# Methane paradox in tropical lakes? Sedimentary fluxes rather than pelagic production in oxic conditions sustain methanotrophy and emissions to the atmosphere

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**Abstract.** Despite growing evidence that methane ($CH_4$) formation could also occur in well-oxygenated surface freshwaters, its significance at the ecosystem scale is uncertain. Empirical models based on data gathered at high latitude predict that the contribution of oxic $CH_4$ increases with lake size and should represent the majority of $CH_4$ emissions in large lakes. However, such predictive models could not directly apply to tropical lakes which differ from their temperate counterparts in some fundamental characteristics, such as year-round elevated water temperature. We conducted stable isotope tracer experiments which revealed that oxic $CH_4$ production is closely related to phytoplankton metabolism, and is a common feature in five contrasting African lakes. Nevertheless, methanotrophic activity in surface waters and $CH_4$ emissions to the atmosphere were predominantly fuelled by $CH_4$ generated in sediments and physically transported to the surface. Indeed, $CH_4$ bubble dissolution flux and diffusive benthic $CH_4$ flux were several orders of magnitude higher than $CH_4$ production in surface waters. Microbial $CH_4$ consumption dramatically decreased with increasing sunlight intensity, suggesting that the freshwater "$CH_4$ paradox" might be also partly explained by photo-inhibition of $CH_4$ oxidizers in the illuminated zone. Sunlight appeared as an overlooked but important factor determining the $CH_4$ dynamics in surface waters, directly affecting its production by photoautotrophs and consumption by methanotrophs.

## 1. Introduction

Emissions from inland waters are an important component of the global $CH_4$ budget (Bastviken et al. 2011), in particular from tropical latitudes (Borges et al. 2015). While progress has been made in evaluating the $CH_4$ emission rates, much less attention has been given to the underlying microbial production (methanogenesis) and loss (methane oxidation) processes. It is generally assumed that $CH_4$ in lakes originates from the degradation of organic matter in anoxic sediments. Because most methanogens are considered to be strict anaerobes and net vertical diffusion of $CH_4$ from anoxic bottom waters is often negligible (Bastviken et al. 2003), physical processes of $CH_4$ transport from shallow sediments are usually invoked to explain patterns of local $CH_4$ concentration maximum in surface waters (Encinas-Fernandez et al. 2016, Peeters et al. 2019, Martinez-Cruz et al. 2020). Indeed, $CH_4$-rich pore water is regularly released from littoral sediment into the water column during resuspension events associated with surface waves (Hofmann et al. 2010).

The view that $CH_4$ is formed under strictly anaerobic conditions has been challenged by several recent studies which proposed that acetoclastic methanogens directly attached to phytoplankton cells are involved in epilimnetic $CH_4$ production (Grossart et al. 2011, Bogard et al. 2014), and are responsible of distinct near-surface peaks of $CH_4$ concentration in certain thermally stratified, well-oxygenated waterbodies (Tang et al. 2016). It has also been showed that Cyanobacteria (Bizic et al. 2020) and widespread marine phytoplankton (Klintzsch et al. 2019) are able to release substantial amount of $CH_4$ during a culture study, and this $CH_4$ production mechanism might be linked to photosynthesis. From a model-based approach, epilimnetic $CH_4$ production was shown to sustain most of the $CH_4$ oxidation in 14 Canadian lakes (DelSontro et al. 2018), and would even

represent up to 90% of the $CH_4$ emitted from a temperate lake (Donis et al. 2017). Further, empirical models based on data gathered in boreal and temperate lakes predict that the contribution of oxic $CH_4$ increases with lake size (Gunthel et al. 2019) and should represents the majority of $CH_4$ emissions in lakes larger than 1 km². Still, aerobic $CH_4$ production has so far only been documented in temperate and boreal lakes so that such predictive models could not directly apply to tropical lakes which differ from their temperate counterparts in some fundamental characteristics, such as year-round elevated water temperature. Primary production, methanogenic and methanotrophic activities, and cyanobacterial dominance are potentially much higher in tropical lakes due to favorable temperature (Lewis 1987, Kosten et al. 2012). It has also been shown that CH4 emissions are positively related to temperature at the ecosystem scale (Yvon-Durocher et al. 2014)

Here, we tested the hypothesis that phytoplankton metabolism could fuel $CH_4$ production in well-oxygenated waters in five contrasting tropical lakes in East Africa covering a wide range of size, depth, and productivity (L. Edward, L. George, L. Katinda, L. Nyamusingere and L. Kyambura). Phytoplankton activity could provide diverse substrates required for $CH_4$ production mediated by methanogenic Archaea, or alternatively $CH_4$ could be directly released by phytoplankton cells. Additionally, the significance of epilimnetic $CH_4$ production at the scale of the aquatic ecosystem was assessed by quantifying CH4 release from sediments, $CH_4$ production and oxidation rates in the water column, and $CH_4$ diffusive and ebullitive emissions to the atmosphere.

## 2. Material and methods

### 2.1. Site description

The sampled lakes cover a wide range of size (<1 to 2300 km²), maximum depth (3-117 m), mixing regimes, phytoplankton biomass and primary productivity (Table S1). Oligotrophic L. Kyamwinga (-0.18054°N, 30.14625°E) and eutrophic L. Katinda (-0.21803°N, 30.10702°E) are stratified, small but deep tropical lakes located in Western Uganda. Neighboring L. Nyamusingere (-0.284364°N, 30.037635°E) is a small but shallow and polymictic eutrophic lake. L. George is a larger (250 km²), hypereutrophic, shallow lake located at the equator (-0.02273°N, 30.19724°E). A single outlet (Kazinga Channel) flows from L. George to the neighboring Lake Edward (-0.28971°N, 29.73327°E), a holomictic, mesotrophic large lake (2325 km²). Water samples from pelagic stations of L. Katinda (0.3km from shore, 14m deep), L. George (2 km from shore, 2.5m deep) and L. Edward (15 km from shore, 20m deep) were collected in April 2017 (rainy season) and January 2018 (dry season). Pelagic sites of L. Kyamwinga (0.5 km from shore, 40m deep) and L. Nyamusingere (0.2 km from shore, 3 m deep) were sampled only once, in April 2017 and January 2018, respectively.

### 2.2. Environmental setting of the study sites

Conductivity, temperature and dissolved oxygen concentration measurements were performed with a Yellow Spring Instrument EXO II multiparametric probe. Samples for particulate organic carbon (POC) concentration were collected on glass fiber filters (0.7 µm nominal pore size) and analyzed with an elemental analyzer coupled to an isotope ratio mass spectrometer (EA-IRMS) (Morana et al. 2015). Pigment concentrations were determined by high performance liquid chromatography (Descy et al. 2016) after filtration of water samples through glass fiber filters (0.7 µm nominal pore size).

Water samples for determination of dissolved $CH_4$ concentration were transferred with tubing from the Niskin bottle to 60 ml borosilicate serum bottles that were poisoned with 200µL of a saturated solution of $HgCl_2$, closed with a butyl stopper and sealed with an aluminum cap. The concentrations of dissolved $CH_4$ was measured with the headspace equilibration technique (20 ml headspace) using a gas chromatograph with flame ionization detection (GC-FID, SRI8610C).

Samples for $\delta^{13}$C-$CH_4$ determination were collected in 60 ml serum bottles following the same procedure than samples for $CH_4$ concentration determination. In the laboratory, $\delta^{13}$C-$CH_4$ was measured as described in Morana et al. (2015). Briefly, a 20ml helium headspace was created in the serum bottles, then samples were vigorously shaken and left to equilibrate

overnight. The sample gas was flushed out through a double-hole needle and purified of non-$CH_4$ volatile organic compounds in a liquid $N_2$ trap, $CO_2$ and $H_2O$ were removed with a soda lime and a magnesium perchlorate traps, and the $CH_4$ was converted to $CO_2$ in an online combustion column similar to that in an elemental analyzer (EA). The resulting $CO_2$ was subsequently preconcentrated in a custom-built cryo-focussing device by immersion of a stainless-steel loop in liquid $N_2$, passed through a micro-packed GC column (HayeSep Q 2 m, 0.75mm ID; Restek), and finally measured on a Thermo Scientific Delta V Advantage isotope ratio mass spectrometer (IRMS). $CO_2$ produced from certified reference standards for $\delta^{13}C$ analysis (IAEA-CO1 and LSVEC) were used to calibrate $\delta^{13}C$-$CH_4$ data. Reproducibility of measurement estimated based on duplicate injection of a selection of samples was typically better than 0.5 ‰, or better than 0.2‰ when estimated based on multiple injection of standard gas.

### 2.3. Diffusive $CH_4$ flux calculation

Surface $CH_4$ concentrations were used to compute the diffusive air-water $CH_4$ fluxes ($FCH_4$) according to eq. (1):

$$FCH_4 = k \times \Delta CH_4 \quad (1)$$

Where k is the gas transfer velocity of $CH_4$ computed from wind speed (Cole & Caraco 1998) and the Schmidt number of $CH_4$ in freshwater (Wanninkhof 1992), and $\Delta CH_4$ is the air-water gradient. Wind speed data were acquired with a Davis Instruments meteorological station located in Mweya peninsula (0.11°S 29.53°E).

### 2.4. $CH_4$ ebullition flux

$CH_4$ ebullition flux was investigated in L. Edward (at 20 m), George (at 2.5 m), and Nyamusingere (at 3 m) only, at the same sampling sites where biogeochemical processes were measured Bubble traps made with an inverted funnel (24 cm diameter) connected to a 60 ml syringe were deployed for a period between 24 h and 48 h at 0.5 m below the water surface (4 replicates). Measurements were performed at sites with water depth of 20 m, 2.5 m and 3 m for L. Edward, George and Nyamusingere, respectively. After measuring the gas volume collected within the trap during the sampling period, the gas bubbles were transferred in a tightly closed 12 ml Exetainer vial (Labco) for subsequent analysis of their $CH_4$ concentration. Variability of the gas volume in the 4 replicates was less than 10%. We used the SiBu-GUI software (McGinnis et al. 2006, Greinert et al. 2009) to correct for gas exchange within the water column during the rise of bubbles and thus obtained the $CH_4$ ebullition and $CH_4$ bubble dissolution fluxes. Calculations were made following several scenarios: two extreme bubble-size scenarios considering a release of many small (3 mm diameter) bubbles or fewer large (10 mm) bubbles, and an intermediate scenario of release of 6 mm diameter bubbles. Delwiche & Hemond (2017) experimentally determined that a large majority (~ 90%) of the bubbles released from lake sediments fall in this size interval.

