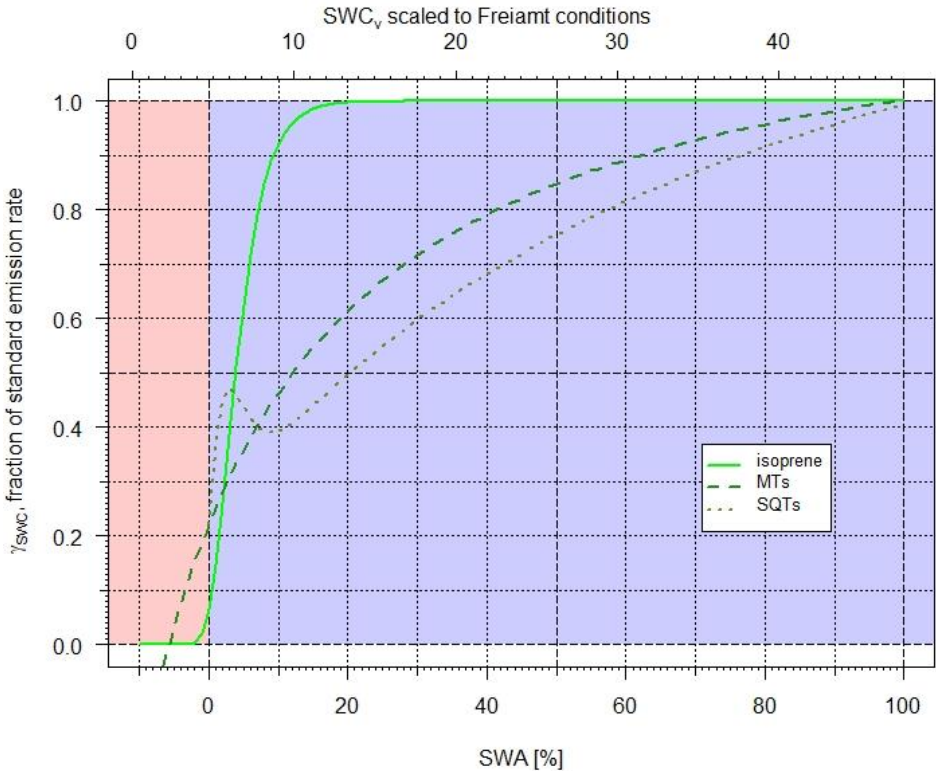




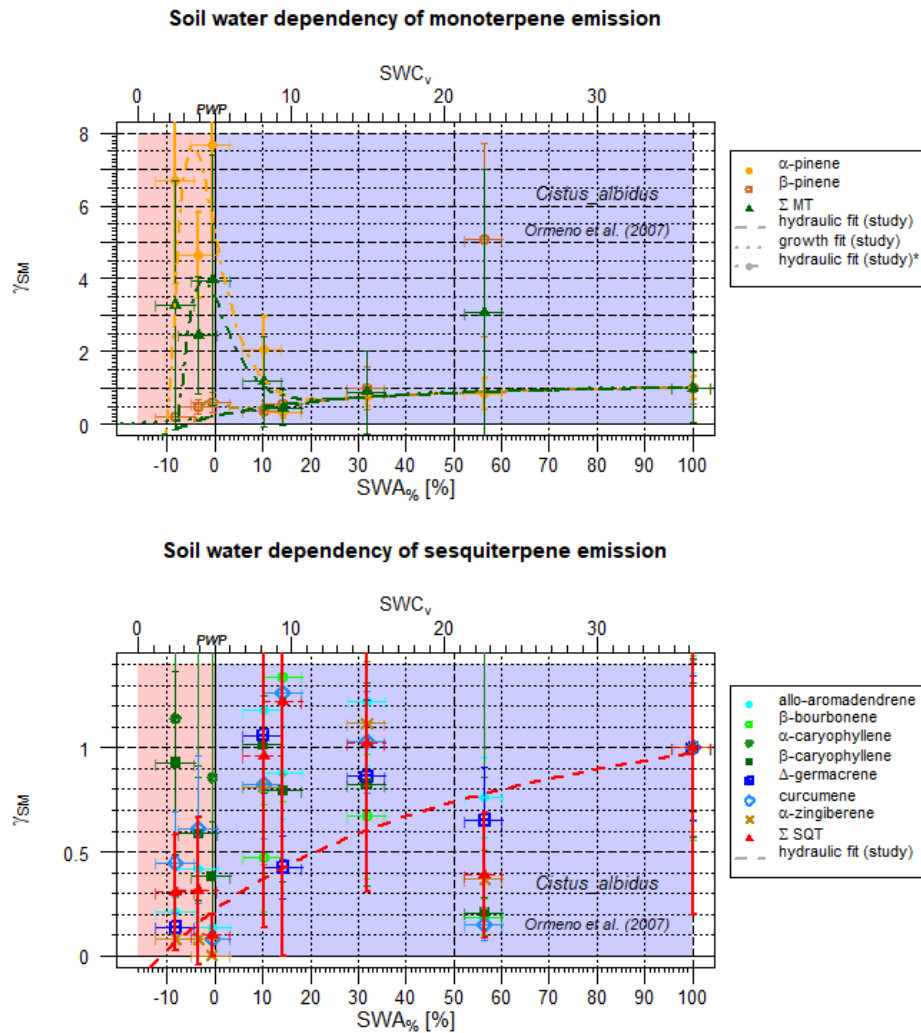
**Figure S2.**

Relative effect of available soil moisture on emission rates of isoprene, MT and SQT at standard conditions otherwise (T=30°C). The local peak of SQT emission rates is based on *Rosmarinus officialis* (Ormeno et al., 2007) and *Cistus ladanifer* (Haberstroh et al., 2018) emission rates and depends on individual species behaviour.



**Figure S3.**

Relative effect of available soil moisture on emission rates of MT and SQT at standard conditions otherwise ( $T=30^{\circ}\text{C}$ ) for *Cistus albidus*. (top) Effect on monoterpene emissions (Ormeno et al., 2007) with different fit options explained in the study; (bottom) effect on sesquiterpene emission rates (Ormeno et al., 2007). The corresponding fitted curves are listed below the figure. Please note, that in case of sesquiterpenes scattering and uncertainty ranges are so large that no distinct pattern is apparent which might be fitted to an assumption other than the hydraulic fit.



