on the isotopic equilibrium with the labeled water instead of an observed derivation of H from water in vivo into microbial biosynthesis. Again, as previous comment, showing the measured ^2H abundances over the length of experimentation could provide insights to clarify these points.

Answer: We agree with the referee who pointed a key issue of the state of H_{dfw} in the soil. We considered it as non-exchangeable hydrogen, either organic or inorganic. You mention the possible occurrence of water in the so-called inorganic HNE. We demonstrated the formation of organic HNE through the incorporation into the lipid fraction (table 3), and the linearity of this pool with the varied and/or nil amounts of added C (fig. 3 and fig. 4). The carbon dependency of hydrogen derived from water confirmed that the measured H_{dfw} (fig. 2) is involved in biological reactions. However, inorganic HNE is not excluded in the measurement of H_{dfw} as discussed in section 4.3.
At the moment of the δ^2H measurement, non-organic H (which might be in the state of inorganic hydroxyl, hydrated ions, water in different states) was estimated from total H measurement of the dry soil as varying from 0 mg/g (podzol) to 6 mg/g (leptosol).
You point out in your comment on Table 3 the difference between the bulk soil and lipids H_{dfw} in one of the soil. In the most clayey soil, the inorganic H_{dfw} pool is in an amount accounting for less than 1/40 of the amount of water-H. In any case, this pool appears very slowly exchangeable with “free water” (the magnitude of kinetic constant if any is in the range of weeks to months in situ at 28°C).
Inorganic H_{dfw} is discussed in detail in section 4.3. The assessment of presence-absence of a water contribution requires additional experiments, for instance warming the soil at different temperature levels. It would take the risk to be too simplistic in such a medium, e.g., that includes dynamic formation/destruction of poorly crystalline minerals, hydrated minerals, smectites saturated with hydrated calcium ions etc. It would require long methodological discussions, and we preferred in this paper the pragmatic estimate of HNE/HE in realistic conditions (ambient temperature and moisture).

Action: We clarified the section 4. lines 345-353 p.10 and lines 390-394 p.11. We also replaced the term "HNE", which is here conceptual and might be confusing" by "H_{dfw}", which corresponds exactly to the measurement.

18. *Table 3. One more noticeable thing in the table is that the results between H (% Hdfw) and C (% Cdfm) are quite consistent in lipids, which do not have exchangeable hydrogen. In the other hand, the proportion of hydrogen derived from the labeled source did not behave similarly in the bulk soil (with exchangeable H). A clear explanation on the treatment of exchangeable hydrogen can convince the reader on a differential isotopic routing of H and C.*

Answer: see comment 17

Calculation of water recycling:

See supplementary material, section 2.4. for equations references.
To calculate the incorporation of the water hydrogen coming from the mineralisation of the added molecule (recycling), we assume that the labeled molecule is completely mineralised in water. The resulting isotopic composition of water in experiment 1(A_{w2}) can be calculated from the isotopic composition of the labelled molecule as follow:

$$A_{w2} = (A_m * H_m)/H_w \quad \text{(SI9)}$$

Then, the amount of non-exchangeable hydrogen that can be derived from this water (H_{dfw2}) can be calculated using the value H_{dfw} calculated in equation (3) :

$$H_{dfw2} = (A_{w2} - A_{tot\_0})/ (A_m - A_{tot\_0}) * H_{dfw} \quad (4)$$

The proportion of deuterium derived from the molecule but incorporated in the soil by the water is given by (H_{dfw2}/H_{dfm})*100 where H_{dfm} is calculated in equation (2).

$\delta^{13}C$ and $\delta^{2}H$ results of the incubation samples:

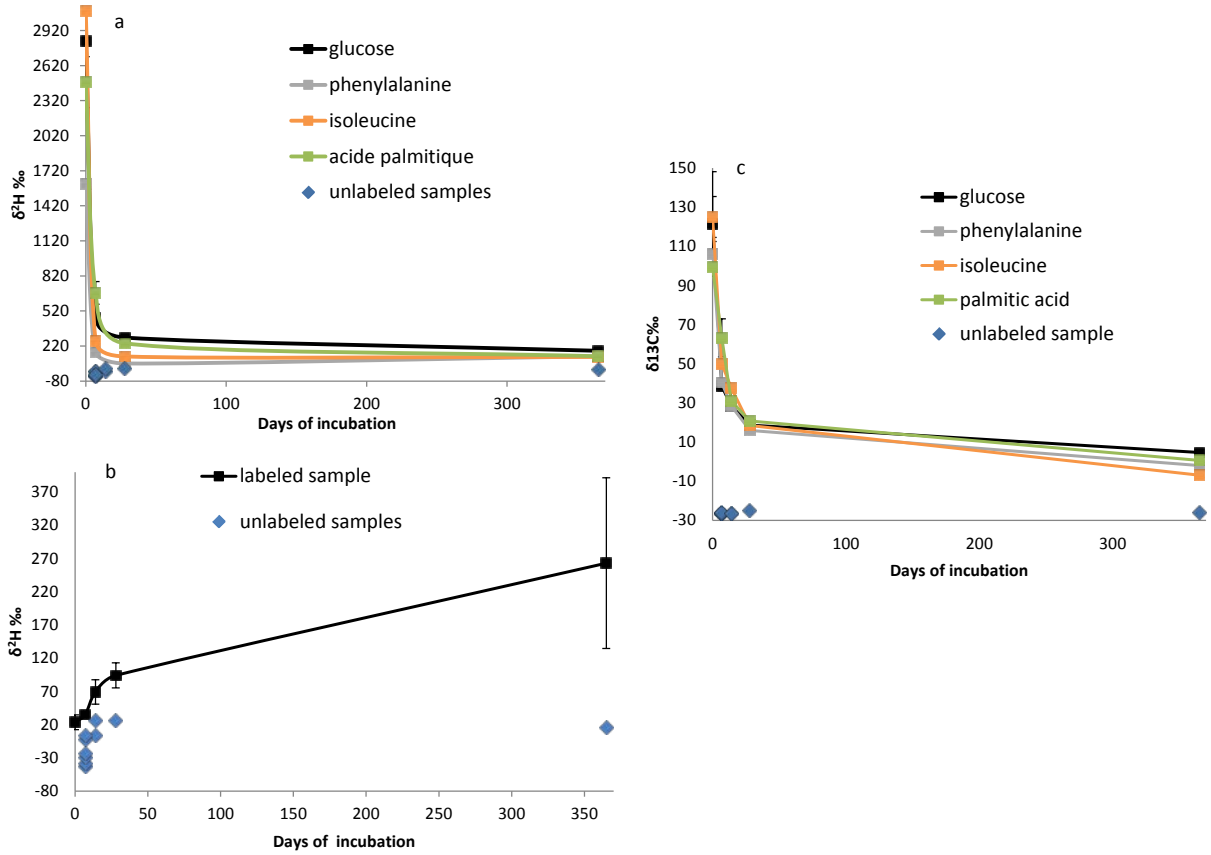

Figure S4.1: Cambisol ^{13}C and ^{2}H isotopic variation a. $\delta^{2}H$ variation through time of the bulk soil that received labeled glucose, phenylalanine, isoleucine and palmitic acid and unlabeled samples. b. $\delta^{2}H$ variation through time of bulk soil that received labeled water and unlabeled samples. c. δ ^{13}C variation through time of the bulk soil that received labeled glucose, phenylalanine, isoleucine and palmitic acid and unlabeled samples. Standard deviations are less than 3 ‰ for unlabeled samples.