### 2.5. Potential $CH_4$ flux across the sediment-water interface

The potential $CH_4$ flux across the sediment-water interface was determined from short-term intact core incubations in L. Edward, L. George and L. Nyamusingere only, at the same sampling sites where biogeochemical processes were measured. $CH_4$ flux was quantified from the change of $CH_4$ concentration in overlying waters at 5 different time steps, every 2 hours. Briefly, in every lake, 2 sediment cores (6 cm wide; ~ 30 cm sediment and 30 cm of water) were collected taking care to avoid disturbance at the sediment-water interface. Cores were kept in the dark until back in the laboratory, typically 6h later. Overlying water was carefully removed with a syringe taking care to avoid any disturbance of the sediments. It was replaced by bottom lake water which had been previously filtered through 0.2µm polycarbonate filters (GSWP, Millipore) in order to remove water column methanotrophs and degassed with helium during 20 minutes in order to remove background $O_2$ and $CH_4$.and the filtered and degassed water was gently returned in the core tubes, on top of the sediments. Core tubes were tightly closed with a thick rubber stopper equipped with two sampling valves. A magnetic stirrer placed ~ 10 cm above the sediments

was allowed to rotate gently in order to homogenize the overlying water layer during the incubation. At each time step, 60 ml of overlying water was sampled by connecting a syringe to the first sampling valve while an equivalent volume of degassed water was allowed to flow through the second valve in order to avoid any pressure disequilibrium. Subsamples of overlying water were transferred into a two 20 ml serum bottles filled without headspace and poisoned with $HgCl_2$. Determination of the dissolved $CH_4$ concentration was performed with a GC-FID following the same procedure as described above. The removal of oxygen might have inhibited the biogeochemical activity of aerobic methanotrophs present at the sediment-water interface, hence we considered that the results gathered from this experiment are representative of a potential (maximal) $CH_4$ flux.

**2.6. Primary production and $N_2$ fixation**

Primary production and $N_2$ fixation rates were determined from dual stable isotope photosynthesis-irradiance experiments using $NaH^{13}CO_3$ (Eurisotop) and dissolved $^{15}N_2$ (Eurisotop) as tracers for incorporation of dissolved inorganic carbon (DIC) and $N_2$ into the biomass. The $^{15}N_2$ tracer was added dissolved in water (Mohr et al. 2010). Incident light intensity was measured by a LI-190SB quantum sensor during day time during the entire duration of the sampling campaign. At each station a sample of surface waters (500 ml) was spiked with the tracers (final $^{15}N$ atom excess ~5%). Three subsamples were preserved with $HgCl_2$ in 12-mL Exetainers vials (Labco) for the determination of the exact initial $^{13}C$-DIC and $^{15}N$-$N_2$ enrichment. The rest of the sample was divided into nine 50-ml polycarbonate flasks, filled without headspace. Eight flasks were placed into a floating incubation device providing a range of light intensity (from 0 to 80% of natural light) using neutral density filter screen (Lee Filters). The last one was immediately amended with neutral formaldehyde (0.5% final concentration) and served as killed control sample. Samples were incubated *in situ* during 2 hours around mid-day just below the surface at lake surface temperature. After incubation, biological activity was stopped by adding neutral formaldehyde into the flasks, and the nine samples were filtered on pre-combusted GF/F filters when back in the lab. Glass fiber filters were decarbonated with HCl fumes overnight, dried, and their $\delta^{13}C$-POC and $\delta^{15}N$-PN values were determined with an EA-IRMS (Thermo FlashHT – delta V Advantage). For the measurement of the initial $^{15}N_2$ enrichment, a 2-ml helium headspace was created, and after 12h equilibration, a fraction of the headspace was injected into the above-mentioned EA-IRMS equipped with a Cu column warmed at 640°C and a $CO_2$ trap. Initial enrichment of $^{13}C$-DIC was also measured.

Photosynthetic ($P_i$) (Hama et al. 1983) and $N_2$ fixation ($N_2fix_i$) (Montoya et al. 1996) rates in individual bottles were calculated, and corrected for any abiotic tracer incorporation by subtraction of the killed control value. For each experiment, the maximum photosynthetic and $N_2$ fixation rates ($P_{max}$, $N_2fix_{max}$) and the irradiance at the onset of light saturation ($I_{k\_PP}$, $I_{k\_N2fix}$) were determined by fitting $P_i$ and $N2fix_i$ to the light intensity gradient provided by the incubator ($I_i$) using the equation (eq. 2) for photosynthesis activity (Vollenweider 1965) and (eq. 3) for $N_2$ fixation (Mugidde et al. 2003).

$$P_i = 2P_{max}\left[\frac{I_i/2I_{k\_PP}}{1+(I_i/2I_{k\_PP})^2}\right] \quad (2)$$

$$N_2fix_i = 2N_2fix_{max}\left[\frac{I_i/2I_{k\_N2fix}}{1+(I_i/2I_{k\_N2fix})^2}\right] \quad (3)$$

**2.7. Determination of $CH_4$ oxidation rates.**

$CH_4$ oxidation rates in surface waters (1m depth) were determined from the decrease of $CH_4$ concentrations measured during short (typically < 24h) time course experiments. Samples for $CH_4$ oxidation rate measurement were collected in 60 mL glass serum bottles filled directly from the Niskin bottle with tubing, left to overflow, and immediately closed with butyl stoppers previously boiled in milli-Q water, and sealed with aluminum caps. The first bottle was then poisoned with a saturated solution of $HgCl_2$ (100 µl) injected through the butyl stopper with a polypropylene syringe and a steel needle and corresponded to the initial $CH_4$ concentration at the beginning of the incubation (T0).

The remaining bottles were incubated in the dark, at in situ (~26°C) temperature during ~12h or ~24h except in L. George and Nyamusingere where the incubation was shorter (~6h). At 4 different times step one bottle was poisoned with 100 µL of $HgCl_2$ and stored in the dark until measurement of the $CH_4$ concentrations with the above-mentioned GC-FID. $CH_4$ oxidation rates were calculated as a linear regression of $CH_4$ concentrations over time ($r^2$ generally better than 0.80) during the course of the incubation. $O_2$ consumption was followed in a parallel incubation of water samples in 60 mL biological oxygen demand bottles (Wheaton) to ensure the samples were incubated under oxic conditions during the full course of this short-term experiment.

### 2.8. Sunlight inhibitory effect on $CH_4$ oxidation

The influence of light intensity on methanotrophy was investigated in L. Edward (April 2017 and January 2018), L. Nyamusingere and L. George (January 2018) by means of a stable isotope ($^{13}CH_4$) labelling experiment. For each experiment, 12 serum bottles (60 mL) were filled with lake surface waters (1 m) as described above. All bottles were spiked with 100 µL of a solution of dissolved $^{13}CH_4$ (50 µmol $L^{-1}$ final concentration, 99% enrichment) added in excess. In L. Edward in January. Half of the bottles were amended with 3-(3,4-dichlorophenyl)-1,1-dimethylurea (DCMU, 0.5 mg $L^{-1}$) in order to inhibit photosynthesis (Bishop 1958) and investigate the hypothetical inhibitory effect of dissolved $O_2$ production by phytoplankton. Two bottles were poisoned immediately with pH-neutral formaldehyde (0.5% final concentration) and served as killed controls. The ten others were incubated during 24h at 26°C in a floating device providing 5 different light intensities (from 0 to 80% of natural light using neutral density filter screens (Lee Filters). For every bottle at the end of the incubation, one 12-mL vial (Labco Exetainer) was filled with the water sample and preserved with 50 µL $HgCl_2$. The rest of the sample (~50 mL) was filtered on a precombusted GF/F filter for subsequent $\delta^{13}$C-POC measurement.

$\delta^{13}$C-DIC and $\delta^{13}$C-POC were determined with an EA-IRMS following the method described in Gillikin & Bouillon (2007) and Morana et al. (2015), respectively. The methanotrophic bacterial production, defined at the $CH_4$-derived $^{13}C$ incorporation rates into the POC pool was calculated as in eq. (4) (Morana et al. 2015):

$$MBP = \frac{POC_t \times (\%^{13}CPOC_t / \%^{13}CPOC_i)}{t \times (\%^{13}CCH_4 / \%^{13}CPOC_i)} \quad (4)$$

Where $POC_t$ is the concentration of POC after incubation, $\%^{13}C\text{-}POC_t$ and $\%^{13}C\text{-}POC_i$ are the final and initial percentage of $^{13}C$ in the POC, t is the incubation time and $\%^{13}C\text{-}CH_4$ is the percentage of $^{13}C$ in $CH_4$ after the inoculation of the bottles with the tracer. Similarly, the methanotrophic bacterial respiration rates, defined as the $CH_4$-derived $^{13}C$ incorporation rates into the DIC pool, were calculated as in eq. (5):

$$MBR = \frac{DIC_t \times (\%^{13}CDIC_t / \%^{13}CDIC_i)}{t \times (\%^{13}CCH_4 / \%^{13}CDIC_i)} \quad (5)$$

Where $DIC_t$ is the concentration of DIC after the incubation, $\%^{13}C\text{-}DIC_t$ and $\%^{13}C\text{-}DIC_i$ are the final and initial percentage of $^{13}C$ in DIC and $\%^{13}C\text{-}CH_4$ is the percentage of $^{13}C$ in $CH_4$ after the inoculation of the bottles with the tracer.

Potential $CH_4$ oxidation rates (MOX) were calculated as the sum of MBP and MBR rates. The fraction (%) of MOX inhibited by light was calculated at every light intensity as (eq.6):

$$MOX_{inihibition}(\%) = (1 - MOX_i / MOX_{dark}) \times 100 \quad (6)$$

Where $MOX_i$ is the potential $CH_4$ oxidation for a given light treatment and $MOX_{dark}$ is the potential $CH_4$ oxidation in the dark.

### 2.9. Determination of pelagic $CH_4$ production rates.

Time course $^{13}C$ tracer experiments were carried out in well oxygenated surface waters at every sampling site. Measurement of the isotopic enrichment of the $CH_4$ during this experiment allowed to estimate production rates of $CH_4$ issued from 3 different precursors: $^{13}C$-DIC ($NaH^{13}CO_3$), $^{13}C_{(1,2)}$-acetate and $^{13}C_{methyl}$-methionine. Serum bottles (60 ml) were spiked with 1 ml of $^{13}C$ tracer solution, or with an equivalent volume of distilled water for the control treatment. $NaH^{13}CO_3$ was added in the bottles at a tracer level (less than 5% of ambient $HCO_3^-$ concentration) while $^{13}C_{(1,2)}$-acetate and $^{13}C_{methyl}$-methionine were added largely in excess (>99% of ambient concentration). Therefore, we assume the $CH_4$ production rates measured from $^{13}C$-DIC could be representative of in-situ rates, but the production rates measured from $^{13}C$-acetate and $^{13}C$-methionine should instead be viewed as potential rates. The exact amount of $^{13}C$-DIC added in the bottles was determined filling a borosilicate 12 ml exetainer vials preserved and analysed for $\delta^{13}C$-DIC as described above.