Fits (MT):

$$\gamma_{SM}(\Sigma \text{ MT, growth fit}) = \exp(-\exp(0.0441 \cdot \exp(1)) \cdot (-5.0 - \text{SWA}_{\%}) + 1)$$

$$\gamma_{SM}(\Sigma \text{ MT, hydr. fit}) = 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.333 / (45 + \text{SWA}_{\%})$$

$$\gamma_{SM}(\Sigma \text{ MT, hydr. fit}^*) = 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.333 / (45 + \text{SWA}_{\%}) + 0.5 \cdot (\text{SWA}_{\%} + 4)^{1.6} \cdot \exp(-0.6 \cdot (\text{SWA}_{\%} + 4))$$

$$\gamma_{SM}(\alpha\text{-pinene, hydr. fit}^*) = 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.333 / (45 + \text{SWA}_{\%}) + 2.6 \cdot (\text{SWA}_{\%} + 10)^{1.6} \cdot \exp(-0.3 \cdot (\text{SWA}_{\%} + 10))$$

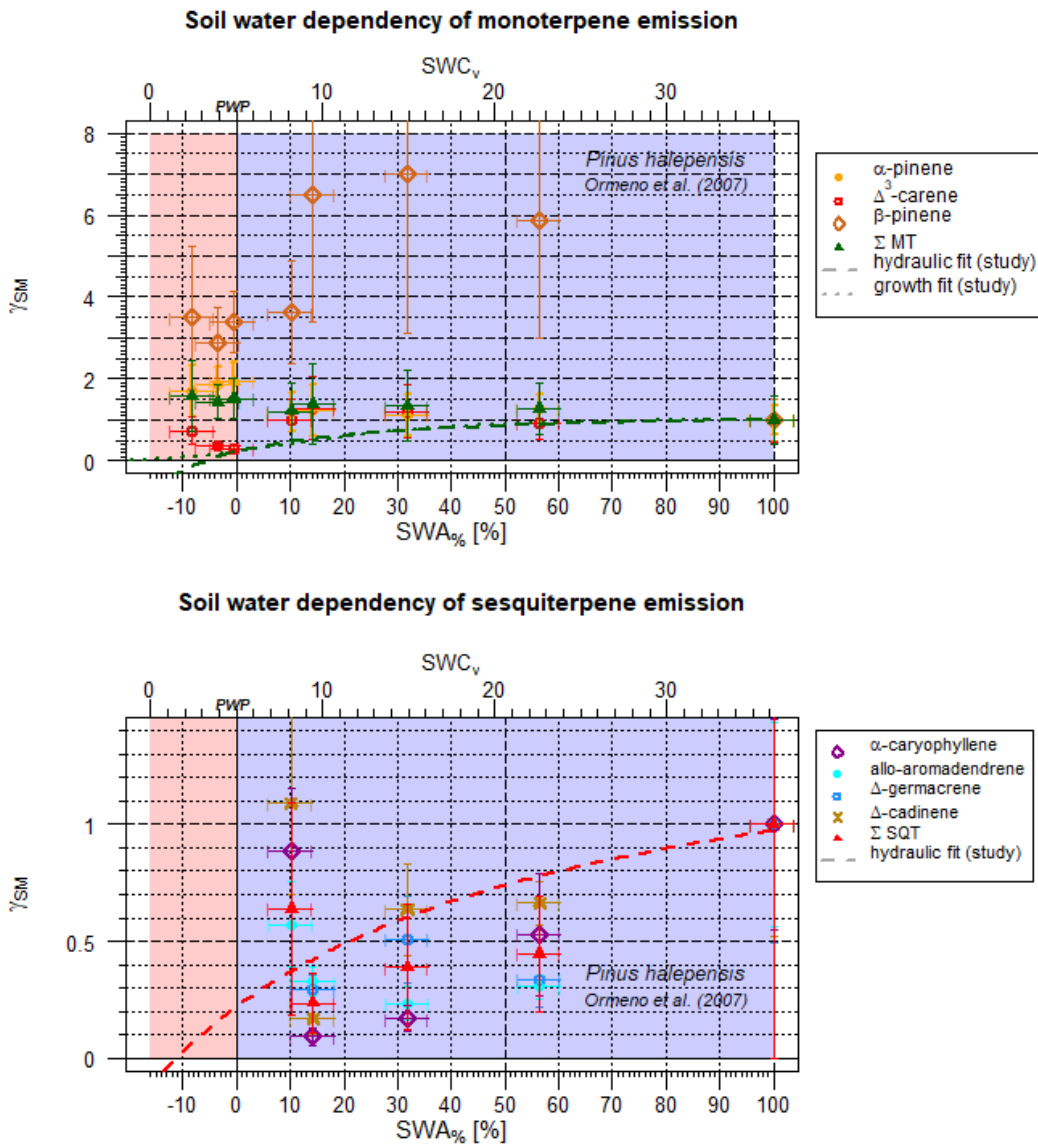
$$\gamma_{SM}(\beta\text{-pinene, hydr. fit}^*) = 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.333 / (45 + \text{SWA}_{\%}) + 1.3 \cdot (\text{SWA}_{\%} + 8)^{1.6} \cdot \exp(-0.3 \cdot (\text{SWA}_{\%} + 8))$$

Fits (SQT):

$$\gamma_{SM}(\Sigma \text{ SQT, hydr. fit}) = 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.6 / (81 + \text{SWA}_{\%})$$

**Figure S4.**

Relative effect of available soil moisture on emission rates of MT and SQT at standard conditions otherwise ( $T=30^{\circ}\text{C}$ ) for *Pinus halepensis*. (top) Effect on monoterpene emissions (Ormeno et al., 2007) with different fit options explained in the study. Please note that  $\gamma_{SM}(\beta\text{-myrcene})$  values are beyond 17 and were discarded from that plot to make others visible; (bottom) effect on sesquiterpene emission rates (Ormeno et al., 2007). Please note, that in case of individual (except  $\beta$ -pinene) and total monoterpenes scattering and uncertainty ranges are so large that no distinct pattern is apparent, which cannot be fitted to an assumption other than no change ( $=1$ ). Values of  $\beta$ -pinene may tentatively indicate a  $\gamma_{SM}$  value maximum around 30%  $\text{SWA}_{\%}$ , but all values except the one at 100%  $\text{SWA}_{\%}$  are insignificantly different!