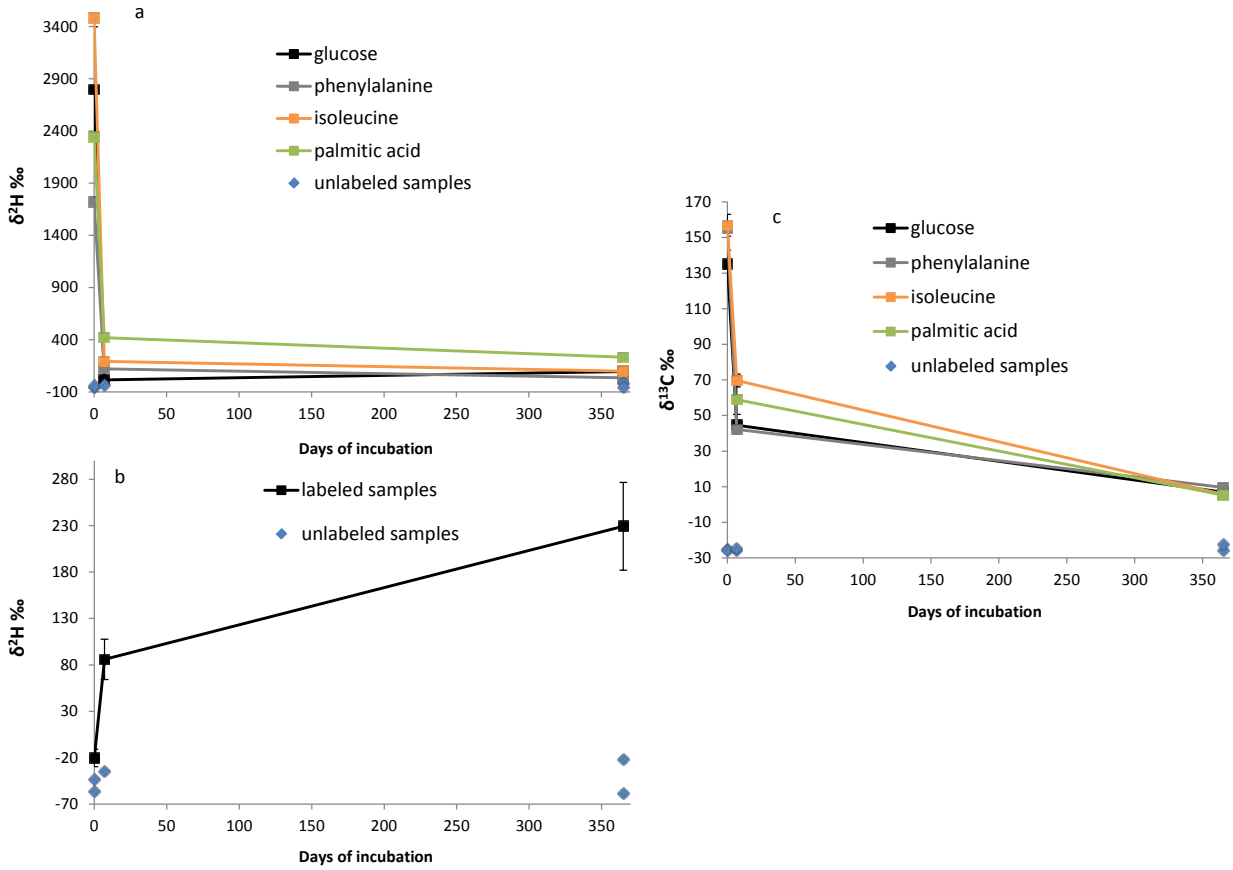

Figure S4.2: Podzol ^{13}C and ^{2}H isotopic variation a. $\delta^{2}H$ variation through time of the bulk soil that received labeled glucose, phenylalanine, isoleucine and palmitic acid and unlabeled samples. b. $\delta^{2}H$ variation through time of bulk soil that received labeled water and unlabeled samples. c. δ ^{13}C variation through time of the bulk soil that received labeled glucose, phenylalanine, isoleucine and palmitic acid and unlabeled samples. Standard deviations are less than 3 ‰ for unlabeled samples.

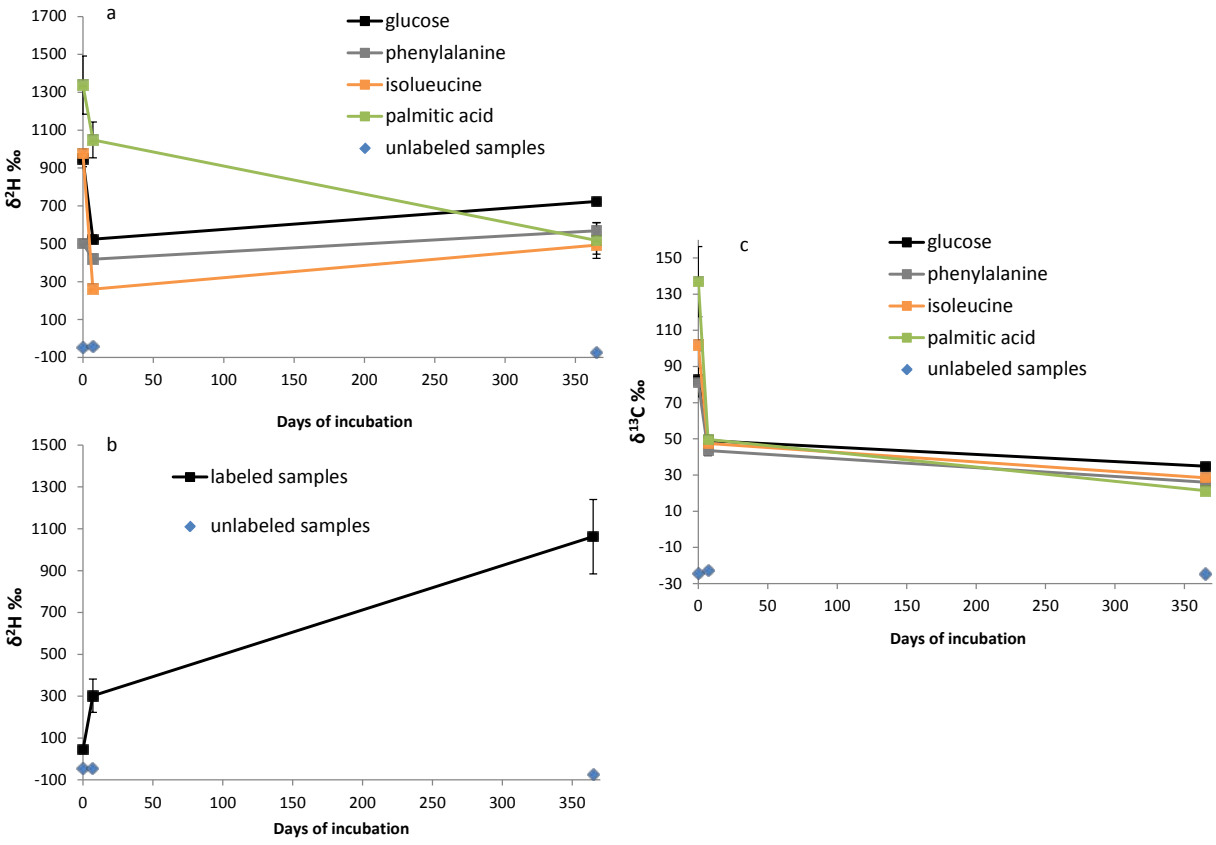

Figure S4.3: Leptosol ^{13}C and ^{2}H isotopic variation a. $\delta^{2}H$ variation through time of the bulk soil that received labeled glucose, phenylalanine, isoleucine and palmitic acid and unlabeled samples. b. $\delta^{2}H$ variation through time of bulk soil that received labeled water and unlabeled samples. c. δ ^{13}C variation through time of the bulk soil that received labeled glucose, phenylalanine, isoleucine and palmitic acid and unlabeled samples. Standard deviations are less than 3 ‰ for unlabeled samples.

Anonymous referee#2

This is an important paper in the context of understanding of hydrogen dynamics in soil organic matter. The paper is well written and sound.

1. *I believe the backline story about Tritium detracts from the main study research outcomes as the context of tritium remains only touched upon, as we get no real concrete data about the concerns raised line 64/65. There is mention of enhancement tritium entering the environment due to historic bomb-testing but no mention that tritium is a radioactive form of H (half-life 12.3 years), unlike 2H and 1H who are stable isotopes, so will over time will dissipate and has done already decline since the bomb-14C peak. The authors should either reduce the tritium context or make it more quantitative.*

   Answer: We decided to reduce the tritium context in the text but we want to keep in readers mind that such a study on hydrogen dynamics in soil organic matter could be used for the prediction of tritium fate.

   Action: We reorganized the paragraph in the introduction about tritium in the lines 60-69 p. 2/3.

2. *The authors may wish to comment on the potential of water in the air (different isotopic H signature) to enter the experimental jars and when opening them to prevent anaerobic conditions occurring in the jar*.

   Answer: By taking the value of the saturation vapor pressure at 28°C (28 g/m^3), the amount of water contained in the headspace jar (0.17 dm^3) was 4.8 mg. The proportion of labeled water vapor lost by the renewal of incubation jar headspace was estimated at 0.7 % the first months and 2 % at one year. The impact of the atmosphere renewal on the isotopic composition was therefore neglected.

   Action: We provided information in section 2.2.3, p. 4/5

3. *Tremendously (line 229) maybe find another word to describe the very large increase.*

   Action: We removed the word that was not appropriated in this case.

4. *For Figure 2 and 3 the scale on the y-axis between the various soil types are different, which makes immediate comparisons difficult. If, the authors want to retain this, maybe indicate in the legends of Figure 2 and 3 that this is the case ħote, the scale of Y-axis varies between the subfigures'*

   Action: We changed the legend of figure 2 and 3 in this way.