The control bottles and the bottles amended with the different $^{13}C$ tracer were incubated under constant temperature conditions (26°C) following three different treatments : (1) one third were incubated under constant light (PAR of ~ 200 µmol photon $m^{-2}$ $s^{-1}$), (2) another third were incubated under the same light intensities conditions but were first amended with DCMU (0.5 mg $L^{-1}$ ; final concentration), an inhibitor of photosynthesis, (3) and the last third were incubated in the dark.

At each time step (typically every 6-12h, 5-time steps), the biological activity was stopped by adding 100 µL of a saturation solution of $HgCl_2$. Bottles were kept in the dark until $CH_4$ concentration measurement and $\delta^{13}C$-$CH_4$ determination as described above. $O_2$ consumption was followed in a parallel incubation of water samples in 60 mL biological oxygen demand bottles (Wheaton) to ensure the samples were incubated under oxic conditions during the full course of the experiment.

The term $CH_{4\_prod}$ (nmol $L^{-1}$ $h^{-1}$) defined as the amount of $CH_4$ produced from a specific tracer during a time interval t (h), was calculated following this equation (eq. 7) derived from Hama et al. (1983):

$$CH_{4\_prod} = \frac{CH_{4\_t} \times (\%^{13}CCH_{4\_t}/\%^{13}CCH_{4\_i})}{t \times (\%^{13}Ctracer/\%^{13}CCH_{4\_i})} \quad (7)$$

Where $CH_{4\_t}$ and $\%^{13}CCH_{4\_t}$, represent the $CH_4$ concentration (nmol $L^{-1}$) and the $\%^{13}C$ atom of the $CH_4$ pool at a given time step, respectively. $\%^{13}CCH_{4\_i}$ represent the $\%^{13}C$ atom of the pool of $CH_4$ at the beginning of the experiment. $\%^{13}Ctracer$ represent the $\%^{13}C$ atom of the isotopically enriched pool of the precursor molecule tested ($NaHCO_3$, methionine or acetate, depending of the treatment). $\%^{13}C$-tracer was assumed constant during the full course of the incubation given the high concentration of ambient DIC in the sampled lakes (~ 2 mmol $L^{-1}$ in L. George, > 6 mmol $L^{-1}$ in the other lakes) and that acetate and methionine were spiked in large excess (>99%).

### 2.10. DNA extraction

Surface water sample for DNA analysis (between 1 L and 0.15 L, depending on the biomass) were first filtered through 5.0 $\mu$m pore size polycarbonate filters (Millipore). The eluent was then subsequently filtered through 0.2 $\mu$m pore size polycarbonate filters (Millipore) to retain free living prokaryotes. Filters were stored frozen (-20°C) immerged in a lysis buffer until processing in the laboratory. Total DNA was extracted from the 0.2 $\mu$m and 5.0 $\mu$m 47 mm filters using DNeasy PowerWater kit (Qiagen) following the manufacturer's instructions. Quality and quantity of the extracted DNA were estimated using the NanoDrop ND-1000 spectrophotometer (ThermoFisher) and the Qubit 3.0 fluorometer (Life technology). Extracted DNA was stored at -20 °C until further use.

### 2.11. Quantification of mcrA via qPCR

Quantification of *mcrA* gene copies was performed by quantitative PCR (qPCR) on the total extracted DNA. The used primer pair consisted of forward primer *qmcrA-F* 5'-TTCGGTGGATCDCARAGRGC-3'and *qmcrA-R* 5'-GBARGTCGWAWCCGTAGAATCC-3' (Denman et al. 2007). The reaction mixture contained 3 µL of total community

DNA extract, 7.5 µL ABsolute qPCR SYBR Green Mix (ThermoFisher, Cat. AB1158B), 0.3 µL of 10 µM forward primer *mcrF*, 0.3 µL of 10 µM reverse primer *mcrR*, 1.5 µL of a 1% w/v Bovine Serum Albumin solution (Amersham Bioscience) and 2.4 µL of Nuclease/DNA-free water. The qPCR was performed in a Rotorgene 3000 (Corbett Research) using the following conditions: 95 °C (15 min) followed by 40 cycles of 20 s at 95 °C, 20 s at 58 °C, 20 s at 72 °C and a final extension step of 5 s at 80 °C. Standard curves were prepared from serial dilutions of a prequantified *mcrA* PCR fragment amplified using primers mcrF and mcrR from a plasmid extract carrying the complete *mcrA* gene using concentrations ranging from $1x10^2$ to $1x10^8$ copies $\mu L^{-1}$. Samples were analyzed in triplicates.

### 2.12. 16S rRNA gene amplicon sequencing

Sequencing of the 16S rRNA gene was done on the total extracted DNA to assess community composition. 16S rRNA gene sequencing was done with the Illumina MiSeq v3 Chemistry following the "16S Metagenomic Sequencing Library Preparation" protocol with the following universal 16S rRNA gene primers targeting the V4 region, forward UniF/A519F-(S-D-Arch-0519-a-S-15) 5'-CAGCMGCCGCGGTAA-3' and reverse UniR/802R-(S-D-Bact-0785-b-A-18) 5'-TACNVGGGTATCTAATCC-3' (Klindworth et al. 2013). Sequenced read quality was checked using FastQC v0.11 (https://www.bioinformatics.babraham.ac.uk/projects/fastqc/). Short reads were trimmed to 250 bp with FastX Toolkit v0.0.13 (http://hannonlab.cshl.edu/fastx_toolkit/) in order to remove trailing Ns and low quality bases. Operational Taxonomical Units (OTU) for each analyzed sample were obtained from the quality trimmed reads using mothur v1.39.5 (Kozich et al. 2013) and following the online MiSeq SOP (https://mothur.org/wiki/MiSeq_SOP - accessed April 2018) using the Silva v128 16S rRNA database with the following parameters: *maxambig = 0 bp*; *maxlength = 300 bp*; *maxhomop = 8*; and classify OTUs to 97% identity. Generated OTU table was used to calculate relative abundances of each OTU per sample.

## 3. Results

### 3.1. Environmental settings

The sampled lakes cover a wide range of size (<1 to 2300 km²), maximum depth (3-117 m), mixing regimes, phytoplankton biomass and primary productivity (Table S1, Fig. S1). Phytoplankton biomass (Chlorophyll-a from 3.6 µg $L^{-1}$ to 190.2 µg $L^{-1}$) was dominated by Cyanobacteria (>95%) in the most productive lakes, while Diatoms (<20%) and Chrysophytes (<40%) also contributed in the less productive ones (Fig. S2). Maximum potential photosynthetic activity ($P_{max}$) varied from 1.5 µmol C $L^{-1}$ $h^{-1}$ in L. Edward to 199.0 µmol C $L^{-1}$ $h^{-1}$ in L. George and was linearly related to chlorophyll a concentration. Light-dependent $N_2$ fixation was detected in every lake with the exception of L. Kyamwinga. No significant $N_2$ fixation rates were measured in the dark. Maximum potential $N_2$ fixation rates ($N_2fix_{max}$) ranged between 1 nmol $L^{-1}$ $h^{-1}$ and 128 nmol $L^{-1}$ $h^{-1}$ and were positively related to $P_{max}$ (Fig. S1).

We detected and quantified the abundance of the archaeal alpha subunit of methyl-coenzyme M reductase gene (*mcrA*), a proxy for methanogens, in the surface waters of each lake. *mcrA* gene copy abundance (*mcrA* copy ng $DNA^{-1}$) ranged between 319 ± 41 (L. Edward) and 7537 ± 476 (L. Katinda) in the fraction of seston < 5 µm, and between 541 ± 19 (L. Edward) and 7968 ± 167 (L. Katinda) in the fraction of seston > 5 µm (Fig. S3). Illumina 16S rRNA gene amplicon sequencing indicated that methanogens accounted for a small fraction of the prokaryotic community in the surface waters of L. Edward (0.01 %), L. Kyamwinga (0.03 %) and L. Nyamunsingere (0.08 %). They represented a substantially higher fraction of the community in L. Katinda (0.38%) and L. George (0.57 %) (Fig. S4). In all lakes, hydrogenotrophic (*Methanomicrobiales* and *Methanobacteriales*) were always more abundant than acetoclastic (*Methanosarcinales*) microorganisms, representing at least 65% of the methanogens (up to 95% in L. Katinda, Fig. S4).

### 3.2. Water column $CH_4$ concentration and $\delta^{13}C$-$CH_4$ patterns

Surface waters were super-saturated in $CH_4$ in all lakes, with surface concentrations (at 1 m) ranging between 78 and 652 nmol $L^{-1}$ (atmospheric equilibrium ~ 2 nmol $L^{-1}$). Diffusive $CH_4$ emissions varied between 0.05 mmol $m^{-2}$ $d^{-1}$ (L. Edward)

and 0.40 mmol $m^{-2}$ $d^{-1}$ (L. Katinda). The benthic $CH_4$ flux across the sediment-water interface was elevated in comparison with the diffusive $CH_4$ emissions in the three lakes where it was measured: L. Edward (0.96 mmol $m^{-2}$ $d^{-1}$), L. George (9 mmol $m^{-2}$ $d^{-1}$) and L. Nyamusingere (5 mmol $m^{-2}$ $d^{-1}$). $CH_4$ ebullition was the dominant pathway of $CH_4$ evasion to the atmosphere in the 3 lakes (L. Edward, L. George, L. Nyamusingere) where ebullitive fluxes were investigated. Ebullitive flux in L. Edward ranged between 0.16 and 0.24 mmol $m^{-2}$ $d^{-1}$ depending of the bubble-size scenario considered (see material & methods), being at least 4 times higher than the diffusive $CH_4$ flux. This discrepancy was even larger in the shallower L. George (13.26-13.9. mmol $m^{-2}$ $d^{-1}$) and L. Nyamusingere (19.03-19.09 mmol $m^{-2}$ $d^{-1}$) where $CH_4$ ebullition appeared ~100 and ~50 times higher than diffusive $CH_4$ emissions. During the ascent of CH4 bubbles to the surface, gas exchange occurs, and we estimated that, depending of the $CH_4$ bubble size considered, between 0.04 and 0.21 mmol $CH_4$ $m^{-2}$ $d^{-1}$ dissolved in the water column of L. Edward during bubble ascent. This bubble dissolution flux ranged between 0.70 – 3.30 mmol $m^{-2}$ $d^{-1}$ and 1.21 – 5.55 mmol $m^{-2}$ $d^{-1}$ in L. George and L. Nyamusingere, respectively.

