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Controls on redox-sensitive trace metals in the Mauritanian oxygen minimum zone

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ABSTRACT

The availability of the micronutrient iron (Fe) in surface waters determines primary production, N_2 fixation and microbial community structure in large parts of the world´s ocean, and thus plays an important role in ocean carbon and nitrogen cycles. Eastern boundary upwelling systems and the connected oxygen minimum zones (OMZs) are typically associated with elevated concentrations of redox-sensitive trace metals (e.g. Fe, manganese (Mn) and cobalt (Co)), with shelf sediments typically forming a key source. Over the last five decades, an expansion and intensification of OMZs has been observed and this trend is likely to proceed. However, it is unclear how trace metal (TM) distributions and transport are influenced by decreasing oxygen (O_2) concentrations. Here we present dissolved (d; <0.2 µm) and leachable particulate (Lp; >0.2 µm) TM data collected at 7 stations along a 50 km transect in the Mauritanian shelf region. We observed enhanced concentrations of Fe, Co and Mn corresponding with low O_2 concentrations (<50 µmol kg^{-1}), which were decoupled from major nutrients and nutrient-like and scavenged TMs (cadmium (Cd), lead (Pb), nickel (Ni) and copper (Cu)). Additionally, data from repeated station occupations indicated a direct link between dissolved and leachable particulate Fe, Co, Mn, and O_2. An observed dFe decrease from 10 to 5 nmol L^{-1} coincided with an O_2 increase from 30 to 50 µmol kg^{-1} and with a concomitant decrease in turbidity. The changes in Fe (Co and Mn) were likely driven by variations in their release from sediment pore water, facilitated by lower O_2 concentrations and longer residence time of the water mass on the shelf. Variations in organic matter remineralization and lithogenic inputs (atmospheric deposition or sediment resuspension) only played a minor role in redox-sensitive TM variability. Vertical dFe fluxes from O_2-depleted subsurface to surface waters (0.08–13.5 µmol m^{-2} d^{-1}) were driven by turbulent mixing and vertical advection, and were an order of magnitude larger than atmospheric deposition fluxes (0.63–1.43 µmol m^{-2} d^{-1}). Benthic fluxes are therefore the dominant dFe supply to surface waters on the continental margins of the Mauritanian upwelling region. Overall, our results indicated that the projected future decrease in O_2 concentrations in OMZs may result in increases in Fe, Mn and Co concentrations.

1. INTRODUCTION

The micronutrient iron (Fe) is essential for phytoplankton growth, but due to biological uptake coupled with a low solubility and low supply rates, the availability of Fe is typically low in open ocean surface waters (Bruland and Lohan, 2006). As a result, Fe limits primary production in high nitrate low chlorophyll regions (Boyd, 2007) and regulates dinitrogen (N_2) fixation in (sub)-tropical waters (Moore et al., 2009). Alongside Fe, other trace metals (TMs) such as cobalt (Co), manganese (Mn), zinc (Zn), cadmium (Cd) and copper (Cu) may (co-)limit phytoplankton growth and influence community composition (Browning et al., 2017; Moore et al., 2013; Morel and Price, 2003; Saito et al., 2008).

Oxygen minimum zones (OMZs) are characterized by stable subsurface oxygen (O_2) minima, which are maintained by a combination of enhanced O_2 consumption in the thermocline and a limited supply of O_2 rich water masses (e.g. Brandt et al., 2015; Karstensen et al., 2008; Wyrtki, 1962). Enhanced O_2 consumption is a result of elevated surface productivity caused by upwelling of nutrient-rich subsurface waters in eastern boundary regions of the oceans through Ekman divergence, and intense remineralization of sinking particles (e.g. Helly and Levin, 2004). Elevated organic matter supply and water column O_2 depletion lead to enhanced benthic release of redox-sensitive elements by influencing sediment diagenetic processes (Noffke et al., 2012; Severmann et al., 2010). Elevated concentrations of sediment derived dissolved Fe, Co and Mn have been associated with lateral offshore advection in O_2 depleted waters in the Arabian Sea, Pacific and Atlantic Ocean (Biller and Bruland, 2013; Hatta et al., 2015; Hawco et al., 2016; Milne et al., 2017; Moffett et al., 2015; Noble et al., 2012).

Oxygen concentrations affect the distribution of redox-sensitive TMs by controlling oxidation rates and influencing microbially mediated redox transformations. The reduced form of redox-sensitive TMs, such as iron (Fe(II)), cobalt (Co(II)) and manganese (Mn(II)), have a higher solubility in aqueous solutions than their oxidized forms (Fe(III), Co(III), Mn(III/IV)) (Liu and Millero, 2002; Stumm and Morgan, 1995). Reduction of these metals occurs to a large extent in anoxic sediment pore waters by microbial induced dissolution of particulate Fe(III) and Mn(III/IV) oxyhydroxides (Burdige, 1993; Chaillou et al., 2002; Froelich et al., 1979). Sediment pore waters are released to overlying bottom waters by diffusion and bio-irrigation and during submarine groundwater discharge (Beck et al., 2007; Elrod et al., 2004; Green et al., 2002). In contact with O_2 and other oxidants (e.g. nitrate (Schlosser et al., 2018) and hydrogen peroxide (Moffett and Zika, 1987)), Fe(II) oxidizes to the poorly soluble Fe(III) species, that are rapidly transformed into amorphous Fe oxyhydroxides or scavenged onto particle surfaces (Moffett and Zika, 1987; Scholz et al., 2016; Wu and Luther, 1994). Mn(II) also oxidizes to insoluble Mn(III/IV) oxides, but due to the slow abiotic oxidation kinetics, especially under low O_2 conditions (e.g. von Langen et al., 1997), biotic oxidation by manganese oxidizing bacteria is the main oxidation mechanism for Mn (Moffett, 1994; Sunda and Huntsman, 1988; Tebo and Emerson, 1986). Co(II) removal is mainly associated with incorporation of Co into Mn oxides by Co co-oxidation (Moffett and Ho, 1996).

Stabilizing mechanisms that prevent removal by scavenging and precipitation of Fe, Co and Mn are organic ligand complexation (Elrod et al., 2004; Liu and Millero, 2002; Oldham et al., 2017; Parker et al., 2007) and adsorption onto small slow sinking or neutral buoyant particles (Lam et al., 2012). Recent studies suggest a potentially important role for dynamic exchange processes between dissolved and particulate phases of Fe, thereby influencing cycling and transport (Achterberg et al., 2018; Fitzsimmons et al., 2017; Labatut et al., 2014; Milne et al., 2017). This was further indicated by Fe isotope studies suggesting an equilibrium isotopic fractionation between dissolved and particulate

phase in deep waters (Labatut et al., 2014) and the concomitant deepening of the dissolved and particulate Fe plume that originated from a hydrothermal vent (Fitzsimmons et al., 2017).

Spatial and seasonal variations in TMs that are released from sediments, as well as ex-situ sediment incubation experiments suggest a direct influence of bottom water and water column O_2 concentrations on the distribution of Fe, Co and Mn (e.g. Biller and Bruland, 2013; Homoky et al., 2012). Differences in benthic TM supply in field studies however are also influenced by a range of other processes as for example sediment type and organic matter supply (Homoky et al., 2016). Ex-situ sediment incubation experiments offer a potential means to disentangle the influence of O_2 concentrations relative to these controls (Homoky et al., 2012). These experiments, however, need to be interpreted within the context of the confined conditions that eliminate potentially important interactions in open systems, such as seawater exchange and mixing. Furthermore, they offer no means to confidently evaluate controls on TM distributions in the pelagic water column.

In an attempt to resolve the controls on TM release and stabilization in OMZs we measured the concentration of a suite of TMs along a 50 km long transect on the Mauritanian shelf in the Eastern Tropical North Atlantic (ETNA). The Mauritanian shelf is associated with a major OMZ (minimum O_2 concentrations below 40 µmol kg^{-1}; Brandt et al., 2015) and is an important Fe source to the North Atlantic Ocean (Milne et al., 2017). Furthermore, atmospheric dust deposition from the Saharan desert can markedly elevate surface water Fe concentrations in the ETNA (Conway and John, 2014; Rijkenberg et al., 2012). Recent observations suggest a decline in O_2 content of the oceans, particularly in the northern and southern eastern Atlantic, and an expansion of OMZs, modulated by the variability of our climate system (Hahn et al., 2017; Schmidtko et al., 2017; Stramma et al., 2008b). These changes may result in changes in TM supply, and a mechanistic understanding of the factors regulating TM release and stabilization in OMZs is therefore urgently needed. The aim of this study was to evaluate the direct influence of variability in water column O_2 concentrations on the distribution of redox-sensitive TMs and to identify responsible control mechanisms. Firstly, we assess the fluxes of dFe in the OMZ to surface waters by advection and diffusive mixing and compared those to the atmospheric deposition flux of dFe. Secondly, we evaluate the importance of redox and non-redox controls on Fe, Co and Mn by focusing on the influence of O_2 and particles on the distribution of dissolved and leachable particulate TMs, including redox-sensitive (Fe, Co and Mn) and nutrient-type and scavenged trace metals (aluminum (Al), lead (Pb), nickel (Ni), Cd and Cu). Thirdly, we determine the influence of variability of the eastern boundary circulation and O_2 concentrations in regulating TM concentrations.

2. METHODS

2.1 Sampling

Samples were collected on RV Meteor cruise M107 in June 2014 during nine deployments at seven locations (two stations were occupied twice) along a cross-shelf transect at 18°20'N on the Mauritanian shelf in the ETNA (Figure 1). The bottom depths of stations varied between 50 m on the shelf to 1136 m furthest off shore. Seawater sampling was carried out using a trace metal clean CTD (TM-CTD, Sea-Bird SBE25) rosette frame equipped with 24 trace metal clean samplers (12 L, Ocean Test Equipment (OTE)). The CTD frame was attached to plastic coated nonconductive steel cable and deployed using a carousel auto-fire module (AFM, Sea-Bird) that closed the bottles at predefined depths. After recovery, the bottles were transferred to a clean-laboratory container and pressurized to 0.2 bar overpressure using filtered N_2 gas. Samples were collected unfiltered for total dissolvable (TD) TM measurements, and filtered using a 0.2 µm cartridge filter (Acropack 500, Pall) for dissolved (d) TMs and iodide. Trace metal samples were collected in acid clean 125 mL low density polyethylene (LDPE) bottles (Nalgene), and iodide samples in opaque 60 mL high density polyethylene (HDPE) bottles (Nalgene). Trace metal samples were acidified to pH 1.9 using ultra clean HCl (UpA, Romil) and stored double-bagged for >6 months before preconcentration and analysis. Samples for iodide measurements were stored frozen at -20°C until analysis.

Samples for the determination of radium isotopes (^{223}Ra; $t_{½}$ = 11.4 d; ^{224}Ra $t_{½}$ = 3.7 d) were obtained using in-situ filtration pumps (Challenger Oceanic) following the procedures described in Charette et al. (2015) and Henderson et al. (2013). Briefly, each in-situ filtration pump was equipped with two particle filters (70 µm; 1 µm) and two Mn dioxide (MnO_2) impregnated cartridges (CUNO Micro Klean III acrylic) on which dissolved Ra adsorbs. The pumped water volumes varied between 1000 L and 1700 L. For the determination of Ra in surface waters (~5 m water depth) about 200–300 L of seawater was pumped into 500 L plastic barrels followed by filtration over MnO_2 coated acrylic fibers (Mn-fibers).