Hydrogen dynamics in soil organic matter as determined by ^{13}C and ^{2}H labeling experiments

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Abstract: Understanding hydrogen dynamics in soil organic matter is important to predict the fate of ^{3}H in terrestrial environments. One way to resolve hydrogen fate and to point out processes is to examine the isotopic signature of the element in soil. However, non-exchangeable hydrogen isotopic signal in soil is complex and depends on the fate of organic compounds and microbial biosyntheses that incorporate water-derived hydrogen. To decipher this complex system and to understand the close link between hydrogen and carbon cycles, we followed labeled hydrogen and labeled carbon all along natural-like soil incubations. We performed incubation experiments with three labeling conditions: 1- $^{13}C^{2}H$ double-labeled molecules in the presence of $^{1}H_2O$, 2- ^{13}C-labeled molecules in the presence of $^{2}H_2O$, 3- no molecule addition in the presence of $^{2}H_2O$. The preservation of substrate-derived hydrogen after one year of incubation (ca. 5% in most cases) was lower than the preservation of substrate-derived carbon (30% in average). We highlighted that 70% of the C-H bonds are broken during the degradation of the molecule which permits the exchange with water hydrogen. Added molecules are used more for trophic resources. The isotopic composition of the non-exchangeable hydrogen was mainly driven by the incorporation of water hydrogen during microbial biosynthesis. It is linearly correlated with the amount of carbon that is degraded in the soil. The quantitative incorporation of water hydrogen in bulk material and lipids demonstrates that non-exchangeable hydrogen exists in both organic and mineral-bound forms. The proportion of the latter depends on soil type and minerals. This experiment quantified the processes affecting the isotopic composition of non-exchangeable hydrogen, and the results can be used to predict the fate of tritium in the ecosystem or the water deuterium signature in organic matter.

1 Introduction

Our knowledge of the nature of soil organic matter (SOM) has made great progress in recent decades: it is now considered to be a continuum of progressively decomposing organic compounds (Lehmann and Kleber, 2015), composed of all the components of living material such as glucides, peptides, lipids, organic acids and phenolic

compounds (Kelleher and Simpson, 2006). Small molecules are associated with each other in supramolecular structures or with mineral particles by weak bonds, including H-bonds (Sutton and Sposito, 2005; Lehmann and Kleber, 2015). Their lifetime in soils is controlled more by sorption or protection than from intrinsic chemical recalcitrance to biodegradation (Schmidt et al., 2011; Basile-Doelsch et al., 2015) with the exception of pyrolytic products. Highly degradable compounds, such as glucides and peptides, contribute to the oldest SOM components, and biomarkers tend to indicate that old SOM was derived more from microbial-derived products than from plant-derived molecules as a result of the mineral protection processes (Rafter and Stout, 1970; Derrien et al., 2007; Bol et al., 2009). Carbon dynamics in this continuum have been widely studied using the natural $^{14}C/^{12}C$ and $^{13}C/^{12}C$ (Rafter and Stout, 1970; Balesdent et al., 1987) ratios and also through labeling experiments (Jenkinson, 1965; Murayama, 1988; Derrien et al., 2004). The results of these experiments have highlighted the different turnover of soil organic pools. Plant material is rapidly decomposed into microbial biomass, and a small portion of both can be protected from biodegradation for decades to centuries, representing the main part of SOM. One part of this organic matter remains stabilized for millennia, especially in deep soil horizons. It is therefore expected that the non-exchangeable hydrogen (NEH) dynamics, bound to carbon in soil, will be controlled by the same processes: organic carbon inherited from vegetation, biodegradation, microbial biosyntheses and stabilization.

Hydrogen has various molecular positions in soil. It can be organic or inorganic and non-exchangeable or exchangeable with available hydrogen. The abiotic exchange of organic hydrogen depends on the strength of the bond and the energy required for exchange (Belot, 1986; Schimmelmann, 1991; Ciffroy et al., 2006; Sauer et al., 2009). Bound to N, O and S, hydrogen is usually exchangeable with ambient water and water vapor (Schimmelmann, 1991; Wassenaar and Hobson, 2000). However, hydrogen bound to carbon is considered to be stable and non-exchangeable due to the strong covalent bonds (Baumgärtner and Donhaerl, 2004; Diabaté and Strack, 1997; Kim et al., 2013). At the ecosystem scale, H bound to C is not exchangeable (Sessions et al., 2004). Hydrogen can also interfere with clay minerals. Interlayer water exchanges with free water within a few hours and is removed after drying (Savin and Epstein, 1970). However, the structural water and the hydroxyl hydrogen of clay are non-exchangeable at room temperature (Savin and Epstein, 1970).

The natural $^{2}H/^{1}H$ ratio of plants and sediments has been used as a proxy to reconstruct past climate and paleoenvironmental conditions, such as temperature, water use efficiency (Epstein et al., 1976; Sessions et al., 2004; Zech et al., 2014; Tuthorn et al., 2015) The isotopic composition of the NEH preserves the initial composition of the plant and registers the rain isotopic composition (Sessions et al., 2004; Schimmelmann et al., 2006; Ruppenthal et al., 2010). The $\delta^{2}H$ of water and exchangeable hydrogen is not stable. Whereas soil organic carbon and nitrogen cycles have been extensively studied, soil organic hydrogen and its recycling in the environment remain poorly understood due to its complex behavior. The total bulk soil, composed of a mixture of non-exchangeable and exchangeable, organic and inorganic hydrogen, makes the hydrogen isotopic composition hard to determine.

The composition and exchanges between these pools can be of great importance when modeling, for instance, tritium fate in the environment. Tritium is a radioactive isotope of hydrogen which was released in large amounts in the atmosphere by nuclear weapon tests in the 60s. Since then, tritium levels have greatly declined because of its relatively short half-life (12.3 years). However the concentration of organically-bound tritium can often be higher than the concentration of water tritium due its longer residence time (Gontier and Siclet, 2011; Kim et al.,

2012; Eyrolle-Boyer et al., 2014; Thompson et al., 2015). Models used to predict the fate and behavior of tritium in the environment often simplify processes linked to the formation and degradation of organic bound tritium. Since tritium behaves as stable hydrogen in the environment, the assessment of the fate and residence time of organic tritium could be improved by quantifying the preservation of organic hydrogen from vegetation, the accumulation of hydrogen from water in the soil and the processes involved in organic matter decomposition and mineralization.

To decipher and quantify the preservation of the organic material and the microbial biosyntheses incorporating water-derived hydrogen, we designed incubation experiments with labeled compounds by assuming that the non-exchangeable hydrogen dynamics are controlled by the carbon dynamics in the soil organic matter. Three scenarios were addressed:

1- $^{13}C^{2}H$-double labeled molecules in the presence of $^{1}H_2O$, 2- ^{13}C-labeled molecules in the presence of $^{2}H_2O$, 3- no molecule addition in the presence of $^{2}H_2O$. The ^{13}C and ^{2}H bulk soil isotopic compositions were analyzed at different times to quantify the processes involved. The isotopic composition of lipids was also analyzed as an indicator of organically bound NEH.

The medium-term ^{13}C and ^{2}H labeling experiments were conducted on different types of soil (clayey leptosol, cambisol and podzol) from 0 to 1 year to highlight and quantify the processes affecting hydrogen based on the carbon dynamics in the soil organic matter and their dependence on the soil properties.

2 Materials and Methods

2.1 Soil sampling

Three soils with contrasting physical and chemical properties were selected for this study (Table 1):

Cambisol: the 0-25 cm surface layer of a cambisol was sampled from an INRA long-term field experiment in the Parc du Château de Versailles, France in March 2014. This soil is a neutral eutric cambisol with a clay composition of 17%, 33% of sand, 50% of silt and a nil carbonate content (Dignac et al., 2005). The plant cover is wheat. After each harvest, wheat residues were returned to the soil, and the first 25 cm were ploughed each fall.

Podzol: the 0-25 cm surface layer of a podzol was sampled from an INRA field experiment in Pierroton (close to Bordeaux, SW France) in May 2014. The lands cover of les Landes de Gascogne is a mixt forest dominated by *Pinus pinaster*. The sampling plot was converted into maize in 1992. The soil is a sandy hydromorphic podzol (Jolivet et al., 2006) with a clay content of less than 5 % and a sand content higher than 90 %. The first 25 cm are ploughed and crop residues are returned to the soil after each harvest.

Leptosol: the 5-10 cm surface layer of a mollic leptosol was sampled from the long-term Ecosystem Research experiment "Oak Observatory at Observatory of Haute-Provence" (O3HP), France in July 2014. The vegetation is dominated by *Quercus pubescens.* This soil is derived from limestone, compact and iron-rich with clay content (mainly smectite) of 54%.

2.2 Soil incubation

2.2.1 Soil preparation

Soil samples were air-dried at 20°C and sieved to 2 mm. Air-dry ~~Residual~~ soil moisture was determined in parallel by oven-drying an aliquot at 105°C during 24h. The residual soil moisture was determined once before adding substrate and water in order to know the exact amount of water at the beginning of incubation.
The residual soil moisture was 5.9 % for the cambisol, 0.8 % for the podzol and 8.5 % for the leptosol.
Thirty-five grams of cambisol and podzol and 30 g of mollic leptosol were transferred into 210 mL incubation jars. Each incubator was then moistened with ~~ulta~~ultra-pure water (δ^2H = - ~~55~~63.8 ~~‰~~ ± 0.5 ‰) at ~~24~~ 0.24 g g^{-1} of dry soil for cambisol (6 ml of ultrapure water added) and leptosol (~~-~~4 ml of ultrapure water added) and ~~10~~ 0.10 g g^{-1} for podzol (3 ml of ultra-pure water added) before pre-incubation at 28°C in the dark for 10 days to re-establish the biological activity to the basal level and to avoid confusion between rewetting-induced and substrate-induced activity.