Vertical patterns of $CH_4$ and stable carbon isotope composition of $CH_4$ ($\delta^{13}C$-$CH_4$) were variable among the different lakes. In L. Kyamwinga and Katinda, higher $CH_4$ concentrations and lower $\delta^{13}C$-$CH_4$ values were observed in the well-oxygenated epilimnion compared to the metalimnion showing a source of relatively $^{13}C$-depleted $CH_4$ to the epilimnetic $CH_4$ pool (Fig. 1). The $CH_4$ concentrations and $\delta^{13}C$-$CH_4$ were homogeneous in the water column of L. Edward that is much larger than the other studied lakes (2300 km², Table S1) and characterized by a higher wind exposure and a substantially weaker thermal stratification (Fig. 1). However, a clear horizontal gradient in $CH_4$ concentration and $\delta^{13}C$-$CH_4$ occurred between the littoral and pelagic zones (Fig. S5). Vertical gradients were also observed at much smaller scale in the near sub-surface (top 0.3 m) in the shallow and entirely well oxygenated L. George and L. Nyamusingere (Fig. 2). In both lakes $CH_4$ concentrations were relatively modest in the hypolimnion (< 50 nmol $L^{-1}$) but increased abruptly in the thermal gradient (0.3 m interval) to reach a surface maximum > 240 nmol $L^{-1}$ (Fig. 2). $\delta^{13}C$-$CH_4$ mirrored this pattern with significantly lower values in surface than at the bottom of the water column indicating that a source of relatively $^{13}C$-depleted $CH_4$ contributed to the higher epilimnitic $CH_4$.

### 3.3. Occurrence of microbial $CH_4$ production in surface waters under oxic conditions

Despite the prevalence of oxic conditions, $^{13}C$-labelling experiments revealed that $CH_4$ was produced in surface waters of each lake with the exception of L. Kyamwinga (Fig. 3). The kinetic of incorporation of $NaH^{13}CO_3$ into the $CH_4$ pool revealed that a substantially higher amount of $CH_4$ was produced from dissolved inorganic carbon (DIC) in illuminated waters, and this mechanism of $CH_4$ formation appears to be related to photosynthesis, as none or only modest quantities of $CH_4$ were produced from $^{13}C$-labelled DIC under darkness or when photosynthesis was inhibited by DCMU (Figs. 3a and S6). Furthermore, $CH_4$ production from DIC appeared strongly correlated ($r^2 = 0.91$) to the photosynthetic activity (Fig. 4a) and $N_2$ fixation rates (Fig 4b), supporting the view that $CH_4$ formation in oxic waters was directly linked to phytoplankton metabolism (Bizic et al. 2020).

Aside from DIC, an appreciable amount of $CH_4$ was generated in all lakes from the sulfur bonded methyl group of methionine when bottles were incubated under light, irrespective of the addition of DCMU (Fig. 3b and S6), that were approximately 4 times higher than in the dark. In addition, a positive relationship between $CH_4$ production from methionine in the light and the photosynthetic activity was found (Fig. 4c).

$^{13}C$-labelled acetate, the substrate of acetoclastic methanogenesis, supported the production of $CH_4$ in all lakes with the exception of L. Kyamwinga, but at much lower rates compared to light-dependent $CH_4$ production from DIC (50 times lower, n=7) or methionine (10 times lower, n=4) (Fig. 3c and S6). $\delta^{13}C$ analysis of the DIC in the bottles spiked with $^{13}C$-labelled acetate showed that the acetate was mineralized at rates of 5-6 orders of magnitude higher than acetoclastic methanogenesis so that added acetate appeared to be used almost exclusively by heterotrophic micro-organisms other than methanogens. Pattern of acetate-derived production of $CH_4$ were similar in light and dark treatments (Figs. 3c and S6) and this mode of $CH_4$ production appeared unrelated to phytoplankton activity (Fig. 4d).

### 3.4. Microbial methane oxidation

Net $CH_4$ oxidation was detected in all 5 investigated lakes ranging from 11 to 5212 nmol $L^{-1}$ $d^{-1}$ (Fig. 5), and was by far the largest loss term of dissolved $CH_4$ at ecosystem scale (8 to 46 times higher than the diffusive emission to the atmosphere). Surface water $CH_4$ turnover times were particularly short in the shallow and eutrophic L. George (2h) and L. Nyamusingere (3h) and slightly longer in the deeper and less productive L. Katinda (11h), L. Kyamwinga (77h) and L. Edward (100h). In all studied lakes, the volumetric $CH_4$ oxidation rates were always much higher than the volumetric $CH_4$ production we measured during the stable isotope tracer experiments, regardless of the $CH_4$ precursors tested. Pelagic $CH_4$ production rates represented 8.5%, 2.6%, 0.2% and 0.1% of $CH_4$ oxidation rates, in L. Edward, George, Katinda and Nyamusingere, respectively.

The influence of light on methanotrophy was investigated in the deep L. Edward, and shallow L. George and L. Nyamusingere, revealing that $CH_4$ oxidation rates decreased dramatically with increasing light intensity (Fig. 5). For instance, when exposed to full sunlight intensity, methanotrophs consumed only 42% (L. Edward), 54% (L. Nyamusingere) or 74% (L. George) of the $CH_4$ they were able to oxidize in the dark. The magnitude of this sunlight-induced inhibition decreased substantially with decreasing sunlight intensities (Fig. 5). In L. Edward (April 2017) sunlight inhibition of methanotrophy followed the same pattern of lower rates at high sunlight intensities in the bottles where $O_2$ production (via photosynthesis) was stopped by DCMU addition.

## 4. Discussion

### 4.1. Mechanisms of $CH_4$ production under aerobic conditions

The results from the stable isotope labelling experiment highlight that only a minimal fraction of the $CH_4$ produced under aerobic conditions originated from acetate in contrast with several earlier studies (Bogard et al. 2014, Donis et al. 2017) which proposed, based on the apparent fractionation factor of $\delta^{13}C\text{-}CH_4$, that acetoclastic methanogenesis linked to phytoplankton production of organic matter would be the dominant biochemical pathway of pelagic $CH_4$ production in oxic freshwaters. Instead, our results support the study of Bizic et al. (2020) and suggest that epilimnetic $CH_4$ production in well-oxygenated conditions was related to DIC fixation by photosynthesis (Fig. 3), and correlated to primary production (Fig. 4a) and $N_2$ fixation (Fig 4b). When normalized to POC concentrations, the average DIC-derived $CH_4$ production rates ($0.08 \pm 0.05$ nmol $mmol_{POC}^{-1}$ $h^{-1}$ n = 7) was remarkably similar to the $CH_4$ production rates recently reported in Cyanobacteria cultures ($0.04 \pm 0.02$ nmol $mmol_{POC}^{-1}$ $h^{-1}$) grown at 30°C, among which the freshwater *Microcystis aeruginosa* (Bizic et al. 2020), the dominant Cyanobacterium species in the tropical lakes investigated in our study (see Fig S2). These $CH_4$ production rates are 2 orders of magnitude higher than rates reported in an axenic culture of the eukaryote *Emiliania huxleyi* ($0.19 \pm 0.07$ pmol $mmol_{POC}^{-1}$ $d^{-1}$) (Lenhart et al. 2016), but they are 4 orders of magnitude lower than typical anoxic $CH_4$ production rates by methanogenic Archaea (Mountford & Asher 1979). Although it seems improbable that $^{13}C$-DIC acted as a direct precursor molecule for the $CH_4$ released by phytoplankton (Lenhart et al. 2016, Klintzsch et al. 2019) $^{13}C$-DIC could have been taken up by phytoplankton cells and then used as a C source for the synthesis of many different organic molecules that may serve as the actual $CH_4$ precursors. Indeed, healthy phytoplankton cells actively release a variety of low molecular weight molecules which are generally highly labile and rapidly consumed (Baines & Pace 1991, Morana et al. 2014). Phytoplankton metabolism could have fuelled $CH_4$ production pathways, at least partially, excreting substrates involved in $CH_4$ production via biochemical processes such as demethylation of a variety of organic molecules like methionine, one of the S-bonded methylated amino acids (Lenhart et al. 2016), trimethylamine (Bizic et al 2018), or methylphosphonate (Yao et al. 2016).

While the source of methylphosphonate in freshwaters is obscure and its actual natural abundance remains to be determined, dissolved free amino acids would represent up to 4% of the DOC produced by phytoplankton and are rapidly consumed by heterotrophic bacteria (Sarmento et al. 2013). Our incubations indeed demonstrated that the methyl group of

methionine was a potential precursor of $CH_4$ in all lakes investigated, in line with recent findings showing that *Emiliania huxleyi* could act as a direct source of $CH_4$ in oxic conditions using methionine as precursor, without involvement of any other micro-organisms (Lenhart et al. 2016). We found that $CH_4$ production from methionine was clearly stimulated under light, even when photosynthetic activity was inhibited by DCMU, while little $CH_4$ from methionine was produced in darkness (Fig. 3b). DCMU notably prevents reduction of plastoquinone at photosystem II and generates singlet oxygen (Petrillo et al. 2014). The mechanism of $CH_4$ production from methionine is still unclear, but its residue in proteins is particularly sensitive to oxidation to methionine sulfoxide by radical oxygen species (ROS) (Levine et al. 1996) so that methionine would act as an effective ROS scavenger and play important protective roles under photooxidative stress conditions, as shown in vascular plants (Bruhn et al. 2012). The side chain of methionine sulfoxide is identical to dimethyl sulfoxide which is known to react with hydroxyl radicals (OH) to form $CH_4$ (Repine et al. 1979). Besides its photoprotective role for phytoplankton, methionine could also be catabolized by a wide variety of microorganisms to methanethiol, which could in turn be transformed to $CH_4$ as shown in Arctic Ocean surface waters (Damm et al. 2010). Nevertheless, occurrence of this latter mechanism in the tropical lakes investigated seems unlikely as this mode of $CH_4$ production would be expected to be insensitive to light irradiance and no $CH_4$ was produced from methionine in the dark during the incubations.