*Fits (MT):*

$$\gamma_{SM}(\Sigma MT, \text{growth fit}) = \exp(-\exp(0.0441 \cdot \exp(1)) \cdot (-5.0 - \text{SWA}_{\%}) + 1)$$

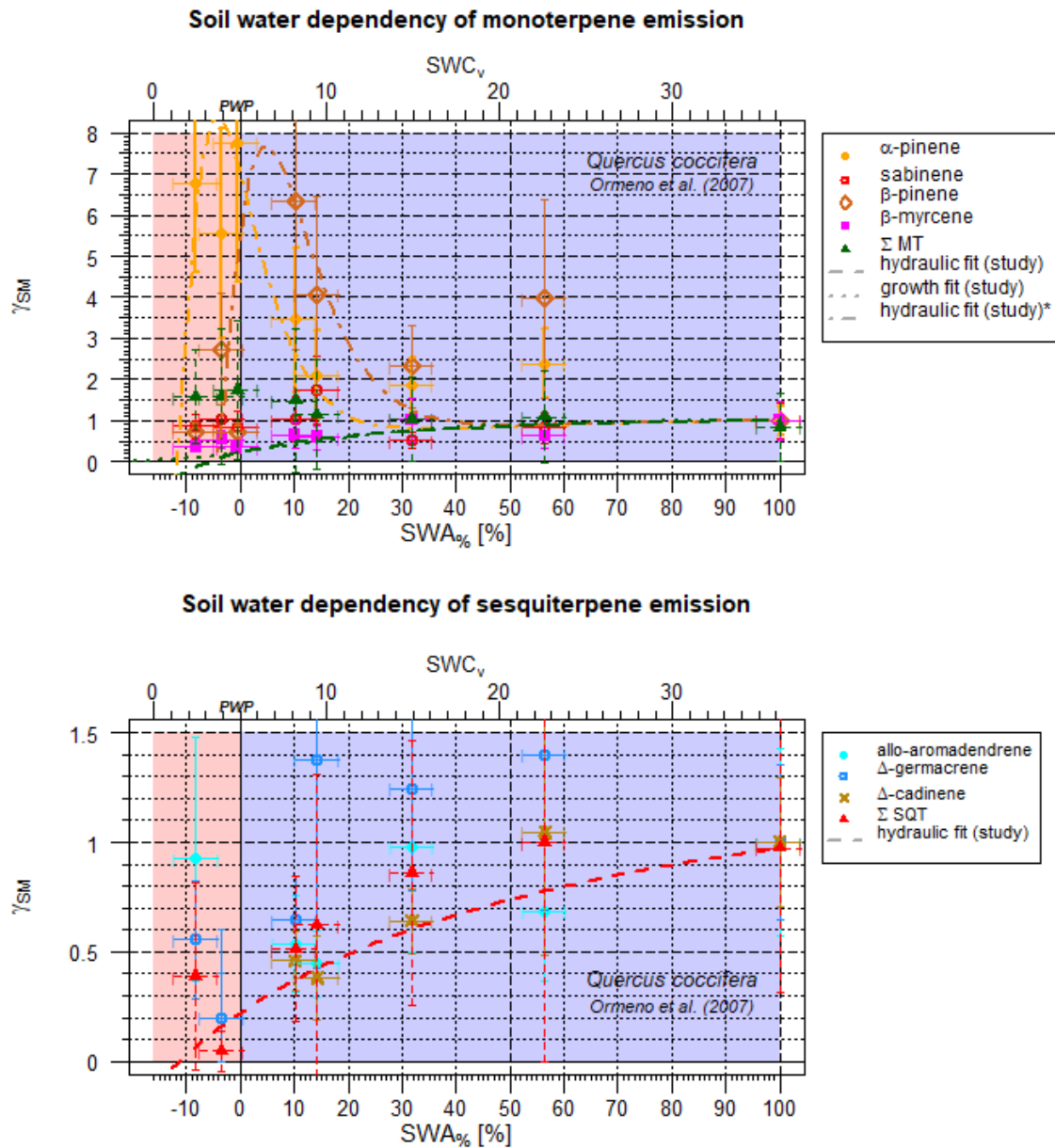
$$\gamma_{SM}(\Sigma MT, \text{hydr. fit}) = 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.333 / (45 + \text{SWA}_{\%})$$

*Fits (SQT):*

$$\gamma_{SM}(\Sigma SQT, \text{hydr. fit}) = 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.6 / (81 + \text{SWA}_{\%})$$

**Figure S5.**

Relative effect of available soil moisture on emission rates of MT and SQT at standard conditions otherwise ( $T=30^{\circ}\text{C}$ ) for *Quercus coccifera*. (top) Effect on monoterpene emissions (Ormeno et al., 2007) with different fit options explained in the study; (bottom) effect on sesquiterpene emission rates (Ormeno et al., 2007). Please note, that in case of total sesquiterpenes scattering and uncertainty ranges are so large that no distinct pattern is apparent which might be fitted to an assumption other than the hydraulic fit.



Fits (MT):

$$\gamma_{SM}(\Sigma \text{MT, growth fit}) = \exp(-\exp(0.0441 \cdot \exp(1) \cdot (-5.0 - \text{SWA}_{\%}) + 1))$$

$$\gamma_{SM}(\Sigma \text{MT, hydr. fit}) = 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.333 / (45 + \text{SWA}_{\%})$$

$$\gamma_{SM}(\alpha\text{-pinene, hydr. fit}^*) = 1.6 \cdot (\text{SWA}_{\%} + 12)^{1.6} \cdot \exp(-0.21 \cdot (\text{SWA}_{\%} + 12)) + 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.333 / (45 + \text{SWA}_{\%})$$