2.2.2 Substrate incubations

Glucose, palmitic acid, phenylalanine and isoleucine were introduced separately in different incubators. They represent the most common primary compounds of the glucide, lipid and protein families found in either plant or microbial matter and contain different functional groups.
^{13}C-labeled and ^{2}H-labeled molecules and $^{2}H_2O$ were provided by Euriso-Top (Cambridge Isotope Laboratories, Inc., Andover, England).
The isotopic abundance of each molecule was adjusted to the desired value by mixing labeled and unlabeled sources. We prepared "$^{13}C^{2}H$" (double-labeling) solutions and "$^{13}C^{1}H$" (mono-labeling) solutions for all molecules. The incubation characteristics are shown in Table S1 (supplementary material). Mixing was performed gravimetrically.
For palmitic acid, the equivalent amount of unlabeled and labeled compound was added to 200 mg of soil and was melted at 70°C. We finely ground the cooled mixture to obtain a homogenized powder that could be added to the incubators. Two powders were prepared: a $^{13}C^{2}H$-enriched powder and a ^{13}C-enriched powder.
Three distinct labeling experiments were performed:

1) Experiment 1: "$^{13}C^{2}H + H_2O$": Double-labeling molecule introduced to the soil with ultra-pure water.
2) Experiment 2: "$^{13}C^{1}H + {}^{2}H_2O$": Mono-labeling molecule introduced to the soil with deuterated water.
3) Experiment 3: "no molecule + $^{2}H_2O$": Only deuterated water introduced to the soil.

The final humidity of the soil was 0.30 g g^{-1}, 0.15 g g^{-1}, and 0.31 g g^{-1} dry weight, respectively, for cambisol, podzol and leptosol.

2.2.3 Incubations

The 300 incubators were incubated at 28°C in the dark, and three were frozen at 0, 7, 14, and 28 days and 1 year for the cambisol and at 0 and 7 days and 1 year for the two other soils. Jars were briefly opened (few seconds for each samples) every two days during the first three weeks and then every week until the end of incubation to keep the system under aerobic conditions. The evaporation of water is very limited during this step. However, this leads to the renewal of water vapor in the jar head space by ambient atmosphere. The isotopic composition

of ^{2}H of the new atmosphere in the jar is depleted compared to the previous one. By taking the value of the saturation vapor pressure at 28°C (28 g/m^{3}), the amount of water contained in the headspace jar (0.17 dm^{3}) was 4.8 mg. The proportion of the lost labeled water was estimated at 0.7% the first months and 2% at one year. The impact of the atmosphere renewal on the isotopic composition was therefore neglected.

Control incubators were prepared for each experiment at each time without any added substrate or deuterated water under the same incubation conditions.

To highlight the link between the NEH and carbon dynamics we initially added three different amounts of labeled glucose to the podzol and we analyzed the results after 7 days of incubation.

2.3 Lipid extraction

Lipids were extracted to isolate the organic non-exchangeable hydrogen. Lipids extractions were performed on samples that had received glucose and had been incubated for 1 year. Between 10 and 15 g of soil were sub-sampled and phosphate buffer, chloroform and methanol were added (0.8:1:2; v:v:v). After 2 min of ultrasonic and 30 min of warming (37°C), samples were centrifuged for 8 min a 2600 tr min-1. Supernatant was retrieved and stored at room temperature while chloroform and methanol (1:2; v:v) were added to the remaining soil and centrifuged again. Supernatant was retrieved and added to the previous aliquot. Twenty ml of NaCl were then added to the supernatant to distinguish two phases. The denser part was collected and dried under nitrogen.

2.4 Isotopic measurements

Prior to analysis, incubated samples were freeze-dried during 28h~~,~~. Samples were then ground to a fine powder and kept in closed tubes under laboratory atmosphere. ~~Labeled and unlabeled samples were kept under the same atmosphere until the final $\delta^{2}H$ measurement.~~

Twenty to fifty-five milligrams of soil were then introduced into a 10 mm tin capsule.

Lipid samples were solubilized in dichloromethane before introducing them to the tin capsules. We let the solvent evaporate before the analysis. The mean isotopic signature of this bulk lipid fraction was measured using the same method as for the soil samples.

The ^{13}C and ^{2}H contents were analyzed simultaneously with a combustion module-cavity ring-down spectroscopy (CM-CRDS) isotope analyzer (Picarro, B2221-i). The organic standards polyethylene (IAEA CH7; $\delta^{13}C$ = - 32.15 ± 0.05‰; $\delta^{2}H$ = - 100.3 ± 2.0‰) and oil NBS-22 ($\delta^{13}C$ = - 30.03 ± 0.05‰, $\delta^{2}H$ = - 119.6 ± 0.6‰) were used to calibrate the measurements. A homemade standard (olive oil) was also used in each run ($\delta^{13}C$ = - 29.0 ± 0.2‰, $\delta^{2}H$ = - 153 ± 5 ‰).

To validate the measurement of highly enriched samples by CM-CRDS, we compared the measurements at initial conditions (time 0 of the incubations, before the degradation of substrate occurred), to the theoretical isotopic composition at initial conditions obtained by calculation. Both the linearity of the $\delta^{2}H$ measurement for enriched samples, and the full recovery of labeled NEH during the drying process were confirmed by the measured vs. theoretical $\delta^{2}H$ of initial labeled soil samples (mixtures of soil and labeled source before incubation), which yielded to a slope of 1.02 and r^{2} = 0.99.

To deal with the ^{2}H memory effect often recorded with CM-CRDS (Koehler and Wassenaar, 2012), 5 repetitions

were done for each sample; the last three were used for interpretation when standard deviation was less than 1.5 ‰ for natural samples and less than 10 ‰ for enriched samples. Moreover, we analyzed samples from the more depleted to the more enriched and ashes were removed from the combustion tube each 45 samples to limit contamination.

The isotopic composition of ^{13}C and ^{2}H are expressed by abundance (A) or as δ (‰)

$A^{13}C = {}^{13}C/({}^{13}C+{}^{12}C)$ and $A^{2}H = {}^{2}H/({}^{2}H+{}^{1}H)$

$\delta‰ = [R_{sample} / R_{standard} - 1]*1000$

where $R= {}^{13}C/{}^{12}C$ or $^{2}H/^{1}H$. The international standard was VPDB for carbon and VSMOW for hydrogen.

2.5 Quantification of NEH derived from the labeled source

2.5.1. Samples equilibration

We performed equilibration of the labeled and unlabeled samples with the same atmosphere to reach the same isotopic composition of exchangeable hydrogen in the dry soils. To do so, before analysis, unlabeled and labeled samples were equilibrated with the laboratory atmosphere for 2 hours after soil grinding (exchanges also occur during the grinding ≈ 20 min and during evaporation by nitrogen flushing in the CM-CRDS introduction line). The differences of $\delta^{2}H$ between unlabeled and labeled samples are a mean to eliminate the contribution of labile hydrogen (short-time exchange) in the final isotopic calculation (see mass calculation below). Unlabeled and labeled sample received exactly the same treatment. Hydrogen that did not exchange during the length of equilibration is considered as non-exchangeable in this study.

2.5.2. Mass balance equations

Table 2 summarizes the different variables used in the mass balance equations and supporting information (S2) is provided for further understanding of calculations.

The carbon isotopic composition of the total bulk soil corresponds to the proportion of molecule-derived (labeled) and soil-derived (unlabeled) carbon (Eq. (1)).

$$C_{dfm} = ({}^{13}A_{tot} - {}^{13}A_{tot\_0})/({}^{13}A_{m} - {}^{13}A_{tot\_0})* C_{tot} \quad (1)$$

Exchangeable hydrogen has the isotopic composition of the atmosphere when the sample is dry. Non-exchangeable hydrogen from the labeled source was estimated based on the simultaneous measurement of labeled and unlabeled samples equilibrated with the same atmosphere using the equations (2) and (3), which attributes all the excess deuterium (difference between the ^{2}H abundances of the labeled sample and unlabeled control) to the NEH derived from the labeled source atoms (see supplementary material S2 for the calculation).

In Experiment 1 (labeled molecule):

$$H_{dfm}= (A_{tot} - A_{tot\_0})/(A_{m} - A_{tot\_0})*H_{tot} \quad (2)$$

In Experiment 2 (labeled water):

$$H_{dfw}= (A_{tot} - A_{tot\_0})/(A_{w} - A_{tot\_0})*H_{tot} \quad (3)$$

The labeled source is highly enriched compared to natural soil or water: $^{13}A_m$=6.08 to 16.08 %, A_m = 2 to 3.5 % and A_w= 0.26 % (supplementary material, Table S1).