### 4.2. Relevance of pelagic $CH_4$ production compared to methanotrophy and $CH_4$ emissions at ecosystem scale

The stable isotope labelling experiments revealed that microbial $CH_4$ oxidation largely exceed the pelagic $CH_4$ production.All of the major sources and sinks of $CH_4$ at ecosystem scale were experimentally determined offshore in three lakes (L. Edward at 20 m depth, George at 2.5 m and Nyamusingere at 3 m) (Fig. 6). In contrast with high latitude lakes where no ebullition could be detected in location with water column depth higher than 2 m (DelSontro et al. 2018), we found that the $CH_4$ ebullition flux dominates over the $CH_4$ diffusion flux in every sampling site. However, due to the dissolution of arising bubbles the contribution of ebullition to the total $CH_4$ emissions might be lower in deeper (> 20 m) location of L. Edward (maximum depth of 113m), as shown elsewhere (DelSontro et al. 2015). Comparison of the $CH_4$ production, consumption and emission fluxes (Fig 6.) show that the, depth-integrated $CH_4$ production rates determined from diverse precursors molecules were modest relative to the diffusive $CH_4$ efflux to the atmosphere and the depth-integrated microbial $CH_4$ oxidation.. In opposition, the combined $CH_4$ bubble dissolution flux and diffusive benthic $CH_4$ flux were several orders of magnitude higher than $CH_4$ production in surface waters, and were sufficient to support the microbial $CH_4$ oxidation and the emissions to the atmosphere(Fig. 6). These results gathered in tropical lakes of various size (from 0.44 to 2300 $km^2$) and depth are in sharp contrast with the estimation of an empirical model (Gunthel et al. 2019) which proposed that mechanisms of oxic $CH_4$ production represents the majority of $CH_4$ emissions in lakes larger than 1 $km^2$. This discrepancy highlights the need to consider the unique limnological characteristics of a vast region of the world that harbours 16% of the total surface of lakes (Lehner & Doll 2004). One of the most distinctive features of tropical aquatic environment is the persistent elevated water temperature in the hypolimnion and at the water-sediment interface which favours methanogenic activity in sediment and decreases $CH_4$ solubility, enhancing bubbles formation.

Epilimnetic $CH_4$ production appeared as a marginal flux at the ecosystem scale and could not explain alone the accumulation of $^{13}C$-depleted $CH_4$ in the epilimnion of most of the lakes of our dataset (Figs. 1, 2), for which we propose a combination of two other alternative mechanisms: dissolution of arising $CH_4$ bubbles in the epilimnion combined with inhibition by light of $CH_4$ oxidation. The partial dissolution of the $CH_4$ bubbles as they rise in the epilimnion would allow a rapid transport of $^{13}C$-depleted $CH_4$ from the sediment, bypassing the hotspot of $CH_4$ oxidation at the sediment-water interface and representing an alternative source of $^{13}C$-depleted $CH_4$ in water column. This mechanism of bubble-mediated transport to the epilimnion would be especially important in shallow lakes as CH4 ebullition is widely variable in function of water column depth (DelSontro et al. 2015). The shallower L. George and L. Nyamusingere were notably characterized by sharp thermal density gradients (Fig. 2) and extreme phytoplankton biomass largely dominated by *Microcystis aeruginosa* (Chlorophyll *a* up

to 190 µg $L^{-1}$). *Microcystis aeruginosa* cells form large aggregates (>1 mm) embedded in a matrix of extracellular polymeric substance that might act as a barrier to trap small $CH_4$ bubbles arising from the sediment (Fig S7). Dissolution of $CH_4$ bubbles could be enhanced at the very near surface due to the entrapments of bubbles at the air-water interface by abundant surface organic films that delay the bubble "burst". The presence of a sharp sub-surface temperature gradient would further enhance $CH_4$ accumulation during day-time near the air-water interface (by blocking vertical redistribution of $CH_4$ by mixing). We hypothesize that this process could be widespread in shallow tropical lakes which are characterized by high productivity and are susceptible to be simultaneous large benthic $CH_4$ sources.

The stable isotope labelling experiment carried out to investigate the inhibitory effect of light on methanotrophic activity demonstrated that $CH_4$ consumption dramatically decreased with increasing light intensities (Fig. 5), as already reported in a tropical reservoir (Dumestre et al. 1999) and Lake Biwa (Murase et al. 2005). The inhibitory effect of light on freshwater methanotrophs remains surprisingly understudied since it was first reported 20 years ago (Dumestre et al. 1999), so that the physiological mechanism of photoinhibition is still not understood. The physiological mechanism of photoinhibition of $CH_4$ oxidation could be related to the fact that the copper-containing methane monooxygenase enzyme and structurally close to the ammonia monooxygenase enzyme, and might be inactivated by ROS produced during photooxidative stress, as shown for ammonium oxidizers (French et al. 2012, Tolar et al. 2016). Altogether, our results emphasize the role of sunlight irradiance as an important, but frequently overlooked, environmental factor driving the $CH_4$ dynamics in lake surface waters and possibly contributing to the occurrence of $^{13}C$ depleted $CH_4$ in surface waters.

## 5. Supplement

Supplementary figures are available on-line on the *Biogeosciences* website.

## 6. Data availability

All data included in this study are available upon request by contacting the corresponding author.

## 7. Author contributions

This study was designed by C. Morana, A.V. Borges & S. Bouillon. All authors participated to samples collection, data acquisition and analysis. The manuscript was drafted by C. Morana and substantially improved with the inputs of all authors. All authors approved the final version of the manuscript.

## 8. Competing interests

The authors declare that they have no conflict of interest.

## 9. Acknowledgments

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## 10. References

Bastviken, D., Ejlertsson, J., Sundh, I., & Tranvik, L.: Methane as a source of carbon and energy for lake pelagic food webs. Ecology, 84(4), 969-981, 2003

Bastviken, D., Tranvik, L. J., Downing, J. A., Crill, P. M., & Enrich-Prast, A.: Freshwater methane emissions offset the continental carbon sink, Science, 331(6013), 50-50, 2011

Baines, S. B., & Pace, M. L: The production of dissolved organic matter by phytoplankton and its importance to bacteria: patterns across marine and freshwater systems. Limnology and Oceanography, 36(6), 1078-1090, doi: 10.4319/lo.1991.36.6.1078, 1991

Bishop, N. I.: The influence of the herbicide, DCMU, on the oxygen-evolving system of photosynthesis. Biochimica et biophysica acta, 27(1), 205-206, doi: 10.1016/0006-3002(58)90313-5, 1958

Bižić, M., Ionescu, D., Günthel, M., Tang, K. W., & Grossart, H. P.: Oxic Methane Cycling: New Evidence for Methane Formation in Oxic Lake Water. Biogenesis of Hydrocarbons, 1-22, 2018

Bižić, M., Klintzsch, T., Ionescu, D., Hindiyeh, M. Y., Günthel, M., Muro-Pastor, A. M., Eckert, W., Urich, T., & Grossart, H. P.: Aquatic and terrestrial cyanobacteria produce methane. Science advances, 6(3), eaax5343, doi: 10.1126/sciadv.aax5343, 2020

Bogard, M. J., Del Giorgio, P. A., Boutet, L., Chaves, M. C. G., Prairie, Y. T., Merante, A., & Derry, A. M.: Oxic water column methanogenesis as a major component of aquatic CH4 fluxes. Nature communications, 5, 5350, doi: 10.1038/ncomms6350, 2014

Borges, A. V., Darchambeau, F., Teodoru, C. R., Marwick, T. R., Tamooh, F., Geeraert, N., Morana, C., Okuku, E. & Bouillon, S.: Globally significant greenhouse-gas emissions from African inland waters. Nature Geoscience, 8(8), 637-642, 2015

Bruhn, D., Møller, I. M., Mikkelsen, T. N., & Ambus, P.: Terrestrial plant methane production and emission. Physiologia plantarum, 144(3), 201-209, doi: 10.1111/j.1399-3054.2011.01551, 2012

Cole, J. J., & Caraco, N. F.: Atmospheric exchange of carbon dioxide in a low-wind oligotrophic lake measured by the addition of SF6. Limnology and Oceanography, 43(4), 647-656, doi: 10.4319/lo.1998.43.4.0647, 1998

Damm, E., Helmke, E., Thoms, S., Schauer, U., Nöthig, E., Bakker, K., & Kiene, R. P.: Methane production in aerobic oligotrophic surface water in the central Arctic Ocean. Biogeosciences, 7(3), 1099-1108, doi: 10.5194/bg-7-1099-2010, 2010

DelSontro, T., McGinnis, D. F., Wehrli, B., and Ostrovsky, I.: Size does matter: Importance of large bubbles and small-scale hot spots for methane transport. Environmental science & technology, doi.org/10.1021/es5054286, 2015

DelSontro, T., del Giorgio, P. A., & Prairie, Y. T.: No longer a paradox: the interaction between physical transport and biological processes explains the spatial distribution of surface water methane within and across lakes. Ecosystems, 21(6), 1073-1087, doi: 10.1007/s10021-017-0205-1, 2018

Delwiche, K. B., and Hemond, H. F.: Methane bubble size distributions, flux, and dissolution in a freshwater lake. Environmental Science & Technology, doi.org/10.1021/acs.est.7b04243, 2017

Denman, S. E., Tomkins, N. W., & McSweeney, C. S.: Quantitation and diversity analysis of ruminal methanogenic populations in response to the antimethanogenic compound bromochloromethane. FEMS microbiology ecology, 62(3), 313-322, doi: 10.1111/j.1574-6941.2007.00394.x, 2007

Descy, J. P., Darchambeau, F., Lambert, T., Stoyneva-Gaertner, M. P., Bouillon, S., & Borges, A. V.: Phytoplankton dynamics in the Congo River. Freshwater Biology, 62(1), 87-101, doi: 10.1111/fwb.1285, 2017

Donis, D., Flury, S., Stöckli, A., Spangenberg, J. E., Vachon, D., & McGinnis, D. F. (2017). Full-scale evaluation of methane production under oxic conditions in a mesotrophic lake. Nature communications, 8(1), 1661, doi: 10.1038/s41467-017-01648-4, 2017

Dumestre, J. F., Guézennec, J., Galy-Lacaux, C., Delmas, R., Richard, S., and Labroue, L.: Influence of light intensity on methanotrophic bacterial activity in Petit Saut Reservoir, French Guiana. Applied and environmental microbiology, doi: 10.1128/AEM.65.2.534-539, 1999

Fernández, J. E., Peeters, F., & Hofmann, H.: On the methane paradox: Transport from shallow water zones rather than in situ methanogenesis is the major source of CH4 in the open surface water of lakes. Journal of Geophysical Research: Biogeosciences, 121(10), 2717-2726, doi: 10.1002/2016JG003586, 2016

French, E., Kozlowski, J. A., Mukherjee, M., Bullerjahn, G., & Bollmann, A.: Ecophysiological characterization of ammonia-oxidizing archaea and bacteria from freshwater. Applied and Environmental Microbiology, 78(16), 5773-5780, doi: 10.1128/aem.00432-12, 2012