2.2 Trace metal analysis

Determination of Co, Mn, Fe, Cd, Pb, Ni and Cu was carried out as described in Rapp et al. (2017). Briefly, samples were preconcentrated using an automated preconcentration device (Sea*FAST*, Elemental Scientific Inc.) equipped with a cation chelating resin (WAKO; Kagaya et al., 2009). Samples were UV-digested prior to preconcentration to breakdown metal-organic complexes, which would cause an underestimation of the determined TM concentrations. Samples were buffered in-line to pH 6.4 ± 0.2 using 1.5 M ammonium acetate buffer, before loading onto the resin. The pH buffer was prepared using an ammonium hydroxide solution (22%, OPTIMA grade, Fisher) and acetic acid (glacial, OPTIMA grade, Fisher) in de-ionized water (MilliQ, Millipore), adjusted to pH 8.5. Retained TMs were eluted from the resin using 1 M distilled HNO_3 and collected in 4 mL polypropylene scintillation vials (Wheaton). The acid was distilled from supra-pure HNO_3 (SpA grade, Romil) using a sub-boiling PFA distillation system (DST-1000, Savillex). Preconcentration was performed within a clean laboratory (ISO 5) and all sample and reagent handling was performed within the same

laboratory in an ISO 3 laminar flow bench with a HEPA filter unit. Preconcentrated samples were analyzed by high resolution inductively coupled plasma-mass spectrometry (HR-ICP-MS, ELEMENT XR, ThermoFisher Scientific) using isotope dilution for Fe, Cd, Pb, Cu and Ni and standard additions for Co and Mn. SAFe reference seawater S and D2 were analyzed with each analytical run and concentrations produced were in good agreement with consensus values (Table 1).

Leachable particulate (Lp) concentrations were calculated as the difference between total dissolvable and dissolved concentrations. The limit of quantification (LOQ) for the Lp concentrations was determined as the sum of the analytical standard deviations of TD and dissolved concentrations. Extended uncertainty calculations were performed using the Nordtest approach (Naykki et al., 2015) accounting for random as well as systematic errors (Rapp et al., 2017). The Lp fraction represents the particulate fraction which is readily dissolvable in the acidified samples during storage at pH 1.9 for 6 months and therefore does not contain any refractory particle components. This more labile fraction of particulate TMs mainly includes TMs in organic/biogenic particles, adsorbed to particle surfaces and TM oxides/oxyhydroxides (Hurst et al., 2010).

2.3 Aluminum measurements

Aluminum concentrations were determined in surface water samples for all stations along the transect and at two stations (3 and 8) for the entire water column. Samples were analyzed for Al according to Hydes and Liss (1976). Acidified samples were buffered with a 2 M ammonium acetate buffer (Romil, UpA) to a pH between 5.1 and 5.2. Buffered samples were spiked with a 2 mg L^{-1} lumogallium (TCI) solution. The lumogallium solution was prepared in 2 M ammonium acetate buffer (Romil, UpA). After spiking, samples were heated up for 1.5 h at 80°C in an oven (Heratherm, Thermo Scientific) and left to cool down overnight at room temperature to allow the formation of a fluorescence Al complex. Samples were measured using a fluorescence spectrophotometer (Cary Eclipse, Agilent). The samples were measured with an excitation and emission wavelength of 465 and 555 nm, respectively. All samples were analyzed in duplicate and the concentrations calculated from the peak heights via standard addition. GEOTRACES reference seaweater (GS) was run with a mean average Al value of 27.76 ± 0.17 nmol L^{-1} (n=4; consensus value 28.2 ± 0.2 nmol L^{-1}).

2.4 Iodide measurements

Frozen samples were defrosted overnight at room temperature prior to analysis for iodide by cathodic stripping square wave voltammetry after Luther et al. (1988). The voltammetry unit consisted of a voltammeter stand (663 VA, Metrohm), an autosampler (863 Compact Autosampler, Metrohm) and an automatic burette (843 Pump Station, Metrohm) for automated spike addition. The system was controlled by Computrace software (797 VA; Metrohm).

2.5 Oxygen, salinity, nutrient, turbidity and chlorophyll fluorescence analysis

Oxygen, salinity, nutrients, turbidity and chlorophyll fluorescence was measured during 62 CTD deployments (including some repeated deployments at the same location) along the 18°20`N transect using a Sea-Bird SBE 9 CTD rosette system equipped with double sensor packages for O_2, salinity and temperature and 24 niskin samplers (10 L; OTE). Turbidity and chlorophyll a fluorescence were measured with single sensor units on the CTD. Oxygen sensor data were calibrated by Winkler titration (Hansen, 2007; Winkler, 1988) on 348 discrete water samples that were collected from the OTE samplers. The O_2 calibration was undertaken using a linear fit with respect to O_2 concentration, temperature, and pressure. An uncertainty of 1.5 µmol kg^{-1} was determined. On-board nutrient measurements of nitrite (NO_2^-), nitrate (NO_3^-), phosphate (PO_4^{3-}) and silicic acid ($Si(OH)_4$) of the discrete water samples were conducted using a QuAAtro autoanalyzer (Seal Analytical) according to Grasshoff et al. (1983).

Apparent Oxygen Utilization (AOU) was calculated as the difference between saturation concentrations of O_2 and measured O_2 concentrations. The saturation concentration of O_2 was calculated after the Weiss methods (Weiss, 1970) using the R package marelac (Soataert et al., 2016), taking into account salinity and temperature.

2.6 Radium analysis

On-board the ship the Mn-cartridges and Mn-fibers were washed with Ra-free tap water and afterwards partially dried with filtered compressed air to remove excess water. The samples were analyzed for ^{223}Ra, ^{224}Ra and ^{228}Th using a Radium Delayed Coincidence Counting System (RaDeCC) (Moore and Arnold, 1996). For the efficiency calibration of the RaDeCC, ^{227}Ac and ^{232}Th standard solutions were used, and the calibration followed the procedure described in Scholten et al. (2010) and Moore and Cai (2013). Counting errors were propagated following Garcia-Solsona et al. (2008). Excess ^{224}Ra ($^{224}Ra_{ex}$), i.e. the ^{224}Ra activity corrected for ^{228}Th-supported ^{224}Ra was calculated by subtracting the ^{228}Th activity from the ^{224}Ra activity. As we measured only the first Mn cartridge and the Mn cartridges do not adsorb radium quantitatively, we report here only $^{224}Ra_{ex}/^{223}Ra$ ratios.

2.7 Turbulence measurements and vertical flux calculations

In order to advance understanding of the role of benthic Fe supply to the high productive surface waters of the upwelling region, vertical diffusive fluxes (eq 1: left term, right hand side) and upwelling induced vertical advective fluxes (eq 1: right term, right hand side) were estimated. At depth in the water on a continental margin, solutes are transferred vertically toward the surface waters by turbulent mixing processes and by vertical advection forced by Ekman divergence (e.g. Steinfeldt et al., 2015):

$$J_Z = K_z \frac{\partial[TM]}{\partial z} + w \cdot \Delta[TM] \qquad (1)$$

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Here, K_z is the turbulent eddy diffusivity in $m^2\ s^{-1}$, ∂[TM]/ ∂z the vertical gradient with depth (z) of the TM concentration [TM] in µmol m^{-4}, Δ[TM] a TM concentration difference in µmol m^{-3} and w represents vertical velocity in m s^{-1}. The equation is solved by vertically integrating the tracer transport budget equation between two vertical layers while ignoring lateral fluxes, changes of w with depth and assuming steady state. Vertical advective fluxes resulting from meso- and submesoscale processes along sloping isopycnals were not considered. The TM-fluxes were evaluated for the depth interval from the upper boundary of the shallow O_2-depleted waters to a depth of increased chlorophyll a fluorescence (8–29 m depth).

Diffusive Fe fluxes were determined by combining TM concentration measurements from the TM-CTD stations with nearby measured microstructure profiles. The microstructure measurements were performed with an MSS90-D profiler (S/N 32, Sea & Sun Technology). The loosely-tethered profiler was optimized to sink at a rate of 0.55 m s^{-1} and equipped with three shear sensors, a fast-response temperature sensor, and an acceleration sensor, two tilt sensors and conductivity, temperature, depth sensors sampling with a lower response time. At TM-CTD stations with bottom depths less than 400 m, 18 to 65 microstructure profiles were available at each station. At deeper stations, the number reduced to 5 to 12 profiles. Standard processing procedures were used to determine the rate of kinetic energy dissipation (ε) of turbulence in the water column (see Schafstall et al. (2010) for detailed description). Subsequently, K_Z values were determined from $K_p = \Gamma\varepsilon N^{-2}$ (Osborn, 1980), where N is stratification and Γ is the mixing efficiency for which a value of 0.2 was used. The use of this value has recently shown to yield good agreement between turbulent eddy diffusivities determined from microstructure measurements and from tracer release experiments performed in our study region (Köllner et al., 2016). The 95% confidence intervals for station-averaged K_ρ values were determined from Gaussian error propagation following Schafstall et al. (2010). Finally, diffusive fluxes were estimated by multiplying station-averaged K_ρ with the vertical gradient of the respective TM solute, implicitly assuming K_Z=K_ρ.

The vertical advective flux by Ekman divergence requires determination of vertical velocity in the water column that varies with depth and distance from the coast line. Recent studies found good agreement between vertical velocities derived from Ekman divergence and from helium isotope equilibrium within the Mauritanian and Peruvian coastal upwelling regions (Steinfeldt et al., 2015) when parameterizing vertical velocities as (Gill, 1982):

$$w = \frac{\tau_y}{\rho f L_r} e^{-x/L_r}$$

where τ_y represents the alongshore wind stress, ρ the density of sea water, x the distance from maximum Ekman divergence taken here as the position at 50 m bottom depth on the shelf and L_r the first baroclinic Rossby radius. The parameterization results from considering the baroclinic response

of winds parallel to a coastline in a two-layer ocean (Gill, 1982). The baroclinic Rossby radius $L_r = f^{-1}\sqrt{g\frac{\rho_2-\rho_1}{\rho}\frac{H_1H_2}{H_1+H_2}}$ ($\rho_{1/2}$ and $H_{1/2}$ is density and thickness of the surface and lower layer, respectively) was found to be 15 km from hydrography collected during the cruise, similar to the values determined by Steinfeld et al. (2015) in the same region. Using average alongshore wind stress from satellite data (0.025 Nm^{-2}, ASCAT winds; Ricciardulli and Wentz, 2016) for June 2014, maximum vertical velocities of $3.7x10^{-5}$ m s^{-1} were determined for the shelf region (50 m water depth), which decayed offshore to $1.7x10^{-6}$ m s^{-1} at the position of the 1000 m isobath at 18°N. As these vertical velocities describe the magnitude of upwelling at the base of the mixed-layer, additional corrections need to be considered for deeper depths. Here, we approximated the vertical decay of *w* as a linear function which diminishes at the ocean floor.

The calculation of the vertical advective flux supplying solutes from the shallow O_2-depleted waters to the chlorophyll *a* maximum requires knowledge of a concentration difference Δ[TM] associated with the upwelling flux. Ideally, the vertical scale of the concentration difference is determined by correlation analysis of vertical velocity fluctuations and concentration variability at different depths (*w*'·[TM]'). As these data are not available, we chose to use the mean vertical concentration differences over a vertical distance of 10 m. Thus, the vertical advective flux F_{az} at each station was estimated from $F_{az} = w(x,z)\cdot\overline{\frac{\partial[TM]}{\partial z}}10\,m$.

2.8 Figures

All figures were produced in R (version 3.4.3). Data gridding in figures 2 and 3 was performed using the Tps function within the fields package in R (Nychka et al., 2016).

3. RESULTS & DISCUSSION

3.1 Oceanographic settings of the study area

The cruise was conducted in June 2014 along a transect crossing a narrow shelf off the Mauritanian coast at 18°20'N. The vertical structure of the OMZ in this region is characterized by a deep OMZ at about 400 m depth, and a shallow OMZ at about 100 m depth (Brandt et al., 2015). Coastal upwelling of nutrient-rich deep water occurs as a result of offshore transport of surface waters caused by a Northeast Trade wind component parallel to the coast. While north of 20°N upwelling persists throughout the year, upwelling south of 20°N, including the Mauritanian upwelling region, undergoes seasonal changes in upwelling strength (Barton et al., 1998), with strongest upwelling occurring between December and April. The seasonal variability is mainly driven by changes in wind forcing associated with the migration of the Intertropical Convergence Zone (Schafstall et al., 2010).