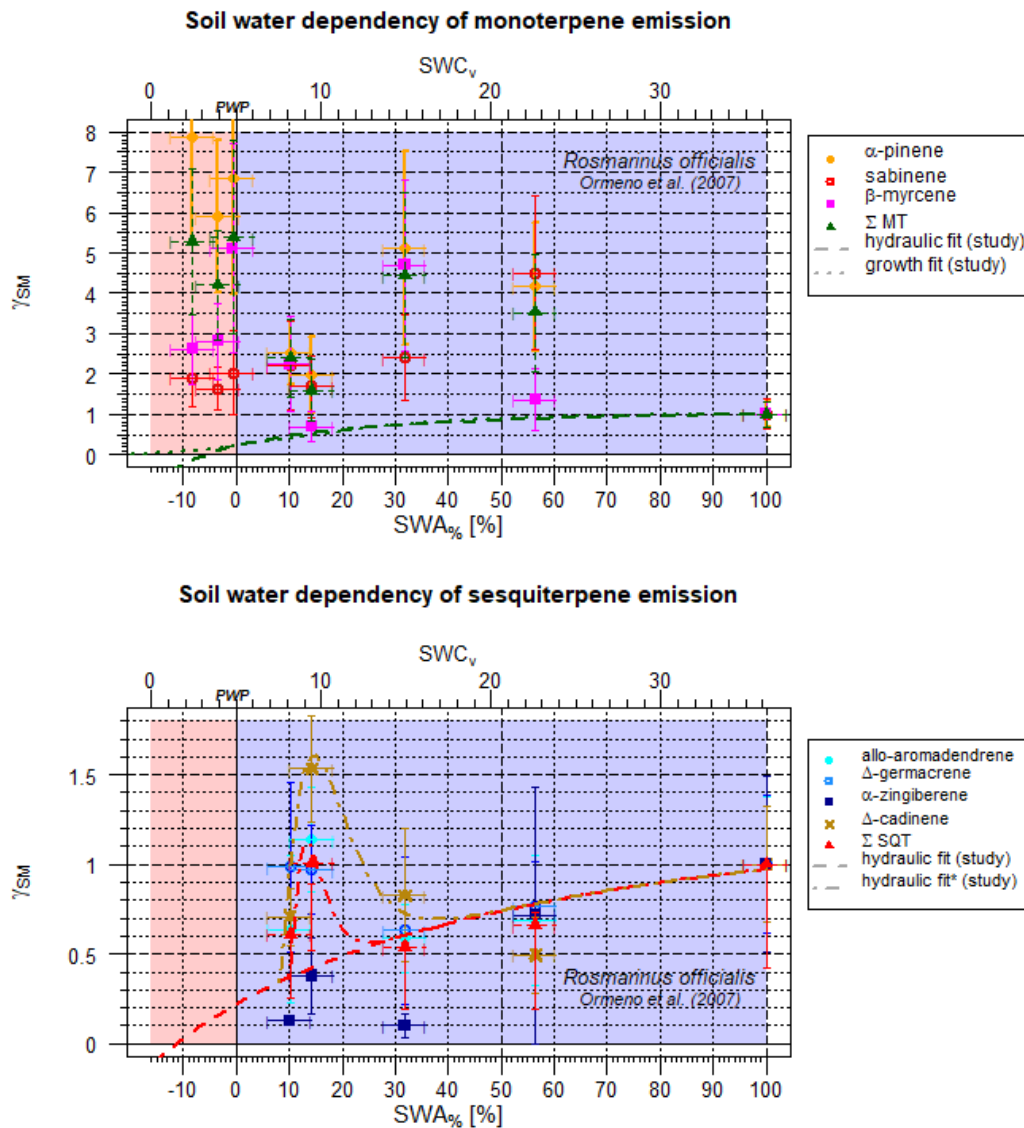
$$\gamma_{SM}(\beta\text{-pinene, hydr. fit}^*) = 1.1 \cdot (\text{SWA}_{\%} + 4.3)^{1.6} \cdot \exp(-0.18 \cdot (\text{SWA}_{\%} + 4.3)) + 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.333 / (45 + \text{SWA}_{\%})$$

Fits (SQT):

$$\gamma_{SM}(\Sigma \text{SQT, hydr. fit}) = 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.6 / (81 + \text{SWA}_{\%})$$

**Figure S6.**

Relative effect of available soil moisture on emission rates of MT and SQT at standard conditions otherwise ( $T=30^{\circ}\text{C}$ ) for *Pinus halepensis*. (top) Effect on monoterpene emissions (Ormeno et al., 2007) with different fit options explained in the study. Please note that  $\gamma_{SM}(\beta\text{-myrcene})$  values are beyond 17 and were discarded from that plot to make others visible; (bottom) effect on sesquiterpene emission rates (Ormeno et al., 2007). Please note, that in case of individual (except  $\beta$ -pinene) and total monoterpenes scattering and uncertainty ranges are so large that no distinct pattern is apparent, which cannot be fitted to an assumption other than no change ( $=1$ ). Values of  $\beta$ -pinene may tentatively indicate a  $\gamma_{SM}$  value maximum around 30%  $\text{SWA}_{\%}$ , but all values except the one at 100%  $\text{SWA}_{\%}$  are insignificantly different!



Fits (MT):

$$\gamma_{SM}(\Sigma \text{MT, growth fit}) = \exp(-\exp(0.0441 \cdot \exp(1) \cdot (-5.0 - \text{SWA}_{\%}) + 1))$$

$$\gamma_{SM}(\Sigma \text{MT, hydr. fit}) = 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.333 / (45 + \text{SWA}_{\%})$$

Fits (SQT):

$$\gamma_{SM}(\Sigma \text{SQT, hydr. fit}) = 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.6 / (81 + \text{SWA}_{\%})$$

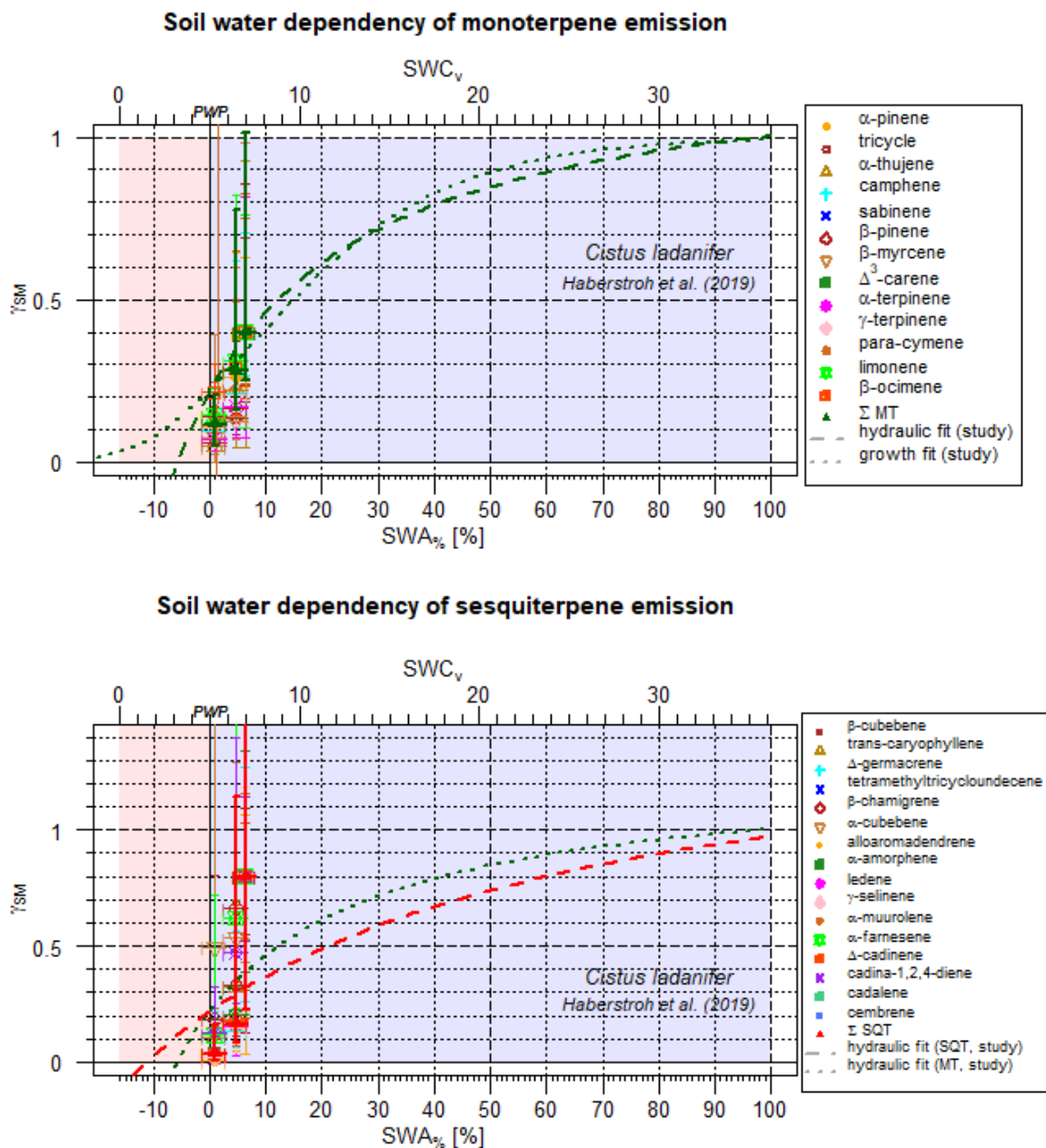
$$\gamma_{SM}(\Sigma \text{SQT, hydr. fit}^*) = 0.3 \cdot (\text{SWA}_{\%} + 10)^{1.6} \cdot \exp(-0.5 \cdot (\text{SWA}_{\%} + 10)) + 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.6 / (81 + \text{SWA}_{\%})$$