The mean ^{2}H abundance of unlabeled soil and water is approximately 0.015 %, and the mean ^{13}C abundance is approximately 1.08 %.

Uncertainties in the element and isotope ratio measurements affect the estimate of the amount of labeled-source–derived carbon or hydrogen atoms. To assess the uncertainty in the calculated values H_{dfm} and C_{dfm}, we calculated the statistical error propagation of the uncertainties of the measured isotopic compositions and the element content of the replicated samples (Supplementary material S3).

3 Results

The raw isotopic composition of the highly enriched samples, ($\delta^{13}C$ and $\delta^{2}H$ measurements) are presented over the incubation period in supplementary material (section S4).

In unlabeled samples $\delta^{2}H$ ranged from - 43 to + 27 ‰ for the cambisol, from -58 to -22 ‰ for the podzol and from -74 to - 43 ‰ for the leptosol with less than 3 ‰ variation.

In the first experiment (labeled molecule), measured $\delta^{2}H$ values of labeled samples at initial conditions ranged from 900 ‰ to 2500 ‰ depending on the soil and the added molecule. They still reached 85 to 576 ‰ after one year of incubation.

$\delta^{2}H$ values of the second experiment (labeled water) ranged from 0 to 260 ‰ for the cambisol, from -20 to 240 ‰ for the podzol and from 40 to 1000 ‰ for the leptosol over the incubation period.

At each time step, as described in the methodology (see section 2.5.2) the difference between the ^{2}H isotopic composition of the labeled sample and the isotopic value of the unlabeled sample analyzed the same day is calculated. The differences between labeled and unlabeled samples are huge.

$\delta^{13}C$ values are between 130 and 110 ‰ at the beginning of the incubation and between 0 and 15 ‰ after one year whereas unlabeled $\delta^{13}C$ values were -26.2 ‰ for the cambisol, -25.4 ‰ for the podzol and -24.7 ‰ for the leptosol with less than 0.3 ‰ variation.

3.1 Comparison of the four substrates mineralization

The fates of labeled C or H atoms are presented as the mass of C or NEH derived from the labeled source, i.e. molecule or water, C_{dfm}, H_{dfm}, H_{dfw}; eq. (1), (2), (3) (dfm and dfw stand for derived from the molecule and derived from water, respectively). We first tested the dependence on time (1 week and 1 year), molecule type and soil type on the basis of a three-way ANOVA of each explained variable. Both C_{dfm} and H_{dfw} were dependent on time ($p<0.001$) and soil ($p<0.001$) but not on molecule. H_{dfm} was dependent on time ($p<0.001$), soil type ($p<0.001$) and molecule ($p<0.001$). The differences in results for H_{dfm} can be explained by the uncertainty in experiments and measurements. Because we found no significant differences between the molecules for C_{dfm} and H_{dfw}, we considered the different molecule incubations as replicates to simplify the results presentation, with only the mean values shown in graphs. The difference in degradation between molecules is therefore contained in the error bars in fig. 1 and 2.

3.2 Carbon mineralization

In fig.1, the results of the carbon and hydrogen derived from the molecule are expressed in percent: the amount of H_{dfm} and C_{dfm} calculated in Eq. (6) and (7) relative to the amount of H_m and C_m in the added molecule. The degradation of the added molecule was very fast. After 7 days, 42 %, 31 % and 53 % of molecule-derived carbon remained in the cambisol, podzol and leptosol, respectively (fig. 1). This trend is in agreement with previous studies (Murayama, 1988; Derrien et al., 2007) and illustrates the almost complete consumption of the substrate in a few days. Approximately 30 to 50 % of the consumed material was converted into microbial products, and the remaining part was used for heterotrophic respiration. During the following months, the mineralization of organic carbon continued due to the partial consumption of the newly formed microbial carbon by the soil food web.

During the incubation, non-labeled carbon (soil-derived carbon) also decreased by 1.8, 2.4 and 4.5 mg g^{-1} within one year in the cambisol, podzol and leptosol, respectively.

3.3 Molecule-derived non-exchangeable hydrogen

After incubation, molecule-derived NEH (H_{dfm}) was considerably lower than molecule-derived carbon for the cambisol and for the podzol (fig. 1): expressed in % of the initial labeled NEH, it was, 12 % and 5 % compared to the 42 % and 31 % recorded for carbon. It is important to note the different fate of the leptosol, where the yield of transfer of NEH reached 55 % after 7 days of incubation (fig. 1). During the following months, H_{dfm} slightly decreased in the three soils (by approximately 6 % ± 5 after one year).

3.4 Incorporation of water hydrogen

Experiment 2 (molecule $^{13}C^{1}H + {}^{2}H_2O$) highlights the incorporation of water hydrogen in the non-exchangeable pool of soil. ~~.~~ For the three soils, the incorporation of hydrogen from water ~~tremendously~~ increased during the first seven days and continued to slowly increase during the incubation year (fig. 2). Respectively, 0.013±0.001, 0.008±0.002 and 0.33±0.07 mg g^{-1} of hydrogen derived from water was found after 7 days of incubation, and 0.06±0.03, 0.023±0.004, and 0.845±0.003 mg g^{-1} was found after one year of incubation for cambisol, podzol and leptosol (fig. 2).

Figure 3 shows the difference in the incorporation of water hydrogen with and without added substrate. The incorporation of water-derived hydrogen was higher when associated with substrate addition. It was twice as high for the podzol after 7 days of incubation. Figure 4 illustrates that this enhancement of incorporation of water hydrogen was linearly dependent on the amount of the substrate added to the soil.

3.5 Isotopic composition of lipids

Carbon and hydrogen isotopic compositions of bulk lipids at 365 days for the control soil are presented in Table 3. The proportions of labeled carbon and hydrogen were calculated as the proportion of the total lipid carbon and hydrogen ($(C_{dfm}/C)_{lipids}$, $(H_{dfm}/H)_{lipids}$, and $(H_{dfw}/H)_{lipids}$) and were compared to the proportion in the bulk soil ($(C_{dfm}/C)_{bulk}$, $(H_{dfm}/H)_{bulk}$, $(H_{dfw}/H)_{bulk}$). The $\delta^{13}C$ and $\delta^{2}H$ of the lipids in the control samples were lower than that of the bulk soil, in agreement with previous work, where the lipid $\delta^{13}C$ was 2-3‰ lower than the bulk $\delta^{13}C$ (Chikaraishi and Naraoka, 2001; Hayes, 2001), and the lipid $\delta^{2}H$ was 150‰ lower than the bulk $\delta^{2}H$ (Sessions et

al., 1999; Chikaraishi and Naraoka, 2001). The average measured H/C ratio of the lipids of the three soils was 2.1 (molar ratio).

The proportion of molecule-derived carbon in the lipids was 1.0, 0.4 and 0.8 % for the cambisol, podzol and leptosol, respectively, compared to the corresponding values of 0.6, 0.4, and 0.4 % in the bulk organic carbon. The proportion of molecule-derived hydrogen was, respectively 0.10, 0.02 and 0.19 % for the cambisol, podzol and leptosol (Table 3). These values were on the same order of magnitude as the molecule-derived hydrogen in the bulk soil. The proportion of labeled water-derived hydrogen was 1.0, 0.4 and 1.1 % of the total hydrogen content in the lipids and 1.5, 0.9 and 6.8 % in the total bulk soil for the cambisol, podzol and leptosol, respectively.

4 Discussion

4.1 Preservation of the organic substrate hydrogen in biosyntheses

The microbial activity is initiated during the first days after the addition of the substrate. The added molecule regardless of its quality is quickly metabolized (fig. 1).

We independently traced the preservation of organic hydrogen (experiment 1) and the incorporation of water-derived hydrogen (experiment 2) during decomposition and biosynthesis. The conservation of organic hydrogen from the initial substrate is very low in both the total and lipid NEH. The carbon-hydrogen bonds are broken during decomposition, and exchange with exchangeable hydrogen can occur. The difference between carbon and hydrogen isotopic fates during the first seven days (fig. 1) reflects the exchange of hydrogen with water during the early stage of degradation. Subsequently, new organic exchangeable hydrogen derived from water can be incorporated into the non-exchangeable pool of organic matter by biological processes. Furthermore, because mineralization of a substrate also results in $^{2}H_2O$ release, one part of the soil organic non-exchangeable ^{2}H may originate from incorporation of this substrate-derived deuterated water into the non-exchangeable pool. Using the assumption that water is a well-mixed isotopic compartment, this amount is between 3 and ~~5~~ 7 % of the residual hydrogen from the organic substrate at 365 days (see supplementary materials for calculation). The isotopic composition of the non-exchangeable organic hydrogen is mainly determined by the water isotopic composition (fig. 2). It is in accordance with the work of Baillif and colleagues that have grown fungi with labeled glucose, water or acetate to trace the incorporation of ^{2}H during fatty acid biosynthesis. They have demonstrated that water is the main donor of hydrogen atoms in the non-exchangeable pool within the biosynthesis cycle (Baillif et al., 2009). Ruppenthal and colleagues have shown as well that precipitation contributes to 80 % of the isotopic composition of non-exchangeable hydrogen (Ruppenthal et al., 2010). Moreover, the incorporation of water hydrogen is favored by the strength of the C-H bond breakage. It can be weak or strong, depending on the enzyme activated for the degradation of the molecule and the position of the bond (Augusti et al., 2006). When C-H breakdown is favored, the surrounding water imprints its hydrogen isotopic signature on the former bounded-H (Augusti et al., 2006). In the present experiment, we show that more than 70 % of the H-C bonds of the initial molecule are broken in the first biodegradation steps (7 days, fig. 1); therefore, the added molecules are used more for energetic and trophic resources than as building blocks in the biosynthesis.