The eastern boundary circulation consists of the Mauritania Current (MC, Fig. 1) flowing poleward at the surface against the equatorward winds and of the Poleward Undercurrent (PUC) flowing in the same direction at depths between 50 and 300 m (Barton, 1989; Klenz et al., 2018; Mittelstaedt, 1983; Peña-Izquierdo et al., 2015). Both currents supply cold, O_2 and nutrient-rich waters of predominantly South Atlantic origin (South Atlantic Central Water, SACW) to the coastal upwelling region (e.g. Mittelstaedt, 1991; Mittelstaedt, 1983; Peña-Izquierdo et al., 2015). In response to the changing winds, the eastern boundary circulation likewise exhibits a pronounced seasonal variability (Klenz et al., 2018; Stramma et al., 2008a). The strongest poleward flow is observed during the relaxation period between May and July when alongshore, upwelling-favorable winds weaken but wind stress curl is at its maximum (Klenz et al., 2018). During the upwelling season in boreal winter, the circulation more closely resembles the classical eastern boundary circulation regime, with a weak poleward undercurrent flowing beneath an equatorward coastal jet. At deeper levels (300–500 m depth), flow was found to be equatorward during both seasons. The shallow (<300 m depth) boundary circulations turn offshore at the southern flank of the Cape Verde frontal zone (CVFZ) (e.g. Tomczak, 1981; Zenk et al., 1991) at about 20°N, separating SACW from more saline and O_2-rich Central Waters formed in the North Atlantic (NACW). The circulation in June 2014 was typical for a relaxation period characterized by little upwelling and a strong poleward flow over the entire shelf between the surface and 250 depth (Klenz et al., 2018).

Meridional sections of water mass properties and O_2 concentrations from around 18°N showed that waters with an enhanced SACW proportion advected from the south as well as NACW coming from the north, have higher O_2 concentrations than the ambient waters (Klenz et al., 2018). The mixture of SACW and NACW waters found in the thermocline particularly during boreal winter, previously identified as a regional water mass and termed the Cape Verde SACW (SACWcv) by Peña-Izquierdo et al. (2015), is a signature of an older water mass with lower O_2 concentrations than those of SACW or NACW due to a longer residence time and O_2 consumption through remineralization. Elevated pelagic oxygen consumption levels at the Mauritanian continental margin were recently determined by Thomsen et al. (2018). During the transition period in May through July upper Central Waters (50–300 m depth) are dominated by SACW accounting for 80–90 % of the water masses in the boundary current region (Klenz et al., 2018).

The SACW transported poleward within the boundary circulation is supplied by the zonal North Equatorial Counter Current (NECC) and North Equatorial Under Current (NEUC), which flow eastward at about 5°N (Brandt et al., 2015) before diverging into a northward and a southward flowing branch in front of the African coast.

As a result of interactions between tidal currents, topography and critically sloping upper continental slope topography (e.g. Eriksen, 1982), the Mauritanian upwelling region is known for elevated nonlinear internal wave activity resulting in enhanced mixing in the water column of the upper slope

and shelf region (Schafstall et al., 2010). Vertical fluxes of nutrients driven by mixing processes are amongst the largest reported in literature (Cyr et al., 2015).

The CTD and microstructure deployments were performed along the east-west transect in the period June 8 to June 27 (2014) (Fig. 1). Oxygen concentrations reached a deep minimum of 40–50 µmol kg^{-1} at about 400 m and a shallow minimum of 30–50 µmol kg^{-1} at about 50–100 m (Fig. 2), which is in agreement with previous studies (Brandt et al., 2015; Thomsen et al., 2018). Mixed layer depths ranged between 10 and 22 m during the cruise. Salinity was highest at the surface (ca. 36.02) and generally decreased with depth to a minimum of 34.71 at around 1000 m. Nitrate (NO_3^-) concentrations in the surface mixed layer varied between 0.1 and 11.3 µmol L^{-1} and phosphate (PO_4^{2-}) between 0.15 and 0.91 µmol L^{-1}. NO_3^- and PO_4^{2-} concentrations increased with depth to a maximum of 47.6 and 3.2 µmol L^{-1}, respectively (Fig. 2).

Over a time period of 19 days, two trace metal stations were reoccupied along the transect at water depths of 170 m (18.23 °N, 16.52 °W, 1st deployment: June 12, 2nd deployment: June 21) and 189–238 m (18.22°N, 16.55°N, 1st deployment: June 24, 2nd deployment: June 26). Minimum O_2 concentrations of 30 µmol kg^{-1} observed before June 15, which increased to 50 µmol kg^{-1} after June 19 or June 24, depending on the location. This oxygenation event that was also captured in ocean glider measurements is discussed in detail by Thomsen et al. (2018). They attributed the change to physical transport of SACW into the region (Thomsen et al., 2018), most likely associated with the observed increase in current speed of the MC flowing northward parallel to the coast line and transporting relatively O_2-rich water while decreasing the residence time of the SACW along the continental margin. Additionally, pelagic oxygen consumption was found to contribute to the variability in oxygen concentrations close to the seafloor (Thomsen et al., 2018).

3.2 Spatial distributions of dissolved and leachable particulate trace metals

Dissolved Fe and LpFe concentrations ranged between 0.97–18.5 nmol L^{-1} and 1.6–351 nmol L^{-1}, respectively (Fig. 3a, b). Surface waters (5–29 m) had lowest dFe (0.97–4.7 nmol L^{-1}) and LpFe (1.6–35.9 nmol L^{-1}) concentrations, whereas highest concentrations were present on the shelf close to the seafloor (up to 18.5 nmol L^{-1} dFe and 351 nmol L^{-1} LpFe). Enhanced concentrations of both Fe fractions at any given station were observed at depths with low O_2 concentrations (30–60 µmol O_2 kg^{-1}). A similar distribution pattern was observed for dCo, with concentrations between 0.069 and 0.185 nmol L^{-1} (Fig. 3c). In contrast, LpCo concentrations varied between below LOQ and 0.179 nmol L^{-1} and were generally highest in surface waters and close to the coast (Fig. 3d). Compared to dFe, the concentration range of dCo was much narrower and enhanced concentrations were observed over a broader depth range and further offshore.

Surface dFe and dCo concentrations were low, presumably due to enhanced biological uptake. No clear increasing trend in dFe and dCo with depth was observed, indicating that processes other than, or

in addition to, remineralization influenced their distributions. Elevated concentrations were found close to the sediments and within low O_2 waters. This suggested a benthic source of Fe and Co under O_2-depleted conditions, and offshore transport along O_2-depleted water filaments, which is in agreement with previous studies (e.g. Hatta et al., 2015; Hawco et al., 2016; Noble et al., 2012). Our sharper onshore-offshore gradient of dFe concentrations compared to dCo in O_2-depleted waters shows that oxidation and removal mechanisms/scavenging rates were faster for Fe than Co (Noble et al., 2012). Previously reported dFe concentrations in coastal regions of the tropical North Atlantic were lower than we observed, between 0.5–6.3 nmol L^{-1} (Hatta et al., 2015; Milne et al., 2017). However, all these samples were collected at a greater distance from the coast. In the near-coastal Oregon and Washington shelf bottom water dFe concentrations were similar to our study under equivalent O_2 concentrations (18.7–42.4 nmol L^{-1} dFe, 42–61 µmol kg^{-1} O_2; Lohan and Bruland, 2008), whereas in the euxinic waters from the Peruvian shelf region, dFe concentrations were more than an order of magnitude higher, exceeding 200 to 300 nmol L^{-1} (Schlosser et al., 2018; Scholz et al., 2016). Similar dCo concentrations to our study were observed in the North and South Atlantic, with highest concentrations of ~0.16 nmol L^{-1} present within O_2-depleted waters (Noble et al., 2012; Noble et al., 2017).

Dissolved Mn concentrations ranged between 0.46–13.8 nmol L^{-1} and LpMn between below LOQ–4.4 nmol L^{-1} (Fig. 3e, f). Highest dMn and LpMn concentrations were observed in surface waters, generally decreasing with depth. Additionally, concentrations were highest on the shelf and decreased offshore. The dMn concentrations were generally elevated within and below the deeper O_2-depleted waters with 0.70–1.34 nmol L^{-1} compared to 0.46–0.91 nmol L^{-1} just above. The increased dMn concentrations within the deeper O_2-depleted waters (~350–500 m depth) indicate a benthic source, similar to Fe and Co, which is in accordance with previous studies (Noble et al., 2012). However, in the shallow O_2-depleted waters (~50–200 m depth), this effect is not resolvable due to high surface concentrations, which were maintained by photo-reduction of Mn oxides to soluble Mn(II) that prevents loss of Mn from solution (Sunda and Huntsman, 1994). Reported dMn concentrations in the North and South Atlantic were lower than in our study, with concentrations <3.5 nmol L^{-1} in surface waters and around 0.5–1 nmol L^{-1} dMn within the OMZ (Hatta et al., 2015; Noble et al., 2012). As for dFe, these lower reported values can also be explained by sampling stations positioned at further distance from the coast and removal of dMn via biological oxidation processes with distance from the source (Moffett and Ho, 1996).

Dissolved Cd and Ni concentrations were lowest in surface waters with 0.022–0.032 nmol Cd L^{-1} and 2.6–2.8 nmol Ni L^{-1}, and showed an increasing trend with depth to maximum values of 0.60 nmol L^{-1} and 5.8 nmol L^{-1}, respectively (Fig. 3g, m). Leachable particulate Cd concentrations were between below LOQ and 0.20 nmol L^{-1}, and LpNi concentrations between below LOQ and 1.7 nmol L^{-1}. A large fraction of Ni (72–100%) was present in the dissolved form. The majority of LpNi samples were

below the LOQ (>70% of the data) and LpNi is therefore not included in Fig. 3. LpCd concentrations were highest close to the coast and decreased offshore (Fig. 3h). In surface waters close to the coast the LpCd fraction was dominant with up to 84.3% of the entire Cd pool (d + Lp). The fraction of LpCd in surface water beyond the shelf break (including stations 2, 1 and 9) contributed still up to 54.3% of the Cd pool, whereas below 50 m only 0–12.8% of TDCd was in the Lp phase beyond the shelf break. In contrast to Fe, Co and Mn, no increases in Cd and Ni were observed near the seafloor and within the O_2-depleted waters indicating that Cd and Ni concentrations are mainly controlled by remineralization of sinking organic matter, which is typical for these two nutrient-like TMs (Biller and Bruland, 2013). Similar distributions with concentrations between 0 and 1000 m water depth ranging from ~2–5.5 and ~0–0.55 nmol L^{-1} for dNi and dCd, respectively, were observed during the GEOTRACES transect GA03_w in the tropical North Atlantic (Mawji et al., 2015; Schlitzer et al., 2018).

Dissolved Cu concentrations in surface waters ranged between 0.63–0.81 nmol L^{-1} (Fig. 3i). Concentrations increased with depth to around 1.37 nmol L^{-1} at 700 m depth close to the seafloor, whereas highest observed concentrations further offshore were 0.95 nmol L^{-1} at the greatest sampled depth of 850 m. These results indicate that in addition to remineralization processes of sinking biogenic particles, the distribution of Cu is influenced by inputs from the seafloor. This is in accordance with previous studies, suggesting that Cu is released from continental shelf sediments under oxic and moderately reducing conditions (Biller and Bruland, 2013; Heggie, 1982), whereas no increase in Cu concentrations near the seafloor was observed at low bottom water O_2 concentrations (O2 <10 µM; Johnson et al., 1988). A decrease in Cu concentrations in the bottom boundary layer was also reported with a seasonal decrease in O_2 in summer from a minimum of 70 µM O_2 in May to 40 µM O_2 in August, suggesting a decrease in sedimentary release of Cu (Biller and Bruland, 2013). In strongly reducing sediments and the presence of H_2S, Cu forms inorganic sulfides and precipitates, which may explain reduced sedimentary Cu release under low bottom water O_2 concentrations (Biller and Bruland, 2013). Therefore, the sediment source of dCu might show a different dependency on bottom water O_2 concentrations than dFe, dCo and dMn explaining the distinct distribution of dCu. Concentrations of LpCu were between below the LOQ to 0.61 nmol L^{-1} with enhanced levels at station 4 close to the coast and at mid depths of the three stations furthest offshore (9, 5 and 2) (Fig. 3j).