$$\gamma_{SM}(\Delta\text{-cadinene, hydr. fit}^*) = 1.6 \cdot (\text{SWA}_{\%} + 8)^{1.6} \cdot \exp(-0.21 \cdot (\text{SWA}_{\%} + 8)) + 0.22 + 0.78 \cdot \text{SWA}_{\%} \cdot 1.6 / (81 + \text{SWA}_{\%})$$



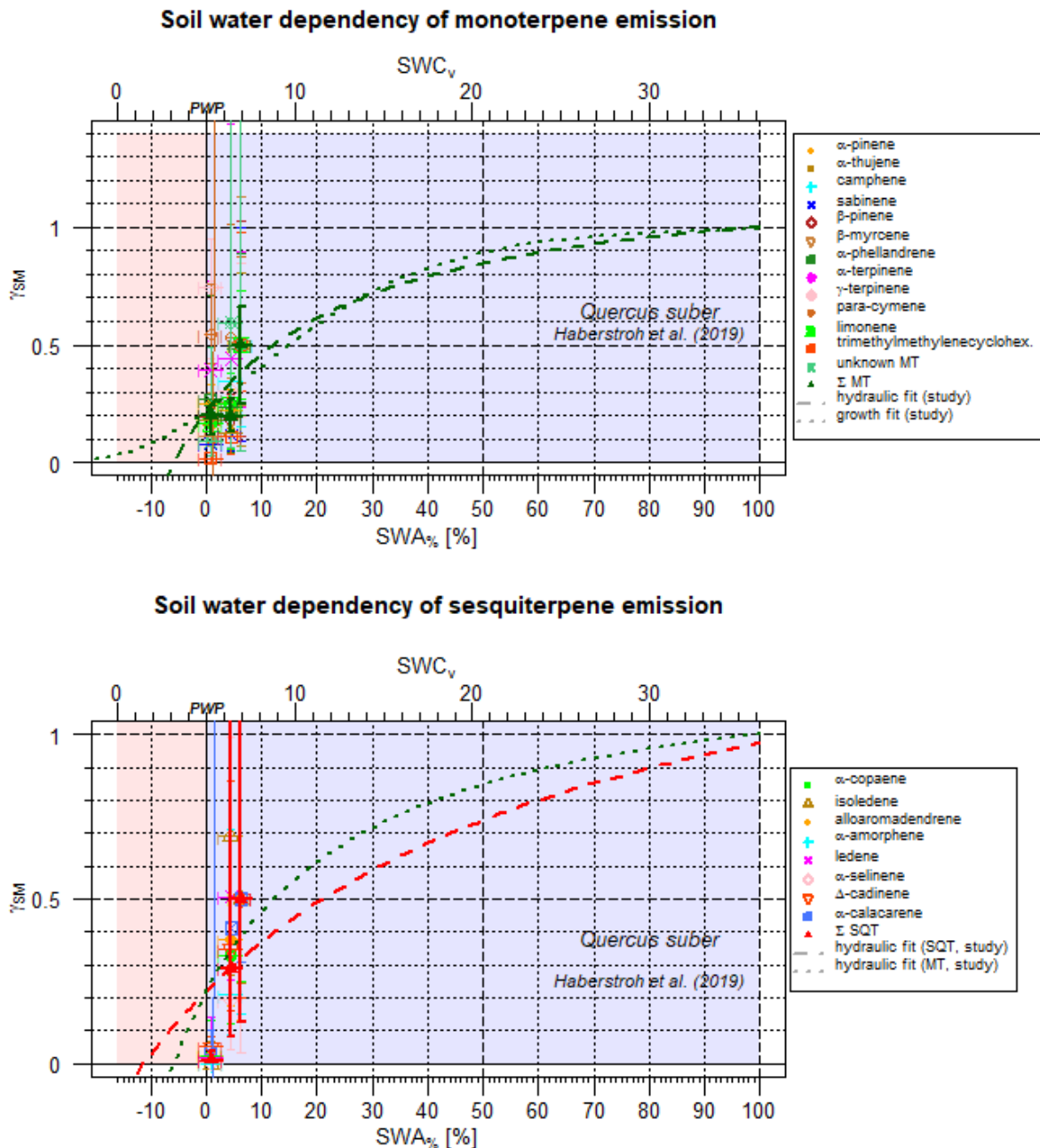
**Figure S7.**

Relative effect of available soil moisture on emission rates of MT and SQT at standard conditions otherwise ( $T=30^{\circ}\text{C}$ ) for *Cistus albidus*. (top) Effect on MT emissions (Haberstroh et al., 2018). Please note that  $\gamma_{\text{SM}}$  (MT) values are derived by assuming the standard emission rates for MT species using the value at  $\text{SWA}_{\%}$  of 6.3% multiplied with a factor of 2.5; (bottom) effect on SQT emission rates (Haberstroh et al., 2018). Please note, that  $\gamma_{\text{SM}}$  (SQT) values are derived by assuming the standard emission rates for SQT species using the value at  $\text{SWA}_{\%}$  of 6.3% multiplied with a factor of 1.25!



**Figure S8.**

Relative effect of available soil moisture on emission rates of MT and SQT at standard conditions otherwise ( $T=30^{\circ}\text{C}$ ) for *Quercus suber*. (top) Effect on MT emissions (Haberstroh et al., 2018). Please note that  $\gamma_{SM}$  values are derived by assuming the standard emission rates for MT species using the value at  $\text{SWA}_{\%}$  of 6.3% multiplied with a factor of 2; (bottom) effect on SQT emission rates (Haberstroh et al., 2018). Please note, that  $\gamma_{SM}$  (SQT) values are derived by assuming the standard emission rates for SQT species using the value at  $\text{SWA}_{\%}$  of 6.3% multiplied with a factor of 2!