Gillikin, D. P., and Bouillon, S.: Determination of delta O-18 of water and delta C-13 of dissolved inorganic carbon using a simple modification of an elemental analyzer-isotope ratio mass spectrometer: an evaluation. Rapid Communications in Mass Spectrometry, doi: 10.1002/rcm.2968, 2007

Greinert, J. and McGinnis, D.F.: Single Bubble Dissolution Model: The Graphical User Interface SiBu-GUI. Environmental Modelling & Software, doi:10.1016/j.envsoft.2008.12.011, 2009

Grossart, H. P., Frindte, K., Dziallas, C., Eckert, W., & Tang, K. W.: Microbial methane production in oxygenated water column of an oligotrophic lake. Proceedings of the National Academy of Sciences, 108(49), 19657-19661, doi: 10.1073/pnas.1110716108, 2011

Günthel, M., Donis, D., Kirillin, G., Ionescu, D., Bizic, M., McGinnis, D. F., Grossart, H. P., & Tang, K. W.: Contribution of oxic methane production to surface methane emission in lakes and its global importance. Nature communications, 10(1), doi: 10.1038/s41467-019-13320-0, 2019

Hama, T., Miyazaki, T., Ogawa, Y., Iwakuma, T., Takahashi, M., Otsuki, A., & Ichimura, S.: Measurement of photosynthetic production of a marine phytoplankton population using a stable 13C isotope. Marine Biology, 73(1), 31-36, doi: 10.1007/BF00396282, 1983

Hofmann, H., Federwisch, L., & Peeters, F.: Wave-induced release of methane: littoral zones as source of methane in lakes. Limnology and Oceanography, 55(5), 1990-2000, doi: 10.4319/lo.2010.55.5.1990 2010

Klindworth, A., Pruesse, E., Schweer, T., Peplies, J., Quast, C., Horn, M., & Glöckner, F. O.: Evaluation of general 16S ribosomal RNA gene PCR primers for classical and next-generation se,quencing-based diversity studies. Nucleic acids research, 41(1), doi: 10.1093/nar/gks808, 2013

Klintzsch, T., Langer, G., Nehrke, G., Wieland, A., Lenhart, K., & Keppler, F.: Methane production by three widespread marine phytoplankton species: release rates, precursor compounds, and potential relevance for the environment. Biogeosciences, 16(20), 4129-4144, doi: 10.5194/bg-16-4129-2019, 2019

Kosten, S., Huszar, V. L., Bécares, E., Costa, L. S., van Donk, E., Hansson, L. A., Jeppesen, E., Kruk, C., Lacerot, G., Mazzeo, N., De Meester, L., Moss, B., Lurling, M., Noges, T., Romo, S., & Scheffer, M.: Warmer climates boost cyanobacterial dominance in shallow lakes. Global Change Biology, 18(1), 118-126, doi: 0.1111/j.1365-2486.2011.02488.x, 2012

Kozich, J. J., Westcott, S. L., Baxter, N. T., Highlander, S. K., & Schloss, P. D.: Development of a dual-index sequencing strategy and curation pipeline for analyzing amplicon sequence data on the MiSeq Illumina sequencing platform. Applied and environmental microbiology, AEM-01043, doi: 10.1128/aem.01043-13, 2013

Lehner, B., Döll, P.: Development and validation of a global database of lakes, reservoirs and wetlands. Journal of Hydrology, 296: 1-22, doi: 10.1016/j.jhydrol.2004.03.028, 2004

Lenhart, K., Klintzsch, T., Langer, G., Nehrke, G., Bunge, M., Schnell, S., & Keppler, F.: Evidence for methane production by the marine algae Emiliania huxleyi. Biogeosciences, 13(10), 3163-3174, doi:10.5194/bg-13-3163-2016, 2016

Levine, R. L., Mosoni, L., Berlett, B. S., & Stadtman, E. R.: Methionine residues as endogenous antioxidants in proteins. Proceedings of the National Academy of Sciences, 93(26), 15036-15040, doi: 10.1073/pnas.93.26.15036, 1996

Lewis Jr, W. M.: Tropical limnology. Annual review of ecology and systematics, 18(1), 159-184, doi: 10.1146/annurev.es.18.110187.001111, 1987

Martinez-Cruz, K., Sepulveda-Jauregui, A., Greene, S., Fuchs, A., Rodriguez, M., Pansch, N., Gonsiorczyk, T., & Casper, P.: Diel variation of CH4 and CO2 dynamics in two contrasting temperate lakes. Inland Waters, 1-15, doi: 10.1080/20442041.2020.1728178, 2020

McGinnis, D. F., Greinert, J., Artemov, Y., Beaubien, S. E., & Wüest, A. N. D. A.: Fate of rising methane bubbles in stratified waters: How much methane reaches the atmosphere?. Journal of Geophysical Research: Oceans, 111(C9), doi: 10.1029/2005JC003183, 2006

Mohr, W., Grosskopf, T., Wallace, D. W., & LaRoche, J.: Methodological underestimation of oceanic nitrogen fixation rates. PloS one, 5(9), e12583, doi: 10.1371/journal.pone.0012583, 2010

Montoya, J. P., Voss, M., Kahler, P., & Capone, D. G.: A Simple, High-Precision, High-Sensitivity Tracer Assay for N (inf2) Fixation. Appl. Environ. Microbiol., 62(3), 986-993. doi: 10.1128/aem.62.3.986-993.1996, 1996

Morana, C., Borges, A. V., Roland, F. A. E., Darchambeau, F., Descy, J. P., & Bouillon, S.: Methanotrophy within the water column of a large meromictic tropical lake (Lake Kivu, East Africa). Biogeosciences, 12(7), 2077-2088, doi: 10.5194/bg-12-2077-2015, 2015

Mountfort, D. O. & Asher, R. A. Effect of inorganic sulfide on the growth and metabolism of Methanosarcina barkeri strain DM. Appl. Environ. Microbiol. 37, 670–675 (1979).

Morana, C., Sarmento, H., Descy, J. P., Gasol, J. M., Borges, A. V., Bouillon, S., & Darchambeau, F.: Production of dissolved organic matter by phytoplankton and its uptake by heterotrophic prokaryotes in large tropical lakes. Limnology and Oceanography, 59(4), 1364-1375.doi:10.4319/lo.2014.59.4.1364, 2014

Morana, C., Darchambeau, F., Roland, F., Borges, A., Muvundja, F., Kelemen, Z., Masilya, P., Descy J-P., and Bouillon, S.: Biogeochemistry of a large and deep tropical lake (Lake Kivu, East Africa: insights from a stable isotope study covering an annual cycle. Biogeosciences, doi:10.5194/bg-12-4953-2015, 2015

Mugidde, R., Hecky, R. E., Hendzel, L. L., & Taylor, W. D.: Pelagic nitrogen fixation in lake Victoria (East Africa). Journal of Great Lakes Research, 29, 76-88, doi: 10.1016/S0380-1330(03)70540-1, 2003

Murase, J., & Sugimoto, A: Inhibitory effect of light on methane oxidation in the pelagic water column of a mesotrophic lake (Lake Biwa, Japan). Limnology and oceanography, 50(4), 1339-1343, doi: doi.org/10.4319/lo.2005.50.4.1339, 2005

Peeters, F., Fernandez, J. E., & Hofmann, H.: Sediment fluxes rather than oxic methanogenesis explain diffusive CH4 emissions from lakes and reservoirs. Scientific reports, 9(1), 1-10, doi: doi.org/10.1038/s41598-018-36530-w, 2019

Petrillo, E., Herz, M. A. G., Fuchs, A., Reifer, D., Fuller, J., Yanovsky, M. J., Simpson, C., Brown, J. W. S., Barta, A., Kalyna, M., & Kornblihtt, A. R.: A chloroplast retrograde signal regulates nuclear alternative splicing. Science, 344(6182), 427-430, doi: 10.1126/science.1250322, 2014

Repine, J. E., Eaton, J. W., Anders, M. W., Hoidal, J. R., & Fox, R. B.: Generation of hydroxyl radical by enzymes, chemicals, and human phagocytes in vitro: detection with the anti-inflammatory agent, dimethyl sulfoxide. The Journal of clinical investigation, 64(6), 1642-1651, doi: 10.1172/jci109626, 1979

Sarmento, H., Romera-Castillo, C., Lindh, M., Pinhassi, J., Sala, M. M., Gasol, J. M., Marassé, C., & Taylor, G. T:. Phytoplankton species-specific release of dissolved free amino acids and their selective consumption by bacteria. Limnology and Oceanography, 58(3), 1123-1135, doi: 10.4319/lo.2013.58.3.1123, 2013

Tang, K. W., McGinnis, D. F., Ionescu, D., & Grossart, H. P.: Methane production in oxic lake waters potentially increases aquatic methane flux to air. Environmental science & technology Letters, 3(6), 227-233, doi: 10.1021/acs.estlett.6b00150, 2016

Tolar, B. B., Powers, L. C., Miller, W. L., Wallsgrove, N. J., Popp, B. N., & Hollibaugh, J. T.: Ammonia oxidation in the ocean can be inhibited by nanomolar concentrations of hydrogen peroxide. Frontiers in Marine Science, 3, 237, doi: 10.3389/fmars.2016.00237, 2016

Vollenweider R.A.: Calculations models of photosynthesis-depth curves and some implications regarding day rate estimates in primary production measurements. Memorie dell'Istituto Italiano di Idrobiologia, 18, 425–457, 1965

Wanninkhof, R.: Relationship between wind speed and gas exchange over the ocean. Journal of Geophysical Research: Oceans, 97(C5), 7373-7382, doi: 10.1029/92JC00188, 1992

Yao, M., Henny, C., & Maresca, J. A.: Freshwater bacteria release methane as a byproduct of phosphorus acquisition. Applied and environmental microbiology, 82(23) 6994-7003, doi: 10.1128/aem.02399-16, 2016

Yvon-Durocher, G., Allen, A. P., Bastviken, D., Conrad, R., Gudasz, C., St-Pierre, A., Thanh-Duc, N., & Del Giorgio, P. A.: Methane fluxes show consistent temperature dependence across microbial to ecosystem scales. Nature, 507, 488-491, doi: 10.1038/nature13164, 2014