Observed dPb concentrations were lowest in the surface waters at 9–14 pmol L^{-1} and increased with depth to 29–86 pmol L^{-1} below 600 m depth (Fig. 3k). Lead is not considered a nutrient-like TM (e.g. Boyle et al., 2014), but our observations indicate a release of Pb from sinking particles following remineralization. The concentration range and depth distribution is similar to reported distributions further offshore at about 21°W (Noble et al., 2015). These authors suggested that increased concentrations of up to 70 pmol L^{-1} between 600 and 800 m depth were related to the influence of Mediterranean Outflow Waters (MOW). Additionally, increased Pb concentrations in proximity to

sediments have been attributed to the benthic release of historic Pb through reversible scavenging from particles and the release of dPb associated with Fe/Mn oxyhydroxides during reductive dissolution of those oxides in anoxic sediments (Rusiecka et al., 2018). The major source of Pb to the ocean is atmospheric dust deposition from anthropogenic emissions (Bridgestock et al., 2016; Nriagu and Pacyna, 1988; Veron et al., 1994) with a recent indication of reduced anthropogenic Pb inputs to surface waters in the eastern tropical Atlantic under the North African dust plume (Bridgestock et al., 2016). Low surface water concentrations on the Mauritanian shelf indicate low atmospheric inputs of Pb to this region. LpPb was below the LOQ–27 pmol L^{-1}, and the distribution of LpPb was similar to that of LpFe, with subsurface maxima within O_2-depleted waters (Fig. 3l) and may indicate increased scavenging of dPb in these layers which might be associated with Fe containing particles.

In general, sediment derived TM concentrations decrease with distance from the shelf and with time that passed since the water mass has been in contact with the sediments due to water mass mixing and removal processes such as precipitation and scavenging (Bruland and Lohan, 2006). Radium isotopes can be used as a tracer for benthic sources. The major source of Ra to the ocean is input from sediments through the efflux of pore water, sediment resuspension, and submarine groundwater discharge (Moore, 1987; Moore and Arnold, 1996; Rama and Moore, 1996). Due to the distinctive half-lives of the different Ra isotopes (e.g. ^{224}Ra ($t_{1/2}$ = 3.66 d) and ^{223}Ra ($t_{1/2}$ = 11.4 d)) and their conservative behaviour in seawater, it is possible to quantify the time that has passed since a parcel of water was in contact with the sediments using the following equation by Moore (2000):

$$\left(\frac{A_{224}}{A_{223}}\right)_{obs} = \left(\frac{A_{224}}{A_{223}}\right)_{i} \frac{e^{-\lambda_{224}\tau}}{e^{-\lambda_{223}\tau}} \qquad (2)$$

solved for water mass age (τ):

$$\tau = \frac{\ln\left(\frac{A_{224}}{A_{223}}\right)_{obs} - \ln\left(\frac{A_{224}}{A_{223}}\right)_{i}}{\lambda_{223} - \lambda_{224}} \qquad (3)$$

where A_{224}/A_{223} is the activity ratio of ^{223}Ra and ^{224}Ra, with the subscript *obs* for the observed seawater ratio and the subscript *i* for the initial groundwater endmember ratio, and λ_{223} and λ_{224} are the decay constants in d^{-1} for ^{223}Ra and ^{224}Ra.

Highest $^{224}Ra_{ex}/^{223}Ra$ activity ratios were observed close to the seafloor (Fig. 3n). The average $^{224}Ra_{ex}/^{223}Ra$ ratio in proximity to the sediment source (< 20 m above seafloor) was 4.1 ± 0.7 and was similar to reported ratios for shelf waters off South Carolina ($^{224}Ra_{ex}/^{223}Ra$ = 4.1 ± 0.7; Moore, 2000). The $^{224}Ra_{ex}/^{223}Ra$ ratios decreased away from their benthic source due to decay ($^{224}Ra_{ex}/^{223}Ra$ = 0–0.5 in surface waters). Ratios close to the seafloor were relatively constant along the transect at bottom depths <600 m, whereas dFe, dCo and dMn concentrations varied largely in the bottom samples. This suggests that factors, which are not influencing the Ra distribution, impacted the distributions of dFe,

dCo and dMn, with a likely influence of enhanced O_2 concentrations reducing sediment release or increasing removal rates of these metals at water depths between 200 and 400 m. At around 800 m bottom depth, $^{224}Ra_{ex}/^{223}Ra$ ratios were slightly elevated and coincided with increased dCo, dFe, dMn and dCu concentrations despite O_2 concentrations >70 µmol kg^{-1}. This suggests that the enhanced TM concentrations at this location were influenced by a strong sediment source which may be related to the presence of a benthic nepheloid layer as indicated by an increase in turbidity in proximity to the seafloor. An elevated $^{224}Ra_{ex}/^{223}Ra$ ratio of 3.5 ± 0.6 was observed at about 16.65°N and 80 m water depth (bottom depth 782 m) and coincided with a local maximum of dFe, dMn and dCo and reduced O_2 concentrations. These observations indicate that the waters with the local maximum of dFe, dMn and dCo have been in relatively recent contact (12–20 days assuming initial pore water $^{224}Ra_{ex}/^{223}Ra$ ratios between 18–38; Moore, 2007) with sediments, likely originated from south of our transect as a result of a strong poleward flow (Klenz et al., 2018), and that the dynamic current system in this region can cause local and short-term variability in the transport of sediment derived TMs.

3.3 Classification of different groups of trace metals based on principal component analysis

Principal Component Analysis (PCA) was performed (using the RDA function within the vegan package in R; Oksanen et al., 2017) to investigate different groups and correlations in the data set. Dissolved TMs (Fe, Mn, Co, Ni, Pb, Cu and Cd), nutrients (silicic acid, nitrate and phosphate), dissolved O_2, Apparent Oxygen Utilization (AOU), depth and iodide concentrations were utilized in the PCA. Radium data were not included in the PCA, as the number of available data points for $^{224}Ra_{ex}/^{223}Ra$ was much lower than for the other parameters. Surface waters shallower than 50 m were excluded from the PCA to remove the influence of local processes in surface waters, such as localized atmospheric deposition and photochemical processes, which in particular influence Mn and iodide distributions. The PCA generated three principal components (PC) with eigenvalues larger than 1, with PC1 explaining 53.6% and PC2 25.5% of the total variance in the dataset (together 79.1%). Inclusion of PC3 in the analysis explained only 6.8% more of the variance.

The first PC group is formed by dCd, dCu, dNi and dPb (Fig. 4), which are associated with depth, AOU, nitrate and phosphate. This indicates that the distribution of Cd, Cu, Ni, and potentially Pb, are controlled by organic matter remineralization processes. This is in agreement with strong Pearson correlations R >0.9 for the relationships of dCd and dNi with depth, nitrate and silicic acid (Supplementary Material, Table S1). Weaker correlations with major nutrients were observed for dPb (R >0.6) and dCu (R >0.4), potentially due to additional remineralization or removal mechanisms for these elements (e.g. prior atmospheric inputs and water mass transport, Pb; sediments, Cu and Pb, and scavenging). The second group of TMs is composed of dFe, dCo and dMn that are associated with elevated iodide and turbidity, and low dissolved O_2 (Fig. 4). Iodide (I^-) is the reduced form of iodine (I), which is typically present as iodate (IO_3^-) in oxygenated subsurface water. Both I forms are present as soluble anions in seawater. Due to a relatively high redox potential (pE ~10), iodine is one of the

first redox-sensitive elements to undergo reduction under suboxic conditions and is therefore a useful indicator for active reductive processes (Rue et al., 1997). Despite their role as micronutrients, Fe, Mn and Co do not correlate with nutrients indicating that processes other than remineralization controlled their distributions.

The anti-correlation with O_2 (also shown in Fig. S1) and correlation with iodide support the notion that Fe, Co and Mn distributions were strongly influenced by water column O_2 concentrations, presumably through: (i) enhanced benthic metal fluxes from anoxic sediments, and (ii) decreased oxidation rates in the overlying water column under O_2-depleted conditions. This is also supported by elevated benthic Fe(II) fluxes observed at the seafloor within the shallow OMZ, with benthic fluxes of 15–27 µmol m^{-2} d^{-1} (Schroller-Lomnitz et al., 2018).

Variability in the redox-sensitive metals, Fe, Mn and Co, were not fully explained by either O_2 or iodide concentrations; Pearson correlations with O_2 were -0.55, -0.61 and -0.58, respectively (Supplementary Material, Table S1). As shown before, other factors such as, for example, water mass mixing and age, the amount and type of particles present, and remineralization all likely impact their dissolved concentrations. Consequently, such a complex chain of factors and processes means that one variable alone is unlikely to explain the behaviour of Fe, Mn, and Co.

3.4 Influence of the different sources of Fe, Mn and Co

The main sources of TMs in our study region are sedimentary release and atmospheric dust deposition (e.g. Rijkenberg et al., 2012). Also release of TMs via organic matter remineralization may have an important influence on the distribution of TMs. In the following, we discuss the relative influence of remineralization, atmospheric dust deposition and sedimentary release on the supply of Fe, Co and Mn to surface waters.

3.4.1 Remineralization

To quantify the influence of remineralization for dFe, we employed dFe to carbon (dFe/C) ratios (carbon was calculated using AOU, with an AOU/carbon ratio of 1.6; Martin et al., 1989). Surface data, where O_2 was over-saturated (due to biological O_2 production), were excluded. Dissolved Fe/C ratios for the entire transect varied between 15 and 74 µmol mol^{-1}. These results agree with those for shelf-influenced waters with dFe/C ratios of 13.3–40.6 µmol mol^{-1} further south at 12°N (Milne et al., 2017). Reported ratios for the North Atlantic, further away from the shelf were lower and ranged between 4 and 12.4 µmol mol^{-1} (Fitzsimmons et al., 2013; Milne et al., 2017; Rijkenberg et al., 2014). To estimate the amount of dFe being derived by remineralization, we assume a dFe/C ratio of 4–12 µmol mol^{-1} from organic matter remineralization, similar to the observed dFe/C ratios in the open ocean close to our study area without a strong shelf influence. These offshore ratios may still be influenced by an atmospheric source of dFe, which would result in an overestimation of dFe/C ratios

from remineralization and thereby an overestimation of the fraction of remineralized dFe. Apart from additional inputs, the dFe/C ratios are influenced by the respective Fe/C stoichiometry in the sinking organic matter and removal of dFe by scavenging. Furthermore, it is not clear if the offshore ratios can be transferred to a location close to the coast, as the balance between remineralization and scavenging processes might be different due to differences in phytoplankton productivity and particle load. Hence, this approach only provides a broad estimate of the relative influence of remineralization on the distribution of dFe in the study area.

We obtain a range between 5 ± 3% and 54 ± 27% for dFe being derived from remineralization processes with lowest values observed on the shelf at 34 m depth at station 4 (5 ± 3%) and highest values estimated beyond the shelf break at Stn 9 at 213 m depth (54 ± 27%) and Stn 2 at 450 m depth (52 ± 26%). However, no clear increase in the contribution of remineralized dFe to total dFe with depth or distance to the coast was observed. For example at depths between 35 and 200 m, our estimates of dFe from remineralization ranged between 10 ± 5% and 51 ± 25% with high values of up to 41 ± 20% at 50 m depth at station 7 close to the coast, whereas relatively low values of 19 ± 9% were observed at 89 m at station 2. These results indicate that, locally, remineralization can be an important control on dFe concentrations, but that the contribution varies largely with additional important controls, often dominating over remineralization.

Similar analysis for dCo/C ratios revealed an increased importance of an additional source close to the shelf. Observed dCo/C ratios ranged between 0.81 and 2.2 µmol mol^{-1}. The larger ratios were observed close to the coast and decreased further offshore. Overall, the observed ratios were somewhat higher than reported cellular ratios of phytoplankton in the North Atlantic of 0.5–1.4 µmol mol^{-1} (Twining et al., 2015). However, relatively constant dCo/C ratios beyond the shelf break (dCo/C: 0.82–1.09 µmol mol^{-1}, stations 2, 5 and 9) that are similar to cellular ratios of phytoplankton suggest a large influence of remineralization on dCo beyond the shelf break, whereas enhanced ratios close to the coast suggest an additional benthic source. Due to the lack of comparable data of offshore dCo/C ratios and the multiple processes influencing this ratio (varying phytoplankton nutrient stoichiometry and scavenging), we did not use these values to estimate the remineralized dCo fraction.