**FOR ALL FIGURES WITH FITS OF INDIVIDUAL SPECIES: THE NUMBER OF DATAPPOINTS IS FAIRLY LOW AND UNCERTAINTIES MAY BE SUBSTANTIAL! SOMETIMES EACH PARAMETER CORRESPONDS TO A SINGLE VALUE! THUS NO RANGES ARE PROVIDED.**



**Table S1.**

Reaction rate constants for selected VOCs with OH and O<sub>3</sub> at T = 25°C provided in cm<sup>3</sup> molecule<sup>-1</sup> s<sup>-1</sup>. The corresponding references or the SAR reference chosen is named on the right. Note, “-” declares the rate constant is either <2e-19 cm<sup>3</sup> molecule<sup>-1</sup> s<sup>-1</sup> i.e. negligible at ambient conditions or not expected. ‘NA’ refers to not available and ‘SAR’ to structure activity relationship methods. Bold marked values have been used for calculation, experimental values are preferred towards calculated ones.

<i>name</i>	<i>k<sub>OH</sub></i> [cm <sup>3</sup> s <sup>-1</sup> ]	<i>k<sub>O3</sub></i> [cm <sup>3</sup> s <sup>-1</sup> ]	<i>reference or method</i>
3-hexen-1-ol	(1.14±0.14)e-10 <b>(1.08±0.22)e-10</b>	(6.4±1.7)e-17	Gibilisco et al., 2015 Atkinson et al., 1995
2-nonenal	(4.35±0.30)e-11	(2.5±0.2)e-18	Gao et al., 2009; Gaona Colman et al., 2015
n-heptanal	(2.96±0.23)e-11	-	Albaladejo et al., 2002
benzaldehyde	(1.26±0.25)e-11	-	IUPAC, 2009
sabinene	1.17e-10	8.6e-17	Atkinson, 1994
2-methyl-2-hepten-6-one	(9±6)e-11	(4.0±0.5)e-16	SAR (Neeb, 2000); Ozonolysis approx. using 2-methyl-2-butene (Greene and Atkinson, 1992)
1,2,4-trimethyl benzene	(5.5±0.8)e-11 <b>(3.4±0.3)e-11</b>	- -	as mesitylene (Alarcon et al., 2015) Hansen et al., 1975
n-octanal	(3.1±0.3)e-11 3.2e-11 <b>3.8e-11</b>	- - -	SAR (Neeb, 2000) EPI Suite, 2017 <b>Karl et al., 2000</b>
para-cymene	<b>(1.51±0.4)e-11</b> (1.6±0.1)e-11	- -	Corchnoy et al., 1990 Alarcon et al., 2013
limonene	1.17e-10	(2.0±0.5)e-16	Shu & Atkinson, 1994; Atkinson, 1994
benzyl alcohol	(2.9±0.4)e-11	-	Bernard et al., 2013
iso-pinocampheol	(0.97±0.16)e-11 <b>1.43e-11</b> 1.77e-11	- - -	SAR (Neeb, 2000) <b>SAR (Kwok &amp; Atkinson, 1995)</b> EPI Suite, 2017
1-methyl-4-(1-methylethenyl) benzene	(9.2±0.9)e-11 <b>5.56e-11</b>	NA <b>1.4e-16</b>	SAR (Neeb, 2000) EPI Suite, 2017
nonanal	<b>(3.6±0.7)e-11</b> 4.5e-11	- -	<b>Bowman et al., 2015</b> Karl et al., 2000
benzene ethanol	(9.0±0.9)e-11 (1.948±0.2)e-12	- -	SAR (Neeb, 2000) EPI Suite, 2017

trans-pinocarveol	(2.8±0.3)e-11	NA 1.50e-17	SAR (Neeb, 2000) EPI Suite, 2017
1-nonanol	(1.9±0.2)e-11	-	SAR (Neeb, 2000)
4-terpineol	(1.9±0.5)e-10	(3.0±0.2)e-16	Wells, 2005
naphthalene	2.2e-11	-	Atkinson et al., 1989
para-cymenol, 2-(4-methylphenyl) propan-2-ol	5.1e-10 <b>(4.9±0.2)e-10</b> 7.4e-12	- - -	Kwok and Atkinson, 1995 SAR (Neeb, 2000) EPI Suite, 2017
methyl-salicylate	(7.4±0.4)e-10	-	SAR (Neeb, 2000)
α-terpineol	(1.9±0.5)e-10	(3.0±0.2)e-16	Wells, 2005
2-(2-hydroxypropoxy) 1- propanol	<b>(3.1±0.6)e-12</b> <b>3.5e-11</b>	- -	SAR (Neeb, 2000) EPI Suite, 2017
3,3'-oxybis 2-butanol	(7.3±1.3)e-12	-	SAR (Neeb, 2000)
2-decenal	(5.3±0.5)e-11 (4.6±0.5)e-11	NA (1.3±0.6)e-18	SAR (Neeb, 2000) EPI Suite, 2017
1-heptadecanol	(5.4±0.5)e-12	-	SAR (Neeb, 2000)
2-undecanal	(2.7±0.3)e-12	-	SAR (Neeb, 2000)
undecanal	(3.6±0.5)e-12	-	SAR (Neeb, 2000)
2-methyl naphthalene	(5.2±0.4)e-11	-	Atkinson & Aschmann, 1986
cis-3-hexenyltiglate	<b>(3.3±0.2)e-10</b> 1.0e-10	NA <b>(2.4±0.4)e-16</b>	SAR (Neeb, 2000) EPI Suite, 2017
dodecanal	(6.0±0.6)e-12	-	SAR (Neeb, 2000)
junipene	(4.7±1.7)e-11	<5e-19	as longifolene
isolongifolene	(1.0±0.1)e-10	2.6e-17	Atkinson, 1997, Richters et al., 2015
tetradecane	(1.8±0.2)e-11	-	Atkinson, 1997
cyclosativen	(3.6±0.4)e-12	-	SAR (Neeb, 2000)
longifolene	(4.7±1.7)e-11	<5e-19	Atkinson, 1994
6,10-dimethyl-5,9- undecadien-2-one	(3.3±0.5)e-10 <b>1.8e-10</b>	NA <b>8.6e-16</b>	SAR (Neeb, 2000) EPI Suite, 2017
2,6-bis (1,1-dimethyl- ethyl)-2,5- cyclohexadiene-1,4- dione	(4.0±0.4)e-10 <b>2.2e-12</b>	NA 3.5e-18	SAR (Neeb, 2000) EPI Suite, 2017
1-dodecanol	<b>(4.1±0.3)e-12</b> 1.8e-11	- -	SAR (Neeb, 2000) EPI Suite, 2017