Deuterium can also accumulate in hydration shells, which have stronger hydrogen bridges than the biomolecule (Baumgärtner and Donhaerl, 2004). The accumulation of deuterium from water occurs in the biomatter during

biological processes but also during the hydration of molecules (Baumgärtner and Donhaerl, 2004; Turner et al., 2009).

4.2 Carbon-driven acquisition of the non-exchangeable hydrogen isotope signature

The rapid mineralization of hydrogen from the molecule is due to biodegradation whereas the rapid incorporation of water during the first seven days of incubation is associated with biosynthesis (figs. 1 and 2). Results obtained on the lipid fraction provide evidence of the formation of organic non-exchangeable hydrogen from H_{dfw} (table 3) Carbon mineralization fosters the formation of non-exchangeable hydrogen from water (figs. 3 and 4). There is also an incorporation of hydrogen from water in the soil in the experiments without substrate. This could be ~~due~~ associated to the microbial transformation~~mineralization~~ of carbon already present in soil (fig. 3) but also to the inorganic fraction that could incorporate hydrogen from water. ~~However, in~~ In fig. 2, exchange of hydrogen with water seems to be continuous: ~~the incorporation of water hydrogen into the non-exchangeable pool~~H_{dfw} continues increasing, following carbon mineralization. The incorporation of water hydrogen in the organic non-exchangeable pool but also in the inorganic fraction is occurring during the length of the incubation.~~.~~

Results of lipids isotopic compositions show that the amount of newly formed NEH (% of H_{dfw} + % of H_{dfm} in lipids; Table 3) is slightly higher than expected from the theoretical organic C-H bond (% of C_{dfm} in lipids; Table 3). This could be due to complete, stoichiometric labeling of newly biosynthesized lipids, i.e., lipids formed on the labeled organic carbon plus a smaller amount of newly synthesized lipids from unlabeled organic matter. The proportion of molecule-derived carbon is higher in the lipid than in the bulk soil, and the lipids are derived mainly from microbial biosynthesis. However, the proportions of H_{dfm} in the lipids and in the bulk soil are of the same magnitude for the three soils (Table 3): hydrogen is derived evenly from the labeled molecule and from the unlabeled soil during lipids biosynthesis.

The proportion of hydrogen derived from the water in lipids is lower than the respective proportion in the bulk soil, which means the proportion ~~of NEH derived from water~~ of H_{dfw} is not necessarily organic. It could be inorganic hydroxyl, hydrated ions or water in different states. The difference is even higher in the clayey soil (leptosol, Table 3) that contains the highest proportion of hydrogen (table 1).~~.~~

In lipids, hydrogen corresponds to the organic, non-exchangeable hydrogen. The hydrogen in lipids formed from water (H_{dfw}) is only organic, whereas in the bulk soil, the ~~non-exchangeable hydrogen~~ H_{dfw} is organic and inorganic. To estimate the proportion of organic non-exchangeable hydrogen, we assume that H_{dfw}/C_{dfm} in lipids (0.20 on average for the three soils based on mass ratio, Table 3) is approximately the same as the H_{dfw}/C_{dfm} in the organic fraction of the bulk soil. Using the measured C_{dfm} in bulk soil, we can then estimate the total organic H_{dfw} as 0.014, 0.015 and 0.028 mg g^{-1}, respectively, for the cambisol, podzol and leptosol. The proportion of inorganic ~~NEH~~ H_{dfw} is therefore 0.046, 0.008 and 0.82 mg g^{-1} for the three soils. The NEH isotopic composition is mainly controlled by the incorporation of water through biosynthesis, but the inorganic ~~NEH~~ H_{dfw} is not negligible, especially in the clayey soil.

4.3 Hydrogen dynamics in different soil types.

The association of organic matter with minerals is known to decrease the decomposition rate of the former (Feng et al., 2013; Jenkinson and Coleman, 2008; Vogel et al., 2014). This result is observed in our experiment by comparing the three soils with increasing clay content and is applicable to both H and C in both bulk soil (fig. 1) and lipids (Table 3). However, the clay content has an important role in the incorporation of water-derived hydrogen beyond this organic matter stabilization effect. In clayey leptosol, the amount of labeled NEH from the molecule (H_{dfm}; fig. 1) is much higher than in the other soils, which may be explained by the preferential use of hydrogen locally near biological reactions. Hydrogen derived from the mineralization of the substrate does not directly exchange with the total pool of water but with a smaller pool. The resulting local water pool has a less negative isotopic signature than the remaining water pool. Water incorporation through biosynthesis could then occur with this smaller pool of ^{2}H-enriched water. Moreover, hydrogen exchange within the whole water pool is slowed by the presence of clay, which accumulates molecule-derived hydrogen in hydroxyl sites. For this reason, ~~water-derived NEH~~H_{dfw} is also much higher in leptosol than in the other soils. ~~The NEH pool in leptosol is bigger (confirmed by the calculation of the total organic NEH) due to the inorganic NEH~~ The non-exchangeable hydrogen pool considered in the present equilibration method contains H in various positions and it may include water molecule acting as non-exchangeable hydrogen. The inorganic H_{dfw} pool is accounting for less than 1/40 of the amount of water-H in the leptosol. This pool appears as very slowly exchangeable with “free water”. In situ, on a short-term dynamics scale, this pool acts as non-exchangeable and is mostly at the hydroxyl position (López-Galindo et al., 2008).

The zonal distribution of organic compounds associated with minerals (Kleber et al., 2007; Vogel et al., 2014) may control the exchange between soil solution and organic compounds at kinetics that differ according to the layer within the organo-mineral interaction zone. The non-exchangeable hydrogen dynamics in soil organic matter are not independent of the mineral structure. The type of clay plays a role in carbon sequestration, depending on the specific surface area of the mineral or aggregate (Vogel et al., 2014; López-Galindo et al., 2008). In leptosol, clays are mainly smectite and have a high specific surface area. The high content of iron and hydroxide present in the leptosol also increases the specific surface area of the aggregates, which increases the organo-mineral association (Baldock and Skjemstad, 2000). Organic carbon cycling itself may be associated with mineral transformation (Basile-Doelsch et al., 2015), which may involve the newly formed hydroxyl.

The short-term dynamics of hydrogen are driven by the incorporation of hydrogen from water by isotopic exchange and by microbial biosynthesis. However, the increase in the incorporation of water hydrogen with the soil clay content suggests that part of the hydrogen is bound to clay or organo-mineral complexes. The production of NEH from water occurs mainly during the first weeks, but slow exchange of water hydrogen continues during the following year. Lopez-Galindo *et al.* observed the same trend, and they related the accumulation rate to the clay mineral properties (López-Galindo et al., 2008).

4.4 Ecosystem-scale production and fate of non-exchangeable hydrogen

In the present experiment, the preservation of non-exchangeable hydrogen from an organic substrate is less than 5 % after one year in soil with a low clay content. Water is the main donor of hydrogen during biosynthesis cycle favored by the breakage of the C-H bonds of the initial substrate. In this work, we showed that 70 % of the C-H

bonds of the initial substrate were broken during biosynthesis. Concerning the fate of tritium in terrestrial ecosystems, the isotopic composition of the organic plant material is a minor determinant of the bulk soil organic matter composition. However, ~~the~~ a better preservation of hydrogen from vegetation could ~~increase~~ arise from higher ~~with the~~ soil clay content and the subsequent ~~by~~ organo-mineral and zonal interactions. Water will be the main donor of organically bound tritium in the soil and the incorporation will be dependent to the carbon mineralization. ~~The lipids~~ Lipid isotopic composition ~~in this experiment have~~ highlighted that i- water derived hydrogen ~~the~~is involved in biosynthesis and ii- the newly formed non-exchangeable hydrogen is not necessary organic. Therefore, the incorporation of tritium from water in NEH pool is dependent on the clay content ~~in addition to the interaction with~~and on the soil hydrodynamics.