The distribution of Mn was not predominantly determined by biological uptake and remineralization processes in our study region. In contrast, dMn/C ratios were largely influenced by photoreduction in the surface (Sunda and Huntsman, 1994), removal via biotic oxidation and formation of Mn oxides at depth (Tebo et al., 2004). Therefore, we did not assess remineralization processes for Mn using dMn/C ratios.

3.4.2 Atmospheric deposition

Aluminum is present as a relatively constant fraction of ~8.15 wt% in the continental crust (Rudnick and Gao, 2006), is supplied to open ocean surface waters mainly by atmospheric deposition (Orians

and Bruland, 1986) and is considered not to be taken up by phytoplankton (apart from a small amount being incorporated into siliceous diatom frustules; Gehlen et al., 2002). Therefore, dAl in the surface mixed layer is used as a tracer for atmospheric deposition to the surface ocean (Measures and Brown, 1996; Measures and Vink, 2000). The atmospheric input in the study region is mainly influenced by North African/Saharan mineral dust with only a small contribution of anthropogenic sources which differ greatly in TM composition and solubilities from mineral dust (Baker et al., 2013; Patey et al., 2015; Shelley et al., 2015). Close to continental shelves, in addition to atmospheric input, Al can also be supplied by sediment resuspension (Menzel Barraqueta et al., 2018; Middag et al., 2012; Moran and Moore, 1991).

Our dAl concentrations in surface water ranged between 30 and 49 nmol L^{-1} and LpAl between 3.4 and 18.2 nmol L^{-1}. Dissolved Al concentrations decreased with depth (Fig. 8), indicating that Al was released by aeolian dust deposition to surface waters and removed through scavenging at depth (Orians and Bruland, 1985). Trace metal (Fe, Co, and Mn) to Al ratios were utilized to investigate the influence of atmospheric dust deposition. We present molar ratios for dissolved (dTM/dAl), total dissolvable (TDTM/TDAl) and leachable particulate (LpTM/LpAl) concentrations. In the surface mixed layer, dFe/dAl molar ratios ranged between 0.019 and 0.114, TDFe/TDAl between 0.236 and 0.826 and LpFe/LpAl between 1.04 and 9.50.

Literature particulate Fe/Al ratios from aerosol samples collected in the remote North Atlantic between 8.7°N and 23°N were in the range of 0.31 ± 0.06 (Buck et al., 2010; Patey et al., 2015) and 0.37 ± 0.02 in the North East Atlantic ~18°N under the Saharan dust plume (Shelley et al., 2015). In contrast, upper crustal material ratios are lower ranging from 0.19 to 0.23 suggesting a slight Fe enrichment of aeolian mineral dust particles (McLennan, 2001; Rudnick and Gao, 2006; Wedepohl, 1995). Lower Fe than Al solubilities from aerosol leach experiments in ultra-high purity water (UHP) and 25% acetic acid (HAc) and seawater have been reported (Baker et al., 2006; Buck et al., 2010; Shelley et al., 2018), but soluble Fe/Al ratios from these experiments varied dependant on the leach medium (UHP: 0.21 ± 0.04, 25% HAc: 0.25 ± 0.04, seawater: 0.051 ± 0.009; Shelley et al., 2018). This indicates that dFe/dAl and LpFe/LpAl ratios in seawater from atmospheric deposition are likely to be lower than particulate ratios of digested aerosol samples in the study region.

Our dFe/dAl ratios at the upper end (dFe/dAl: 0.114) are larger than aerosol leaches in seawater indicating a potential additional input of dFe, whereas our lower dFe/dAl ratios than reported ratios in aerosol leaches suggest removal of dFe by biological uptake or scavenging. Our LpFe/LpAl are all larger than reported ratios in aerosol leaches and total aerosol ratios, which shows that there is an additional source of LpFe or transfer of sediment-derived dFe onto the particulate phase by biological uptake or sorption to particles. Total dissolvable ratios comprise both dissolved and leachable particulate phases, thereby being independent of the phase transfer from dissolved to particulate phase (via biological uptake or sorption). The lower end of total dissolvable ratios (TDFe/TDAl: 0.236) were

close to the total ratios in aerosol samples, suggesting that atmospheric deposition represented an important source of Fe and Al to the surface ocean. At the upper end, ratios were much larger (TDFe/TDAl: 0.826) than aerosol ratios and indicate an additional benthic source of Fe.

These interpretations only apply, however, if residence times of dissolved and particulate Fe and Al phases supplied via atmospheric deposition are similar. This is difficult to assess, as estimated residence times for both elements are dependent on input and removal rates and vary largely between locations. Overall, our Fe/Al ratios suggest that atmospheric deposition is an important source of Fe to surface waters with an additional contribution of benthic inputs. However, uncertainties in solubilities and residence times cause a high uncertainty in the interpretation of the role of atmospheric deposition.

Observed dCo/dAl ratios in the upper 50 m were 0.001–0.004, TDCo/TDAl ratios were slightly higher at 0.003–0.005 and LpCo/LpAl ratios were 0.006–0.020. Cobalt is present in the upper continental crust in a much smaller molar fraction than Fe (Co/Al: 0.000071–0.000097; McLennan, 2001; Rudnick and Gao, 2006; Wedepohl, 1995). However, ratios in aerosol samples under the North African dust plume were slightly higher (Co/Al: 0.00016 ± 0.00002; Shelley et al., 2015) than crustal ratios and solubility of Co from these aerosol samples was much higher than Al solubility resulting in soluble Co/Al of 0.0021 ± 0.0009 in UHP (Shelley et al., 2018). The soluble ratios also varied largely depending on the leach medium and might therefore also vary from the actual aerosol solubility in seawater at our study site. Our ratios of all fractions were larger than total aerosol ratios and mostly higher than soluble ratios from aerosol leaches. This indicates that an additional benthic source of Co likely contributed to the Co present in surface waters.

Dissolved Mn/dAl ratios in the upper 50 m ranged between 0.082 and 0.347, and TDMn/TDAl between 0.083 and 0.256. The ratios are much larger than upper crustal ratios (Mn/Al: 0.0032–0.0037; McLennan, 2001; Rudnick and Gao, 2006; Wedepohl, 1995) but similar to the soluble ratios of Mn/Al from aerosols in UHP (Mn/Al: 0.24 ± 0.09; Shelley et al., 2018) indicating that a large amount of Mn may be derived from atmospheric deposition. However, these ratios are heavily overprinted by the long residence time of Mn in surface waters due to photoreduction. Therefore, it is not possible to reliably estimate the contribution of atmospheric Mn deposition based on the Al data.

Atmospheric dFe fluxes were calculated using the dAl inventory in the surface mixed layer, a residence time of dAl of 0.65 ± 0.45 years as reported for the Canary Current System (Dammshäuser et al., 2011), and a ratio of 0.31 for dust derived dissolved Fe/Al (Buck et al., 2010). This approach assumes that dAl is only supplied to the surface ocean via atmospheric deposition. Vertical fluxes of Al from sediment resuspension are unlikely to largely contribute to concentrations of dAl in surface waters here as dAl concentrations were decreasing with depth. Mean atmospheric dFe fluxes of the individual stations were 0.63–1.43 µmol m^{-2} d^{-1} (Fig. 5, Supplementary Table S2), values similar to reported fluxes close to our study region of 2.12 µmol m^{-2} d^{-1} further north between 22.5–25°N and

26.5–27.5°W (Rijkenberg et al., 2012) and 0.120 nmol m^{-2} d^{-1} around 20°N close to the African coast (Ussher et al., 2013). The uncertainty in the residence time of dAl, however, creates a large uncertainty in calculated fluxes resulting in a lowest flux of 0.37 µmol m^{-2} d^{-1} when using the largest estimated residence time of 1.1 years and a highest flux of 4.65 µmol m^{-2} d^{-1} when using the shortest estimated residence time of 0.2 years.

3.4.3 Vertical trace element fluxes to surface waters

The vertical fluxes (diffusive and advective) of dFe from the top of the shallow O_2-depleted waters (between 23 and 89 m depending on station) into surface waters were determined to assess the potential Fe contribution to phytoplankton growth. A detailed summary of calculated fluxes, the contribution of diffusive and advective term and uncertainties for dFe for all stations is given in Supplementary Information Table S2. Closest to the shelf (bottom depth: 50 m) mean dFe fluxes were 13.5 µmol m^{-2} d^{-1}. Further offshore, vertical dFe fluxes decreased to 0.16 µmol m^{-2} d^{-1} (station 2, bottom depth: 1136 m, 77 km offshore) (Fig. 5). However at station 5 higher dFe fluxes were observed (dFe: 1.3 µmol m^{-2} d^{-1}) than at stations 9 (closer to the shelf) and 2 (further offshore). At station 5, eddy diffusivity was determined from only 5 microstructure profiles and was unusual high at this station. Therefore, the enhanced vertical fluxes are likely caused by a rare elevated mixing event and do not represent a long-term average. Between repeat stations 3A and 3B, mean fluxes decreased from 2.3 (Stn 3A) to 1.35 µmol m^{-2} d^{-1} (Stn 3B), which was partly caused by a difference in the vertical concentration gradient of dFe and partly by a change in diffusivity.

On the shelf (station 4, bottom depth: 45 m), dFe fluxes were dominated by vertical advective rather than diffusive fluxes due to the strong upwelling velocity on the shelf (Table S2). At the continental slope stations (stations 3, 7 and 8, bottom depth: 90–400 m), fluxes were dominated by high diffusive fluxes, which were around 3 times larger than the advective flux term. Further offshore (stations 2 and 9, bottom depth: >400 m) the contribution of advective and diffusive fluxes were similarly low except for station 5 with particularly strong vertical mixing. Similar vertical dFe (16 µmol m^{-2} d^{-1}) to the upper water column were reported on the shelf at 12°N (Milne et al., 2017). Although, in the study region atmospheric fluxes of dFe were enhanced relative to global averages (Mahowald et al., 2009) with mean fluxes of 0.63–1.43 µmol m^{-2} d^{-1}, our vertical dissolved Fe fluxes from the shallow O_2 depleted waters of 0.95–13.5 µmol m^{-2} d^{-1} exceeded atmospheric fluxes at all stations apart from station 2 (0.16 µmol m^{-2} d^{-1}) furthest offshore and potentially station 9 (0.08 µmol m^{-2} d^{-1}), where no atmospheric fluxes were determined. The weaker influence of atmospheric deposition in this region close to the coast is in accordance with previous studies that demonstrated sediments to be the major contributor to the Fe inventory in the coastal region of the eastern tropical Atlantic, whereas the importance of atmospheric inputs increases further offshore (Milne et al., 2017). Our vertical advective fluxes are likely lower than the annual average and also lower than usually during the relaxation period as upwelling favourable winds were particularly low in June 2014.

Dissolved Co fluxes ranged between 2 and 113 nmol m^{-2} d^{-1}. These values are lower than reported upwelling fluxes of dCo of 250 nmol m^{-2} d^{-1} for this region (Noble et al., 2017), but are larger than atmospheric deposition fluxes of 1.7 nmol m^{-2} d^{-1} (Shelley et al., 2015). Fluxes of dMn are downwards from surface waters to O_2 depleted waters due to higher concentrations in surface waters.

3.5 Removal mechanisms and particle interactions

In the top 50 m of the water column a large part of the LpTMs may be part of living biological cells (e.g. phytoplankton) or organic detritus. Additionally, LpTMs may be part of lithogenic phases from Saharan dust and sediment particles, or authigenic phases. Authigenic phases are formed in-situ by TM adsorption onto particle surfaces or by the formation of amorphous TM oxides and hydroxides (e.g. FeO(OH) in the mineral structure of goethite) (Sherrell and Boyle, 1992).