$\alpha$ -bergamotene	(4.5±0.5)e-10 <b>1.8e-10</b>	NA <b>8.6e-16</b>	SAR (Neeb, 2000), ring eff. not incl. EPI Suite, 2017
pentadecane	2.1e-11	-	Atkinson, 1997
$\alpha$ -farnesene	(2.19±0.11)e-10	(5.9±3.4)e-16	Kim et al., 2011; Kourtchev et al., 2011
tridecanal	(6.0±0.6)e-12	-	SAR (Neeb, 2000)
1,2,3,4-tetrahydro-1.6-dimethyl-4-(1-methylethyl)-naphthalene	(9.0±0.9)e-11	-	SAR (Neeb, 2000)
1-butylhexyl benzene	(9.0±0.9)e-11	-	SAR (Neeb, 2000)
cadina-1(10),6,8-triene	<b>(9.0±0.9)e-11</b> 2.8e-11	- -	SAR (Neeb, 2000) EPI Suite, 2017
1,6-dioxacyclo-dodecane-7,12-dione	(3.0±0.2)e-12	-	SAR (Neeb, 2000)
3-ethyl-tridecane	(2.0±0.5)e-11	-	Approximated by similar alkanes (Atkinson, 1997)
cis-3-hexenyl benzoate	<b>(4.2±0.2)e-10</b> 6.1e-11	NA <b>(1.3±0.5)e-16</b>	SAR (Neeb, 2000); EPI Suite, 2017
hexadecane	2.3e-11	-	Atkinson, 1997
santalol	(3.8±0.3)e-11	-	SAR (Neeb, 2000)
1-butylheptyl benzene	(9.0±0.9)e-11	-	SAR (Neeb, 2000)
4-methyl hexadecane	(3.0±0.3)e-11	-	SAR (Neeb, 2000)
dotriacontane	(3.0±0.3)e-11	-	SAR (Neeb, 2000)
hexadecahydro-pyrene	(5.7±0.4)e-12	-	SAR (Neeb, 2000)
5.6-bis(2.2-dimethylpropylidene)-decane	(1.9±0.2)e-10	-	SAR (Neeb, 2000)
2,3-dihydro-1,1,3-trimethyl-3-phenyl 1H indene	<b>(7.8±1.9)e-11</b> 1.1e-11	- -	SAR (Neeb, 2000) EPI Suite, 2017
octadecane	(2.3±1.1)e-11 2.2e-11 <b>(2.6±0.3)e-11</b>	- - -	SAR (Neeb, 2000) EPI Suite, 2017 app. by sim. alkanes (Atkinson, 1997)
isopropyl myristate	(5.5±0.4)e-12	-	SAR (Neeb, 2000)
6,10,14-trimethyl 2-pentadecanone	(4.4±0.3)e-12	-	SAR (Neeb, 2000)

oxacycloheptadecan-2-one	(5.0±0.2)e-12	-	SAR (Neeb, 2000)
1-octanol	(2.5±0.3)e-12	-	SAR (Neeb, 2000)
trimethyl (2-methyl-1-propenylidene) cyclopropane	<b>(1.1±0.1)e-10</b> <b>1.1e-10</b>	NA <b>4.1e-17</b>	SAR (Neeb, 2000) EPI Suite, 2017
2-nonenal	(5.9±0.6)e-11 <b>(4.4±0.3)e-11</b> (4.4±0.1)e-11	NA NA <b>(1.3±0.5)e-18</b>	SAR (Neeb, 2000) <b>Gao et al., 2009</b> <b>EPI Suite, 2017</b>
3-cyclohexen-1-ol 4-methyl-1-(1-methylethyl), terpinen-4-ol	<b>(1.0±0.1)e-10</b> <b>1.0e-10</b>	NA <b>4.3e-16</b>	SAR (Neeb, 2000) EPI Suite, 2017
trans-chrysanthenyl acetate	<b>(2.8±0.3)e-11</b> 9.2e-11	NA <b>4.3e-16</b>	SAR (Neeb, 2000) EPI Suite, 2017
dodecyl oxirane	(5.5±0.3)e-12	-	SAR (Neeb, 2000)
hexatriacontane	(1.00±0.05)e-11	-	SAR (Neeb, 2000)
2-dodecanone	(3.0±0.2)e-12	-	SAR (Neeb, 2000)
1,8-dimethyl naphthalene	(8.4±0.6)e-11	-	SAR (Neeb, 2000)
3-methyl tetradecane	(2.1±0.2)e-11	-	Approx. by using Atkinson, 1997
10-methyl eicosane	(3.6±0.5)e-11	-	Approx. by using Atkinson, 1997
2-methyl pentadecane	(2.3±0.3)e-11	-	Approx. by using Atkinson, 1997
3-methyl pentadecane	(2.3±0.3)e-11	-	Approx. by using Atkinson, 1997
3-methyl heptadecane	(2.8±0.3)e-11	-	Approx. by using Atkinson, 1997
1,1'-oxybis octane	(9.5±0.6)e-12	-	SAR (Neeb, 2000)
n-eicosane	(3.4±0.4)e-11	-	Approx. by using Atkinson, 1997
heneicosane	(3.6±0.5)e-11	-	Approx. by using Atkinson, 1997
trans-β-ocimene	(2.5±0.2)e-10 NA	NA <b>(3.7±0.9)e-16</b>	Atkinson et al., 1986; Gaona-Colman et al., 2016