In our work, the isotopic composition of the NEH pool is determined by comparing labeled samples with unlabeled samples equilibrated under the same atmosphere. This method includes inorganic NEH. A proportion of the inorganic, non-exchangeable hydrogen should be taken into account in the prediction of the dynamics of hydrogen and tritium.

The main finding of the work is that the long-term fate of hydrogen in terrestrial environments and by extension the fate of tritium will depend on the status of soil carbon dynamics.

The proportion of NEH associated with minerals is itself partially related to the carbon dynamics.

Both the carbon dynamics and the incorporation of inorganic hydrogen in soils should therefore be taken into account in a conceptual model for the prediction of the long-term fate of hydrogen, and thereafter of tritium, in soil organic matter. The results of the present study can be used for the parameterization of the carbon-hydrogen coupling in such prediction models.

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Supplementary material includes the summary of incubation characteristics (Table S1), mass balance calculation (S2) ~~and~~ , propagation error calculation (S3) and $\delta^{13}C$ and $\delta^{2}H$ results of the incubation samples (S4)~~.~~

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Tables and Figures

Table 1: Pedologic, geographic information and carbon and hydrogen isotopic composition of the three bulk soils.

Soil	Latitude, longitude	Sampling depth (cm)	plant cover	% clay	% silt	% sand	C mg g^{-1}	H mg g^{-1}	N mg g^{-1}	pH	Mean bulk $\delta^{13}C$ (‰)	Mean bulk $\delta^{2}H$ (‰)
Cambisol	48° 48' 0.2 N; 2° 06' 33 W	0-25	wheat	17	50	33	12	3.2	1.27	6.8	-26 ± 0	-53±3
Podzol	44° 44' 33 N; 0° 47' 37 W	0-25	maize	5	3	92	20	2.3	0.84	5.5	-26 ± 0	-51±3
Leptosol	43° 56' 06 N; 5° 42' 06 W	5-10	oak forest	54	37	9	39	13.7	3.59	7.5	-25 ± 0	-52±2

Table 2: definition of the variable used in calculations

	Variables	Definition
Quantity (mg g^{-1} of dry soil)	H_{tot}	Total amount of hydrogen in the soil
	C_{tot}	Total amount of carbon in the soil
	H_{dfw}	Amount of non-exchangeable hydrogen derived from water
	H_{dfm}	Amount of non-exchangeable hydrogen derived from molecule
	C_{dfm}	Amount of carbon derived from molecule
	H_m	Initial amount of non-exchangeable hydrogen in the added molecule
	C_m	Initial amount of carbon in the added molecule
Abundance	$^{13}A_{tot\_0}$	^{13}C abundance of the unlabeled experiment (control)
	A_{tot_0}	^{2}H abundance of the unlabeled experiment (control)
	$^{13}A_{tot}$	^{13}C abundance of the total bulk soil
	A_{tot}	^{2}H abundance of the total bulk soil
	$^{13}A_m$	Initial ^{13}C abundance of the labeled molecule
	A_m	Initial ^{2}H abundance of the labeled molecule
	A_w	Initial ^{2}H abundance of the labeled water

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Table 3: $\delta^{13}C$, $\delta^{2}H$ of bulk soil and lipids at 365 days of incubation for cambisol, podzol and leptosol and the proportion of carbon and hydrogen derived from the labeled-source. In brackets the concentration in mg g^{-1} of carbon and hydrogen derived from the labeled-source

	$\delta^{13}C$ ‰		$\delta^{2}H$ ‰		% and concentration (mg g^{-1}) of the labelled source in the bulk soil			% and concentration (mg g^{-1}) of the labelled source in lipids		
	Bulk	lipids	bulk	lipids	% Cdfm (mg g^{-1})	% Hdfm (mg g^{-1})	% Hdfw (mg g^{-1})	% Cdfm (mg g^{-1})	% Hdfm (mg g^{-1})	% Hdfw (mg g^{-1})
					Cambisol					
Experiment 1	5.3 ± 0.9	19 ± 3	179 ± 2	133 ± 43	0.6 % (0.054 ± 0.009)	0.082 % (0.0031 ± 0.0003)	-	1 % (3.3 ± 0.6)	0.1 % (0.13 ± 0.01)	-
Experiment 2	5 ± 1	17 ± 11	422 ± 15	36 ± 6	0.6 % (0.055 ± 0.009)	-	1.5 % (0.06 ± 0.03)	1 % (4.2 ± 0.8)	-	1 % (0.68 ± 0.09)
					Podzol					
Experiment 1	13.2 ± 0.7	7 ± 5	94 ± 39	-106 ± 42	0.4 % (0.07 ± 0.01)	0.07 % (0.0017 ± 0.0007)	-	0.4 % (1.8 ± 0.9)	0.02 % (0.05 ± 0.02)	-
Experiment 2	7.0 ± 0.7	6 ± 7	259 ± 25	-38 ± 55	0.4 % (0.07 ± 0.01)	-	0.9 % (0.023 ± 0.004)	0.4 % (1.4 ± 0.1)	-	0.4 % (0.4 ± 0.2)
					Leptosol					
Experiment 1	22 ± 3	97 ± 2	723 ± 10	458 ± 65	0.4 % (0.15 ± 0.03)	0.24 % (0.029 ± 0.003)	-	0.8 % (5.9 ± 0.2)	0.19 % (0.18±0.01)	-
Experiment 2	25 ± 1	76 ± 20	798 ± 58	23 ± 5	0.4 % (0.14 ± 0.03)	-	6.8 % (0.8 ± 0.1)	0.8 % (4.8 ± 0.8)	-	1.1 % (0.95 ± 0.03)

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List of figures:

Figure1: Percentage of non-exchangeable hydrogen and carbon derived from the added molecule during one year of incubation for cambisol, podzol and mollic leptosol. The grey part corresponds to the results from 0 to 28 days. The line corresponds to the mean value calculated at each time for all molecules experiments.

Figure 2: Concentration of non-exchangeable hydrogen derived from the water H_{dfw} and derived from the molecule H_{dfm} for cambisol, podzol and mollic leptosol from 0 to 365 days. The line corresponds to the mean value calculated at each time for all molecules experiments. Note that the scale of Y axis is different for the three soils.

Figure 3: Concentration of non-exchangeable hydrogen derived from water with and without addition of substrate for the cambisol, podzol and leptosol from 0 to 28 days of incubation. Note that the scale of Y axis is different for the three soils.

Figure 4: Amount of substrate-derived H and C at 7 days of incubation versus the concentration of substrate as carbon (0.14, 0.29 and 0.43 mg g^{-1}). The concentration C in experiment 1 is equal to 0.43 mg g^{-1}.

Fig. 1

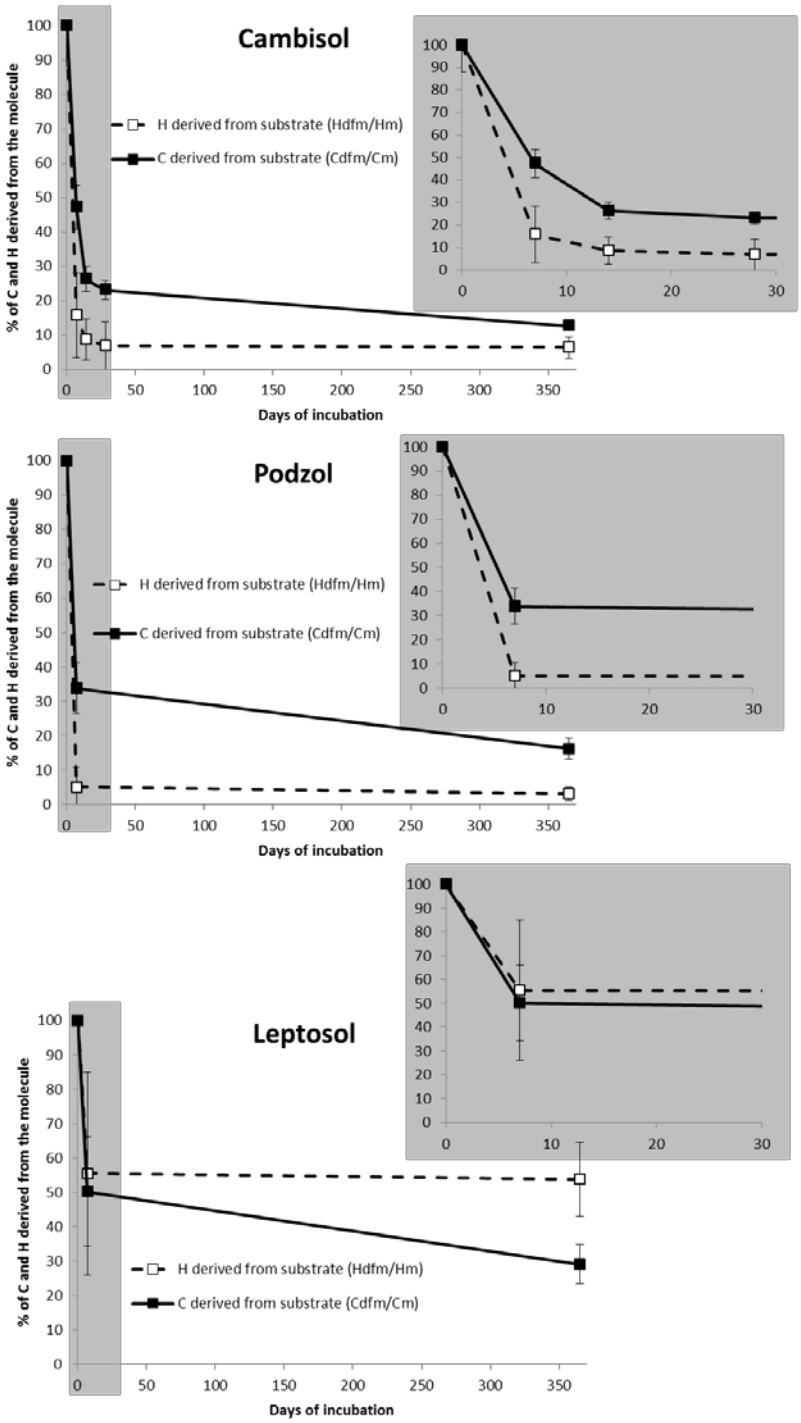
Cambisol
Podzol
Leptosol
% of C and H derived from the molecule
Days of incubation
H derived from substrate (Hdfm/Hm)
C derived from substrate (Cdfm/Cm)