Iron was mainly present in the size fraction >0.2 µm with TDFe concentrations being 0.44–44.5 times higher than dFe (<0.2 µm) (Fig. 6a). To investigate the influence of particle load on the distribution between dissolved and particulate phases, the fraction of Lp (Lp/TD) TMs and Lp concentrations are plotted against turbidity for Fe, Co and Mn (Fig. 6b, c). A low fraction of LpFe of around 60% was observed at lowest turbidity. As turbidity increases from 0.1 to 0.2 NTU, the LpFe fraction increased to >90%. This suggests that the fraction of LpFe is tightly coupled to the particle load. Iron adsorption onto particles has been demonstrated to be reversible with a constant exchange between dissolved and particulate fractions (Abadie et al., 2017; Fitzsimmons et al., 2017; John and Adkins, 2012; Labatut et al., 2014). Furthermore, offshore transport of acid-labile Fe particles originating from reductive dissolution processes from continental shelf sources was observed in the North Pacific (Lam and Bishop, 2008) and may contribute to the bioavailable Fe pool. Therefore an important fraction of Fe may be transported offshore adsorbed to particles and can enter the dissolved pool by cycling between dissolved and particulate phases.

Manganese and Co mainly occurred in the dissolved form. The LpCo fraction ranged between 0 and 75%, and the fraction and concentration of LpCo, showed linear increases with turbidity, indicating an influence of particle load on Co size fractionation, similar to Fe. In contrast to Fe and Co, the fraction of LpMn varied between 3 and 40%, and did not show a correlation with turbidity, whereas LpMn concentrations showed an increase with turbidity. This indicates that an increased presence of particles coincided with enhanced LpMn levels, but that the particle load did not substantially influence the distribution between dMn and LpMn phases and that particles did not contribute to the dMn fraction. This suggests that particles did not play a major role in transport of dMn, which agrees with a study on hydrothermal vent plumes, where the distribution of the dMn plume was decoupled from the distribution of the particulate Mn plume (Fitzsimmons et al., 2017).

The increase in LpFe concentrations with increasing turbidity was weaker in the surface waters compared to water depths below 50 m (Fig. 6c). This suggests a large additional LpFe source at depth

with either a higher Fe content of particles or the presence of different sizes of particles causing different responses in turbidity measurements. The large additional LpFe source at depths is likely associated with benthic dFe inputs, with a subsequent transfer to the particulate phase by adsorption or oxidation with subsequent formation of Fe(oxihydr)oxides. Enhanced turbidity at depth may also indicate sediment resuspension, which would result in the release of TM-containing particles from sediments and enhanced release of dTMs from sediment pore water. The effect of sediment resuspension is discussed in more detail below (section 3.6.2).

In contrast to Fe, the increase in LpCo and LpMn concentrations with turbidity was similar in surface waters and below and suggests less variability in the composition of the particulate Co and Mn phase throughout the water column with a potentially weaker influence of sediment release on the distribution of particulate Mn and Co. A weaker influence of sediment release might be influenced by a weaker release of Co and Mn from sediments in the dissolved form and slower oxidation rates compared to Fe, in particular for Co (Noble et al., 2012), resulting in a slower conversion into the particulate phase. Such an interpretation based on turbidity data alone, however, is very hypothetical and would require further investigation of particulate TM species composition in this area.

3.6 Temporal variability in redox-sensitive trace metals

Large temporal changes in O_2, turbidity and redox-sensitive TMs were observed within a short time scale of a few days at two repeat stations, station 3A/3B and station 8A/8B (Fig. 7).

Station 3 and 8 were sampled twice with a period of nine days between both deployments for station 3 (Fig. 7a) and two days for station 8 (Fig. 7b). At station 3, O_2 concentrations in the upper 50 m were very similar between both deployments, whereas below 50 m O_2 increased from 30 µmol kg^{-1} during the first deployment to 50 µmol kg^{-1} nine days later. At the same time, turbidity below 50 m had decreased from 0.35 to below 0.2, and dFe concentrations from a maximum of 10 nmol L^{-1} to 5 nmol L^{-1} nine days later. In addition, dMn and dCo concentrations decreased from 5 to 3 nmol L^{-1} and 0.14 to 0.12 nmol L^{-1}, respectively. Particularly large changes were also observed for LpTM concentrations with a decrease from 147–322 nmol L^{-1} to 31–51 nmol L^{-1} for LpFe, from 0.066–0.114 nmol L^{-1} to 0.015–0.031 nmol L^{-1} for LpCo and from 1.24–2.64 to 0.16–0.54 for LpMn. In contrast, no changes in water mass properties (T/S) occurred below 50 m (Fig. 7a).

Similar changes in O_2 and turbidity were observed at station 8. During the first deployment a local minimum in O_2 below 30 µmol kg^{-1} was present between 105 m and 120 m water depths which coincided with a maximum in turbidity of 0.4 (Fig. 7b). In contrast O_2 concentrations and turbidity during the second deployment were relatively constant (50–60 µmol kg^{-1} O_2 and turbidity 0.2) below 50 m. At depth of the local O_2 minimum and turbidity maximum, concentrations of dFe, dMn and dCo were elevated during the first deployment with concentrations of 9.4 ± 2.1 nmol dFe L^{-1}, 3.7 ± 0.6

nmol dMn L^{-1} and 0.145 ± 0.033 nmol dCo L^{-1} in comparison to 4.6 ± 1.0 nmol dFe L^{-1}, 2.6 ± 0.5 nmol dMn L^{-1}, and 0.122 ± 0.028 nmol dCo L^{-1} at similar depth during the second deployment.

3.6.1 Remineralization

We compared the results of the redox-sensitive TMs to other nutrient-like TMs and PO_4. For both repeat stations only small changes in dCd (Stn 3A: 0.107–0.231 nmol L^{-1}; Stn 3B: 0.135–0.150 nmol L^{-1}) and PO_4 (Stn 3A: 1.59–1.85 µmol L^{-1}; Stn 3B: 1.55–1.71 µmol L^{-1}) concentrations were observed below 50 m (Fig. 8), suggesting that only a small fraction of dFe under lower O_2 conditions was supplied by more intense remineralization of biogenic particles in the water column.

A weak influence of remineralization processes on the variability in dFe concentrations was confirmed by substantially higher dFe/C ratios at lower O_2 concentrations (40–72 µmol mol^{-1} at Stn 3A compared to 33–41 µmol mol^{-1} at Stn 3B, both below 50 m water depth). Assuming a dFe/C ratio of around 12 (see section 3.4.1) from remineralization, only about 0.25 nmol L^{-1} of the difference in dFe concentrations between repeated deployments can be explained by the difference in remineralization, suggesting that most of the difference in dFe between deployments was caused by changes in source inputs, such as enhanced sediment release during lower bottom water O_2 concentrations, or slower removal by oxidation under lower O_2 conditions.

In contrast, dCo/C ratios were similar between repeat deployments within the OMZ (0.90–1.04 at Stn 3A and 0.92–1.06 µmol mol^{-1} at Stn 3B). Thus, changes in remineralization could be a reason for the changes in observed dCo concentrations during repeated deployments, indicating that the sensitivity of dCo sediment input or change in oxidation rates is low at an O_2 shift from 30 to 50 µmol kg^{-1}.

Similar to Fe, higher dMn/C ratios were observed at lower O_2 concentrations (3.4–5.5 µmol mol^{-1} at Stn 3A compared to 2.1–2.9 µmol mol^{-1} at Stn 3B). These results indicate that other processes than remineralization are also important for the change in dMn concentrations. An additional factor compared to Fe, might involve changes in intensity of photoreduction which may be influenced by differences in surface turbidity observed at station 3 (lower dMn/C and higher surface turbidity during second deployment). This, however, cannot explain the changes in dMn/C at station 8, where a higher surface turbidity coincided with a higher dMn/C ratio at the local minimum in O_2.

3.6.2 Atmospheric dust deposition and sediment resuspension

Within the OMZ at station 3 and 8, dAl concentrations ranged between 10 and 15 nmol L^{-1}, and LpAl concentration between 1.2 and 11.1 nmol L^{-1} and no substantial changes were observed between deployments (Fig. 8). As lithogenic material has a high Al content, no substantial changes in Al concentrations signify that lithogenic inputs did not differ much between the deployments. Consequently neither increased atmospheric input, nor sediment resuspension are likely to explain the differences in turbidity and redox-sensitive TM concentrations. Hence, changes in turbidity may

mainly have been caused by biogenic particles, such as resuspended organic matter (Thomsen et al., 2018). This finding can be confirmed by substantial changes in TM/Al ratios observed during the deployments (Table 2 and Fig. S2). The Fe/Al ratios in the solid phase of underlying sediments during the cruise were 0.23–0.30 (Schroller-Lomnitz et al., 2018) with Mn/Al ratios of 0.0015–0.0020 (Schroller-Lomnitz, pers. com.). Slight increases in LpAl towards the sediment indicate some influence of sediment resuspension on the TM distribution. Overall much higher TM/Al ratios compared to ratios in the sediments and aerosol samples from this region (Fe/Al: 0.37 ± 0.02, Co/Al: 0.00016 ± 0.00002, Mn/Al: 0.0061 ± 0.0002; Shelley et al., 2015), suggest a large additional source of Fe, Co and Mn in the OMZ close to the shelf. This again points towards a large influence of benthic release of Fe, Co and Mn from sediment pore waters and subsequent partial adsorption to particle surfaces.

3.6.3 Other possible causes for TM variability

From the comparison above, we can conclude that the variations in Fe concentrations during repeated deployments were not caused by increased remineralization or changes in lithogenic inputs from atmospheric deposition or sediment resuspension. The large changes in the Lp fractions must therefore be of biogenic or authigenic origin. If all LpCo would be present in biogenic particles of suspended phytoplankton cells, at our observed maximum of 0.114 nmol L^{-1} LpCo at station 3A we would expect around 4.6 nmol L^{-1} LpFe in sinking phytoplankton, using an average Fe/Co ratio in phytoplankton of 40 (Moore et al., 2013) (observed ratios close to our study area were 20–40; Twining et al., 2015). However, LpFe concentrations were 322 nmol L^{-1} and thereby 70 times larger than our estimate in biogenic particles (4.6 nmol L^{-1}), revealing that the majority of LpFe must be authigenically formed. Altogether our results suggest that changes in particle load as indicated by changes in turbidity do not comprise a major source of dFe, moreover a sink of previously dissolved Fe. Therefore, higher dissolved and Lp concentrations during the first deployment with lower O_2 concentrations must be caused by a stronger benthic source of dissolved Fe.

It is not possible to extract from our data whether the stronger benthic source under low O_2 conditions is directly driven by lower O_2 concentrations in surface sediments and in the water column resulting in higher benthic Fe fluxes and slower oxidation rates in the water column, or by a longer residence time of the water mass on the shelf. However, increased benthic fluxes are in accordance with previous findings from ex-situ sediment incubation experiments, where Fe fluxes increased with decreasing O_2 concentrations (Homoky et al., 2012). Therefore, we hypothesize that with a reduction of bottom water O_2 concentrations from 50 to 30 µmol kg^{-1}, drastically more Fe is effectively released from the sediments by diminished oxidation rates at the sediment-water interface, and that a large fraction gets directly adsorbed onto particles. Therefore, particles do not compose a major source of Fe here, but may play an important role in Fe offshore transport.

Due to much lower changes in concentrations of dissolved and LpCo, and the additional effect of photoreduction and strong scavenging for Mn, we were unable to resolve the main mechanisms for changes in Co and Mn concentrations with changes in O_2 and turbidity. Nevertheless, due to their similar redox-sensitive behavior and distribution in OMZs, it is likely that they are also affected by reduced O_2 conditions. The magnitude of response however, is much lower.