### References:

- Alarcón, P., Strekowski, R., Zetzsch, C., 2013. Reversible addition of the OH radical to p-cymene in the gas phase: kinetic analysis assuming formation of a single adduct. Part 1. *Phys. Chem. Chem. Phys.*, **15**, 20105-20114, doi: 10.1039/c3cp53040j.
- Alarcón, P., Bohn, B., Zetzsch, C., 2015. Kinetic and mechanistic study of the reaction of OH radicals with methylated benzenes: 1,4-dimethyl-, 1,3,5-trimethyl-, 1,2,4,5-, 1,2,3,5- and 1,2,3,4-tetramethyl-, pentamethyl- and hexamethylbenzene. *Phys. Chem. Chem. Phys.*, **17**, 13053-13065, DOI: 10.1039/c5cp00253b.
- Albaladejo, J., Ballesteros, B., Jimenez, E., Martin, P., Martinez, E., 2002. A PLP-LIF kinetic study of the atmospheric reactivity of a series of C-4-C-7 saturated and unsaturated aliphatic aldehydes with OH. *Atmos. Environ.* **36**, 3231-3239, doi: 10.1016/S1352-2310(02)00323-0.
- Atkinson, R., Aschmann, S.M., 1986. Kinetics of the reactions of naphthalene, 2-methylnaphthalene, and 2,3-dimethylnaphthalene with OH radicals and with O<sub>3</sub> at 295 ± 1 K. *Int. J. Chem. Kinet.* **18**, 569-573.
- Atkinson, R., Aschmann, S.M., Pitts Jr., J.N., 1986. Rate constants for the gas-phase reactions of the OH radical with a series of monoterpenes at 294 ± 1 K. *Int. J. Chem. Kinet.* **18**, 287-299.
- Atkinson, R., Baulch, D.L., Cox, R.A., Hampson, R.F., Kerr, J.A., Troe, J., 1989. Evaluated kinetic and photochemical data for atmospheric chemistry.3. IUPAC subcommittee on gas kinetic data evaluation for atmospheric chemistry. *J. Phys. Chem. Ref. Data*, **18**, 881-1097, doi: 10.1063/1.555832.
- Atkinson, R., 1994. Gas-phase tropospheric chemistry of organic compounds. *J. Phys. Chem. Ref. Data*, Monograph **2**, 1-216.
- Atkinson, R., Arey, J., Aschmann, S. M., Corchnoy, S. B., and Shu, Y. H., 1995. Rate Constants for the gas-phase reactions of cis-3-Hexen-1-ol, cis-3-hexenylacetate, trans-2-hexenal, and linalool with OH and NO<sub>3</sub> radicals and O<sub>3</sub> at 296±2 K and OH radical formation yields from the O<sub>3</sub> reactions. *Int. J. Chem. Kinet.* **27**, 941–955, doi: 10.1002/kin.550271002.
- Atkinson, R., 1997. Gas-phase tropospheric chemistry of volatile organic compounds 1. Alkanes and alkenes, *J. Phys. Chem. Ref. Data*, **26**, 215-290, 1997.
- Bernard, F., Magneron, I., Eyglunent, G., Daële, V., Wallington, T.J., Hurley, M.D., Mellouki, A., 2013. Atmospheric Chemistry of Benzyl Alcohol: Kinetics and Mechanism of Reaction with OH Radicals. *Environ. Sci. Technol.*, **47**, 3182–3189, doi: 10.1021/es304600z.
- Bowmann, J.H., Barkett, D.J., Shepson, P.B., 2015. Atmospheric chemistry of nonanal. *Environ. Sci. Technol.* **37**, 2218-2225, doi: 10.1021/es026220p.
- Corchnoy, S.B., Atkinson, R.A., 1990. Kinetics of the gas-phase reactions of OH and NO<sub>3</sub> radicals with 2-carene, 1,8-cineole, p-cymene, and terpinolene. *Environ. Sci. Technol.* **24**, 1497-1502.
- Gao, T., Andino, J.M., Rivera, C.C., Marquez, M.F., 2009. Rate constants of the gas-phase reactions of OH radicals with *trans*-2-hexenal, *trans*-2-octenal, and *trans*-2-nonenal. *Int. J. Chem. Kinet.* **41**, 483-489.
- Gaona-Colman, E. Blanco, M.B., Barnes, I., Teruel, M.A., 2016. Gas-phase ozonolysis of beta-ocimene: Temperature dependent rate coefficients and product distribution. *Atmos. Environ.* **147**, 46-54, doi: 10.1016/j.atmosenv.2016.09.053

- Gibilisco, R.G., Blanco, M.B., Bejan, I., Barnes, I., Wiesen, P., Teruel, M.A., 2015. Atmospheric sink of (E)-3-hexen-1-ol, (Z)-3-hepten-1-ol, and (Z)-3-octen-1-ol: Rate coefficients and mechanisms of the OH-radical initiated degradation. *Environ. Sci. Technol.* **49**, 7717–7725, DOI: 10.1021/es506125c
- Gill, K. J., Hites, R.A., 2002. Rate constants for the gas-phase reactions of the hydroxyl radical with isoprene, a- and b-pinene, and limonene as a function of temperature. *J. Phys. Chem. A* **106**, 2538-2544.
- Greene, C.R., Atkinson, R.A., 1992. Rate constants for the gas-phase reactions of O<sub>3</sub> with a series of alkenes at 296±2 K. *Int. J. Chem. Kinet.* **24**, 803-811, doi: 10.1002/kin.550240905.
- Hansen, D.A., Atkinson, R., Pitts Jr., J. N., 1975. Rate constants for the reaction of OH radicals with a series of aromatic hydrocarbons. *J. Phys. Chem.* **17**, 1763-1766.
- Karl, T., Jordan, A., Fall, R., Crutzen, P.J., Mandl, M., Staudinger, M., Lindinger, W., 2000. *Diurnal variation of the concentration of OH during the 27<sup>th</sup> November 1999 at the Sonnblick observatory (3106 m a.s.l.)*. Poster presentation at Ionicon conference, Innsbruck, 2000.
- Kim D., Stevens, P.S., Hites, R.A., 2011. Rate constants for the gas-phase reactions of OH and O<sub>3</sub> with β-ocimene, β-myrcene, and α- and β-farnesene as a function of temperature. *J. Phys. Chem. A*, **115**, 500–506.
- Kourtchev, I., Bejan, I., Sodeau, J., Wenger, J.C., 2011. Gas-phase reaction of OH radicals with (E)-β-farnesene at 296±2 K: Rate coefficient and carbonyl products. *Atmos. Environ.* **46**, 338-345.
- Kwok, E.S.C., Atkinson, R., 1995. Estimation of hydroxyl radical reaction rate constants for gas-phase organic compounds using a structure-reactivity relationship: an update. *Atmos. Environ.* **29**, 1685–1695.
- IUPAC, 2009. *IUPAC Task Group on Atmospheric Chemical Kinetic Data Evaluation*, <http://iupac.pole-ether.fr>.
- Neeb, P., 2000. Structure-reactivity based estimation of the rate constants for hydroxyl radical reactions with hydrocarbons. *J. Atmos. Chem.*, **35**, 295–315.
- Richters, S., Herrmann, H., Berndt, T., 2015. Gas-phase rate coefficients of the reaction of ozone with four sesquiterpenes at 295±2 K. *Phys. Chem. Chem. Phys.* **17**, 11658-11669, doi: 10.1039/c4cp05542j
- Shu, Y., Atkinson, R., 1994. Rate constants for the gas-phase reactions of O<sub>3</sub> with a series of terpenes and OH radical formation from the O<sub>3</sub> reactions with sesquiterpenes at 296±2 K. *Int. J. Chem. Kinet.* **26**, 1193.
- Wells, J.R., 2005. Gas-phase chemistry of α-terpineol with ozone and OH radical: Rate constants and products. *Environ. Sci. Technol.*, **39**, 6937-6943, doi: 10.1021/es0481676.
- Zhao, Y., Sun, H., Qin, M., Wanga, R., Xu, L., 2015. Kinetics and mechanism for OH-initiated gas-phase chemistry of α-terpineol. *RSC Adv.* **5**, 95096–95103, DOI: 10.1039/c5ra14781f.