Fig. 2

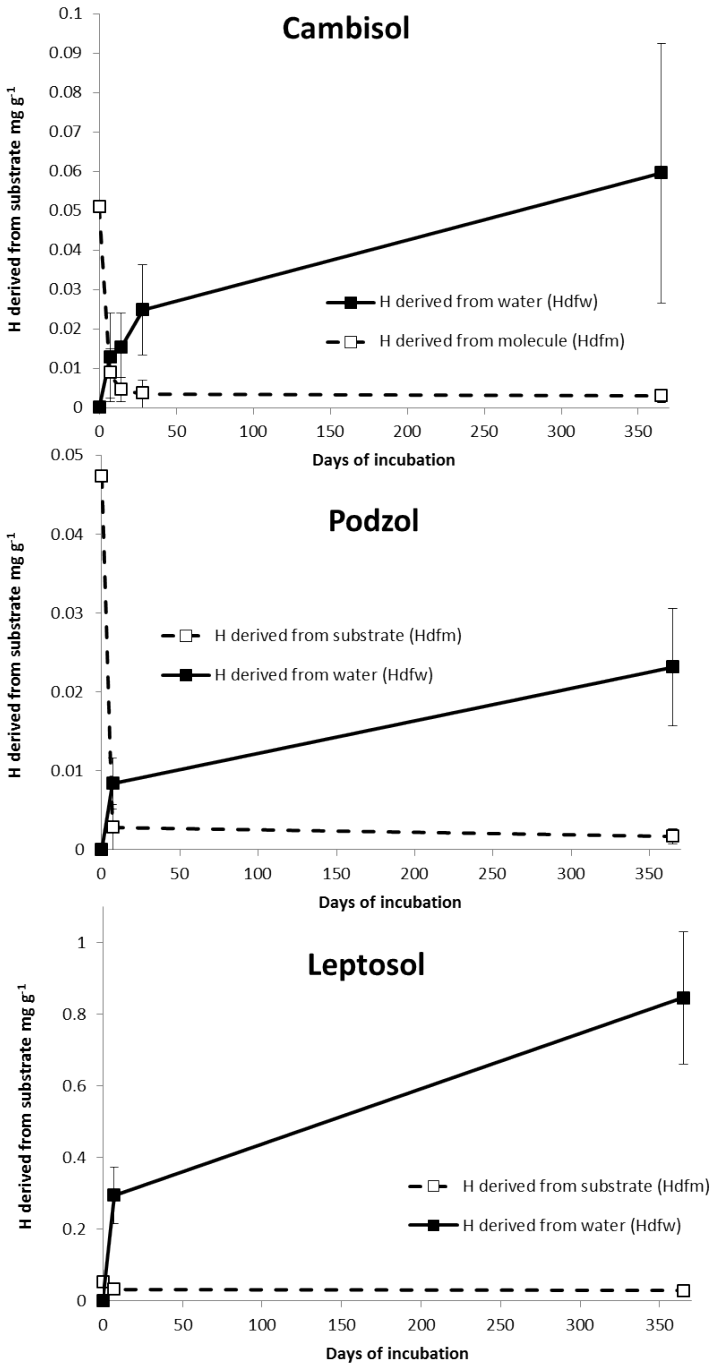
Cambisol
H derived from substrate mg g-1
Days of incubation
H derived from water (Hdfw)
H derived from molecule (Hdfm)
Podzol
H derived from substrate mg g-1
Days of incubation
H derived from substrate (Hdfm)
H derived from water (Hdfw)
Leptosol
H derived from substrate mg g-1
Days of incubation
H derived from substrate (Hdfm)
H derived from water (Hdfw)

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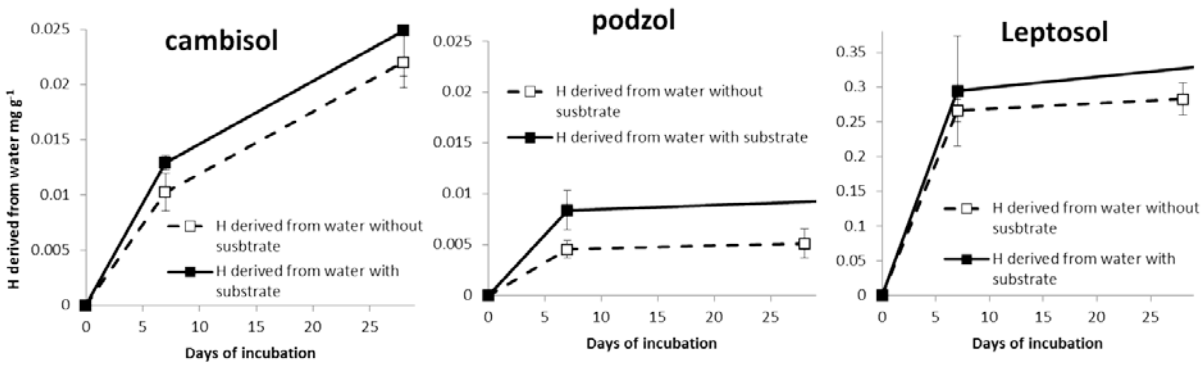
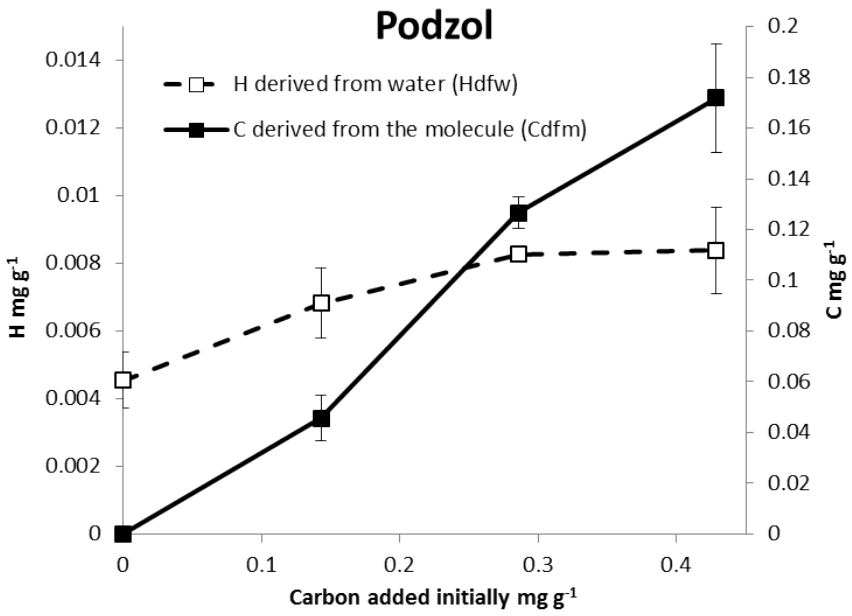
Fig.3


cambisol
podzol
Leptosol
H derived from water mg g-1
Days of incubation
H derived from water without susbtrate
H derived from water with substrate

615 Fig.4


Podzol
H derived from water (Hdfw)
C derived from the molecule (Cdfm)
H mg g-1
C mg g-1
Carbon added initially mg g-1
0
0.002
0.004
0.006
0.008
0.01
0.012
0.014
0.02
0.04
0.06
0.08
0.1
0.12
0.14
0.16
0.18
0.2
0.1
0.2
0.3
0.4