4. CONCLUSION

Sediments are an important source of Fe, Co and Mn to OMZ waters in the Mauritanian shelf region. Remineralization and atmospheric deposition appear less important than benthic sources for dFe, with vertical fluxes exceeding atmospheric fluxes but gaining importance with distance from shelf. We showed that changes in O_2 concentrations from 30 to 50 µmol kg^{-1} had a substantial influence on dissolved and LpFe concentrations and to a lesser extent on Co and Mn concentrations by decreasing the sediment source strength. The presence of a large part of sediment-derived Fe in the leachable particulate phase highlights the importance of offshore particle transport on the Fe inventory, including the dissolved form by reversible scavenging. To our knowledge, this is the first field study that demonstrated strong short-term variability in redox-sensitive TMs over a few days to be directly linked to changes in O_2. These findings demonstrate that projected long-term changes in oceanic O_2 concentrations will impact biogeochemical cycles and have important implications for global TM distributions and their process parameterisations in biogeochemical models. Current models do not account for small changes in O_2 on TM distributions and benthic TM fluxes. Determining the processes involved and quantifying the effect of O_2 will be crucial for the implementation into current modeling approaches. Not all processes could be resolved in this study, including the influence of the residence time of the water masses on the shelf compared to the direct influence of O_2, and it is unclear whether the changes observed on a small scale are readily transferable to a global scale. Therefore, we suggest further investigations on short-term variability of O_2 and particle load in the Mauritanian and other dynamic OMZs including water column TM measurement in combination with benthic TM fluxes and more detailed analysis of amount and types/composition of present particles.

Data availability. The CTD sensor and nutrient bottle data are freely available at https://doi.pangaea.de/10.1594/PANGAEA.860480 and https://doi.pangaea.de/10.1594/PANGAEA.885109 respectively. According to the SFB754 data policy (https://www.sfb754.de/de/data, all remaining data (trace metal data set) associated with this manuscript will be published at PANGEA (www.pangea.de, search projects:sfb754) upon publication of this manuscript.

Author contributions. IR analyzed the trace metal concentrations and drafted the manuscript. EPA and MG designed the project and CS carried out the trace metal sampling at sea. J-LMB oversaw, and BW

carried out, the aluminium sample analysis. MD carried out the microstructure measurements at sea, oversaw the calculation of the vertical flux estimates and contributed to the writing of the manuscript. JL carried out the processing of microstructure data and calculation of the eddy diffusivity. JS, BG and PR carried out the radium isotope analysis and their interpretation. IR and MG oversaw, and FW carried out, the iodide analysis. All co-authors commented on the manuscript.

Competing interests. The authors declare that they have no conflict of interest.

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Table 1. Analyzed reference seawater, procedural blanks and detection limits (three times the standard deviation of the blank). Mean values and standard deviation for Cd, Pb, Fe, Ni, Cu, Mn and Co and available consensus values (± 1 standard deviation), n = number of measurements.

	SAFe S (nmol L^{-1}) n=11	SAFe S consensus value (nmol L^{-1})	SAFe D2 (nmol L^{-1}) n=7	SAFe D2 consensus value (nmol L^{-1})	Blank (pmol L^{-1})	Detection limit (pmol L^{-1})
Cd	0.003 ± 0.002	0.001	1.089 ± 0.043	1.011 ± 0.024	2.2 ± 0.3	0.8
Pb	0.050 ± 0.003	0.049 ± 0.002	0.028 ± 0.001	0.029 ± 0.002	0.4 ± 0.2	0.6
Fe	0.091 ± 0.009	0.095 ± 0.008	1.029 ± 0.038	0.956 ± 0.024	68 ± 10	29
Ni	2.415 ± 0.086	2.34 ± 0.09	9.625 ± 0.175	8.85 ± 0.26	112 ± 20	59
Cu	0.514 ± 0.037	0.53 ± 0.05	2.176 ± 0.152	2.34 ± 0.15	14 ± 3	9.3
Co	0.005 ± 0.001	0.005 ± 0.001	0.048 ± 0.003	0.047 ± 0.003	2.7 ± 0.8	2.5
Mn	0.814 ± 0.033	0.810 ± 0.062	0.437 ± 0.029	0.36 ± 0.05	14 ± 6	17

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Table 2. TM/Al ratios of different fractions for the repeated deployments at station 3 within the OMZ below 50 m water depth.

Parameter	Stn 3A	Stn 3B
dFe/dAl	0.38–0.79	0.35–0.37
TDFe/TDAl	4.00–13.42	1.83–2.81
LpFe/LpAl	10.00–29.50	3.64–8.59
dCo/dAl	0.009–0.011	0.009–0.011
TDCo/TDAl	0.009–0.010	0.006–0.008
LpCo/LpAl	0.007–0.011	0.001–0.005
dMn/dAl	0.26–0.45	0.19–0.21
TDMn/TDAl	0.26–0.32	0.12–0.17
LpMn/LpAl	0.14–0.28	0.02–0.09

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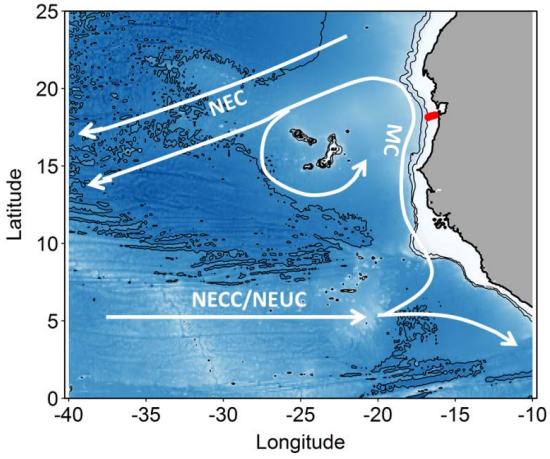

Figure 1. Map of the study area. Stations along the transect during M107 (June 2014) are displayed in red circles and major currents in white lines (adapted from Brandt et al. 2015). MC = Mauritania Current; NEC = North Equatorial Current; NECC = North Equatorial Countercurrent; NEUC = North Equatorial Undercurrent.

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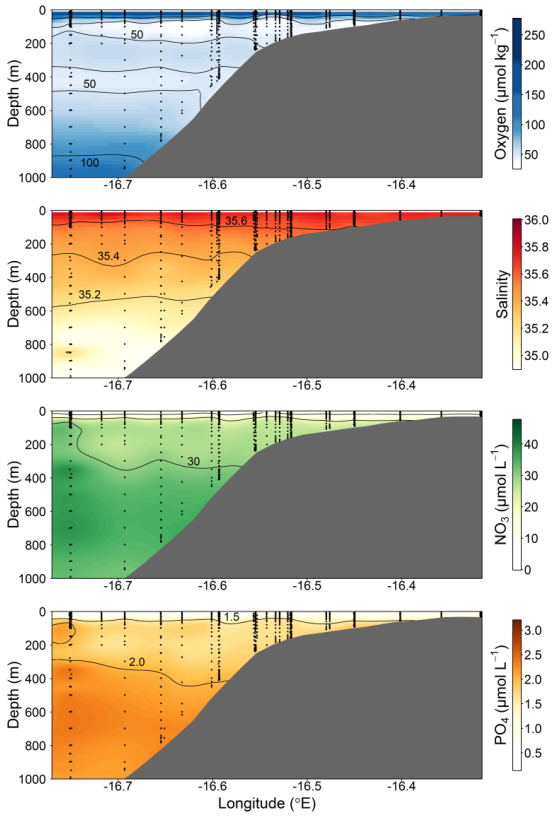

Figure 2. Section plots of oxygen (µmol kg^{-1}), salinity (PSU), NO_3 (µmol L^{-1}) and PO_4 (µmol L^{-1}) along the transect off the Mauritanian coast in June 2014.

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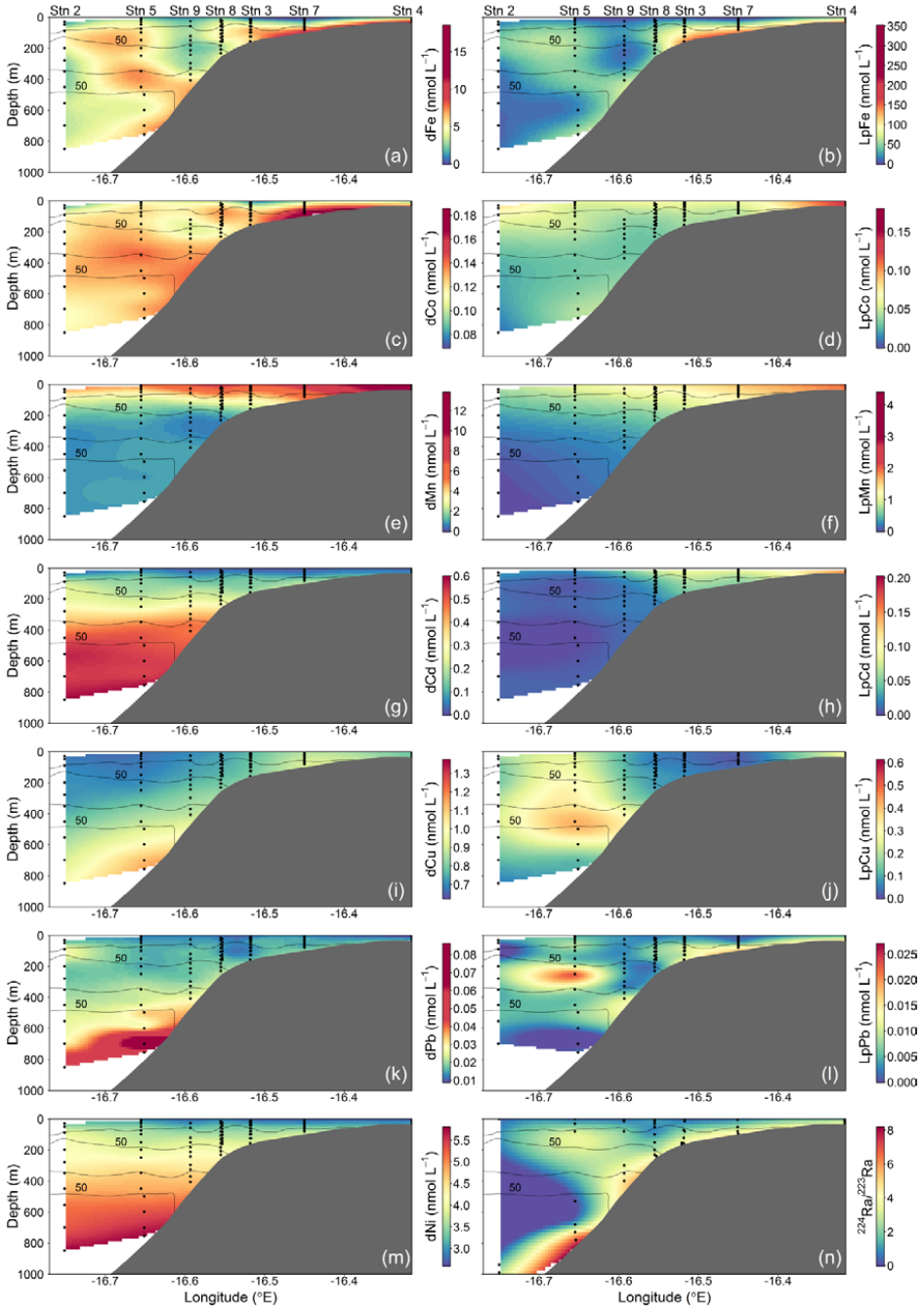

Figure 3. Spatial distributions of dissolved (d) and leachable particulate (Lp) trace metals and $^{224}Ra/^{223}Ra$ across the Mauritanian shelf at 18°20'N in June 2014. Each sample location is indicated as black dot and oxygen contours at 50 µmol kg^{-1} enclosing the upper and lower OMZ are displayed as black contour lines.

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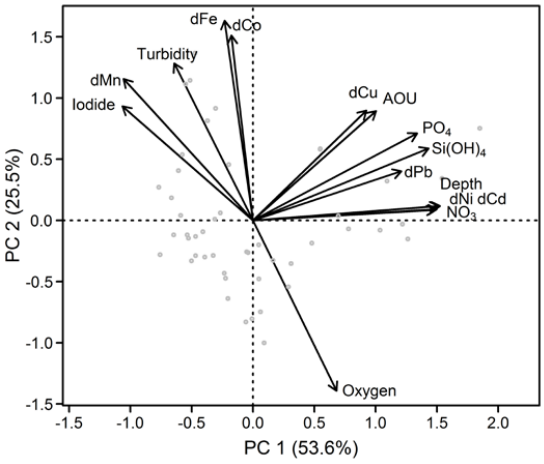

Figure 4. Principal component analysis of the Mauritanian shelf data set. Principal component loadings for each variable are indicated by black vectors. Component scores of each sample are indicated as grey circles. Loadings/scores have been scaled symmetrically by square root of the eigenvalue.