**11. Figure captions**

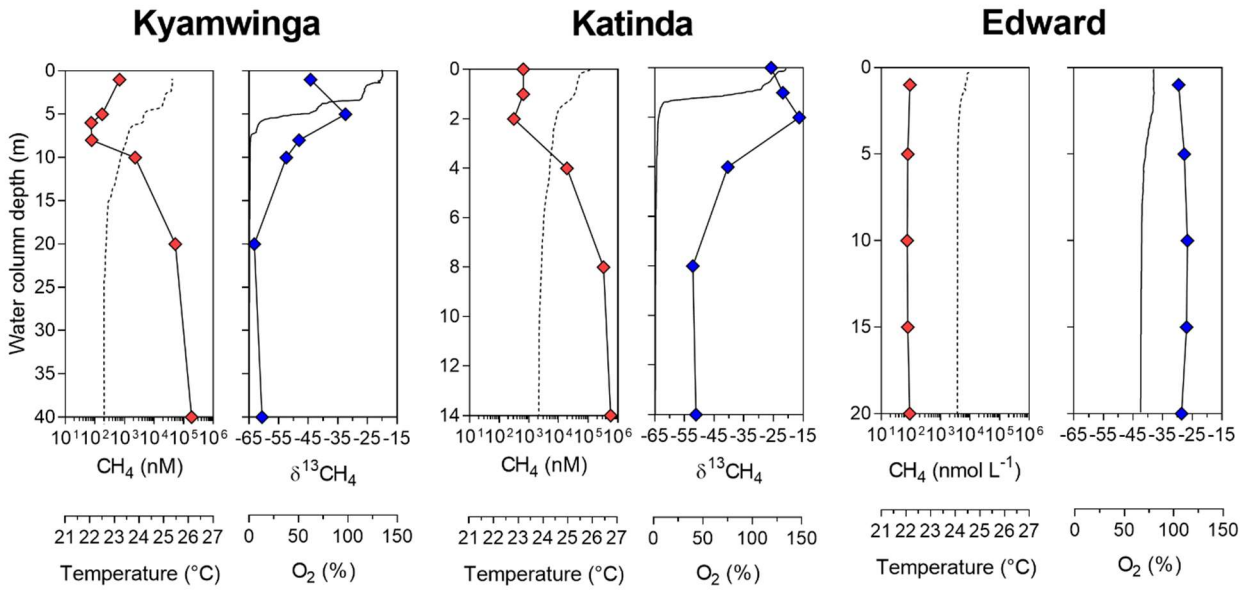
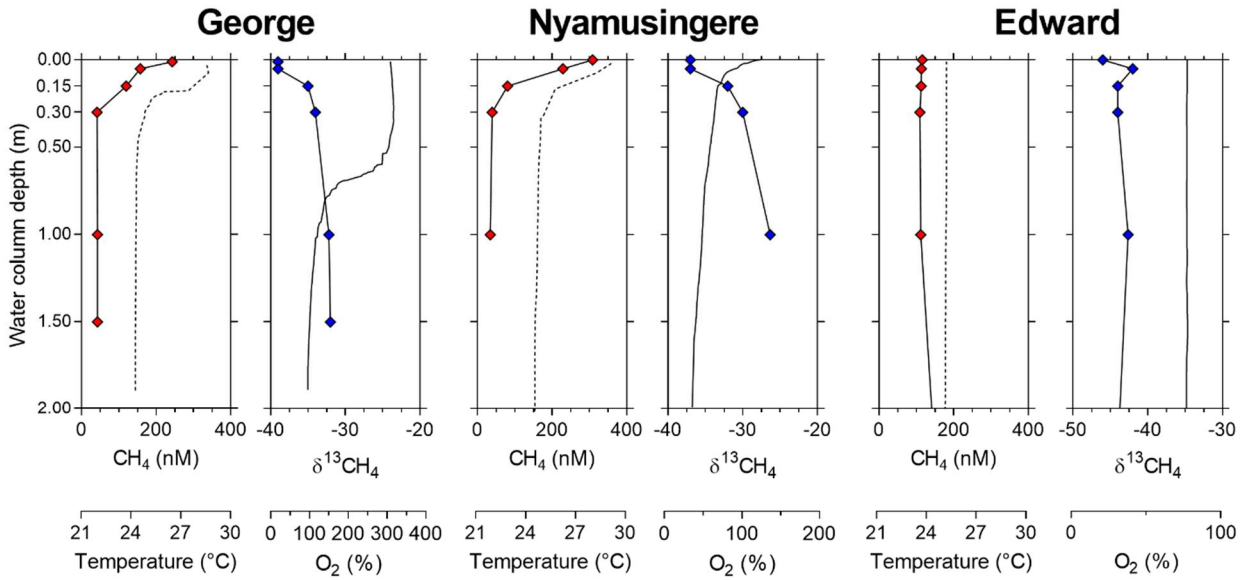



**Figure 1. Depth profile.** Depth profile of the temperature (°C ; dashed line), $CH_4$ concentration (nmol $L^{-1}$ ; red symbols), dissolved oxygen saturation (%, solid line) andstable isotope carbon composition of $CH_4$ ($\delta^{13}$C-$CH_4$, ‰ ; blue symbols) in Lake Kyamwinga (left), Lake Katinda (middle), and Lake Edward (right).




650 **Figure 2. Depth profile, focus on the surface.** Depth profile of the temperature (°C ; dashed line), $CH_4$ concentration (nmol $L^{-1}$ ; red symbols), dissolved oxygen saturation (%, solid line) and stable isotope carbon composition of $CH_4$ ($\delta^{13}$C-$CH_4$, ‰ ; blue symbols) in Lake George (left), Lake Nyamusingere (middle), and the surface waters (0-2 m) of Lake Edward (right).

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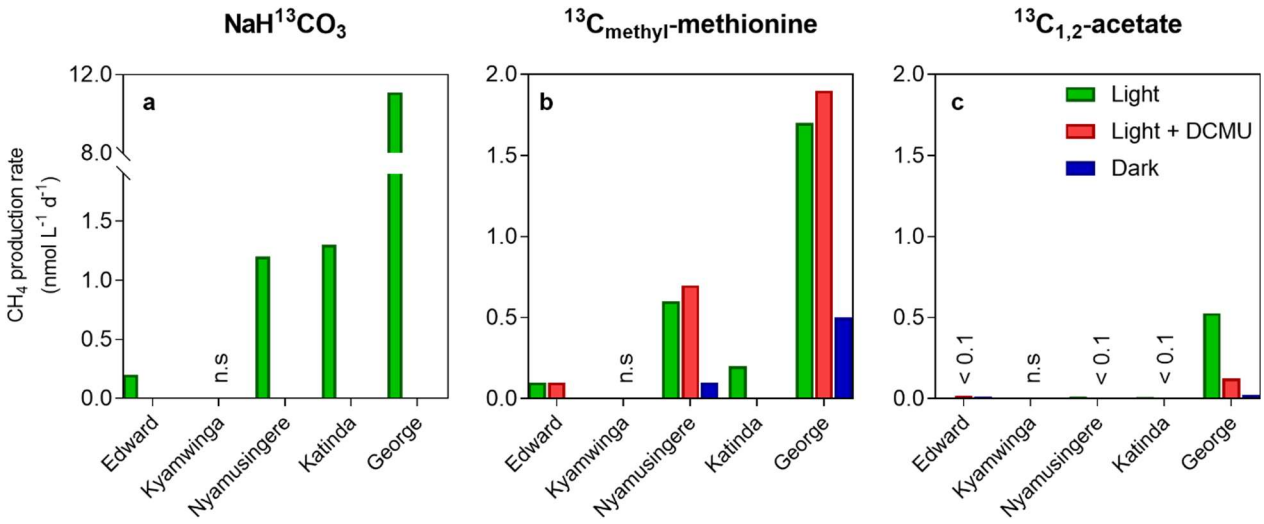


**Figure 3. Tracer experiments show $CH_4$ production in well-oxygenated surface waters.** $CH_4$ production rates (nmol $L^{-1}$ $d^{-1}$) from dissolved inorganic carbon (a), the methyl group of methionine (b) and acetate (c) measured in the surface waters (0.3 m) of a variety of African tropical lakes. Green, grey and dark bars respectively represent rates measured under light, light in presence of a photosynthesis inhibitor (DCMU), or darkness. Values showed for L. Edward, L. George and L. Katinda are the average of 2017 and 2018 sampling campaign measurement. n.s = not significant, < 0.1 = below 0.1 nmol $L^{-1}$ $d^{-1}$.