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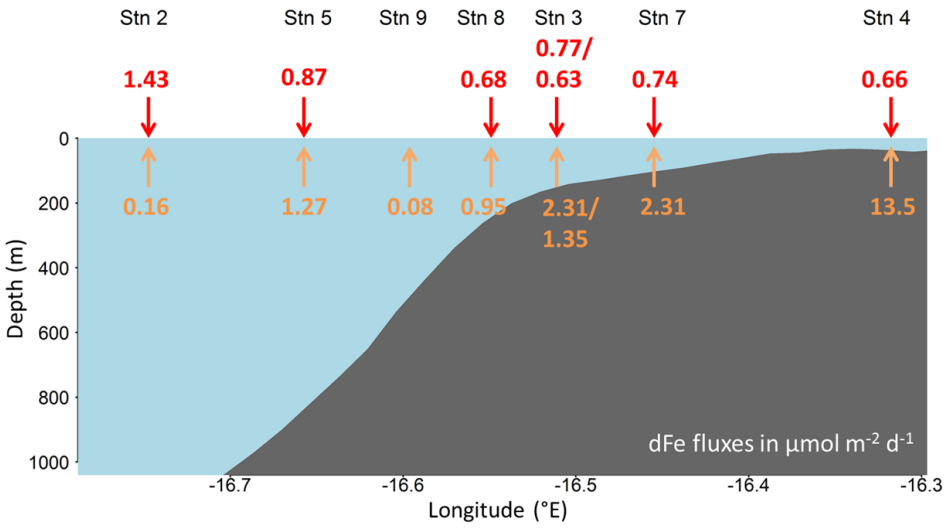

Figure 5. Atmospheric dFe fluxes (red) and vertical dFe fluxes (orange) in µmol m^{-2} d^{-1} along the transect at 18°20'N in June 2014.

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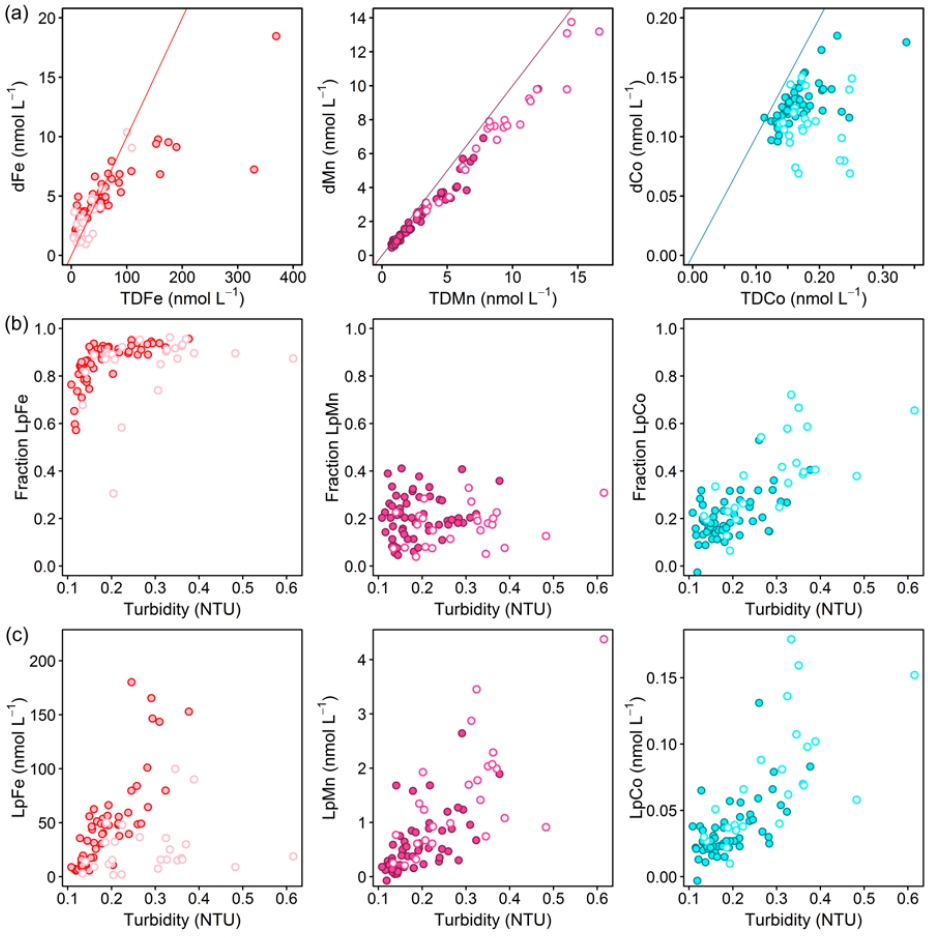

Figure 6. (a) Dissolved against total dissolvable trace metal concentrations for Fe (left; red line: TDFe = 10*dFe), Mn (middle; purple line: TDMn = dMn) and Co (right; turquoise line: TDCo = dCo). (b) Fraction of leachable particulate trace metals (Lp/TD) against turbidity and (c) Leachable particulate concentrations against turbidity for Fe (left), Mn (middle) and Co (right). Filled circles display all data points below 50 m depth, open circles at depths shallower than 50 m.

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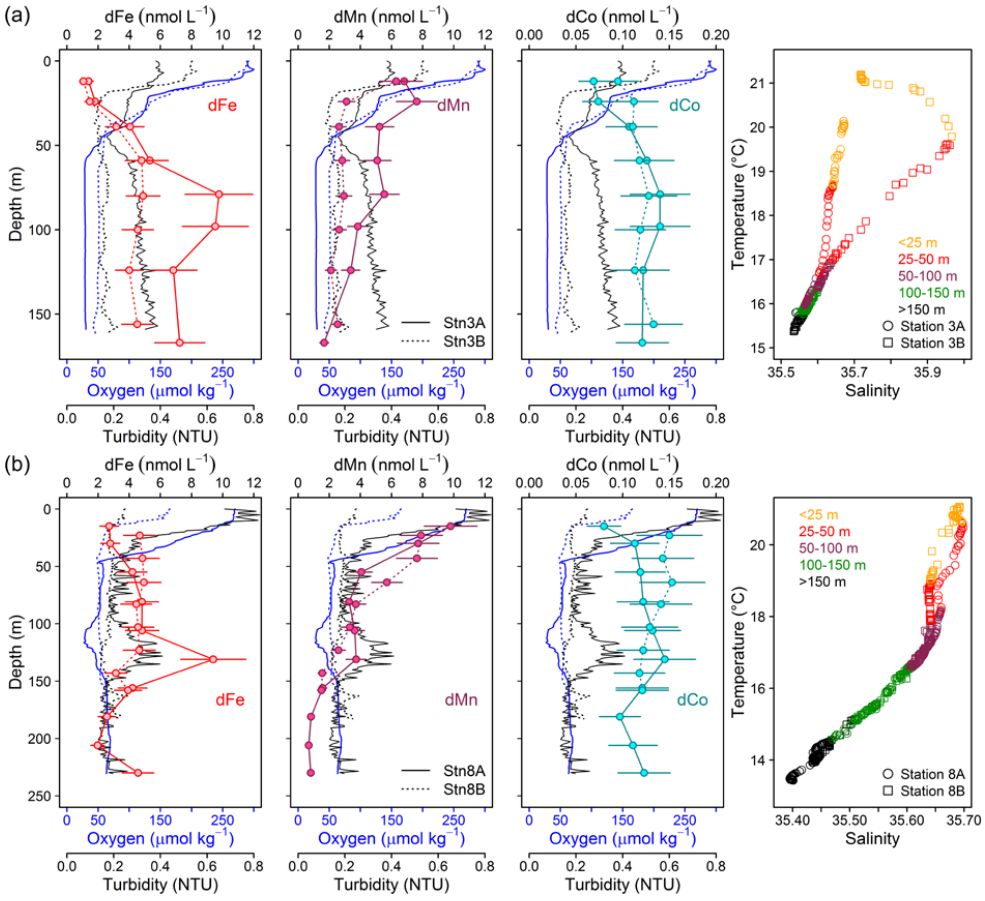

Figure 7. Repeat stations: oxygen concentration, turbidity and dissolved trace metals (Fe, Mn and Co) and temperature vs salinity plots. First deployment displayed as solid line and second deployment displayed as dashed line. (a) Station 3 (18.23°N, 16.52°W, 170 m water depth, 9 days between deployments). (b) Station 8 (18.22°N, 16.55°N, 189–238 m water depth, 2 days between deployments).

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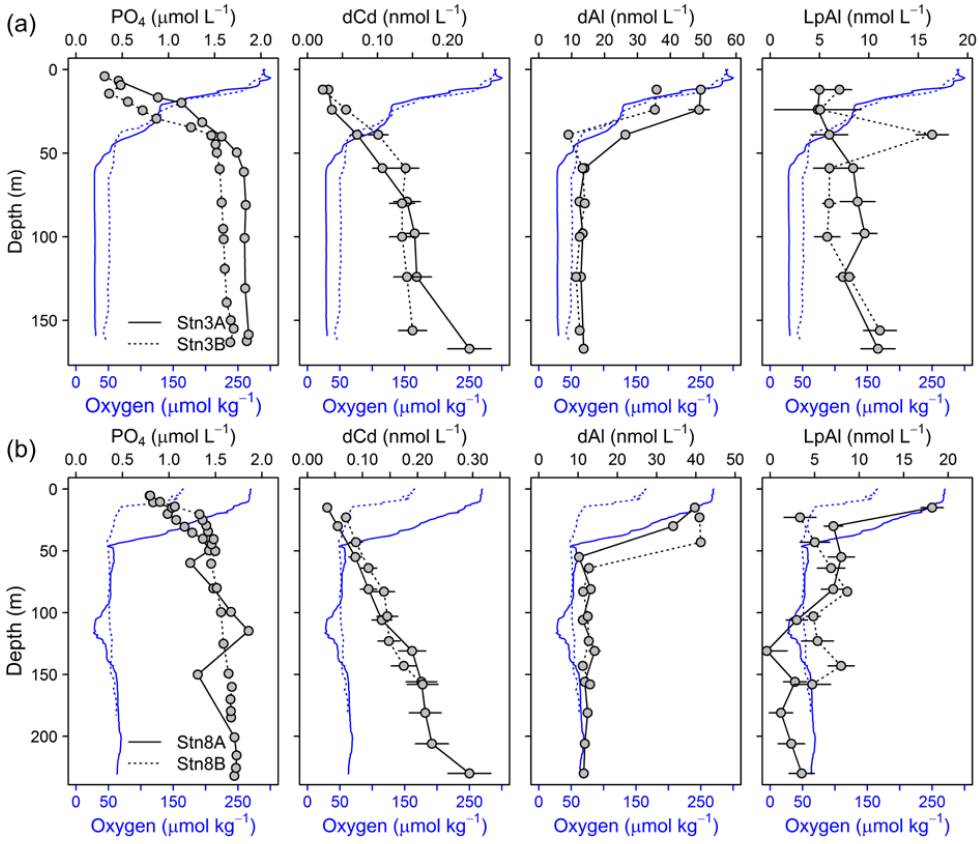

Figure 8. Depth profiles of dCd, PO_4, dAl and LpAl of repeat stations. First deployment displayed as solid black line and second deployment displayed as dashed black line. Oxygen concentrations are indicated as blue solid line for the first deployment and dashed blue line for the second deployment. (a) Station 3 (18.23°N, 16.52°W, 170 m water depth, 9 days between deployments and (b) Station 8 (18.22°N, 16.55°N, 189–238 m water depth, 2 days between deployments).