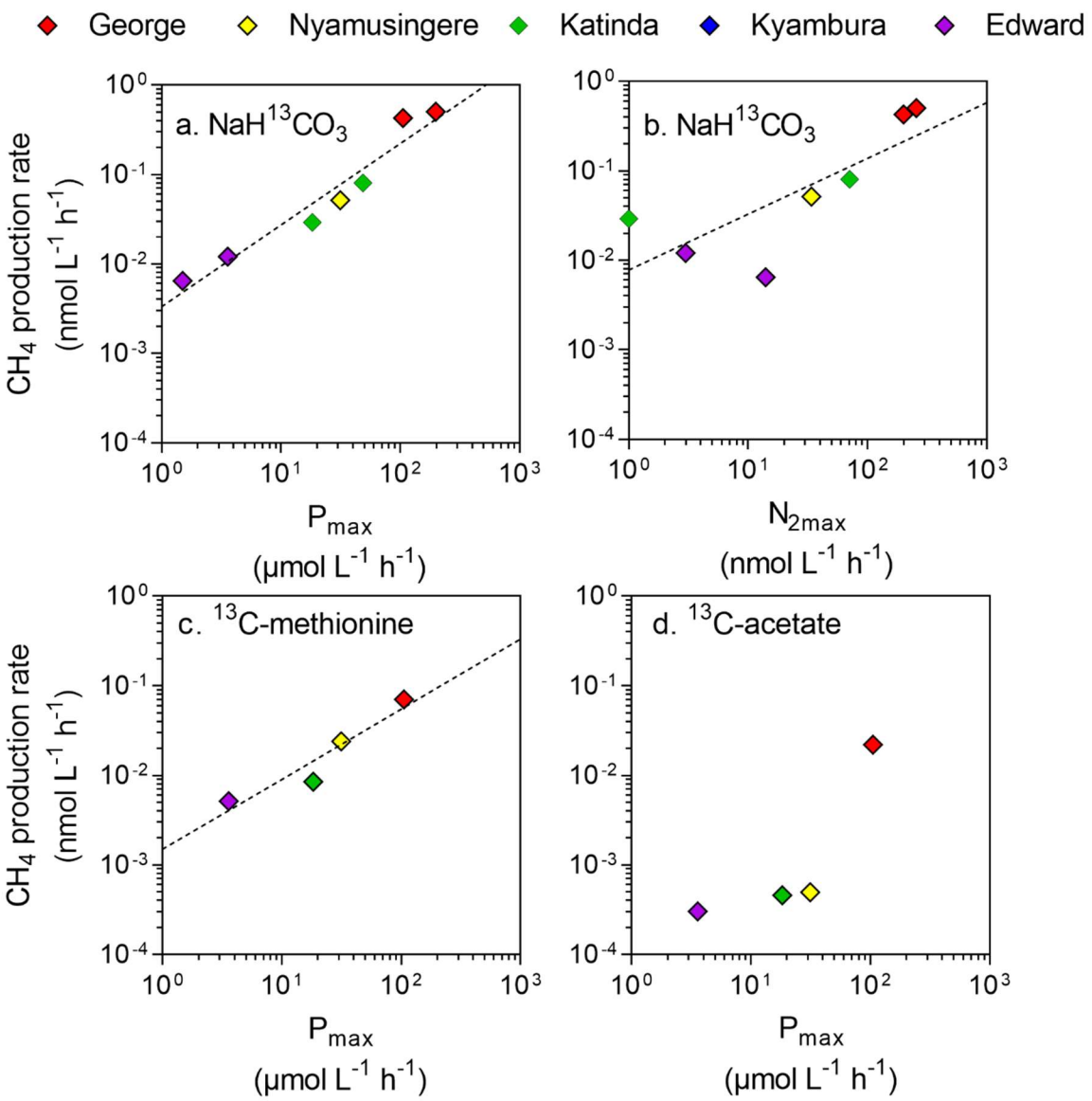


**Figure 4. Direct link between $CH_4$ production and phytoplankton metabolism.** Relationship between the maximum photosynthetic activity ($P_{max}$, µmol C $L^{-1}$ $h^{-1}$) or maximum nitrogen fixation rates ($N_{2max}$, nmol $L^{-1}$ $h^{-1}$) and surface $CH_4$ production rates (nmol C $L^{-1}$ $h^{-1}$) from dissolved inorganic carbon (a, b), methyl group of methionine (c), and acetate (d).

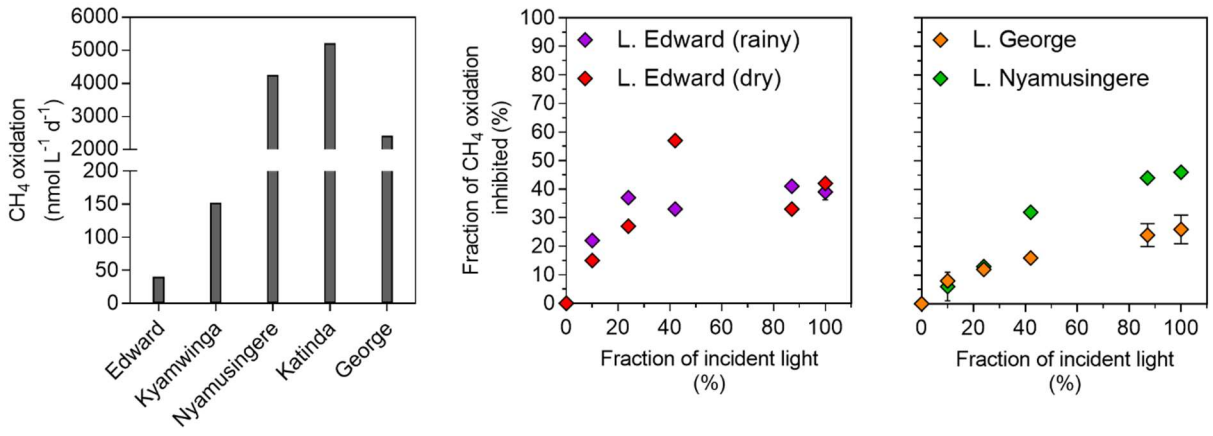


**Figure 5. Light inhibition patterns of $CH_4$ oxidation in surface waters.** Left panel: $CH_4$ oxidation rates (nmol $L^{-1}$ $d^{-1}$) measured in the surface waters (0.3 m) in the dark of a variety of African tropical lakes. Right panel: relationship between illumination (fraction of incident sunlight irradiance, %) and $CH_4$ oxidation inhibition (fraction of $CH_4$ oxidation in the dark inhibited at a given irradiance, %) in Lake Edward, Lake George and Lake Nymusingere. Symbols represent the mean, and error bars represent the maximum and minimum of duplicate experiments

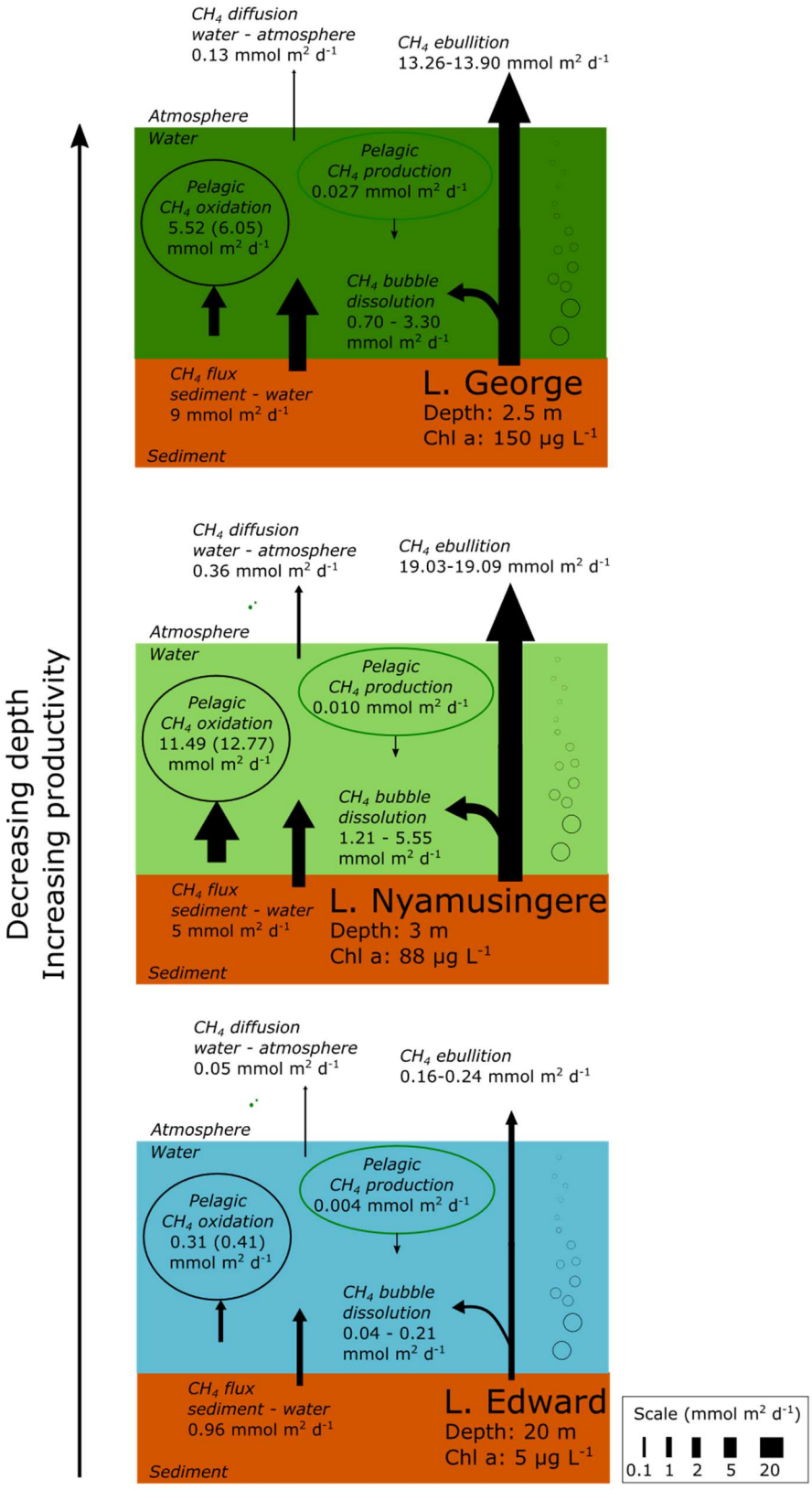
Decreasing depth
Increasing productivity
CH4 diffusion
water - atmosphere
0.13 mmol m2 d-1
CH4 ebullition
13.26-13.90 mmol m2 d-1
Atmosphere
Water
Pelagic
CH4 production
0.027 mmol m2 d-1
Pelagic
CH4 oxidation
5.52 (6.05)
mmol m2 d-1
CH4 bubble
dissolution
0.70 - 3.30
mmol m2 d-1
CH4 flux
sediment - water
9 mmol m2 d-1
L. George
Depth: 2.5 m
Chl a: 150 µg L-1
Sediment
CH4 diffusion
water - atmosphere
0.36 mmol m2 d-1
CH4 ebullition
19.03-19.09 mmol m2 d-1
Atmosphere
Water
Pelagic
CH4 production
0.010 mmol m2 d-1
Pelagic
CH4 oxidation
11.49 (12.77)
mmol m2 d-1
CH4 bubble
dissolution
1.21 - 5.55
mmol m2 d-1
CH4 flux
sediment - water
5 mmol m2 d-1
L. Nyamusingere
Depth: 3 m
Chl a: 88 µg L-1
Sediment
CH4 diffusion
water - atmosphere
0.05 mmol m2 d-1
CH4 ebullition
0.16-0.24 mmol m2 d-1
Atmosphere
Water
Pelagic
CH4 production
0.004 mmol m2 d-1
Pelagic
CH4 oxidation
0.31 (0.41)
mmol m2 d-1
CH4 bubble
dissolution
0.04 - 0.21
mmol m2 d-1
CH4 flux
sediment - water
0.96 mmol m2 d-1
L. Edward
Depth: 20 m
Chl a: 5 µg L-1
Sediment
Scale (mmol m2 d-1)
0.1 1 2 5 20

**Figure 6. Epilimnitic $CH_4$ production is a marginal source of $CH_4$ compared to sedimentary sources and $CH_4$ sinks in several contrasting African lakes.** Summary of the different $CH_4$ flux experimentally measured in L. Edward, L. George and L. Nyamusingere. Values of $CH_4$ oxidation in brackets are values not considering $CH_4$ photoinhibition. Pelagic $CH_4$ production are values determined from $NaH^{13}CO_3$ (~5% final enrichment) and $^{13}$C-acetate (99% final enrichment), as described in the methods section. $^{13}$C-labelling experiment carried out under constant light irradiance. $CH_4$ flux at the water-air and sediment-water interface were determined experimentally as described in the Methods. $CH_4$ bubble dissolution and $CH_4$ ebullition flux were determined using the SiBu-GUI software (Greinert & McGinnis 2009); minimum and maximum represents the values obtained from two extreme bubble-size scenarios considering a release of many small (3 mm diameter) bubbles or fewer large (10 mm) bubbles.